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The Radiochemistry of Uranium



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The Radiochemistry of Uranium

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Issuance Date: March 1962

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FOREWORD

The Subcommittee on Radiochemistry is one of a number of subcommittees working under the Committee on Nuclear Science within the National Academy of Sciences - National Research Council. Its members represent government, industrial, and university laboratories in the areas of nuclear chemistry and analytical chemistry.

The Subcommittee has concerned itself with those areas of nuclear science which involve the chemist, such as the collection and distribution of radiochemical procedures, the establishment of specifications for radiochemically pure reagents, availability of cyclotron time for service irradiations, the place of radiochemistry in the undergraduate college program, etc.

This series of monographs has grown out of the need for up-to-date compilations of radiochemical information and procedures. The Subcommittee has endeavored to present a series which will be of maximum use to the working scientist and which contains the latest available information. Each monograph collects in one volume the pertinent information required for radiochemical work with an individual element or a group of closely related elements.

An expert in the radiochemistry of the particular element has written the monograph, following a standard format developed by the Subcommittee. The Atomic Energy Commission has sponsored the printing of the series.

The Subcommittee is confident these publications will be useful not only to the radiochemist but also to the research worker in other fields such as physics, biochemistry or medicine who wishes to use radiochemical techniques to solve a specific problem.

W. Wayne Meinke, Chairman
Subcommittee on Radiochemistry

INTRODUCTION

This volume which deals with the radiochemistry of uranium is one of a series of monographs on radiochemistry of the elements. There is included a review of the nuclear and chemical features of particular interest to the radiochemist, a discussion of problems of dissolution of a sample and counting techniques, and finally, a collection of radiochemical procedures for the element as found in the literature.

The series of monographs will cover all elements for which radiochemical procedures are pertinent. Plans include revision of the monograph periodically as new techniques and procedures warrant. The reader is therefore encouraged to call to the attention of the author any published or unpublished material on the radiochemistry of uranium which might be included in a revised version of the monograph.

CONTENTS

I.	General Reviews of the Inorganic and Analytical Chemistry of Uranium	1
II.	General Reviews of the Radiochemistry of Uranium	3
III.	Table of Isotopes of Uranium	3
IV.	Review of Those Features of Uranium Chemistry of Chief Interest to the Radiochemist	5
	A. Metallic Uranium	5
	1. Preparation	5
	2. Physical properties	6
	3. Chemical properties	6
	B. Compounds of Uranium	7
	C. The Chemistry of Uranium in Solution	14
	1. Oxidation states	14
	2. Complex ion formation	21
	3. Non-aqueous solutions of uranium	30
	D. Separation of Uranium	39
	1. Precipitation	40
	2. Solvent extraction	60
	Ethers, esters, ketones, and alcohols	63
	Organophosphorus compounds	122
	Amines and quaternary ammonium salts	169
	Carboxylic acids	180
	Chelating agents	182
	3. Ion exchange	202
	Anion exchange	204
	Cation exchange	222
	4. Chromatography	227
	5. Volatilization	231
	6. Electrochemical methods	232
	7. Pyrometallurgical processes	235
	E. Determination of Uranium	236
	1. Counting techniques	236
	2. Sample preparation	241
	3. Activation analysis	248

F. Dissolution of Uranium Samples	252
1. Metallic uranium	252
2. Alloys of uranium	254
3. Compounds of uranium	254
4. Meteorites, minerals, and ores	255
5. Biological samples	256
6. Air dust samples	256
V. Collection of Detailed Procedures	257

The Radiochemistry of Uranium

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III. Table of Isotopes of Uranium^a.

Isotope	Half-Life	Type and Energy of Radiation (Mev)	Method of Preparation
U ²²⁷	1.3 min	α 6.8	Th ²³² (α ,9n)
U ²²⁸	9.3 min	α (~80%) 6.67 EC(~20%)	Th ²³² (α ,8n); ~2% daughter 36 min Pu ²³²
U ²²⁹	58 min	EC(~80%) α (~20%) 6.42	Th ²³² (α ,7n); ~0.1% daughter 20 min Pu ²³³
U ²³⁰	20.8 day	α 5.884 (67.2%) 5.813 (32.1%) 5.658 (0.7%)	Th ²³² (α ,6n); ~15% daughter 17.7 day Pa ²³⁰ ; ~6% daughter 9.0 hour Pu ²³⁴
U ²³¹	4.3 day	EC (99+%) α (5.5x10 ⁻³ %) 5.45	Th ²³² (α ,5n); Pa ²³¹ (d,2n); 3 x 10 ⁻³ % daughter 26 min Pu ²³⁵
U ²³²	74 year	α 5.318 (68%) 5.261 (32%) 5.134 (0.32%)	Th ²³² (α ,4n); daughter 1.31 day Pa ²³² ; daughter 2.85 year Pu ²³⁶ ; U ²³³ (n,2n)

Table of Isotopes of Uranium (Continued)

Isotope	Half-Life	Type and Energy of Radiation (Mev)	Method of Preparation
U ²³³	1.626 x 10 ⁵ year _b	α 4.816 (83.5%)	daughter 27.0 day
		4.773 (14.9%)	Pa ²³³
		4.717 (1.6%)	
		4.655 (0.07%)	
		4.582 (0.04%)	
U ²³⁴ (U _{II})	2.48 x 10 ⁵ year	α 4.768 (72%)	natural radioactivity
		4.717 (28%)	0.0056%; daughter 1.175 min Pa ^{234m} (UX ₂); daughter 6.66 hour Pa ²³⁴ (UZ); daughter 86.4 year Pu ²³⁸ ; U ²³³ (n,γ)
U ^{235m}	26.5 min	IT	daughter 24,360 year Pu ²³⁹
U ²³⁵ (AcU)	7.1 x 10 ⁸ year	α 4.559 (6.7%)	natural radioactivity
		4.520 (2.7%)	0.720%;
		4.370 (25%)	daughter 26.5 min
		4.354 (35%)	U ^{235m} ; daughter
		4.333 (14%)	410 day Np ²³⁵ ;
		4.318 (8%)	daughter 23.7 min
U ²³⁶	2.39 x 10 ⁷ year	α 4.499	Pa ²³⁵
			51% daughter Np ²³⁶ c; daughter 6,580 year Pu ²⁴⁰ ; U ²³⁵ (n,γ)
U ²³⁷	6.75 day	β ⁻ 0.248	daughter 11 min Pa ²³⁷ d; 4x10 ⁻³ % daughter Pu ²⁴¹ ; U ²³⁸ (n,2n); U ²³⁶ (n,γ)

Table of Isotopes of Uranium (Continued)

Isotope	Half-Life	Type and Energy of Radiation (Mev)	Method of Preparation
$U^{238}(U_I)$	4.51×10^9 year	α 4.195	natural radioactivity
U^{239}	23.54 min	β^- 1.21	99.276% $U^{238}(n,\gamma)$; $U^{238}(d,p)$
U^{240}	14.1 hour	β^- 0.36	daughter $\sim 7.6 \times 10^7$ year Pu^{244} ; 2nd order neutron capture on U^{238}

^a Data concerning half-lives, radiations and branching ratios, unless otherwise noted has been obtained from the "Table of Isotopes" by D. Strominger, J. M. Hollander and G. T. Seaborg, Reviews of Modern Physics, 30, No. 2, Part II, April, 1958. This compilation may be consulted for more complete information on the isotopes and for references to the original literature.

^b Ya. P. Dokuchayev and I. S. Osipov, Atomnaya Energiya, 6, 73 (1959).

^c J. E. Gindler and R. K. Sjoblom, J. Inorg. Nuclear Chem., 12, 8 (1959).

^d The half-life of Pa^{237} has been reported recently to be 39 ± 3 min, K. Takahashi and H. Morinaga, Nuclear Physics, 15, 664 (1960).

IV. Review of Those Features of Uranium Chemistry of Chief Interest to the Radiochemist.

A. Metallic Uranium

1. Preparation. Uranium metal may be prepared by several methods:¹ the reduction of uranium oxides with carbon in an arc-melting furnace; reduction of uranium oxides with magnesium, aluminum, calcium or calcium hydride; the reduction of uranium halides with alkali or alkaline-earth metals; electrolytic reduction of uranium halides; and the thermal decomposition of uranium iodide.

2. Physical properties. Metallic uranium exists in three allotropic forms:^{2,3} the orthorhombic alpha form, stable below 663°C; the tetragonal beta form which exists between 663°C and 770°C; and the body-centered cubic form which exists at higher temperatures (> 770°C). The physical properties of the metal as compiled by Grainger⁴ are given in Table I. Because of the method of preparation, impurities may be contained in the metal which alter its properties. Also, a number of the physical characteristics depend upon anisotropic and structural effects, eg. thermal expansion. Therefore, if physical properties are pertinent to an experiment or design, it is best to determine them individually for the metal used.

The changes wrought in metallic uranium by radiation and thermal cycling may be considerable. The results of reactor radiation on the metal are: dimensional instability, surface roughening and pimpling, warping, high hardness, extreme brittleness, cracks and porosity, broadened x-ray diffraction lines, and decreased thermal and electrical conductivity.³ Thermal cycling growth is similar in many respects to that caused by radiation damage. However, differences exist, the fundamental difference being in the mechanism of growth. (The reader is directed to reference 3 for more detailed discussion of this subject.)

3. Chemical properties. Uranium is a highly reactive metal. A potential of +1.80 volts for the half-cell reaction, $U \rightarrow U^{+3} + 3e$, places it below beryllium and above hafnium and aluminum in the electromotive force series.⁵ The metal forms intermetallic compounds with Al, Be, Bi, Co, Cu, Ga, Au, Fe, Pb, Mn, Hg, Ni, Sn, Ge, In, Ir, Pd, Pt, Tl, and Zn;² solid solutions with Mo, Ti, Zr, and Nb.² It reacts at varying temperatures with H_2 , B, C, Si, N_2 , P, As, O_2 , S, Se, F_2 , Cl_2 , Br_2 , I_2 , H_2O , $HF(g)$, H_2S , NH_3 , NO, $HCl(g)$, N_2O_4 , CH_4 , CO, CO_2 .^{1,2} In air, at room temperature, massive uranium tarnishes to form a yellow and eventually a black oxide

coating. Finely divided powder may burn spontaneously. In boiling water, massive uranium corrodes slowly with the formation of uranium dioxide and hydrogen. The reaction products with steam are uranium oxide and hydride. The dissolution of uranium metal is discussed in section IV-F.

Table I. Physical Properties of Uranium Metal^a

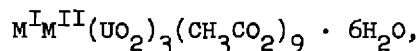
Density (high purity)	19.05 ± 0.02 gm/cm ³
Density (industrial uranium)	18.85 ± 0.20 gm/cm ³
Melting point	1.132 ± 1°C.
Boiling point	3,818°C.
Heat of fusion	4.7 kcal/mole
Vapor pressure (1,600°C.)	10 ⁻⁴ mm
Thermal conductivity (70°C.)	0.071 cal/cm-sec-°C.
Electrical resistivity (25°C.)	35 x 10 ⁶ ohm/cm ³
Mean coefficient of linear thermal expansion (random orientation 25-100°C.)	16 x 10 ⁻⁶ /°C.
Specific heat (25°C.)	6.65
Enthalpy (25°C.)	1,520 cal/mole
Entropy (25°C.)	12.0 cal/mole/°C.

^a L. Grainger, reference 4.

IV-B. Compounds of Uranium

Uranium combines with most elements to form a large number and variety of compounds. "Gmelins Handbuch der Anorganischen Chemie,"⁶ which surveys the literature through the year 1935, describes several hundred compounds. Katz and Seaborg² describe some of the more recently prepared compounds, principally of organic character, such as chelates, alkoxides, amides, mercaptides, and π -cyclopentadienyl compounds.

The oxidation states of uranium in the combined form vary from II to VI. Divalent uranium compounds reported are UO and US. Trivalent uranium compounds are more numerous and include the hydride, nitride, sesquisulfide, halides and borohydride. Uranium (III) sulfate $UH(SO_4)_2$ has also been reported.⁷ A large number of tetravalent compounds are known varying in complexity from the oxide and simple binary salts to more complicated organic structures. Complex salts such as $3(CN_3H_6)_2CO_3 \cdot U(CO_3)_2 \cdot 4H_2O$ and $2(NH_4)_2C_2O_4 \cdot U(C_2O_4)_2 \cdot 6H_2O$ form an important group of uranium (IV) compounds. Complex salts are formed also with halide, sulfite, sulfate, and phosphate ions. Inorganic compounds of pentavalent uranium are UF_5 , UCl_5 , $UCl_5 \cdot SOCl_2$,⁸ $UCl_5 \cdot PCl_5$,⁶ and $UF_5 \cdot xHF$.⁶ $UOCl_3$ has been reported as an intermediate compound in the chlorination of uranium oxides with carbon tetrachloride.¹ Uranium (V) alkoxides have been prepared.² Also, the compounds $(C_5H_6N)_2 UOCl_5$ and $UOCl_3 \cdot EtOH$ have been reported.² Hexavalent uranium is represented by UF_6 , UCl_6 , UO_3 , uranates, and uranyl (UO_2^{2+}) compounds. Uranyl compounds are the most numerous uranium compounds and vary in type from simple salts to complex organic arrangements. Complex salts are formed with halide, iodate, nitrate, carbonate, cyanide, acetate, oxalate, sulfate, phosphate, arsenate, chromate and vanadate ions. Triple acetate salts of the form



where M^I is an alkali metal (Li, Na, or K) and M^{II} is a divalent metal (Mg, Ni, Zn, etc.), are used in analytical separations of uranium. Addition compounds, such as $UO_2(NO_3)_2 \cdot 2CH_3COC_4H_9$, represent a large number of uranyl compounds.

Uranates and peruranates are important in the analytical chemistry of uranium. Uranates have the general formula $xM^I O \cdot yUO_3$ or $xM^{II} O \cdot yUO_3$. They may be prepared by different methods.^{6,10} However, in usual analytical procedures, they are precipitated

from a uranyl solution by the addition of a soluble metal hydroxide, NH_4OH , NaOH , $\text{Ca}(\text{OH})_2$, etc. The uranates are insoluble in water but dissolve in acids.

Peruranates⁶ are formed when uranyl solutions containing hydrogen peroxide are made alkaline. The composition of the peruranates depends upon the concentration of the alkali and peroxide. The following groups have been identified: $\text{M}_2\text{U}_2\text{O}_{10} \cdot x\text{H}_2\text{O}$, $\text{M}_2\text{UO}_6 \cdot x\text{H}_2\text{O}$, $\text{M}_6\text{U}_2\text{O}_{13} \cdot x\text{H}_2\text{O}$, and $\text{M}_4\text{UO}_8 \cdot x\text{H}_2\text{O}$. The peruranates are generally soluble in water. The least soluble are those of the $\text{M}_2\text{U}_2\text{O}_{10} \cdot x\text{H}_2\text{O}$ group. The peruranates are soluble in dilute mineral acids.

Table II lists a number of uranium compounds together with their behavior in different solvents. The compounds listed are primarily binary compounds or simple salts. The order in which they appear is the order in which they may be found in "Gmelins Handbuch der Anorganischen Chemie." ⁶

Table II. Uranium Compounds and Their Solvents.

Compound	Solvent
UH_3	s. HNO_3 (vigorous), conc. HClO_4 , hot conc. H_2SO_4 , a. $+\text{H}_2\text{O}_2$ i. alk., liq. NH_3
UO_2	s. HNO_3 , aq. reg., conc. H_2SO_4 , slowly converted to U(IV)-salts in hot fum. HCl
U_3O_8	s. HNO_3 ; heated to redness U_3O_8 is only v. sl. s. dil. HCl and H_2SO_4 , more s. conc. a., s. hot conc. H_2SO_4 ; HF forms s. UO_2F_2 and i. UF_4
UO_3	s. mineral a.
$\text{UO}_2 \cdot x\text{H}_2\text{O} [\text{U}(\text{OH})_4 \cdot (x-2)\text{H}_2\text{O}]$	s. dil. a.
$\text{U}_3\text{O}_8 \cdot x\text{H}_2\text{O}$	s. a.
$\text{UO}_3 \cdot 2\text{H}_2\text{O}$	s. a., converted to $\text{UO}_3 \cdot \text{H}_2\text{O}$ in boiling H_2O

Table II. - Continued

Compound	Solvent
$\text{UO}_3 \cdot \text{H}_2\text{O}$, $[\text{H}_2\text{UO}_4, \text{UO}_2(\text{OH})_2]$	s. a., warm conc. $\text{UO}_2(\text{NO}_3)_2$ soln.
$\text{UO}_4 \cdot 2\text{H}_2\text{O}$	solubility in H_2O : 20°C -.0006 g/100 ml, 90°C -.008 g/100 ml; d. HCl; alk. hydroxides form UO_3 and s. peruranates
U_3N_4	s. HNO_3 ; i. conc. HCl, H_2SO_4
$\text{UO}_2(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ^a	solubility in H_2O : 0°C -170.3 g/100 ml, 60°C -soluble in all proportions; s. al., ether, acetone, dil. a.
UF_3	s. h. HClO_4 , h. HNO_3 , h. H_2SO_4 , H_3BO_3 + mineral a.
UF_4	i. H_2O ; s. fum. HClO_4 , HNO_3 + H_3BO_3 ; metathesized to U(IV)-hydroxide by heating with NaOH
UF_5 , U_2F_9 , U_4F_{17}	d. H_2O forms s. UO_2F_2 and i. UF_4
UF_6	s. H_2O -vigorous reaction, CCl_4 , CHCl_3 ; v. s. $\text{C}_2\text{H}_2\text{Cl}_4$; d. alcohol, ether
UO_2F_2	s. H_2O , alcohol; i. ether, amyl alcohol
UCl_3	s. H_2O , HCl, glac. acetic a.; i. CCl_4 , CHCl_3 , acetone, pyridine
UCl_4	s. H_2O , $\text{C}_2\text{H}_5\text{OH}$, acetone, ethyl acetate, ethyl benzoate; i. ether, CHCl_3 , benzene
UCl_5	s. H_2O (d.); absolute alcohol, ethyl benzoate, trichloroacetic acid, ethyl acetate, benzonitrite, CS_2 , SOCl_2
UCl_6	d. H_2O ; s. CCl_4 , CHCl_3
UOCl_2	s. H_2O
UO_2Cl_2	s. H_2O : 18°C -320 g/100 ml; s. alcohol, ether
UBr_3	s. H_2O

^a di- and tri-hydrates are also well established.

Table II. - Continued

Compound	Solvent
UBr_4	s. H_2O , acetone, methyl- and ethyl-acetate, pyridine; i. ether
UOBr_2	s. H_2O
UO_2Br_2	s. H_2O , alcohol, ether
UI_3	s. H_2O
UI_4	s. H_2O
$\text{UO}_2(\text{IO}_3)_2 \cdot \text{H}_2\text{O}$	v. sl. s. H_2O : 18°C - α form, 0.1049 g/100 ml, β form, 0.1214 g/100 ml; cold ppt. s. HNO_3 and H_3PO_4 , i. a. after previously heating to boiling temp.; s. alk. carbonates
US	v. difficultly s. conc. HCl , dil. HNO_3
U_2S_3	+0 aq. reg., conc. HNO_3
US_2	d. steam, HNO_3 ; s. hot conc. HCl
UO_2S	sl. s. H_2O ; s. dil. a., alcohol, $(\text{NH}_4)_2\text{CO}_3$; i. absolute alcohol
$\text{UO}_2\text{SO}_3 \cdot 4\text{H}_2\text{O}$	i. H_2O ; s. aq. or alcoholic SO_2 solution
$\text{U}(\text{SO}_4)_2 \cdot 9\text{H}_2\text{O}$	s. dil. H_2SO_4
$\text{U}(\text{SO}_4)_2 \cdot 8\text{H}_2\text{O}$	hydrolyzes in H_2O with separation of basic sulfate, $\text{UOSO}_4 \cdot 2\text{H}_2\text{O}$; s. dil. mineral a., acetic a.
$\text{U}(\text{SO}_4)_2 \cdot 4\text{H}_2\text{O}$	hydrolyzes in $\text{H}_2\text{O}(\text{d.})$; s. dil. H_2SO_4 , HCl
$\text{UOSO}_4 \cdot 2\text{H}_2\text{O}$	s. a.
$\text{UO}_2\text{SO}_4 \cdot 3\text{H}_2\text{O}$	s. H_2O : 15.5°C - 20.5°C g/100 ml, 100°C - 22.2°C g/100 ml; s. mineral a.
$\text{UO}_2\text{SO}_4 \cdot \text{H}_2\text{O}$	s. H_2O
USe_2	ignites with HNO_3 ; chemical properties similar to US_2
UO_2Se	d. H_2O ; s. cold HCl - forms UO_2Cl_2

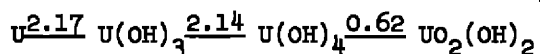
Table II. - Continued

Compound	Solvent
	and H ₂ Se; reacts violently with HNO ₃ - Se is first formed and is then oxidized
UO ₂ SeO ₃ · 2H ₂ O	1. H ₂ O; s. HCl
UB ₂	s. aq. reg., HNO ₃ , HF
UB ₄	s. cold HF, cold HCl, HNO ₃ , conc. H ₂ O ₂ , reduces conc. H ₂ SO ₄
UB ₁₂	1. hot conc. HCl, HF; slowly s. hot conc. H ₂ SO ₄
U(BH ₄) ₄	d. H ₂ O, alcohol
UC ₂	d. H ₂ O, dil. HCl, dil. HNO ₃ , dil. H ₂ SO ₄ ; reacts vigorously with heated conc. a.
UO ₂ CO ₃	s. a.
UO ₂ (HCO ₂) ₂ · H ₂ O	s. H ₂ O: 15°C-420 g/100 ml; s. methyl alcohol; sl. s. formic a.; 1. ethyl alcohol, ether, acetone, CS ₂ , CCl ₄ , CHCl ₃ , benzene, petroleum ether
UO ₂ (HCO ₂) ₂ · UO ₃ · 3H ₂ O	less s. H ₂ O than neutral salt; more s. formic a. than neutral salt
UO ₂ (CH ₃ CO ₂) ₂ · 2H ₂ O	s. H ₂ O: 15°C-7.694 g/100 ml; v. s. alcohol; 1. ether
U(C ₂ O ₄) ₂ · 6H ₂ O	1. H ₂ O, dil. a.; s. warm conc. HCl, conc. HNO ₃
UO ₂ C ₂ O ₄ · 3H ₂ O	sl. s. H ₂ O: 14°C-0.8 g/100 ml, 100°C- 3.3 g/100 ml; s. mineral a., H ₂ C ₂ O ₄ and alk. oxalate solutions
U(C ₄ H ₄ O ₆) ₂ · 2H ₂ O	1. H ₂ O, organic solvents; s. tartaric a., tartrates, conc. a.
UO ₂ (C ₄ H ₄ O ₆) · 4H ₂ O	sl. s. H ₂ O: 17°C-3.28 g/100 cc solution
UO ₂ (CNS) ₂ · 8H ₂ O	s. H ₂ O, ethyl and amyl alcohol, acetone, ether

Table II. - Continued

Compound	Solvent
USl_2	1. cold or hot conc.: HCl, HNO_3 , H_2SO_4 , aq. reg.; s. conc. HF; converted to silicate and uranate by molten alk. and alk. carbonates at red heat
U_3P_4	d. boiling conc. HNO_3 , aq. reg., alk. hydroxide
$U(H_2PO_2)_4 \cdot xH_2O$	1. H_2O , dil. a.; s. conc. a., 50% H_3PO_2
$UO_2(H_2PO_2)_2$	1. H_2O , dil. H_2SO_4 ; s. HNO_3
$U(HPO_3)_2 \cdot 4H_2O$	1. H_2O , dil. a.; s. conc. a., 50% H_3PO_3
UO_2HPO_3	1. H_2O , dil. a.; s. conc. a.
$U_3(PO_4)_4$	1. H_2O ; attacked by a., esp. HNO_3
$UH_2(PO_4)_2 \cdot 2H_2O$	s. conc. HCl
$(UO_2)_3(PO_4)_2 \cdot xH_2O$	1. H_2O , acetic a.; s. mineral a.
$UO_2HPO_4 \cdot xH_2O$	1. H_2O ; s. mineral a., xs. $(NH_4)_2CO_3$
UP_2O_7	1. H_2O , cold a.
$(UO_2)_2P_2O_7 \cdot 5H_2O$	1. H_2O , alcohol, ether; s. xs. $Na_4P_2O_7$, HNO_3
$U(PO_3)_4$	1. H_2O , HCl, HNO_3 , H_2SO_4
$UO_2(PO_3)_2$	
U_3As_4	s. HNO_3
$U_3(AsO_4)_4$	s. HNO_3
$UH_2(AsO_4)_2 \cdot 3H_2O$	s. HCl
$UH_2(AsO_4)_2 \cdot 2H_2O$	1. H_2O ; s. dil. a., esp. arsenic a.
$UO_2HASO_4 \cdot 4H_2O$	1. H_2O , acetic a.
$(UO_2)_2As_2O_7$	
$5UO_2 \cdot 3Sb_2O_5 \cdot 15H_2O$	s. aq. reg., hot conc. HCl, d. HNO_3
$UO_3 \cdot V_2O_5 \cdot xH_2O$	
$UO_3 \cdot V_2O_5 \cdot H_2O$	
$2UO_3 \cdot V_2O_5$	
$UO_2CrO_4 \cdot xH_2O$	s. H_2O

Basic solution:



Tripositive uranium, U^{+3} . Evidence for the existence of U^{+3} comes from the reversibility of the U(III)/U(IV) couple. Solutions may be prepared by the dissolution of a uranium trihalide or by the electrolytic reduction of a uranium (IV) or (VI) solution. Chloride, bromide, iodide, perchlorate and sulfate solutions of uranium(III) have been reported.¹¹ They are deep red in color and unstable, with oxidation of U^{+3} to U^{+4} occurring and hydrogen being evolved. Strongly acidified solutions¹² or those kept at low temperatures¹⁰ appear to be more stable.

Tetrapositive uranium, U^{+4} . The existence of the U^{+4} ion in solution has been confirmed by measurement of the acid liberated on dissolving UCl_4 ¹³ and by solvent extraction studies of U(IV) with thenoyltrifluoroacetone^{14a} and acetylacetone.^{14b} Uranous solutions may be prepared by dissolution of a water-soluble salt: the chloride, bromide, iodide, or sulfate; by dissolution of uranium or a uranium compound in an appropriate solvent, e.g., uranium metal in sulfuric or phosphoric acid; or by reduction of a uranyl solution by chemical, electrochemical or photochemical means. The solutions are green in color. They are stable in the absence of air but are oxidized by oxygen. Uranium(IV) undergoes hydrolysis with evidence in the first stages for the formation of the mononuclear species, UOH^{+3} .^{13,15-18} Polymeric species also are formed which apparently are not in equilibrium with the monomer.^{13,15,19} Hietanen¹⁷ found that in addition to the monomeric species, a polymer of the type $U[(OH)_3U]_n^{4+n}$ could account for the hydrolysis of uranium(IV) to good approximation. Table III, based primarily upon the data compiled by Bjerrum, Schwarzenbach, and Sillén,²⁰

Table III. Hydrolysis of U^{4+} and UO_2^{2+} Ions^a

Ion	Method	T	Medium	Log of equilibrium constant, remarks	Reference	
U^{4+}	magnet	20	var	*p_2 -2.30	21	
	sp	25	$C(NaClO_4)$	*K_1 -1.63(C=2), -1.56(C=1), -1.50(C=0.5)	13	
	sp	10-43	$0.5(NaClO_4)$	*K_1 -1.90(10°), -1.47(25°), -1.00(43°)	22	
	sp	10-43	→ 0	*K_1 -1.12(10°), -0.68(25°), -0.18(43°)	13,21	
				$\Delta^*H_1 = 11.7$, $\Delta^*S_1 = 36(25^\circ)$		
	sp	15-25	$0.19(HClO_4)$	*K_1 -1.38(15.2°), -1.12(24.7°)	16	
		25	0 corr	$\Delta^*H_1 = 10.7$, $\Delta^*S_1 = 33$, $\Delta S_1 = 52$	16	
	gl, red	25	$3(Na)ClO_4$	*K_1 -2.0, $^*p_{3n,n+1}$ -1.2---3.4n	17	
	sol	25	0 corr	K_{S_5} -3.77["U(OH) ₄ (S)"]	23	
	sp	25	$2(ClO_4^-)$	*K_1 -1.68, *K_1 -1.74 $^*K_1 : U^{4+} + H_2O \rightleftharpoons UOH^{3+} + H^+$ $^*K_1 : U^{4+} + D_2O \rightleftharpoons UOD^{3+} + D^+$ $^*p_2 : U^{4+} + 2H_2O \rightleftharpoons U(OH)_2^{2+} + 2H^+$ $^*p_{3n,n+1} : (n+1)U^{4+} + 3nH_2O \rightleftharpoons U(OH)_3U_1^{n+1} + 3nH^+$	18	
	UO_2^{2+}	qh	25-26	0 corr	*K_1 -4.3	49
		gl	20-25	var	$^*p_{22}$ -5.87	37
sol		?	var	*K_1 -4.50, $^*p_{22}$ -4.95 ev UO_2NO_3OH , $UO_2(NO_3)_3OH^{2-}$	38	
UO_2^{2+}	qh	25	0 corr	*K_1 -4.09	250	
	gl	15	$C[Ba(NO_3)_2]$	$^*p_{22}$ -5.97(C=0.6), -5.72(C=0.06)	39,52,53	
	qh, gl	20	$1(Na)ClO_4$	*K_1 -4.70(?, see ref. 43), ev polyn opx	42	
		20	$1(Na)ClO_4$	$^*p_{2n,n+1}$ 0.30---6.35n	42, cf 43	
		20	$1(Na)ClO_4$	$^*p_{2n,n+1}$ 0.30---6.40n	42, cf 44	
	gl, fp, sp	25?	$0.15(NaClO_4)$	$^*p_{22}$ -5.94, $^*p_{43}$ -14.29 $K_a[U_3O_8(OH)^{2-n}] : -3.55(n=0)$, -6.5(n=1), -7.4(n=2), -11.0(n=3), -11.4(n=4)	40	
	gl, fp, sp	25?	0.1 ClO_4	$^*p_{22}$ -5.94, $^*p_{43}$ -12.90	40	
	p(H ₂ O)	25	$2UO_2(ClO_4)$ + UO_3	ev $(UO_2OH)_2^{2+}$, not UO_2OH^+	44	
	sol	25	0 corr	$^*K_{S_0}$ 6.04, $^*K_{S_1}$ 1.90, K_{S_3} -3.60, K_{S_4} -3.77 *K_1 -4.14, ev polyn opx	48	
	sp	?	var	*K_1 -4.19	50	
	diat	25	$0.1(Na)ClO_4$	*K_1 -4.2, *K_2 -5.20	51	
	gl	25	0 corr	$^*p_{22}$ -5.06, $^*p_{24}$ -1.26	45	
	gl	25	$0.347[Ba(ClO_4)_2]$	*K_1 -5.40, $^*p_{22}$ -5.82	46	
gl	25-40	$0.0347[Ba(ClO_4)_2]$	*K_1 -5.82(25°), -5.10(40°), $\Delta^*H_1=20.8$ $^*p_{22}$ -6.15(25°), -5.92(40°), $\Delta^*H_{p_{22}}=-6.7$	46		
UO_2^{2+}		25	$0.0347[Ba(ClO_4)_2]$	$\Delta^*S_1 = 4.3$, $\Delta^*S_{p_{22}} = -6$	46	
	qh, gl	25	$1(Na)ClO_4$, $0.4UO_2^{2+}$	$^*p_{12}$ -3.66, $^*p_{22}$ -6.02	47	
	qh, gl	25	$3(Na)ClO_4$, $1.4UO_2^{2+}$	$^*p_{12} = 3.68$, $^*p_{22} = 6.3$, $^*p_{43} \times -12.6(qh)$, $^*p_{43} = 12.9(gl)$ $^*p_{2n,n+1} : (n+1)UO_2^{2+} + 2nH_2O \rightleftharpoons$ $UO_2[(OH)_2UO_2]_n^{2+} + 2nH^+$ etc.	47	

^a After J. Bjerrum, G. Schwarzenbach, and L. G. Sillén, reference 20.

Table III (Continued)

TABLE NOTATION

The notation used in this and in the tables on complex ion formation is patterned after that used by Bjerrum, Schwarzenbach, and Sillén, reference 20. The explanation of the table which follows is taken primarily from part II of reference 20 and this should be consulted for further details concerning notation. Part I of reference 20 may be consulted for a description of the various methods by which equilibrium constants are determined.

Column one, "Ion," refers to the central ion M about which the complex is formed.

Column two, "Method," refers to the method by which the constants were measured. The abbreviations used are:

ani ex	anion exchange	pH	pH method, not specified
cfug	centrifuge or ultracentrifuge	pol	polarography
col	calorimetry	prep	preparative work
cond	conductivity	qh	quinhydrone electrode
dist	distribution between two phases	red	emf with redox electrode
emf	emf, not specified	sol	solubility
fp	freezing point	sp	spectrophotometry
gl	glass electrode	X-ray	X-ray diffraction
i ex	ion exchange	ΔG	combination of thermodynamic data
magnet	magnetic susceptibility	Pt	emf measurement with Pt electrode
$p(K_2O)$	partial pressure of substance indicated	$Ag_2O_2O_4$	emf measurement with $Ag_2O_2O_4$ electrode
		?	method not known to compilers

Column three, "T," gives the temperature in °C, "RT" indicates room temperature, and "?" is used if the temperature is unknown to the compilers.

Column four, "Medium," denotes the nature of the medium to which the equilibrium constants refer. The concentrations given in terms of moles per liter or moles per kilogram are not distinguished. Water is the solvent unless otherwise stated.

Symbols used are:

-> 0	constants extrapolated to zero ionic strength.		with the ion shown in parentheses as the inert cation.
0 corr	constants corrected to zero ionic strength by application of some theoretical or empirical formula.	$0(OClO_4^-)$	measurements made at a series of perchlorate concentrations.
0.1	an ionic strength of 0.1 mole per liter.	dil	dilute solution, concentration usually not more than 0.01 mole per liter.
1 $NaClO_4$	constant concentration of the substance stated (1 mole per liter $NaClO_4$).	var	ionic medium varied, and in some cases no special attempt was made to control the ionic strength.
1($NaClO_4$)	ionic strength held constant at value stated (1 mole per liter) by addition of the inert salt indicated.	HCl var	the medium was mainly aqueous HCl at various concentrations.
1($NaClO_4$)	measurements made at a series of ionic strengths (I) with $NaClO_4$ as the inert salt.	EtOH	ethanol as solvent.
1(Na) $OClO_4$	concentration of the anion ($OClO_4^-$) held constant at the value stated (1 mole per liter)	org	various organic solvents.
		50% MeOH	50% methanol-water as solvent.

Column five, "Log of equilibrium constant, remarks." The following conventions are used: " K_1 6" means " $\log_{10} K_1 = 6$ "; " $K_2 > 3$ " means " $\log_{10} K_2 > 3$ "; i.e., the "log" and "=" symbols are omitted. Concentrations [] are usually expressed as moles per liter, but is not distinguished from moles per kilogram. Pressures p are in atmospheres.

The equilibrium constants refer to various types of reactions indicated on the next page.

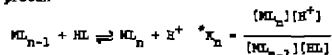
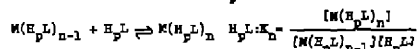
Table III (continued)

1. Consecutive or step-wise constants: K

a. Addition of ligand (L)



b. Addition of protonated ligand (HL) with elimination of proton

c. Addition of protonated ligand (H_pL)

d. Addition of central atom (M)

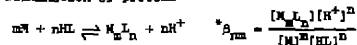
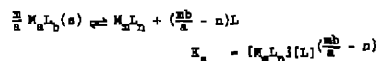
2. Cumulative or gross constants: β

In β_{nm} and ${}^*\beta_{nm}$ the subscripts n and m denote the composition of the complex M_nL_m formed. When $m = 1$, the second subscript is omitted.

a. Addition of central atoms (M) and ligands (L)

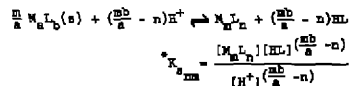


b. Addition of central atoms (M) and protonated ligands (HL) with elimination of protons

3. Solubility constants: K_s a. Solid M_aI_b in equilibrium with free ions in solutionb. Solid M_aI_b in equilibrium with complex M_nL_m and ligand L in solution

In K_{snm} the subscripts n and m denote the composition of the complex M_nL_m formed in solution. When $m = 1$, the second subscript is omitted.

c. Protonated ligand reacts with the elimination of proton



4. Acidic and basic constants:

a. When L is hydroxide (OH^-), HL is water and *K_n is the nth acid dissociation constant for the hydrolysis of a metallic ion.

b. The use of H^+ as the central atom to represent protolytic constants is illustrated by 1.d. above.

c. Other acidic constants are denoted by K_n , followed by parentheses enclosing the formula of the species donating the proton.

d. Basic constants are denoted by K_b follow, if necessary, by parentheses enclosing the formula of the species accepting the proton.

5. Special constants:

a. K (equation)

The equation defines the reaction to which K refers.

b. ${}^*K, {}^*K_{snm}$

The corresponding reaction is given in parentheses after the constant when the latter is first used for a particular ligand or central atom, or the reaction is given immediately below the equilibrium constants for a particular ligand or central atom.

c. (formula)

The formula gives the composition of the complex in terms of the species from which it is formed. Species with negative subscripts are eliminated in the formation of the complex.

d. K_s (formula)

The formula gives the composition of the solid phase in terms of the species with which it is in equilibrium in solution. Species with negative subscripts are eliminated in the formation of the solid.

Heat content and entropy changes are included in column five. ΔH is usually written in kilocalories and ΔS in calories per degree. They are related to the corresponding cumulative equilibrium constants as follows: $\beta_n, \Delta H_n; \beta_{nm}, \Delta H_{nm}; {}^*\beta_{nm}, \Delta^*H_{nm}$. Where the symbol K is used for the equilibrium constant, H or S is given the same superscript or subscript as the corresponding K: e.g., $K_n, \Delta H_n; K_{s0}, \Delta H_{s0}; {}^*K_{snm}, \Delta^*H_{snm}$; etc.

Other abbreviations used in column five are:

ev M_nL_m	evidence for the existence of the complex M_nL_m
opt	complex
cat	cationic
ani	anionic
unch	uncharged
polyn	polymolecular
?	authors' doubt expressed in reference given
(?)	compilers' doubt

Column six, "References," list the references as they are found at the end of this work. A reference such as "6, of 25" indicates that calculations have been made in reference 25 based upon data in reference 6.

summarizes the results of several studies on the hydrolysis of the uranium(IV) ion.

Pentapositive uranium, UO_2^+ . The existence of uranium(V) ion in solution has been confirmed by polarographic measurements.²⁴⁻²⁶ Support for the UO_2^+ ion comes from the reversibility of the U(V)/U(VI) couple²⁷ and from infrared²⁸ and crystallographic^{29,30} studies of uranium and transuranic elements. Solutions of UO_2^+ may be prepared by dissolution of UCl_5 ³¹ or by reduction of a uranyl solution, electrolytically or with U(IV) ions, hydrogen, or zinc amalgam.¹⁵ The formation of U(V) is an intermediate process in the photochemical reduction of U(VI) in a sucrose solution.³² The solutions are unstable and disproportionate to U(VI) and U(IV). The rate of disproportionation is second order in uranium(V) concentration and first order in acid concentration.^{33,15} The UO_2^+ ion is most stable in the pH range of 2 to 4.¹⁵ It is oxidized to the uranyl ion by molecular oxygen, Fe(III) and Ce(IV).¹⁵

Hexapositive uranium, UO_2^{+2} . A number of physical-chemical measurements as well as crystallographic, infrared and Raman spectra studies support the existence of U(VI) ion as UO_2^{+2} .^{2,10} Uranyl solutions are easily prepared by dissolution of water-soluble salts: the nitrate, fluoride, chloride, bromide, iodide, sulfate, and acetate. Other water-soluble uranyl salts include those of other organic acids: the formate, propionate, butyrate, and valerate; and certain double salts such as potassium uranyl sulfate, sodium uranyl carbonate, sodium uranyl chromate, etc. Uranyl solutions may be prepared also by dissolution of a uranium(VI) compound in an appropriate solvent, by dissolution of a lower valence uranium compound in an oxidizing medium, or by oxidation of lower valence uranium ions already in solution. Uranyl solutions are yellow in color. They are the most stable of uranium solutions. As indicated in preceding paragraphs, the

uranyl ion may be reduced by reducing agents³⁴ or by electrochemical or photochemical means. The degree of dissociation of uranyl salts in aqueous solution varies. Uranyl perchlorate is apparently completely dissociated;¹⁰ whereas, uranyl fluoride is undissociated and tends to form dimers (see section on complex ion formation - IV-C2).^{35,36} Hydrolysis of the uranyl ion has been the subject of extensive investigation. Considerable evidence has been adduced for the formation of polymeric species of the type $UO_2(UO_3)_n^{2+}$.³⁷⁻⁴⁷ According to Sutton,⁴⁰ formation of polymers beyond the trimer $U_3O_8^{2+}$ is negligible. However, the trimer itself may undergo further hydrolysis with the formation of $U_3O_8(OH)^+$, $U_3O_8(OH)_2$, and eventually anionic species.^{40,41} Ahrlund,⁴² in his original paper, proposed the formation of the monomer $UO_2(OH)^+$ as well as polynuclear species. In a reappraisal of the work, Ahrlund, Hietanen, and Sillén⁴³ stated that there was no certain indication of mononuclear complexes being formed. Rather, the experimental data was explained on the basis that complex ions of the type $UO_2[(OH)_2UO_2]_n^{2+}$ were formed. From the data it was not possible to distinguish between a limited mechanism in which n varied from 1 to 3 or 4 or an unlimited mechanism in which n assumed all integral values. The authors were inclined to prefer the latter. Kraus⁴⁹ suggested that reactions leading to the formation of polymers may have a less positive value of ΔH than the reaction leading to the formation of the monomer. Consequently, the latter process might be identified more readily at high temperatures than at room temperature. This is apparently the case as was shown by Hearne and White⁴⁶ who determined the enthalpy change to be 20.8 kcal/mole for the monomeric reaction (UO_2OH^+ formed) and 6.7 kcal/mole for the dimeric reaction ($U_2O_5^{2+}$ formed). Table III summarizes much of the

data available on the hydrolysis of the uranyl ion. Included in the table are values of the equilibrium constant $*K_1$, the constant for the formation of the monomeric species. This constant has been evaluated by at least seven groups of investigators^{25c,38,42,48-51} exclusive of Hearne and White⁴⁶. The values obtained agree very well ($\log *K_1 = -4.09$ to -4.70). However, the experimental conditions^{25c,38,50} and assumptions^{25c,38} used in some of the evaluations have been questioned.⁴⁶ Also, the re-evaluation of Ahrland's⁴² work already has been mentioned, and Rydberg⁵¹ has proposed an explanation for not detecting polynuclear species in his experiments. None-the-less, one must concur with Rydberg⁵¹ who wrote, "---it seems remarkable that the same constants should be obtained for a fictive mononuclear hydrolysis product with different U(VI) concentrations and so different methods of investigation---."

2. Complex ion formation. The ability of uranium to form complex ions in solutions is of considerable importance in its analytical separation and determination. Hydrolysis, mentioned in the previous section, is but a special case of complex ion formation. Numerous complexes have been reported.³⁴ However, the amount of quantitative data for the various ligands is rather limited and often contradictory.

Tripositive uranium. Evidence has been reported for uranium(III) cupferrate⁵⁵ and uranium(III) chloro⁵⁶ complexes.

Tetrapositive uranium. Inorganic complexes of uranium(IV) which have been recognized through the formation of complex salts include the fluoride, chloride, sulfate, sulfite, and phosphate.³⁴ Table IV lists the equilibrium constants and thermodynamic data available for some of the uranium(IV) complexes in aqueous solution. In addition, a carbonate complex, possibly $U(CO_3)_5^{6-}$, has been found to be stable in solutions of excess carbonate or bicarbonate ions.⁶³

Table IV. Complex Formation with U ⁴⁺ Ions - Inorganic Ligands ^a					
Complexing agent	Method	T	Medium	Log of equilibrium constant, remarks	Reference
Thiocyanate; SCN ⁻	red	20	1(NaClO ₄), 0.6H ⁺	K ₁ 1.49, K ₂ 0.46, K ₃ 0.23	57
	dist	10	2(Na)ClO ₄ , 1H ⁺	K ₁ 1.78, K ₂ 0.52	58
	"	25	"	K ₁ 1.49, K ₂ 0.62	58
	"	40	"	ΔH ₁ = -5.7, ΔH ₂ = -1.8, ΔS ₁ = -10, ΔS ₂ = 9.7	58
	"	40	"	K ₁ 1.30, K ₂ 0.68	58
Phosphate, PO ₄ ³⁻	sol	35	var	ev UAl ₆ (HL) ₆ ⁶⁺ or UAl ₈ (HL) ₈ ¹²⁺	59
Sulfate, SO ₄ ²⁻	dist	25	2(HClO ₄)	*K ₁ 2.53, *K ₂ -0.13	14a
	"	"	"	*K ₁ 2.41, *K ₂ 1.32	14a, of 60
	"	"	"	K ₁ 3.24, K ₂ 2.18[K ₁ (H ⁺)1.125](?)	14a, of 60, 61
	"	10	2(Na)ClO ₄ , 1H ⁺	*K ₁ 2.63, *K ₂ 1.34	58
	"	25	"	*K ₁ 2.52, *K ₂ 1.35	58
	"	40	"	ΔH ₁ = -3.2, ΔH ₂ = 0.9, ΔS ₁ = 0.7, ΔS ₂ = 9.3	58
Fluoride, F ⁻	dist	25	2(NaClO ₄), 1H ⁺	*K ₁ 2.38, *K ₂ 1.38	58
Chloride, Cl ⁻	sp	25	0.5(NaClO ₄)	*K ₁ > 6, *β ₂ > 8	58
	"	25	0 corr	K ₁ 0.85	13
	red	20	1(Na)ClO ₄ , 0.6H ⁺	K ₁ 0.30	57
	dist	10	2(NaClO ₄), 1H ⁺	K ₁ 0.52	58
	"	25	"	K ₁ 0.26; or K ₁ 0.08, K ₂ -0.02	58
	"	40	"	K ₁ 0.18; or K ₁ -0.04, K ₂ -0.06	58
and ex	25	HCl var	ev and opx in >5.5M HCl	62	
Bromide, Br ⁻	red	20	1(Na)ClO ₄ , 0.6H ⁺	K ₁ 0.18	57

^aAfter J. Bjerrum, G. Schwarzenbach, and L. G. Sillén, reference 20.

Column one denotes the complexing ligand (L). The notation is explained following Table III.

Numerous organic complexes are formed with the uranium(IV) ion: the acetate, oxalate, tartrate, malate, citrate, lactate, glycolate, etc.³⁴ However, the amount of quantitative data available on their formation is very meager. Tishkoff⁶⁴ has calculated dissociation constants for acetate complexes on the basis of the oxygenated uranium(IV) ion UO²⁺ being formed. The formation constants measured for acetylacetone, thenoyltrifluoroacetone and ethylenediamine tetraacetic acid complexes are given in Table V.

Pentapositive uranium. Although it appears that uranium(V) complexes should be formed in the reduction of uranium(VI) ions

in complexing media,⁶³ little data is available.

Hexapositive uranium. Inorganic uranium(VI) complexes which have been identified through the formation of crystalline salts include the fluorides, chlorides, nitrates, sulfates, carbonates, cyanides, and phosphates.³⁴ Uranyl solutions with these anions present have been studied. The results are listed in Table VI. A number of discrepancies appear in the data. For example, evidence for some complexing of the uranyl ion with nitrate^{74,76,77} and chloride^{67,76,77,105-107} ions has been reported by some investigators; but a complete lack of evidence has been reported by others.⁵⁰ Day and Powers⁷⁷ have pointed out that the constants calculated by them are concentration constants rather than activity constants. Consequently, the small complexing effect may be caused by the variation of activity coefficients with a change in medium. Other investigators, however, who have corrected their results to apply to pure aqueous solutions have found some complexing to occur with the chloride ion.^{67,105-107}

The type of complex formed between uranyl and fluoride ions also is subject to some question. Ahrlund and co-workers^{102,104} have determined equilibrium constants for the formation of complexes UO_2F^+ , UO_2F_2 , $UO_2F_3^-$, and $UO_2F_4^{2-}$ and found no evidence for the dimerization of UO_2F_2 for uranyl ion concentrations less than 0.1M. Day and Powers,⁷⁷ however, found no evidence for the formation of complexes beyond UO_2F^+ ; and Johnson, Kraus and Young³⁶ have reported the dimerization of UO_2F_2 in solutions not very different from those investigated by Ahrlund, et al.¹⁰⁴

Numerous organic complexes have been reported.³⁴ Much of the quantitative data is summarized in Table VII.

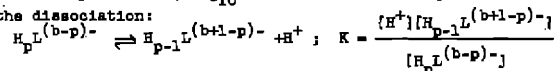
There is often disagreement between different investigators concerning the nature of the complexing ligand. Uranyl-oxalate complexes serve as an example. The oxalate ion, $C_2O_4^{2-}$, has been proposed by some investigators^{146,49} as the complexing ligand; the bi-oxalate ion, $HC_2O_4^-$, by others^{147,116}; and Heidt⁴⁹ has

(Text continues on page 30.)

Table V. Complex Formation with U^{4+} Ions - Organic Ligands ^a							
Complexing agent	Method	T	Medium	pK of H_pL	Log of equilibrium constant, remarks	Reference	
$C_5H_8O_2$ acetylacetone:HL	dist	25	$0.1(ClO_4^-)$	8.82	K_1 8.6, K_2 8.4, K_3 6.4 K_4 6.1	14b, 111	
	"	"	"	8.82	K_1 9.02, K_2 8.25, K_3 6.52 K_4 5.98, β_2 17.27, β_3 23.79 β_4 29.77	14b, of. 112, 113	
	sp	25	0.1		K_1 7.2	114	
$C_8H_5O_2F_3S$ thanyltrifluoro- acetone:HL	sp	25	0.1		K_1 7.2	114	
$C_{10}H_{16}O_6N_2$ ethylenediamine- tetraacetic acid: H_4L	sol	25	$(H_2SO_4)(?)$	H_4L	2.21	K_1 25.6	115
				H_3L^-	2.77	$pK[H_2L^{2-}] \rightleftharpoons H_4L + 2H^+$ 11.91	
				H_2L^{2-}	6.16		
				HL^{3-}	10.26		

^aAfter Bjerrum, Schwarzenbach, and Sillén, reference 20.

The data, with the exception of reference 111, has been compiled by J. Gindler. Column one lists the empirical formula, the name of the ligand, and a formula of the type H_pL which defines the entity L in terms of which the equilibrium constants are expressed. The ligands are placed in order of their empirical formula according to Beilstein's system. Column five lists the pK values ($-\log_{10}K$) of the acid-base equilibria involving the ligands and refer to the dissociation:



The notation is explained following Table III.

Table VI. Complex Formation with UO_2^{2+} Ions - Inorganic Ligands ^a						
Complexing agent	Method	T	Medium		Log of equilibrium constant, remarks	Reference
Cyanoferrate(II), $Fe(CN)_6^{4-}$	sol	25	var		K_{FeO} -13.15	65
Thiocyanate, SCN^-	sp	20	$1(NaClO_4)$		K_1 0.76, K_2 -0.02, K_3 0.44	66
	sp	25	$\rightarrow 0$		K_1 0.93	67
Carbonate, CO_3^{2-}	prep		solid		ev $UO_2L_3^{4-}$	68
	sol	sp?	25?	0 corr	K_3 3.78	69 quoted in ref. 70
	sol	sp	25?	0 corr	$K[UO_2(OH)_2H_2O(s) + CO_2(g) \rightleftharpoons UO_2CO_3(s) + 2H_2O(l)]$	63
					$K_{Fe_2}/K_{FeO} K_1^2(H^+) : UO_2CO_3(s) + 2HCO_3^- \rightleftharpoons UO_2(CO_3)_2^{2-} + CO_2(g) + H_2O$	63
	ΔG	25	0 corr		β_2 14.6, β_3 18.3	63
	sp		-2		K_3 -3.5, ev $(UO_2)_2(OH)_3L^-$	70
	sol	26	$\rightarrow 0$		$K_{Na}[Na_4^+(UO_2L_3)^{4-}]$ -2.8 to -2.0	70
				$K_3/K_{FeO} K_1^2(H^+) : UO_2(CO_3)_2^{2-} + 2HCO_3^- \rightleftharpoons UO_2(CO_3)_3^{4-} + CO_2(g) + H_2O$		
	sol	25	$1(NH_4Cl)$		β_3 22.77	71
	sp		$0.4(KNO_3)$		$^{\dagger}K[UO_2(CO_3)_3^{4-} + H_2O_2 \rightleftharpoons UO_2(CO_3)_2OOH^{3-} + HCO_3^-]$ 2.0	72
	sp		$\rightarrow 0$		$^{\dagger}K$ 2.2	
	sp		var		$K[UO_2(CO_3)_2OO^{4-} + H^+ \rightleftharpoons UO_2(CO_3)_2OOH^{3-}]$ 10.6	73
Nitrate, NO_3^-	sp	25	$5.38(NaClO_4), 2H^+$		K_1 -0.68	74
	sp	25	$7(NaClO_4), 2H^+$		K_1 -0.57	74
	X-ray		solid		ev $UO_2L_3^-$ in $RbUO_2L_3(s)$	75
	qh	20	$1(NaClO_4)$		K_1 -0.3	76
dist	10-40		$2(NaClO_4)$		K_1 -0.52(10°), -0.62(25°), -0.77(40°)	77

Table VI. - Continued

Complexing agent	Method	T	Medium	Log of equilibrium constant, remarks	Reference	
	sp	?	var	no ev cpx: 150-fold excess L	50	
	sp	22-28	org	ev UO_2L_3	76	
	sp		Me_2CO	K_3 3.6 in Me_2CO	79	
	cond	25	$EtOH$	$K_2(?)$ 3.15, $K_3(?)$ 1.39 in $EtOH$	80	
	cond	25	Me_2CO	$K_2(?)$ 3.96, $K_3(?)$ 2.46 in Me_2CO	80	
Phosphate, PO_4^{3-}	sp	25	$HClO_4$	$K[UO_2^{2+} + H_3L \rightleftharpoons UO_2H_{1+x}L^{x+} + (2-x)H^+]$ 1.58	81	
	sol	25	$HClO_4$	$K[UO_2^{2+} + H_3L \rightleftharpoons UO_2H_{1+x}L^{x+} + (2-x)H^+]$ 1.57 $x = 1$ or 2	82,83 cf. 83	
			$HClO_4^?$	$K[UO_2^{2+} + 2H_3L \rightleftharpoons UO_2(H_2L)_2 + 2H^+]$ 1.18		
			$HClO_4^?$	$K[UO_2^{2+} + 3H_3L \rightleftharpoons UO_2(H_2L)_2H_3L + 2H^+]$ 2.30 ^c		
			$HClO_4$	$K_B[UO_2HL(s) + 2H^+ \rightleftharpoons UO_2^{2+} + H_3L]$ -2.85		
			$HClO_4$	$K_B[UO_2HL(s) + xH^+ \rightleftharpoons UO_2H_{1+x}L^{x+}]$ -1.29 $x = 1$ or 2		
			H_3L var	$K_B[UO_2HL(s) + H_3L \rightleftharpoons UO_2(H_2L)_2]$ -1.7		
			H_3L var	$K_B[UO_2HL(s) + 2H_3L \rightleftharpoons UO_2(H_2L)_2H_3L]$ -0.55		
			H_3L dil	$K_B[(UO_2)_3L_2(s) + 6H^+ \rightleftharpoons 3UO_2^{2+} + 2H_3L]$ -6.15		
			H_3L dil	$K_B[(UO_2)_3L_2(s) + H_3L + 3H^+ \rightleftharpoons 3UO_2H_2L^+]$ -2.40		
			H_3L dil	$K_B[(UO_2)_3L_2(s) + 4H_3L \rightleftharpoons 3UO_2(H_2L)_2]$ -2.89		
			H_3L dil	$K_B[(UO_2)_3L_2(s) + 7H_3L \rightleftharpoons 3UO_2(H_2L)_2H_3L]$ 0.53		
			solid:	" $UO_2HL(s)$ " = $UO_2HPO_4(H_2O)_4(s)$		
		sol	25	var	ev $UO_2H_2L^+$, $UO_2H_3L^{2+}$, $UO_2(H_2L)_2$	84,85
		ani ex	20	0 corr	H_2L^- : K_1 3.00, K_2 2.43, K_3 1.90 H_3L : $K_1 < 1.88$, β_2 3.88, β_3 5.23 $K_B[UO_2(H_3L)_3^{2+}]$ -0.53, -2.27, -1.48 $K_B[UO_2(H_3L)_2^{2+}]$ -1.52, -1.18 (H^+ : K_{13} 2.12)	86
		sol	19-20	var	$K_B(UO_2^{2+}NH_4^+L^{3-})$ -26.36 $K_B(UO_2^{2+}K^+L^{3-})$ -23.11 $K_B(UO_2^{2+}HL^{2-})$ -10.67, (H^+ : K_1 12.44, K_{12} 6.71, K_{13} 1.96)	87
		sp	25	var	ev $UO_2H_2L^+$, $UO_2H_3L^{2+}$, $UO_2(H_2L)_2$, $UO_2H_5L_2^+$	88
		sp	25	1.07($NaClO_4$)	$\beta(UO_2^{2+}H_{-1}^+H_3L)$ 1.19 ^a	89,90
		dist	25	1.07($NaClO_4$)	$\beta(UO_2^{2+}H_{-2}^+(H_3L)_2)$ 1.34 ^b $\beta(UO_2^{2+}H_{-2}^+(H_3L)_3)$ 1.01 ^c	89,90
		ani ex	20	0 corr	$K_{1j}[UO_2^{2+} + jH_3L \rightleftharpoons UO_2^{2+}(1-jH_2L^-; jH_3L)^{2+j-1} + (1-j)H^+]$: K_{10} 0.9 ^a , $K_{11} < 1.8$, K_{20} 1.3 ^b , K_{21} 2.4, K_{22} 3.9, K_{30} 1.1, K_{31} 2.5 ^c , K_{32} 4.8, K_{33} 5.3 H_2L^- : K_1 3.0, β_2 5.5, β_3 7.4 H_3L : $K_1 < 1.8$, β_2 3.9, β_3 5.3 (same as K_{11} , K_{22} , K_{33}) (H^+ : K_{13} 2.10)	91
	sp	25	1(ClO_4^-)	K_{1j} : K_{10} 0.7, K_{11} 0.8, K_{20} 0.4, K_{21} 1.4 (H^+ : K_{13} 1.68)	88, cf. 91	
	sp	25	0 Corr(?)	K_{1j} : K_{10} 1.1, K_{11} 1.2, K_{20} 1.3, K_{21} 2.2 (H^+ : K_{13} 2.10)	88, cf. 91	

^a same equilibrium reaction, different notation^b same equilibrium reaction, different notation^c same equilibrium reaction, different notation

Table VI. - Continued

Complexing agent	Method	T	Medium	Log of equilibrium constant, remarks	Reference
Polyphosphate, [F _n O _{3n+1}] ⁽ⁿ⁻²⁾⁻	pH	25	dil	n ~ 5 : K ₁ 3.0, or β ₂ 6.0, or β ₃ 9.5	92
Arsenate, AsO ₄ ³⁻	sol gl	20	var	K ₈ (UO ₂ ²⁺ HL ²⁻)-10.50, K ₈ (UO ₂ ²⁺ L ₁ ⁺ L ³⁻)-18.82 K ₈ (UO ₂ ²⁺ Na ⁺ L ³⁻)-21.87 K ₈ (UO ₂ ²⁺ K ⁺ L ³⁻)-22.60, K ₈ (UO ₂ ²⁺ NH ₄ ⁺ L ³⁻)-23.77 (H ⁺ : K ₁ 11.53, K ₁₂ 6.77, K ₁₃ 2.25)	93
Sulfite, SO ₃ ²⁻	sp		var	ev UO ₂ L ₂ ²⁻ , strong opx	94
	sol	25	(NH ₄) ₂ SO ₃ var	K ₈ -8.59 β ₂ 7.10	95
Sulfate, SO ₄ ²⁻	sp	25	2.65(NaClO ₄), 2H ⁺	*K ₁ 0.70	74
		25	3.5(?), 2H ⁺	K ₁ 1.83[K ₁ (H ⁺)1.125]	74, cf. 61
	qh	20	1(NaClO ₄)	K ₁ 1.70, K ₂ 0.84, K ₃ 0.86 β(UO ₂ ²⁺ L ₂ ²⁻ As ⁻) 3.78 β(UO ₂ ²⁺ L ₂ ²⁻ As ⁻) 4.60	96
	sp	20	1(NaClO ₄)	K ₁ 1.75, K ₂ 0.90	96
	cond	25	0 corr	K ₁ 3.23?	97
	dist	10	2(NaClO ₄)	K ₁ 1.80, K ₂ 0.96[K ₁ (H ⁺)1.01]	77
	dist	25	2(NaClO ₄)	K ₁ 1.88, K ₂ 0.97[K ₁ (H ⁺)1.08]	77
				ΔH ₁ = 2.3, ΔH ₂ = -0.9, ΔS ₁ = 16, ΔS ₂ = 2	
	dist	40	2(NaClO ₄)	K ₁ 1.98, K ₂ 0.93[K ₁ (H ⁺)1.17]	77
	ani ex		var	ev UO ₂ L ₃ ⁴⁻ , U ₂ O ₅ L ₃ ⁴⁻	98
	ani ex	25	var	ev UO ₂ L ₂ ²⁻ , UO ₂ L ₂ ²⁻ , UO ₂ L ₃ ⁴⁻ , U ₂ O ₅ L ₃ ⁴⁻	99
	sp	25	→ 0	K ₁ 2.96, K ₂ -1	67
Fluoride, F ⁻	ani ex	25	HCl var	*K ₁ 1.18	100
				K ₁ 4.32	100, cf. 101
	sp		var	K ₁ 5.5, β ₄ -8	101
	sp	-0	UO ₂ F ₂ var	*K ₄₂ {2UO ₂ F ₂ ⇌ (UO ₂ F ₂) ₂ } 10.18	35
	qh	20	1(NaClO ₄)	K ₁ 4.59, K ₂ 3.34, K ₃ 2.56, K ₄ 1.36 [K ₁ (H ⁺)2.94]	102
	cond	25	0 corr	K ₂ -4.4, ev other opx	103
	dist	10-40	2(NaClO ₄)	*K ₁ 1.74(10°), 1.42(25°), 1.32(40°)	77
		25	2(NaClO ₄)	Δ*H ₁ = -5.14, Δ*S ₁ = -12	77
	dist	25	6(NaClO ₄)	*K ₁ 1.42(C=2), 1.43(C=1), 1.38(C=0.5), 1.57(C=0.25), 1.71(C=0.05)	77
	cfug	0-30	UO ₂ F ₂ var	*K ₄₂ 0.48(0°), 0.85(30°)	36
	qh	20	1(Na)ClO ₄	K ₁ 4.54, K ₂ 3.34, K ₃ 2.57, K ₄ 1.34 [K ₁ (H ⁺)2.93], no ev polyn opx for < 0.1M-UO ₂ ²⁺	102, 104 of 104
Chloride, Cl ⁻	qh	20	1(NaClO ₄)	K ₁ -0.10	76
	sp	20	1(NaClO ₄)	K ₁ -0.30?	76
	pol sp gl	25	0 corr	K ₁ 0.38	105
	dist	10-40	2(NaClO ₄)	K ₁ -0.24(10°), -0.06(25°), 0.06(40°)	77
		25	2(NaClO ₄)	ΔH ₁ = 3.8, ΔS ₁ = 12	77
	ani ex	25	HCl var	ev ani opx in > 0.5M - HCl	106
		25	0 corr	K ₁ -0.1, K ₂ -0.82, K ₃ -1.70	106, cf. 107
	sp	25	→ 0	K ₁ 0.21	67
	sp	?	var	no ev opx : 150-fold excess L ⁻	50
	sp		Me ₂ CO	ev UO ₂ L ⁺ , UO ₂ L ₂ ²⁻ , UO ₂ L ₃ ⁴⁻	108
Perchlorate, ClO ₄ ⁻	sp	25	2-6 ClO ₄ ⁻	no ev opx	109
Bromide, Br ⁻	qh	20	1(NaClO ₄)	K ₁ -0.30	76

Table VI. - Continued

Complexing agent	Method	T	Medium	Log of equilibrium constant, remarks	Reference
	sp	25	→ 0	K_1 -0.20	67
Iodate, IO_3^-	sol	25	0.2(NH ₄ Cl)	K_{HO} -7.01, β_2 2.73, β_3 3.67	110
	sol	60	0.2(NH ₄ Cl)	K_{HO} -6.65, β_2 2.74, β_3 3.44	110

^a After Bjerrum, Schwarzenbach, and Sillén, reference 20.

Data which appeared in the literature prior to the middle of 1957 has been compiled mostly by the above authors. Subsequent data has been compiled by J. Gindler.

The notation is explained following Table III.

Table VII. Complex Formation with UO_2^{2+} Ions - Organic Ligands ^a						
Complexing agent	Method	T	Medium	pK of H_2L	Log of equilibrium constant, remarks	Reference
$\text{C}_2\text{H}_2\text{O}_4$ oxalic acid: H_2L	Pt	25	0 corr	H_2L 1.27 HL^- 4.29	K_1 5.82, K_2 4.74, $\text{H}_2\text{L} : K_1$ 2.57	49
	cat ex	25	0.16HClO ₄	H_2L 1.28 HL^- 3.75	$\text{HL}^- : K_1$ 3.40, K_2 2.56	116
	cat ex	25	1HClO ₄	H_2L 1.28	$\text{HL}^- : K_1$ 2.83, K_2 -1.85	116
	cat ex	25	2HClO ₄	H_2L 1.28	$\text{HL}^- : K_1$ 2.89, K_2 -1.85	116
$\text{Ag}_2\text{C}_2\text{O}_4$		25	0.069		$^{\dagger}K[(\text{UO}_2)_2\text{L}_3^{2-} + 2\text{L}^{2-} \rightleftharpoons (\text{UO}_2)_2\text{L}_5^{6-}]$ 14.42	117
$\text{Ag}_2\text{C}_2\text{O}_4$		25	0.022		$^{\dagger}K$ 4.60, $^{\dagger}K[(\text{UO}_2)_2\text{L}_3^{2-} + \text{L}^{2-} \rightleftharpoons 2(\text{UO}_2\text{L}_2)^{2-}]$ 11.32	117
					$K[2(\text{UO}_2\text{L}_2)^{2-} + \text{L}^{2-} \rightleftharpoons (\text{UO}_2)_2\text{L}_5^{6-}]$ 13.28	
$\text{Ag}_2\text{C}_2\text{O}_4$		25	0.008		$^{\dagger}K$ -0.48	117
	sol	20	var	H_2L 0.97 HL^- 4.19	K_1 6.77, K_2 5.23, β_2 12 $K[(\text{UO}_2)^{2+} + \text{H}_2\text{L} = \text{UO}_2\text{L} + 2\text{H}^+]$ 11.60 $K[(\text{UO}_2)^{2+} + 2\text{H}_2\text{L} = \text{UO}_2\text{L}_2^{2-} + 4\text{H}^+]$ 11.68	118
	sol	20	HClO ₄ , > 1.5H ⁺		$K_8[(\text{UO}_2)^{2+}\text{L}^{2-}(\text{H}_2\text{O})_3]$ -8.66	118
	sol	20	HNO ₃ , > 1.5H ⁺		$K_8[(\text{UO}_2)^{2+}\text{L}^{2-}(\text{H}_2\text{O})_3]$ -8.52	118
	cat ex	25	1HClO ₄		K_1 6.58 graphically, 6.40 analytically, β_2 10.74	116, of 118
	cat ex	25	2HClO ₄		K_1 6.92	116, of 118
	sol	25	→ 0		K_1 6.00, K_2 5.08, 4.44 (preferred value)	119
	sp gl		0.312		K_3 2.1	120
	sp gl		0.05		$K[2\text{UO}_2\text{L} + \text{H}_2\text{O}_2 \rightleftharpoons (\text{UO}_2)_2(\text{OO})^{2-} + 2\text{H}^+]$ -1.62	120
	sp gl		0.312		$K[2\text{UO}_2\text{L}_2^{2-} + \text{H}_2\text{O}_2 \rightleftharpoons (\text{UO}_2)_2(\text{OO})\text{L}_4^{6-} + 2\text{H}^+]$ -3.66	120
$\text{C}_2\text{H}_4\text{O}_2$ acetic acid: HL	emf sp	20	1NaClO ₄	4.59	K_1 2.38, K_2 1.98, K_3 1.98	121
	pol		NaL-HL buffer		K_1 2.63, K_2 2.03, K_3 1.60	64
	cat ex	25	0.16(NaCl)		K_1 2.38, β_3 6.38	122
	ion ex	RT	0.5(NaNO ₃)		K_3 1.22, β_3 5.89	123
$\text{C}_2\text{H}_4\text{O}_3$ glycolic acid: HL	gl	20	1NaClO ₄	3.58	K_1 2.42, K_2 1.54, K_3 1.24	124
	cat ex	25	0.16(NaCl)		K_1 2.78, K_2 1.30	116
	pH	25	0.16-0.19		K_1 2.75, K_2 1.52	116
$\text{C}_2\text{H}_3\text{O}_2\text{Cl}$	emf	20	1NaClO ₄	2.66	K_1 1.44, K_2 0.85, K_3 0.51	125

Table VII. - Continued

Complexing agent	Method	T	Medium	pK of H _p L	Log of equilibrium constant, remarks	Reference
chloroacetic acid:HL	sp	20	1NaClO ₄	2.66	K ₁ 1.38, K ₂ 0.80, K ₃ 0.37	125
C ₂ H ₅ O ₂ N glycine:EL	dist	25	0.45(NaCl)	H ₂ L ⁺ 2.32 pH 1.90	K ₁ 1.43	122
C ₂ H ₆ ON ₂ glycine amide:L	pH	25	0.15-0.25	HL ⁺ 8.06	K ₁ 5.15	126
C ₃ H ₆ O ₃ lactic acid:HL	sp	RT			ev UO ₂ ²⁺ /L cpx in ratio 1/1 pH 3.5	127
	sp,pH	RT			ev UO ₂ ²⁺ /L cpx in ratio 1/1 pH 6; presupposes dimerization	128
C ₃ H ₇ O ₃ N serine:HL	dist	25	0.45(NaCl)	H ₂ L ⁺ 2.21 pH 2.05	K ₁ 0.87	116
C ₄ H ₆ O ₅ malic acid:H ₂ L	sp	RT			ev UO ₂ ²⁺ /malate cpx in ratio 1/1 and 2/1 pH 3.5	127
	sp,pH	RT			ev UO ₂ ²⁺ /malate cpx in ratio 1/1 as dimer acid solution; 3/3 and 3/2 pH ~8	128
C ₄ H ₆ O ₆ tartaric acid:H ₂ L	sp	RT			ev UO ₂ ²⁺ /tartrate cpx in ratio 1/1, 2/1, and 3/1 pH 3.5, 4.6	127
	sp,pH	RT			ev UO ₂ ²⁺ /tartrate cpx in ratio 1/1 as dimer acid solution; 3/3 and 3/2 pH ~8	128
C ₅ H ₈ O ₂ acetylacetone:HL	gl	30	→ 0	8.95	K ₁ 7.74, K ₂ 6.43	129
	gl	10	→ 0	9.10	K ₁ 7.94, K ₂ 6.53	129
	dist	25	0.1(ClO ₄ ⁻)	8.82	K ₁ 6.8, K ₂ 6.3 K[UO ₂ ²⁺ ·L ⁻ + HL ⇌ UO ₂ L(HL) ⁺] 8.7 K[UO ₂ ²⁺ ·2L ⁻ + HL ⇌ UO ₂ L ₂ (HL)] 14.8	51, 111
	sp	25	EtOH		HL : K ₁ 2.43	80
C ₆ H ₈ O ₆ ascorbic acid:(H ₂ L)	sp,pH	20-21	0.1(NaClO ₄)	H ₂ L 4.07 pH 2-3	HL ⁻ : K ₁ 2.48	130
C ₆ H ₈ O ₇ citric acid:H ₃ L	sp	RT	(NaCl)		ev UO ₂ ²⁺ /citrate cpx in ratio 1/1 and 2/1 pH 2-6	131
	pol	30			ev UO ₂ ²⁺ /citrate cpx in ratio 1/1 present as dimer pH 4.6	132
	sp,pH	RT			ev UO ₂ ²⁺ /citrate cpx in ratio 3/3 and 3/2 pH 8	128
	sp,pH	?	var		HL ³⁻ : K ₁ 3.165 pH 4-7	133
	cond				ev UO ₂ HL ⁻ ⇌ UO ₂ L ₂ ²⁻ + H ⁺ pH 4.6 ev UO ₂ ²⁺ /citrate cpx in ratio 2/3 pH 7-9	
	gl	25	0.15(?)	H ₄ L 2.94 H ₃ L ⁻ 4.34 H ₂ L ²⁻ 5.62	HL ³⁻ : K ₁ 8.5	134
C ₆ H ₉ O ₂ N ₃ histidine:HL	pH	25	0.15-0.25	H ₂ L ⁺ 6.17 HL 9.20	K ₁ 7.17	126
C ₇ H ₆ O ₂ salicylaldehyde:HL	sp	?	50% EtOH		K ₁ 1.81 uranyl acetate, acetic acid present, pH 3	135
			var		β ₂ 2.63 uranyl nitrate, pH 5	
C ₇ H ₆ O ₃ salicylic acid:H ₂ L	sp				K ₁ 13.4	136
	dist	25	0.1(NaClO ₄)	H ₂ L 2.82 HL ⁻ 13	HL ⁻ : K ₁ 2.2 K[UO ₂ ²⁺ + HL ⁻ + OH ⁻ ⇌ UO ₂ (HL)(OH)] 12.1	138

Table VII. - Continued

Complexing agent	Method	T	Medium	pK of H_pL	Log of equilibrium constant, remarks	Reference
$C_7H_6O_4$ kojic acid:HL	gl	30	50% dioxan	9.40	K_1 10.1, K_2 7.4	139
$C_7H_6O_6S$ 5-sulfosalicylic acid: H_3L	sp	25	≈ 0.015	H_2L^- 2.86	HL^{2-} : K_1 3.89	140
$C_7H_7O_2N$ salicylamide:HL	sp			8.09	K_1 6.40, K_2 4.97	141
$C_7H_{11}O_2N_3$ histidine methyl ester:L	pH	25	0.15-0.25	H_2L^{2+} 5.38 HL^+ 7.33	K_1 5.76	126
$C_8H_6ON_2$ 8-hydroxycinnoline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ 1.77 HL 8.84	K_1 8.68, K_2 7.16	142
$C_8H_6ON_2$ 5-hydroxyquinoxaline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ < 1 HL 9.29	K_1 8.40, K_2 7.51	142
$C_8H_6ON_2$ 8-hydroxyquinazoline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ 3.30 HL 9.59	K_1 8.99, K_2 7.70	142
$C_8H_8O_3$ methoxybenzoic acid:HL	dist	25	0.1(NaClO ₄)	3.89	$K(UO_2^{2+} + L^- + OH^- \rightleftharpoons UO_2(L)(OH))$ 11.9	138
C_9H_7ON 8-hydroxyquinoline (oxine):HL	gl dist	20 25	50% dioxan 0.3 NaClO ₄ 0.1	H_2L^+ 4.48 HL 10.80	K_1 11.25, K_2 9.64 β_3 23.76	142 143
$C_9H_8ON_2$ 8-hydroxy-4-methylcinnoline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ 2.59 HL 9.00	K_1 9.00, K_2 7.30	142
$C_9H_7O_4NS$ 8-hydroxyquinoline-5-sulphonic acid: H_2L	pH, sp	25	0.1 KNO ₃	H_2L 3.84 HL^- 8.35	K_1 8.52, K_2 7.16 $K(UO_2(OH)L_2^{3-} + H^+ \rightleftharpoons UO_2L_2^{2-})$ 6.68 $K(UO_2(OH)L_2^{3-} + 2H^+ \rightleftharpoons 2UO_2L_2^{2-})$ 11.7 $K(2UO_2(OH)L_2^{3-} \rightleftharpoons (UO_2(OH)L_2^{3-})_2)$ 11.7	144
$C_{10}H_9ON$ 8-hydroxy-2-methylquinoline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ 5.01 HL 11.01	K_1 -9.4, K_2 -8	142
$C_{10}H_9ON$ 8-hydroxy-5-methylquinoline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ 4.71 HL 11.11	K_1 11.25, 9.52	142
$C_{10}H_9ON$ 8-hydroxy-6-methylquinoline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ 4.76 HL 10.71	K_1 10.89, K_2 9.26	142
$C_{10}H_9ON$ 8-hydroxy-7-methylquinoline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ 4.26 HL 11.31	K_1 11.28, K_2 9.78	142
$C_{10}H_{10}ON_2$ 8-hydroxy-2:4-dimethylquinazoline:HL	gl	20	50% dioxan	H_2L^+ 3.15 HL 10.14	K_1 8.77, K_2 7.33	142
$C_{13}H_{13}ON$ 1:2:3:4-tetrahydro-9-hydroxy-acridine:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ 5.49 HL 11.39	K_1 10.10, K_2 8.20	142
$C_{15}H_{12}CN_2$ 8-hydroxy-4-methyl-2-phenylquinazoline:HL	gl	20	50% dioxan 0.3 NaClO ₄	H_2L^+ < 1 HL 11.33	K_1 8.53, K_2 7.85	142
$C_{22}H_{23}O_9N_3$ ammonium aurintricarboxylate (aluminon reagent): $(NH_4)_2H_2L$	sp	25	var		$(NH_4)_2HL^{2-}(\text{?})$: K_1 4.77	145

^a After Bjerrum, Schwarzenbach, and Sillén, reference 20.

Data which appeared in the literature prior to 1956 has been compiled mostly by the above authors. Subsequent data has been compiled by J. Gindler.

The notation is explained following Tables III and V.

stated that a complex is formed with undissociated oxalic acid, $H_2C_2O_4$. In more recent work, Moskvina and Zakharova¹¹⁸ conclude that complexes may be formed with both $C_2O_4^{2-}$ and $HC_2O_4^-$ ions and that the amount of each formed will depend upon its stability and the conditions of the experiment.

The composition of a complex is sometimes decided upon by comparison with complexes having similar ligands. For example, Hök-Bernström,¹³⁸ in her work with salicylic acid, H_2A , and methoxybenzoic acid, HB, was able to show that complexes of the type $UO_2(H_2A)(H^+)_{-1}$, $UO_2(H_2A)(H^+)_{-2}$, and $UO_2(HB)(H^+)_{-2}$ were formed in the aqueous phase. (The negative subscripts indicate that H^+ was eliminated in the formation of the complex.) The experimental data for methoxybenzoic acid was approximated by assuming only the complex $UO_2(B)(OH)$. For salicylic acid, the complexes $UO_2(HA)^+$ corresponding to $UO_2(H_2A)(H^+)_{-1}$ and $UO_2(HA)(OH)$ or UO_2A corresponding to $UO_2(H_2A)(H^+)_{-2}$ were postulated. It was not possible to distinguish between the latter two. However, from the similarity of the distribution curves found for the two acids, it was suggested that the salicylate complexes are formed by HA^- ligands.

A vast amount of work other than that listed in Table VII has been done on the preparation and identification of organic uranyl complexes. Some of the complexing agents studied recently include dihydroxy-maleic acid,¹⁴⁸ triose-reductone^{149,150} (enoltartronaldehyde), reductic acid¹⁵¹ (cyclopentene-2-diol-2,3-one-1), complexones^{152,153} (iminodiacetic acid and its derivatives), xanthates and dithiocarbamates,¹⁵⁴ protoporphyrin,¹⁵⁵ o-cresotic acid,¹⁵⁶ miricitrine,¹⁵⁷ dialkylphosphoric acids,^{158,159} and pyrazolone derivatives.¹⁶⁰

3. Non-aqueous solutions of uranium.

Solubility studies. A number of uranium salts are soluble in organic solvents. Uranyl nitrate is the notable example. As the hexahydrate, this salt is soluble in a

variety of ethers, esters, ketones, alcohols, aldehydes, and substituted hydrocarbons.¹⁶¹⁻¹⁶⁵ The following generalizations have been made¹⁶⁵ concerning its solution in organic solvents:

(1) In a given homologous series, the solubility decreases as the molecular weight of the solvent increases.^{161,165}

(2) Solutions occur with:¹⁶⁵

Ethers: aliphatic
 ethylene glycol
 diethylene glycol
 saturated cyclic

Acetals

Ketones: aliphatic
 aromatic
 alicyclic
 mixed aliphatic-aromatic

Alcohols: aliphatic
 alicyclic

Various esters

Nitrogen-containing solvents: nitriles
 aromatic bases.

(3) Solutions do not occur with:

Hydrocarbons¹⁶¹

Ethers:¹⁶⁵ aromatic
 unsaturated cyclic

Sulfur-containing solvents.¹⁶⁵

Glueckauf¹⁶⁶ has made the phenomenological observation that a plot of the solubility of uranyl nitrate against the oxygen-to-carbon ratio in the solvent molecule results in a single curve for alcohols and ethers; but in a double curve for ketones; one for symmetric and one for asymmetric ketones.

Ionization. The quantity $(\Lambda\eta/\text{constant})^*$ has been used by McKay and co-workers¹⁶⁷⁻¹⁶⁹ to estimate the degree of ionization of uranyl salts in organic solvents. By this criterion, uranyl nitrate in concentrations of 0.01 - 1M is

* Λ = molar conductivity. η = viscosity. The constant 60 is used for 1:1 - electrolytes; 120 for 1:2 - electrolytes.

substantially unionized in water-saturated solutions of ethers, ketones, alcohols and tributyl phosphate.* Only in saturated diethyl cellosolve and in isobutyl alcohol is ionization in excess of 10%. The large amount of water which dissolves in the latter solvent may account for this.

In tributyl phosphate, the dissociation of uranyl nitrate increases as its concentration in the organic phase is decreased. At $10^{-5}M$ the salt is approximately 40 per cent dissociated,¹⁶⁹ i.e., $\Lambda\eta/120 \approx 0.4$. Uranyl perchlorate at this concentration is almost completely ionized.¹⁶⁹ Ion association occurs at higher concentrations, but significantly less than for uranyl nitrate. At approximately $0.01M$, the association of uranyl perchlorate is maximum ($\Lambda\eta/120$ has a minimum value of ~ 0.1) in the concentration range 10^{-5} to $1M$. The ionization of this salt may well be associated with the amount of water contained within the tributyl phosphate since the electrical conductivity is decreased by dehydration.¹⁶⁹

Jezowska - Trzebiatowska and co-workers⁸⁰ have measured the molar conductivity of uranyl nitrate in organic solvents that contain only water from the hexahydrated uranyl salt.** The conductivity was found to be low and to decrease with a decrease of the dielectric constant of the solvent.

Conductivity measurements of UCl_4 in methyl alcohol indicated the salt to be somewhat dissociated.⁸⁰ The dissociation was found to increase on addition of tributyl phosphate.

Kaplan, Hildebrandt, and Ader¹⁷⁰ have classified into

* Solvents tested other than tributyl phosphate: diethyl ether, diethyl cellosolve, dibutyl carbitol, methylisobutyl ketone, isobutyl alcohol, and isoamyl alcohol.

**Solvents tested: methyl alcohol, ethyl alcohol, acetone, ethyl-methyl ketone, methyl isobutyl ketone, acetylacetone, stannous chloride in acetone.

types the absorption spectra of uranyl nitrate in a number of solvents and solvent mixtures.* Differences between types were attributed to a series of hydrated and solvated nitrate complexes, UO_2^{2+} , UO_2NO_3^+ , $\text{UO}_2(\text{NO}_3)_2$, and $\text{UO}_2(\text{NO}_3)_3^-$. The relative concentrations of the complexes depend upon the nature of the solvent, its water content, and the concentration of added nitrates. It is interesting to note that the absorption spectrum of uranyl nitrate in tributyl phosphate (0.016 - 1.6 M) is characteristic of the complex $\text{UO}_2(\text{NO}_3)_2$ and indicates little ionization.¹⁶⁸ A similar spectrum is given by uranyl nitrate in methyl isobutyl ketone (0.02 M).⁷⁸ Uranyl perchlorate in methyl isobutyl ketone (0.02 M), however, exhibits a spectrum characteristic of the uranyl ion, UO_2^{2+} . These results appear to be in general agreement with those obtained through conductivity-viscosity measurements.¹⁶⁷⁻¹⁶⁹

The trinitratouranyl complex $\text{UO}_2(\text{NO}_3)_3^-$ has been studied by a number of workers.^{78,79,171-173} It is formed by the addition of a second soluble nitrate to a solution of uranyl nitrate in a non-aqueous solvent such as anhydrous nitric acid,¹⁷¹ dinitrogen tetroxide,¹⁷² acetone,^{78,79} methyl isobutyl ketone,⁷⁸ dibutyl ether,¹⁷³ etc. Kaplan and co-workers⁷⁸ also report that the complex is formed in 16 M nitric acid, but that its formation is far from complete. The negative character of the complex has been demonstrated by electrolytic transference experiments.⁷⁸ Its composition has been deduced by the isolation of solid compounds from solutions of the type described above^{79,171,172} and by the similarity of the absorption spectrum of such solutions with that of crystalline cesium uranyl nitrate, $\text{CsUO}_2(\text{NO}_3)_3$.^{78,79}

* Spectra classified: uranyl nitrate in water, acetone-water, dioxane-water, n-propanol-water, ethanol, chloroform + 0.7% ethanol, pyridine, acetic acid, ethylacetate, tetraethylene glycol dibutyl ether, nitroethane, methyl isobutyl ketone, cyclohexanone; uranyl perchlorate in methyl isobutyl ketone.

The stability of the complex depends upon the nature of the associated cation as well as the nature of the solvent and the presence of water in the solvent. The general order of the solvents with respect to stability of the trinitrate complex is: ketone > ether > alcohol > water.⁷⁸

The formation of chlorouranyl complexes in non-aqueous solvents has been reported by Vdovenko, Lipovskii, and Nikitina.¹⁰⁸ The complexes UO_2Cl^+ , UO_2Cl_2 , $UO_2Cl_3^-$ were formed by the addition of pyridine hydrochloride or hydroxylamine hydrochloride to a solution of uranyl perchlorate or uranyl chloride in acetone. The stability of the trichlorouranyl complex was found to be dependent upon the amount of water present in the solvent. A compound was separated and identified as $(C_5H_5NH)_2UO_2Cl_4$.

Hydration. In partition studies of uranyl nitrate between aqueous solution and organic solvent (alcohols, esters, ethers and ketones) it is generally found that the water content of the organic phase increases with uranyl nitrate concentration.^{163,167*} For alcohols, the relation between the water content M_w and uranyl nitrate concentration M_u appears to be a complex function.^{**} For esters, ethers, and ketones the relation is linear except possibly at high values of M_u . This relation may be expressed

$$M_w = M_w^0 + hM_u.$$

The quantities are expressed in terms of molalities of the dry solvent. M_w^0 is the solubility of water in the pure solvent; h is a constant. The slope of the line, h, repre-

* The water content of alcohols may decrease initially as the uranyl nitrate concentration is increased from 0 to 0.1-0.2 molal.¹⁶⁷

** Katzin and Sullivan¹⁶³ report a linear relation between M_w and M_u for isobutyl alcohol. McKay and Mathieson¹⁶⁷ point out that if the data of Katzin and Sullivan at low M_u are given significance, then a more complicated relationship between the two quantities exists.

sents the degree of hydration. For many of the linear solvents h is very nearly 4.0.^{163,167} This has been interpreted by Katzin¹⁷⁴ to mean that the species $UO_2(H_2O)_4(NO_3)_2$ is extracted. McKay,¹⁷⁵ however, considers this to be a mean hydration number; that a series of hydrates are present ranging from the di- to the hexahydrate; that these hydrates are in equilibrium with each other and are of comparable stability. The latter view is supported by isopiestic measurements.¹⁷⁶

Infrared measurements on ethereal and ketonic solutions* of uranyl nitrate, indicate two molecules of water to be strongly bound to the uranyl nitrate and the remaining water molecules to be more weakly bound.^{177,178}

The extraction of uranyl nitrate from an aqueous system into tributyl phosphate (TBP) causes the displacement of water from the organic phase.¹⁶⁸ The displacement is roughly linear with h being -2.¹⁶⁸ This is in agreement with the formulae $TBP \cdot H_2O$ and $UO_2(NO_3)_2 \cdot 2TBP$.¹⁶⁸ Uranyl perchlorate, however, apparently does carry some water into tributyl phosphate.¹⁶⁹ Whether this water is associated with free uranyl ions or unionized $UO_2(ClO_4)_2$ is undetermined.

Solvation. The isolation of solvated uranium salts, in particular uranyl nitrate, is reported in the literature.^{6, 163, 180-182} In phase studies of ternary systems: uranyl nitrate, water, organic solvent, Katzin and Sullivan¹⁶³ have concluded that uranyl nitrate in aqueous solution is largely hexasolvated, subject to the activity. As organic molecules are dissolved, 2, 3, 4 and perhaps 6 water molecules may be displaced, depending upon the electron-donor capabilities of the organic molecules. The total solvation is a function of the activity levels of the water and organic molecules.

* Solvents studied: diethyl ether, acetone, methylethyl ketone.

If a particular configuration is stable enough, it may survive as a crystalline solid. The particular stability of the final two water molecules is in agreement with the results of Ryskin and co-workers^{177,178} obtained through infrared absorption measurements. The ability of solvents to displace water is in the order: alcohols > ethers > ketones.¹⁶³ This general order of solvate stability is confirmed by heat of solution measurements.¹⁸³ It is in agreement also with the order of base (electron-donor) strengths of the solvents determined by other means.^{163,183} Methyl isobutyl ketone is anomalous in that it behaves stronger toward uranyl nitrate than its base strength would indicate.^{163,183} Tributyl phosphate, according to heat measurements, competes with water almost as well as diethyl ether and isobutyl alcohol.¹⁸³

The order of solvents with respect to solvate stability is opposite to that with respect to the stability of the trinitratouranyl complex. This suggests a competition between solvent molecule and nitrate ion for coordination with the uranyl ion.¹⁷¹

Feder, Ross and Vogel¹⁸⁴ have studied the stability of molecular addition compounds with uranyl nitrate. The compounds were prepared by shaking uranyl nitrate dehydrate with various addenda in an inert solvent: benzene and/or 1,2-dichloroethane. 1:1 molecular addition compounds were found with uranyl nitrate and ethyl alcohol, n-dodecyl alcohol, tetrahydrofuran, propylene oxide, mesityl oxide, tributyl phosphate, and N,N-dibutylacetamide. 1:2 compounds were observed with uranyl nitrate and acetone, methyl isobutyl ketone, cyclohexanone, ethyl acetate, 2,4-dimethyltetrahydrothiopene 1,1-dioxide, β -chloroethylacetate, ethyl chloroacetate, ethyl cyanoacetate, diethyl ether, allyl alcohol, ethylene chlorohydrin, and acetonitrile. Formation constants were determined from changes in the solubility of uranyl nitrate.

It was shown that the stability of the addition molecules, for those addenda having similar functional groups, was in agreement with the base strength of the addend; i.e., the more stable the molecule, the greater the base strength of the addend.

The average number of solvate molecules n associated with uranyl nitrate in its partition between water and various organic solvents* has been studied by McKay^{185,186} and co-workers. The value of n was found to vary with the uranyl nitrate concentration of the organic phase. For most of the solvents studied, n varied between 1 and 4. For cyclohexane, considerably larger values were found for low uranyl nitrate concentrations.

A saturated solution of uranyl nitrate hexahydrate in tributyl phosphate corresponds closely to the unhydrated disolvated compound $UO_2(NO_3)_2 \cdot 2TBP$.^{168,187} Evidence for the existence of the single species is:^{168,187}

- (1) The solubility is not appreciably temperature dependent over the range 0-50°C.
- (2) On freezing and rewarming a saturated solution, a sharp melting-point of $-6.0 \pm 0.5^\circ C$ is observed.
- (3) The mole ratio of uranyl nitrate to TBP approaches the value 1:2 asymptotically under a variety of conditions.
- (4) The effect of inert diluents for the TBP on uranyl nitrate partition coefficients supports a 1:2 formula, i.e., the partition coefficient of uranyl nitrate varies as the square of the TBP concentration.

The experimental conditions under which Feder, Ross and Vogel¹⁸⁴ reported the formation of $UO_2(NO_3)_2 \cdot TBP$ were considerably different from those of Healy and McKay.¹⁶⁸

* Organic solvents studied: diethyl ether, diisopropyl ether, diethyl cellosolve, dibutyl cellosolve, dibutyl carbitol, penta-ether, isoamylacetate, methyl isobutyl ketone, cyclohexanone.

Jezowska - Trzebiatowska, et al⁸⁰ report that the absorption spectrum of uranyl nitrate in tributyl phosphate gives no indication of the formation of a stable complex. Attempts to identify a complex species in the concentration range 0.02-0.06M were unsuccessful.

Hesford and McKay¹⁶⁹ report evidence for the formation of $UO_2(ClO_4)_2 \cdot 2TBP$ under certain conditions. From a 10.3M aqueous perchloric acid solution, the partition coefficient of uranium varies as the square of the TBP concentration in benzene. Under other conditions, other solvates may be formed.

Jezowska - Trzebiatowska, et al⁸⁰ report the formation of a 1:1 complex between UCl_4 and TBP in methyl alcohol.

Tributyl phosphine oxide, like tributyl phosphate, forms an anhydrous disolvate with uranyl nitrate.¹⁸⁷ Healy and Kennedy¹⁸⁸ report a number of other neutral organophosphorus solvents which form solvates with uranyl nitrate. Most, but not all, of the solvates reported are anhydrous. All of the solvents extract uranium from aqueous solution in proportion to the square of the solvent concentration (in benzene). However, not all solutions of the solvent in benzene and saturated in uranyl nitrate give mole ratios of solvent to uranium of 2:1. For the two diphosphanates and one pyrophosphate studied the mole ratios were 1:1. This may be indicative of chelation or polymer formation. The mole ratio in triphenyl phosphate was ~22:1. This is probably the result of the solvent being unable to displace water from the coordination sphere of the uranyl ion.¹⁸⁸

Solvate formation between uranyl salts and acid organophosphorus compounds, eg. mono- and di-alkyl phosphoric acids has been the subject of some investigation.^{80,158,159,188} 1:1 complexes between uranyl nitrate and mono- and di-butyl phosphate and mono- and di-amyl phosphate in ethyl alcohol have been reported.⁸⁰ In explanation of distribution data

in conjunction with isopiestic and viscosity measurements, Baes, Zingaro and Coleman¹⁵⁸ have hypothesized that uranium (VI) is extracted from aqueous perchlorate solutions into n-hexane solutions of di-(2-ethylhexyl)-phosphate, HL, as the species $UO_2(HL)_2L_2$. As the uranium concentration of the organic phase is increased, there is strong evidence that polymerization occurs.¹⁵⁸ Similar conclusions have been made by Dyrssen¹⁵⁹ on the basis of the distribution of uranium (VI) between aqueous perchlorate solution and dibutyl phosphate, HK, in chloroform. In hexone, the species $UO_2(HK)_2K_2$ and UO_2K_2 have been identified.¹⁵⁹ The extraction of uranium (VI) by dibutyl phosphate from aqueous nitrate solutions into benzene has been studied by Healy and Kennedy.¹⁸⁸ In addition to the species $UO_2(HK)_2K_2$, the polyuranyl species $(UO_2K_2)_x \cdot 2HK$ and the nitrated species $UO_2(NO_3)_2 \cdot 2HK$ have been postulated to explain the shape of the extraction curve as a function of nitric acid concentration.

It has been postulated that the formation of mixed solvates or solvated chelates enhances the extraction of uranium into certain solvent mixtures. These systems are discussed in a later section on solvent extraction.

IV-D Separation of Uranium

A number of review articles have been written on the analytical chemistry of uranium.^{6,34,189-200} These, together with many texts on chemical analysis (see, for example, references 201-209), serve well as guides to the separation and purification of the element. More specialized surveys have been made by Hecht²¹⁰ on the quantitative micro-analysis of uranium-bearing minerals, and by Lawrowski²¹¹ on separation processes for the recovery of nuclear fuels.

Two general techniques are available for the separation of uranium. (1) Uranium is removed from solution in the presence

of contaminants by precipitation, solvent extraction or some alternative method. (2) Uranium is kept in solution and contaminants are removed. These techniques are facilitated by the fact that uranium is reasonably stable in two oxidation states, (IV) and (VI), and that complex formation may be effected to prevent the removal of either uranium or contaminant from solution.

In the following paragraphs, the separation of uranium by precipitation, solvent extraction, and ion exchange are described in some detail. Reference is made also to other methods of separation: chromatography, electrodeposition, volatilization and pyrometallurgy.

1. Precipitation. In classical systems of analysis, uranium is a member of the third group of elements.* That is, it is not precipitated by hydrochloric acid or by hydrogen sulfide in an acidic solution, but it is precipitated by ammonium hydroxide or ammonium sulfide (see references 204, 206, 208, 213). Unfortunately, for the separation of uranium, many other elements also are precipitated by the same reagents. However, there exists a large number of reagents capable of precipitating uranium over a wide range of pH. These, combined with judicious use of the two oxidation states and/or the complexing ability of uranium, may be used to provide reasonably pure uranium deposits.

Precipitants. With the advent of nuclear energy as a source of power, numerous precipitants have been investigated in an effort to find one specific for the separation and/or determination of uranium. None have been found to date. Ware²¹⁴ has summarized early work using organic rea-

* In the system outlined by Noyes and Bray,²⁰² uranium is precipitated in the sixth group with ammonium hydroxide and is converted to the sulfide with hydrogen sulfide. In the system of West and Parks,²¹² uranium is precipitated in the fifth (basic benzoate) group.

gents as precipitants. She²¹⁵ and Bailey²¹⁶ have investigated some of the more promising ones. Rodden and Warf³⁴ have discussed the use of many reagents, both inorganic and organic, and have described procedures for the use of many of them. The latter precipitants, i.e., those for which procedures have been given by Rodden and Warf,³⁴ are denoted by a dagger (†) in the following discussion.

Inorganic precipitants. The reagents are listed alphabetically according to anion.

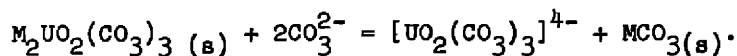
Arsenates. † Arsenic acid and ammonium, sodium and potassium arsenate precipitate uranium as uranyl metal arsenate. Silver, titanium, zirconium, thorium and lead interfere. Separation is made from the alkali metals, alkaline earths, aluminum, iron (II), and rare earths, including trivalent cerium.³⁴

Carbonates. ^{6,63,201,204,206,217-224} Precipitation of uranium with ammonium, sodium, or potassium carbonates is not very satisfactory. Highly soluble carbonate-uranyl complexes are formed. Under proper conditions, the metal uranyl tricarbonate salts $M_4UO_2(CO_3)_3$ are formed. The solubilities of the respective ammonium, sodium, and potassium salts in water are 50(15°C),²²⁴ 150(RT),²²⁵ and 71(18°C)²¹⁸ grams per liter. The solubility of the potassium salt in a 50% solution of potassium carbonate is 0.200 grams per liter.²²¹ The solubility of the sodium salt is decreased by increasing temperature and by increasing sodium salt concentration.²²³

Težak²¹⁷ has studied the precipitation of uranium by ammonium and sodium carbonate. From a 0.043N uranyl nitrate solution, precipitation was observed to be maximum in the region of 0.1N precipitant concentration. Two maxima were observed for ammonium carbonate; one for sodium carbonate. Above and below these fairly narrow regions of precipitant concentration, uranium enters into solution.

The uranium (IV) salt, $\text{Na}_6\text{U}(\text{CO}_3)_5 \cdot 11\text{H}_2\text{O}$ is precipitated from reduced carbonate solutions at high uranium and carbonate concentrations.⁶³

Barium uranyl carbonate salts are reported to be very insoluble.⁶³ However, in the presence of carbonate solutions the alkaline earth salts are unstable according to the reaction,⁶³



Tezak²¹⁷ has found the precipitation of uranium to be nearly complete when the barium:uranium ratio is greater than 600 and the excess carbonate is less than four times the barium concentration.

A suspension of barium carbonate may be used to precipitate uranium.^{34,226} Ammonium salts interfere. A suspension of basic zinc carbonate may be similarly used.^{34,227} Iron, aluminum and thorium also precipitate.

Cyanides. Alkali cyanides form a yellow precipitate when added to uranyl solutions.³⁴

Ferrocyanides.[†] The addition of potassium ferrocyanide to a uranyl salt solution causes the formation of a deep-red precipitate or suspension, depending upon the concentration of uranium. The reaction is used much in qualitative analysis for the identification of uranium. However, it is little used for quantitative separation. The separation is poor and there are many interferences.³⁴ Separation can be made from beryllium in a weakly acidic sulfate solution.²²⁸

Fluorides.[†] Hydrogen fluoride precipitates uranium (IV) as the tetrafluoride. The precipitate is gelatinous and difficult to filter.³⁴ Separation is made from metals complexed by fluoride ions, eg., tantalum and zirconium. Uranium may be reduced to the (IV)-state with zinc in a solution made slightly acidic.¹⁹⁶

The double fluorides, eg., NaUF_5 , are sparingly soluble even in the presence of strong acids.¹⁹⁶ Separation can be made under these conditions from Mo, Ti, Ni, Co, Mn, Cu, Fe (II), and V (III). Aluminum precipitates as the double salt, Na_2AlF_6 . Iron (III) precipitates in part. Reduction to uranium (IV) may be done in the presence of fluorides with iron (II).^{197,229} Rongalite ($\text{Na}_2\text{H}_2\text{S}_2\text{O}_4 \cdot 2\text{CH}_2\text{O} \cdot 4\text{H}_2\text{O}$) also has been used to effect reduction.²³⁰

Hydroxides. The addition of a metal hydroxide to a solution of uranyl salts results in the formation of the metal uranate. It has commonly been assumed that the metal diuranate, $\text{M}_2\text{U}_2\text{O}_7$, is precipitated by ammonium, sodium, or potassium hydroxide. However, experimental evidence indicates that the composition of the precipitate depends upon the conditions which exist during precipitation and upon the subsequent treatment, such as washing, which it receives.²³¹⁻²³⁷

Ammonium hydroxide[†] precipitates uranium quantitatively at pH 4 or greater.³⁴ The presence of ammonium salts and macerated filter paper facilitate precipitation. Separation is made from alkali metals, alkaline earths, and cations forming ammonia complexes. Repeated precipitations may be necessary to give sufficient separation. Phosphorus, vanadium, silicon, boron, aluminum, iron and other elements of the ammonium hydroxide analytical group also are precipitated.²¹² Complexing agents: carbonate, oxalate, citrate, tartrate, fluoride, etc., interfere.

Precipitation with alkali metal hydroxides is similar to that with ammonium hydroxide. Uranium may be precipitated in the presence of carbonate with sodium or potassium hydroxide of sufficient concentration. Carbonate ion interference may be removed by heating.

Pale green gelatinous $\text{UO}_2 \cdot \text{H}_2\text{O}$ is precipitated from uranium (IV) solutions by ammonium and alkali metal hydroxides.¹⁹

Iodates.† Uranium (IV) is precipitated from an acid solution by potassium iodate.³⁴ Separation can be made from copper, molybdenum, and reduced vanadium.²³⁸ Aluminum in amounts up to fifty times that of uranium does not interfere. Larger amounts of aluminum and divalent iron in any concentration cause incomplete precipitation. Titanium, zirconium, cerium (IV), and thorium precipitate with iodate.³⁴

Mercuric oxide. Uranium is precipitated when a suspension of mercuric oxide is boiled in an aqueous solution containing ammonium chloride.³⁴ Separation is made from alkali metals and alkaline earths. Hydroxy acids interfere.

Peroxides.† Hydrogen peroxide precipitates uranium peroxide, $\text{UO}_4 \cdot x\text{H}_2\text{O}$, from slightly acidic solutions. The reaction occurs in the pH range 0.5-3.5. The optimum range is 2.0-2.5. Hydrogen ions released with the formation of uranium peroxide are neutralized with ammonia or ammonium acetate. Complete precipitation requires an excess of hydrogen peroxide. Quantitative separation may be effected by freezing the solution, allowing it to stand, and filtering at 2°C. The separation from most elements is good since it is done from an acidic solution.^{34,239} Plutonium, thorium, hafnium, zirconium, and vanadium also precipitate. Iron interferes by catalytically decomposing hydrogen peroxide. Small quantities of iron may be complexed with acetic, lactic, or malonic acid. Low yields may result from the use of malonic acid. Ammonium, potassium, and alkaline earths retard the rate of precipitation. Complexing ions such as oxalate, tartrate, sulfate, and fluoride in large quantities, also interfere. Fluoride ion may be complexed with aluminum.^{239a}

Phosphates. Phosphoric acid and sodium monohydrogen phosphate precipitate UO_2HPO_4 from uranyl solutions. Uranyl ammonium phosphate, $\text{UO}_2\text{NH}_4\text{PO}_4$, is precipitated by $(\text{NH}_4)_2\text{HPO}_4$ or Na_2HPO_4 in the presence of ammonium acetate.[†] Precipitation is made in the pH range 1.2-2.3, 1.7 being optimum. It is not very selective. Zirconium, bismuth, and thorium precipitate under similar conditions. Alkali metals are retained. Separation is made from vanadium. Both UO_2HPO_4 and $\text{UO}_2\text{NH}_4\text{PO}_4$ are soluble in mineral acids.³⁴

Phosphate precipitation of uranium (IV) is more selective. It is made from dilute hydrochloric or perchloric acid solutions. Separation is made from manganese, iron, vanadium and most other elements. Zirconium, thorium, and, to a smaller extent, titanium and tin precipitate.¹⁹⁶⁻¹⁹⁸ Aluminum interferes by the formation of soluble complexes with uranium and phosphate ions.⁵⁹ With sulfate and aluminum present, uranium is precipitated in a narrow pH-range around one. At higher pH, the soluble aluminum-uranium-phosphate complex is formed; at lower pH, the soluble uranium-sulfate complex. Chromium in excess of 0.2 gram per 100 milliliters causes incomplete precipitation.¹⁹⁶ Large amounts of fluoride ion prevent precipitation.¹⁹⁶

Sodium hexametaphosphate $[(\text{NaPO}_3)_6]$ also precipitates uranium (IV) from acid solutions.²⁴⁰ Adherence to fairly stringent conditions allows for complete precipitation. A 3N HClO_4 solution of uranium (IV) is heated to 60-70°C. If more than 2 mg. of uranium are to be precipitated, a freshly prepared 2% hexametaphosphate solution is added until its concentration in the precipitating medium is 0.30-0.35 per cent. To precipitate smaller amounts of uranium, a 0.5 per cent solution of thorium chloride is added as carrier and the hexametaphosphate added until it is in excess 25 per cent with respect to the thorium, i.e., molar ratio of $\text{Th}:\text{PO}_3^-$

is 1:5. Coagulation is improved by heating in a water-bath for ten to fifteen minutes after precipitation. Under these conditions separation is made from V (III) and (IV), Fe, Cu, and other di- and tri-valent metals. Incomplete precipitation occurs with increased or decreased acidity -- probably because of enhanced solubility of the compound and complex formation, respectively. Precipitation from sulfuric acid is incomplete because of uranium-sulfate complex formation. Under certain conditions both uranium (IV) and (VI) form complexes with hexametaphosphate.

Hypophosphoric acid ($H_4P_2O_6$), sodium dihydrogen hypophosphate ($Na_2H_2P_2O_6$), and sodium pyrophosphate ($Na_2P_2O_7$) precipitate uranium (IV) from acid solutions.[†] Other tetravalent metals, Ti, Zr, and Th, also precipitate. Separation is made from uranium (VI) and trivalent metals in general.³⁴

Phosphites. Sodium hypophosphite (NaH_2PO_2) and ammonium thiosulfate or sulfurous acid precipitate uranium from a boiling, dilute acid solution.²⁴¹ Zirconium and titanium precipitate under similar conditions. These elements may be separated prior to uranium by boiling with sodium hyposulfite alone. Elements forming acid-insoluble sulfides are removed with hydrogen sulfide before adding sodium hypophosphite and ammonium thiosulfate.

Sulfates. Uranium (IV) sulfate is practically insoluble in 47 per cent perchloric acid. Precipitation is made in a sulfuric acid medium. Uranium is reduced on a mercury cathode and concentrated perchloric acid is then added.

Sulfides. Ammonium sulfide[†] or polysulfide precipitates brown, amorphous uranyl sulfide. Numerous other elements are precipitated under similar conditions.²¹³ Complexing agents such as carbonate, pyrophosphate, and citrate interfere.³⁴ Uranium (IV) salts are precipitated as $UO_2 \cdot H_2O$ by ammonium sulfide.¹⁹⁰

Hydrogen sulfide bubbled through a nearly neutral solution of uranyl salts containing hexamethylene tetramine precipitates uranium in a readily filterable, crystalline form of "uranium red."[†] Separation is made from alkali metals and alkaline earths.^{34,197}

Vanadates. Ammonium metavanadate precipitates ammonium uranyl vanadate from uranyl solutions buffered with ammonium acetate.³⁴ Uranovanadic acids are precipitated at pH 2.2-6.5.²⁴² Within these limits, the composition does not depend upon the hydrogen ion concentration. It does depend upon the vanadium : uranium ratio present in solution. Compounds corresponding to the formulae $H[UO_2(OH)_2VO_3] \cdot H_2O$, $H[UO_2(OH)(VO_3)_2] \cdot 2H_2O$, and $H[UO_2(VO_3)_3] \cdot 4H_2O$ have been identified.²⁴² Ammonium salts of these acids have been synthesized in the presence of ammonium chloride.²⁴² Ammonium uranyltrimetavanadate is the least soluble. However, its formation is a long process at room temperature. Heating greatly accelerates its rate of formation.

Organic precipitants. Organic precipitating reagents are listed alphabetically.

3-Acetyl-4-hydroxycoumarin (3-acetyl benzotetronic acid). An alcoholic solution of the reagent added to a uranyl salt solution forms a pale yellow precipitate insoluble in ethanol.²⁴³ Precipitation occurs between pH 1.5 and 7. Below pH 1.5 the reagent precipitates. The thorium complex is soluble in alcohol, but precipitates from an aqueous solution at pH 2-4. Lanthanum and cerium (III) do not interfere when present in amounts ten times that of uranium. Cerium (IV) interferes even in small amounts.

Acradine. Uranium (IV) and (VI) are precipitated by the reagent with the addition of ammonium thiocyanate.²¹⁵ Iron (III), cobalt, copper, zinc, cadmium, mercury, and bismuth precipitate.^{214,216}

Aldehyde ammonia precipitates UO_3 .²¹⁶

Alizarin and Alizarin Red S (sodium alizarin sulfonate).

Uranium is precipitated slowly by the reagents when the uranyl ion concentration is less than 20 micrograms per milliliter.²¹⁵

Aluminon (ammonium salt of aurintricarboxylic acid).

The reagent precipitates both uranium (IV) and (VI) from sulfate solutions at pH 3.5.²¹⁶

Amines. Ammine salts, generally of the form $\text{UO}_2(\text{Amine})_2 \text{X}_2$, where X is an acid radical such as acetate, chloride, nitrate, etc. have been prepared from acetanilide, anti-pyrene, bromoantipyrine, diethylaniline, exalgin, nitrosoantipyrine, p-nitrosodimethylaniline, phenacetin, pyramidon, pyridine, quinaldine, and quinoline. Mono-, tri- and tetra-ammine salts also have been formed. The salts are generally prepared in anhydrous chloroform or amyl alcohol solutions. However, some of the more stable salts may be precipitated from aqueous or alcoholic solutions.^{214,244}

2-Amino pyridine precipitates UO_3 .²¹⁶

Ammonium benzoate. Uranyl ion is precipitated by the reagent from slightly acidic solutions heated to boiling. A 0.05 N solution of the reagent containing about 2.5% NH_4OH is boiled separately and added in an excess of three to four times the uranium present. Carbonate ion prevents quantitative precipitation.^{189,245}

Ammonium dithiocarbonate precipitates uranium (VI).

Derivatives also are formed with Al, Mn(II), Fe(II), Co, Ni, Cu(II), Zn, Ag, Sn(IV), Pb, and Bi.²⁴⁶

AnthragalloI forms brown precipitates or solutions with U^{4+} , UO_2^{2+} , Fe^{3+} , Cu^{2+} , and MoO_4^{2-} .²¹⁵

Anthranilic acid. Uranium (IV) is precipitated from a solution of the reagent buffered with ammonium acetate. The reagent added to a 0.1M $\text{UO}_2(\text{NO}_3)_2$ solution forms a heavy

yellow precipitate. The amount of precipitate is increased by the addition of 1M acetic acid; decreased by the addition of 1M sodium acetate. Fifty micrograms of uranyl ion per drop of solution gives no observable precipitate. In acetic acid-sodium acetate buffered solutions, slightly soluble salts are formed with the reagent and Mn, Co, Ni, Cu, Zn, Cd, Hg(II), and Pb.²¹⁵

Arsonic acids. Benzenearsonic acid[†] precipitates uranium (IV) in a weakly acidic solution, pH 1-3. Titanium and cerium (IV) are partially precipitated. Thorium, zirconium, hafnium, tin (IV), niobium, and tantalum are quantitatively precipitated.³⁴

Arsonilic acid (p-aminobenzenearsonic acid) precipitates uranium (VI) in a weakly acidic solution, pH 1-4 or greater.²⁴⁷ At pH 2.1 or greater the precipitation is quantitative as evidenced by negative ferrocyanide tests of the filtrate. Other ions which precipitate from neutral or slightly acidic solutions include Cu²⁺, Zn²⁺, Cd²⁺, and U⁴⁺. With the addition of sodium acetate, aluminum and ferric ions also precipitate.²¹⁵

Other substituted arsonic acids which give difficulty soluble uranyl salts are 3-nitro-4-hydroxybenzene- and methane-arsonic acids.³⁴

Bis-benzenephosphonic acid. Tests on 50-150 mg/l uranium (VI) in sulfuric acid solutions in the presence of approximately 100- to 1000-fold excess ferrous, sulfate, aluminum, magnesium, and phosphate ions gave nearly 99% precipitation with the reagent. Optimum conditions for precipitation are pH 1, 25°C, and 10/1 molar ratio of reagent to uranium.²⁴⁸

Benzenesulfinic acid precipitates uranium (IV) in acidic solutions. Iron (III) and the tetravalent ions Ti, Sn, Ce, and Th also precipitate.²¹⁴

Benzopurpin precipitates uranium (IV) and (VI).²¹⁶

Benzoylacetone precipitates uranium (VI).²¹⁶

5-Bromo-7-carboxy-8-hydroxyquinoline precipitates uranium (VI), copper, zinc, cadmium, mercury, and lead.²⁴⁹

5-Carboxy-8-hydroxyquinoline precipitates uranium (VI) in solutions buffered with acetic acid and sodium acetate. Iron, copper, zinc, cadmium, mercury, and lead precipitate.²⁴⁹

7-Carboxy-8-hydroxyquinoline precipitates uranium (VI) in ammoniacal tartrate solutions.

Catechol forms compounds with tetravalent uranium, silicon, titanium, zirconium, and thorium.²¹⁴ Catechol combined with pyridine precipitates hexavalent uranium.²¹⁴

Cresotinic acid in the presence of sodium acetate precipitates uranium (IV) from solution. Aluminum and iron (III) also precipitate. Separation is made from Cr, Fe(II), Co, Ni, Cu, Zn, Cd, and Mo.^{215,216}

Cupferron (ammonium nitrosophenylhydroxylamine).[†] Uranium (IV) is precipitated from acidic solutions by the reagent. Good separation is made from other elements if this precipitation follows one in which uranium was kept in the hexavalent state. Ions which are precipitated by cupferron from acidic solutions include Ti, V, Fe, Ga, Zr, Nb, Sn, Sb, Hf, and Ta. Ions which are not precipitated under such conditions include the alkali metals, alkaline earths, Be, B as borate or fluoborate, Al, P, Cr, Mn, Ni, Zn, and U(VI). Precipitation is usually made in a sulfuric acid medium but hydrochloric or organic acids may be used. Nitric acid should be avoided; also perchloric acid if the precipitate is to be ignited. The presence of a reducing agent, hydroxylamine or sodium hydrosulfite, facilitates complete precipitation of uranium (IV). The cupferrate may be filtered or extracted with an organic solvent such as chloroform.^{34,250}

Hexavalent uranium is precipitated by cupferron from neutral solutions.²⁵⁰

Dibenzoyl methane forms a yellow precipitate with uranium (VI).²¹⁶

3,5-Dibromosalicylaldehyde precipitates U(VI), Co, Ni, Hg(II), and Pb.²⁵¹

4,4'-Dihydroxy-3,5,3',5'-tetra(hydroxymethyl)-diphenylmethane precipitates U(VI), Mn, Fe(III), Cu(II), and Hg(II).²¹⁶

Dimethylammonium dimethyldithiocarbamate forms a red precipitate with uranium (VI).²¹⁶

Diphenyl thiocarbamide precipitates uranium (VI) from neutral solutions. Copper (III), silver, lead, and bismuth also precipitate with the reagent.²¹⁶

Dipropylamine forms a yellow precipitate with uranium.²¹⁶

Disalicylaldehyde diimine precipitates uranium (IV) and (VI). Most heavy metals are precipitated by the reagent.³⁴

Ethanolamine precipitates UO_3 .²¹⁶

Ethylenediamine and uranyl nitrate form an insoluble double salt, $UO_2SO_4(H_2SO_4)_2NH_2CH_2CH_2NH_2$, in alcoholic-sulfuric acid solution.²⁵² Double salts of the same type are formed with piperazine and dimethylpiperazine. Siemssen²⁵³ observed that a solution of ethylenediamine added to a uranium solution gives a bright yellow crystalline precipitate that is soluble in excess reagent.

Ethylenediamine tetracetic acid. Uranium is precipitated when a uranyl acetate solution is boiled with solid reagent.²⁵⁴

Gallic acid precipitates U(IV), U(VI), Fe(III), Cu(II), and Zn.²¹⁵

Guaiacol. A brown precipitate results from the reaction of the potassium salt of guaiacol and uranyl acetate in an aqueous solution.²⁵⁵

Hexamethylene tetramine (urotropine)[†] is a weaker base than ammonium hydroxide and does not absorb carbon dioxide.

This reduces the likelihood of carbonate interference and of alkaline earth carbonate precipitation. Uranium is precipitated when the reagent is boiled in a uranyl solution that contains ammonium ion and no excess acid.³⁴ Ions that form stable complexes with uranium interfere. Separation can be made from alkali metals, alkaline earths, Mn, Co, Ni, and Zn. Zr, Ti, Fe, Al, Ce(IV), Th, and some other elements precipitate.

A double salt, $UO_2SO_4 \cdot H_2SO_4 \cdot (CH_2)_6N_4$, is formed with the reagent and an excess of sulfuric acid and uranyl salt.²⁵⁶

a-Hydroxyacetophenone forms a white precipitate with hexavalent uranium.²¹⁶

1-Hydroxyacridine (1-acridol or benzoxine) precipitates uranium (VI) in neutral solutions. Calcium precipitates from neutral solutions; Mg, Ca, and Ba from alkaline solutions; Cr(III), Mn(II), Fe(II and III), Cu(II), Zn, Cd, Hg(I and II), Te(II), and Pb from solutions containing acetic acid and sodium acetate. Al, Sn(II), and Bi do not precipitate.²⁵⁷

1-Hydroxyanthraquinone forms slightly soluble complexes with uranyl, cobalt, cupric, nickel, magnesium, and manganese ions.²¹⁵

1-Hydroxy-3-methoxyxanthone may be used to separate uranium, thorium, ceric salts and cerite earths.²⁵⁸ The ceric salts and cerite earths are not precipitated by the reagent. Thorium is precipitated at pH 2.6-4.0. Uranium (uranyl ion) precipitates at higher pH.

8-Hydroxyquinoline. Tetravalent uranium is precipitated by the reagent with the addition of ammonium acetate. The precipitation of hexavalent uranium is almost quantitative in the pH range 7-9 from carbonate-free ammonium acetate buffer.²⁵⁹ Iron, cobalt, nickel, copper, cadmium, and chromium are precipitated by the reagent.²¹⁵

8-Hydroxyquinoline (oxine).[†] Hexavalent uranium is

precipitated as $\text{UO}_2(\text{C}_9\text{H}_6\text{NO})_2 \cdot \text{C}_9\text{H}_7\text{NO}$ from weakly acidic or basic solutions.³⁴ Quantitative recovery has been reported over the pH range 4.1-13.5. A large number of other elements are precipitated by oxine including Mg, Al, Cr, Fe, Co, Ni, Cu, Zn, Cd, Mo, Bi, and Th.^{215,260-263} Uranium can be precipitated in the presence of small amounts of complexing agents: fluoride, hydroxylamine, oxalate, lactate, and tartrate.³⁴ Separation from small amounts of phosphate also can be made at pH 10-12 using an excess of oxine. Ammonium carbonate interferes. Tetravalent uranium and oxine form a brownish-yellow deposit.²¹⁵

Isatin- β -oxine (β -isatoxime).[†] Uranyl and mercuric ions are precipitated by the reagent from weakly acidic solution. Precipitation is incomplete but can be made quantitative by increasing the pH with sodium acetate. A number of other elements precipitate under these conditions including Fe(II), Co, Ni, Ag, Hg(I), and Pb.²⁶⁴ Separation can be made from Mn(II), Zn, and alkaline earth ions.²⁶⁵ In alkali tartrate solutions, uranium can be separated from cobalt and nickel.²⁶⁶

Isojuglone. The sodium salt of this reagent and uranyl acetate form a carmine-red precipitate after washing with ethanol. Iron, cobalt, nickel, zinc, cadmium, mercury, and lead are precipitated by the reagent.²⁶⁷

Isonitroso-N-phenyl-3-methylpyrazolone. Uranyl nitrate or acetate forms a reddish-orange precipitate with a 1% solution of the reagent in a 50% alcoholic solution. Precipitation is quantitative with the addition of sodium acetate. Mercury (I) and (II), copper (I) and (II) and uranyl ions precipitate in acidic media (nitrate or sulfate). In acetate solutions, Ag, Cd, Ni, Co, Zn, Cu(II), and UO_2^{2+} ions precipitate. By reducing the acidity with sodium acetate, salts of Ag, Pb, Bi, Cd, Mn, Ni, Co, Fe(II), and Fe(III) can be precipitated from

nitrate solutions. Salts of Tl(I), Sb(III), Sn(II), Al, Cr(III), and the alkaline earths do not precipitate.²⁶⁸

Lauramidine hydrochloride. This reagent has been tested for the separation of uranium from phosphate solutions.²⁴⁸ At pH 2.45, 75% of the uranium was precipitated.

N-Lauryl-lauramidine. This reagent also has been tested for the separation of uranium from phosphate solutions. At pH 2.45, 85% of the uranium was precipitated.

Mercapto-acetic acid forms a greenish-white precipitate with tetravalent uranium.^{215,216}

Methylamine precipitates uranium (VI).²¹⁴

Methyl red causes uranium (VI) and aluminum to precipitate.²¹⁵

Morpholine precipitates uranium (IV) and (VI) as well as a number of other metal ions.²¹⁵ A 1 mg per ml solution of uranyl nitrate shows only a yellow color with the reagent. No precipitate is formed.²¹⁵

β -Naphthoquinoline in the presence of thiocyanate ion precipitates uranium (VI), mercury, bismuth, copper, cadmium, nickel, cobalt, zinc, and iron (III) from sulfuric or nitric acid solutions.²⁶⁹

Neo-cupferron (ammonium α -nitrosophthyl hydroxylamine) is similar to cupferron in its application. Uranium (IV) is precipitated by the reagent.²¹⁴

Nitrilotriacetic acid forms derivatives with uranium (VI), iron (III), nickel, and copper (II).²⁷⁰

m-Nitrobenzoic acid precipitates uranium (IV).³⁴

o-Nitrosohydroxylaminophenyl p-toluenesulfonate forms a yellow precipitate with hexavalent uranium. Many other metallic ions are precipitated by the reagent including Al, Cr, Fe(III), Co, Ni, Cu(II), Cd, La, Ce, Hg(II), Bi, Pb, and Th.²⁷¹

α -Nitroso- β -naphthol [†] deposits uranium (VI) as a very fine, yellow-orange³⁴ to brown²¹⁵ precipitate. Precipitation

is made in the pH range 4.0-9.4.²⁷² Metals such as iron, cobalt, nickel, and copper are precipitated from slightly acid solutions. Molybdenum as molybdate ion, zinc, and uranium (IV) form colored solutions.²¹⁵ Aluminum, chromium, and cadmium give no visible reaction.²¹⁵ The uranium compound can be extracted with amyl alcohol.³⁴

β -Nitroso- α -naphthol precipitates uranyl ion from slightly acidic solution. Iron, cobalt, nickel, copper, zinc, and molybdate ions also are precipitated by the reagent. Aluminum, chromium, cadmium, and uranium (IV) give no visible reactions. The precipitation of uranium (VI) is most nearly complete in an acetate buffered solution.²¹⁵

Oleic acid is a precipitant of uranium (VI).²¹⁶

Oxalic acid[†] precipitates uranium (IV) from acidic solution.³⁴ Strongly complexing organic compounds and fluoride, sulfate, and large amounts of phosphate ions interfere. Uranium is precipitated from 2-3N hydrochloric acid media. At lower acidities other metal oxalates precipitate, eg., Fe(II), Zn, Cu. At higher acidities the solubility of uranium (IV) oxalate increases. Immediate filtration of the precipitate may result in losses up to 1% of the uranium to the filtrate. Recovery of uranium may be made more quantitative by chilling the solution and allowing it to stand. Small amounts of manganese, iron, and nickel may be carried with the precipitate. Niobium, the rare earths, and thorium precipitate under similar conditions. If uranium is reduced on a mercury cathode prior to precipitation, no cations in moderate amounts interfere except rare earths and thorium.

Precipitation of uranium can be made in cold 1N nitric acid solutions.²³⁰ The uranium content should be less than 70 grams per liter. Enough oxalic acid is added to give a 10% excess of the amount theoretically required to precipitate $U(C_2O_4)_2$. The uranium then is reduced to the (IV)-state by

adding sufficient rongalite ($\text{Na}_2\text{H}_2\text{S}_2\text{O}_4 \cdot 2\text{CH}_2\text{O} \cdot 4\text{H}_2\text{O}$) to give a 7-10% excess of 1 mole of rongalite per mole of uranium.

Phenanthrene quinone monoxime precipitates uranium (IV) and (VI), aluminum, iron, cobalt, nickel, copper, and zinc.²¹⁵

Phenoxarsinic acid precipitates hexavalent uranium.²⁴⁸

Phthiocol precipitates U^{4+} , UO_2^{2+} , Zn^{2+} , and MoO_4^{2-} ions.²¹⁵

Picrolic acid precipitates tetra- and hexa-valent uranium and most other metallic ions.²¹⁵

Piperazine. (See ethylenediamine).

Pyridine[†] does not absorb carbon dioxide like ammonium hydroxide does. This reduces the possibility of carbonate interference or of alkaline earth precipitation in a uranium separation. Ammonium nitrate facilitates uranium precipitation. Sulfate ion hinders it. Separation can be made from alkali metals, alkaline earths, Mn, Co, Ni, Cu, and Zn. Zr, Ti, Fe, Cr, Al, and others are precipitated by the reagent.³⁴

Pyrogallol and pyridine combine to form a derivative with hexavalent uranium.²¹⁴

Quinaldic acid[†] forms a yellow, amorphous precipitate with uranyl ion.²¹⁶ Precipitation is made from a neutral or weakly acidic (pH 2-3) solution in the presence of ammonium chloride. The reagent precipitates a number of metals including copper, zinc, cadmium²⁷³ and uranium (IV).²¹⁵ Uranyl ion is not precipitated in the presence of alkali tartrate²⁷⁴ or a high concentration of acetate ion.²¹⁵

Quinizarin (1,4-dihydroxy-anthraquinone) precipitates uranium (IV) and (VI), iron and copper.²¹⁵

Rhodizonic acid forms a blue-black precipitate with tetravalent uranium. In neutral solutions, Ag, Hg(I and II), Tl, Pb, Cu(II), Cd, Bi, Zn, Sr, Ba, Fe(II), and UO_2 (II) ions are precipitated. At pH 2.8, Ag, Hg(I), Tl, Pb, Cd, Ba, and Sn(II) are precipitated.²⁷⁵

Salicylic acid. The sodium salt of the reagent forms a greenish-white precipitate with uranium (IV). Under conditions tested, Al, Cr, Fe, Co, Ni, Zn, Cd, MoO_4^{2-} , and UO_2^{2+} ions were not precipitated.²¹⁵

Sebacic acid precipitates uranium (IV).³⁴

Sodium acetate precipitates sodium uranyl acetate from neutral or weakly acidic solutions of uranyl salts.¹⁸⁹ The method is not very useful for the precipitation of traces of uranium. The solubility of sodium uranyl acetate in a solution 5M in sodium nitrate, 1M in acetic acid, and 0.5M in sodium acetate is about 100 mg per liter.²⁷⁶ Neptunium (VI) and plutonium (VI) also precipitate under these conditions. The addition of sodium acetate and zinc acetate to a neutral or weakly acidic uranyl salt solution precipitates the triple salt, sodium zinc uranyl acetate.¹⁸⁹

Sodium diethyldithiocarbamate precipitates tetravalent uranium, aluminum, iron, cobalt, nickel, copper, and cadmium.²¹⁵ Hexavalent uranium may be precipitated when both uranyl and reagent concentrations are sufficiently large.^{246,277}

Sodium ethyl xanthate forms an orange precipitate with uranium (VI).²¹⁶

Strychnine in the presence of fluoride ion precipitates hexavalent uranium as $7(\text{C}_{21}\text{H}_{22}\text{O}_2\text{N}_2\text{HF}) \cdot 6(\text{UO}_2\text{F}_2) \cdot 2\text{HF}$. The solubility of the precipitate in water at 25°C is 47.5 mg/100 ml; in 60% alcoholic solution at 25°C, 30 mg/100 ml.²⁷⁸

Tannic acid (digallic acid)[†] and tannin (a glucose ester of tannic acid)[†] react with uranium (VI) to give a deep-brown precipitate.³⁴ Elements arranged according to decreasing ease of precipitation by tannin are Ta, Ti, Nb, V, Fe, Zr, Hf, Th, U, Al.²⁷⁹ The position of chromium in this series is uncertain. Tantalum, titanium, and niobium may be separated by tannin in a slightly acidic oxalate solution. Uranium and others are precipitated by adding more tannin and by making

the solution ammoniacal. Uranium may be precipitated from such solutions in the presence of carbonate, acetate, or tartrate ions.²⁰⁹

Thioisamine. Uranium and cadmium are precipitated when an alkaline solution containing these elements is boiled with the reagent.²⁸⁰

Carriers. Trace amounts of uranium may be removed from solution by the use of gathering agents or carriers. The choice of a particular agent depends upon the conditions under which precipitation is to be made and upon subsequent chemistry to which the precipitate is to be subjected. Rodden and Warf³⁴ have described the application of several carriers: ferric, aluminum, and calcium hydroxide. The use of barium carbonate and thorium hexametaphosphate has been mentioned in the section on inorganic precipitants. Magnesium oxide and thorium peroxide have been used.³⁴ The oxide and salts of antimony,^{281,282} calcium fluoride,³⁴ and the phosphates of zirconium,¹⁹⁷ bismuth,²⁸³ and thorium^{197, 240} have been used to carry uranium from reduced solutions. Uranium (IV), in general, should behave similarly as neptunium (IV) and plutonium (IV). These are carried by lanthanum fluoride, ceric and zirconium iodates, ceric and thorium oxalates, barium sulfate, zirconium phosphate, and bismuth arsenate.²⁷⁶ Uranium (VI) does not carry with these agents providing the concentration of either carrier or uranium is not too large.

Complexes. The precipitation of uranium in normally precipitating media is inhibited by the formation of soluble complexes.³⁴ Carbonate ion is a very efficient complexing agent of uranyl ion. In ammonium hydroxide solution, uranium can be separated from iron, titanium, zirconium, and aluminum with carbonate ion present. In ammonium sulfide solutions, carbonate ion makes possible the separation

of uranium from manganese, iron, cobalt, zinc, and titanium. Ammonium carbonate prevents the precipitation of uranium with phosphate. Precipitation with sodium carbonate makes possible the separation of uranium from beryllium, manganese, iron, cobalt, nickel, zinc, titanium, zirconium, and the alkaline earths.

Sodium peroxide facilitates the separation of uranium and other metals with sodium carbonate. The addition of the peroxide alone to acid solutions of iron, cobalt, rare earths, titanium, zirconium, hafnium, and thorium causes their precipitation while uranium, if present, remains in solution.

Uranium does not precipitate with tannic acid in slightly acidic solution with oxalate ion present. Titanium, niobium, tin, tantalum, and tungsten are precipitated under such conditions. Oxalate ion also interferes in the precipitation of uranium by ammonia.

Tartrate, citrate, and malate ions prevent the precipitation of uranium by ammonium hydroxide or sulfide.³⁴

Salicylic acid and hydroxylamine have both been used to complex uranium in separations from rare earth elements.³⁴ Hydroxylamine has been used in separations between uranium and beryllium, aluminum, iron, and thorium.³⁴

Complexing agents that form weak complexes with uranium and relatively strong complexes with other metallic ions make separation possible between the two: uranium is precipitated by a suitable reagent; the other ions remain in solution. Ethylenediaminetetracetic acid (complexone II) and its disodium salt (complexone III) have been used successfully in this respect. Uranium has been precipitated with ammonia in the presence of complexones without interference from Al, Cr, Mn, Fe, Co, Ni, Cu, Zn, Cd, La, Ce, Hg, Pb, Bi, and the alkaline earths.²⁸⁴ The recovery of uranium is not entirely quantitative since the complexing agent increases the solubility of the ammonium uranate.^{196,285} The absorp-

tion of impurities in the precipitate may necessitate dissolution and reprecipitation of the uranium.¹⁹⁶ Beryllium and titanium follow the uranium chemistry.²⁸⁶

Quantitative recovery of uranium from the aforementioned cations: Al, Cr, Mn, Fe, etc., can be made with ammonium monohydrogen phosphate, $(\text{NH}_4)_2\text{HPO}_4$, in the presence of ethylenediaminetetracetic acid.^{285,287} Beryllium and titanium again interfere. Small amounts of titanium may be complexed with hydrogen peroxide before the addition of other reagents.²⁸⁷

Sen Sarma and Mallik²⁸⁸ have studied the separation of uranium from other elements using 8-hydroxyquinoline (oxine) as precipitant and complexone III as complexing or masking agent. It was found that complexone had no masking action on uranium in the pH range 5-9. In a solution buffered with acetic acid and ammonium acetate at pH ~5.3 quantitative separation was reported between uranium and Al, Mn, Fe(III), Co, Ni, Cu, Zn, Zr, Cd, rare earths, Pb, Bi, Th, and P_2O_5 . In ammoniacal medium at pH ~8.4, a similar separation was made from V_2O_5 , MoO_3 , and WO_3 . Steele and Taverner,¹⁹⁸ however, were unable to duplicate the above results.

2. Solvent extraction. The solubility of uranyl nitrate in organic solvents has long been recognized.¹⁸⁵ The ability of diethyl ether to extract this salt has been used in systems of analysis for many years. However, it is only within recent years (starting in the 1940's) that widespread use has been made of solvent techniques as a means of separating and purifying inorganic substances in general ²⁸⁹⁻²⁹⁹ and uranium in particular.^{191,192,194,197-199,300-305}

The conditions under which uranium may be extracted are many and varied. In the present paper, extraction from aqueous solution is considered. However, extraction from solid phases³⁰⁵⁻³⁰⁷ and slurries³⁰⁵ has been investigated

and a favorable uranium partition has been found. Conditions which affect the extraction of uranium from aqueous solution by organic solvent are the composition of the aqueous phase, the nature of the organic phase, the temperature, and the time of equilibration. In the aqueous phase, such factors as uranium, acid, common anion, foreign anion, and foreign cation concentration must be considered. The nature of the organic phase depends upon the type and concentration of solvent and diluent. If the organic phase is not initially barren, its concentration of uranium, acid, etc., affects partition.

Because of the number of variables and the large number of uranium solvents, one cannot consider, in a volume of this size, each solvent in the light of each variable. Indeed, the behavioral relation between solvent and the aforementioned variables is known for only a few well-studied solvents. The purpose of the present paper is to provide information on the conditions best-suited for the quantitative extraction of uranium or for the separation of uranium from interfering elements. This is done as much as possible in graphic or tabular form.

The solvents are divided into five general classifications: 1) ethers, esters, ketones, and alcohols; 2) organo-phosphorous compounds; 3) amines and quaternary ammonium salts; 4) carboxylic acids; 5) chelating agents. Dialkylphosphoric acids, eg., dibutyl phosphate, are classified as organophosphorus compounds rather than chelating agents. Carboxylic acids are classified as such, although some may also be considered chelating agents, eg., salicylic acid. A number of extractants may serve also as diluents or secondary solvents for other extractants. Such systems are described under the primary extractant. For example, a cupferron-hexone system is described under "cupferron" rather than under "hexone."

In the discussion, the terms "extractant" and "solvent" are often used interchangeably. "Diluent" is used to describe a secondary solvent rather than the term "inert solvent." The choice of diluent may appreciably affect the partition of uranium. A number of terms that are frequently used are defined below.

Partition or extraction coefficient:

$$\alpha = \frac{C_o}{C_a} = \frac{\text{concentration of a substance in the organic phase}}{\text{concentration of the same substance in the aqueous phase}}$$

Percentage extracted:

$$P = \frac{\alpha}{1+\alpha} \times 100, \text{ when equal volumes of both phases are present after shaking.}$$

Mass ratio:

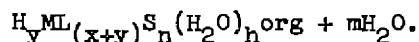
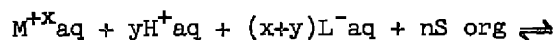
$$\mu = \frac{M_o}{M_a} = \frac{\text{amount of a substance in the organic phase}}{\text{amount of the same substance in the aqueous phase}} = \alpha \frac{V_o}{V_a}$$

Separation factor:

$$\beta = \frac{\frac{\text{concentration of substance A in the organic phase}}{\text{concentration of substance B in the organic phase}}}{\frac{\text{concentration of substance A in the aqueous phase}}{\text{concentration of substance B in the aqueous phase}}} = \frac{\alpha_A}{\alpha_B}$$

Equilibrium laws. The physical chemical principles involved in the solvent extraction of uranyl nitrate have been summarized in references 308-312. Detailed methods of treating the various equilibria involved have been devised.^{111, 313-315} A more simple approach, adapted from a paper by Carleson,³¹¹ is herewith presented.

It may be assumed that within a certain concentration range an average uranium complex is extracted. The complex is representative of a whole set of complexes and may be written $H_y M^{(+x)} L_{(x+y)} (H_2O)_h \cdot (S)_n \cdot M^{+x}$, in this case, may be U^{+4} or UO_2^{+2} . L, as written, is a singly, negatively charged ligand. It may be more highly charged. S represents a solvent molecule. The subscripts y, h, and n need not be integers. The reaction for the extraction mechanism may be written



The thermodynamic equilibrium constant for the reaction is

$$K = \frac{[H_yML_{(x+y)}S_n(H_2O)_h] \text{ org} \{H_2O\}^m}{[M^{+x}] [H^{+}]^y [L^{-}]^{(x+y)} [S]_{\text{Org}}^n \cdot f(\gamma)}$$

where $\{ \}$ and $[]$, respectively, represent the activity and concentration of a quantity in the aqueous phase unless otherwise identified by the symbol "org." $f(\gamma)$ represents the product of the activity coefficients. The partition coefficient is approximated by

$$\alpha \approx \frac{[H_yML_{(x+y)}S_n(H_2O)_h] \text{ org}}{[M^{+x}]}$$

The relation between the partition coefficient and equilibrium constant is

$$\log \alpha = y \log[H^{+}]aq + (x+y)\log[L^{-}]aq + n \log[S]org - m \log \{H_2O\} + \log f(\gamma) + \log K.$$

Information concerning the extracted species may be obtained by measuring the partition coefficient while varying the concentration of only one of the quantities. A knowledge of the activity coefficients is then required or the product of the activity coefficients in both phases must be kept constant.

As stated previously, the above approach to solvent extraction is a simplified version. It represents only an average extracted species. Among other things, it does not consider the effect of water activity in the organic phase, solvent activity in the aqueous phase, complex formation between the various components in either phase, or the formation of polynuclear species. These effects may be large or small depending upon the solvent, aqueous medium, and uranium concentration involved.

ETHERS, ESTERS, KETONES, AND ALCOHOLS. Uranyl nitrate is extracted by many polar solvents which contain donor oxygen atoms such as ethers, esters, ketones, and alcohols. 316-320

Extraction from water solutions is small unless the uranium concentration is appreciable. This is shown in figures 1-5 in which the data of McKay and co-workers,^{176,185} Warner,^{321,322} and Vesely, et al³²³ are plotted. In general, it has been observed³⁰¹ that:

1) the extraction coefficient of uranium decreases when the number of carbon atoms increases for a given homologous series of organic solvents,

2) for a molecule with a given number of carbon atoms and a given chemical functional group, solvents with straight chains are more efficient extractants than those with branched chains,

3) one or more double bonds in a molecule increases the efficiency,

4) primary alcohols are more efficient than secondary ones,

5) the coefficient of extraction increases with the solubility; but there is no well-defined relation between the two.

Evidence considered in the section on non-aqueous solutions indicates that uranium is extracted from aqueous nitrate solutions as hydrated, solvated uranyl nitrate, $UO_2(NO_3)_2(H_2O)_n$. Under appropriate conditions, the hydrated, solvated trinitrate-uranyl complex may be extracted. The relationship between partition coefficient and equilibrium constant for the extraction mechanism shows the extraction of the former species to be favored by large free nitrate and free solvent concentrations and by small water activity.

Effect of nitric acid. The addition of nitric acid to the aqueous phase favors the extraction of uranium by preventing or decreasing the hydrolysis of uranyl ion and by increasing the nitrate ion concentration.³²⁴ Nitric acid is extracted also by the organic solvents. This requires

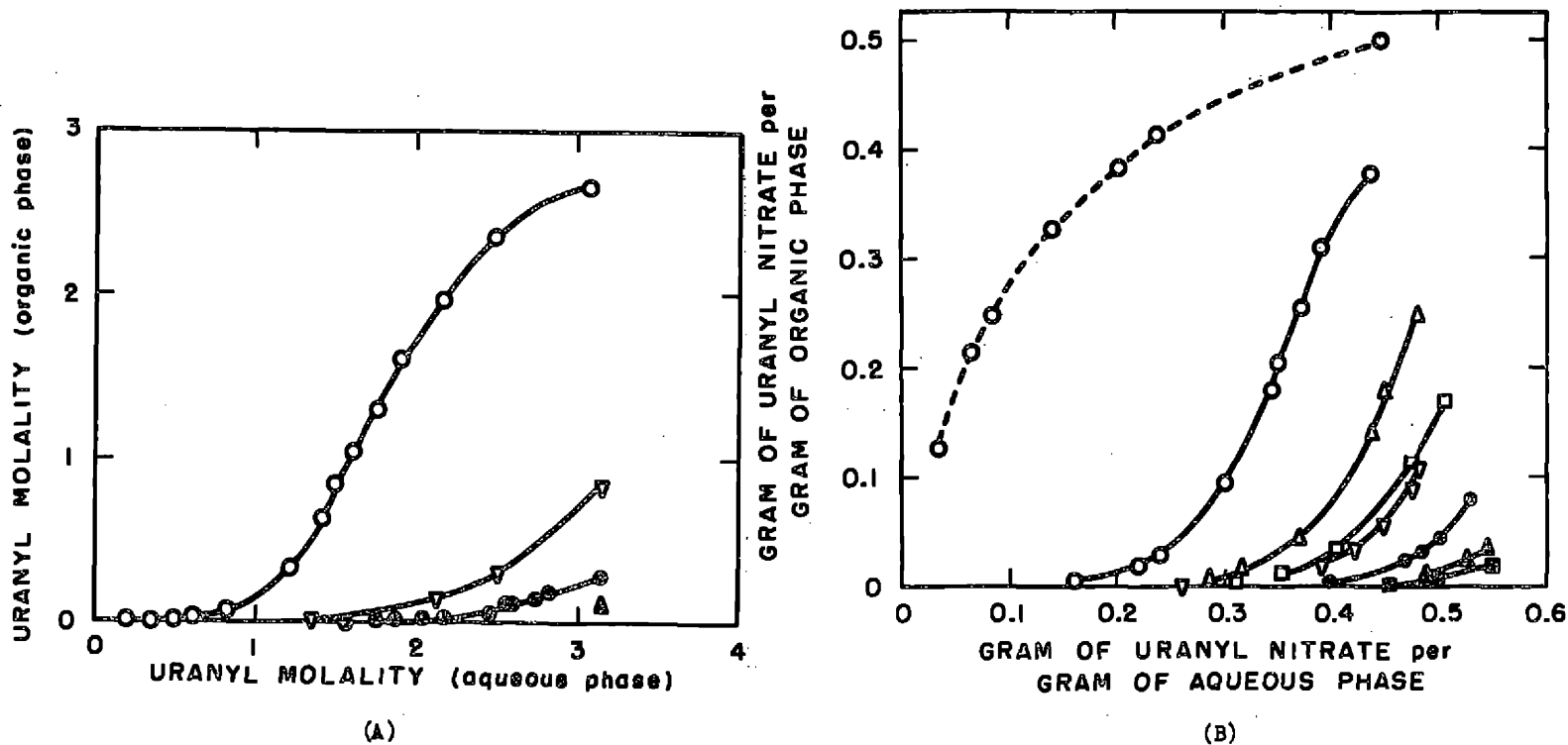
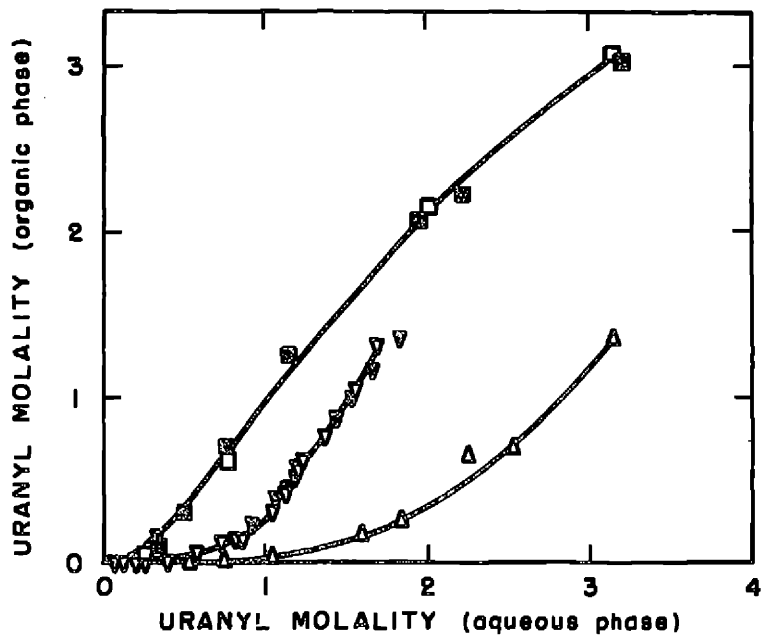


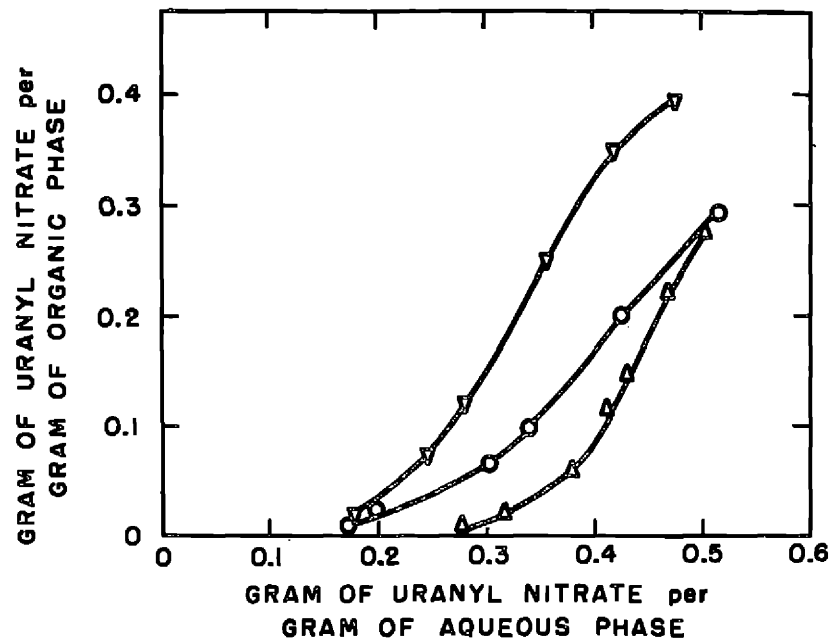
Figure 1. Partition of uranyl nitrate between water and simple ethers. O, diethyl ether. Δ , ethyl-n-propyl ether. \square , ethyl-n-butyl ether. ∇ , di-isopropyl ether. \odot , di-n-butyl ether. \triangle , di-n-hexyl ether. \square , 2,2'-dichlorethyl ether.

1-A. After E. Glueckauf, H. McKay, and A. Mathieson, reference 185. Temperature, 25°C except for diethyl ether: first three points at 25°C, last point at 20°C, remainder at 18°C.

1-B. After R. K. Warner, reference 321. Dashed curve represents the partition of uranyl nitrate between diethyl ether and a saturated ammonium nitrate solution. Temperature, 20°C.



(A)



(B)

99

Figure 2. Partition of uranyl nitrate between water and complex ethers. O, phenyl cellosolve. Δ , dibutyl cellosolve. ∇, ∇ , dibutyl carbitol. \square, \square , pentaether.

2-A. Open symbols, after E. Gleuckauf, H. McKay, and A. Mathieson, reference 185. Solid symbols, after A. Gardner, H. McKay, and D. Warren, reference 176. Temperature, 25°C.

2-B. After R. K. Warner, reference 321. Temperature 20°C.

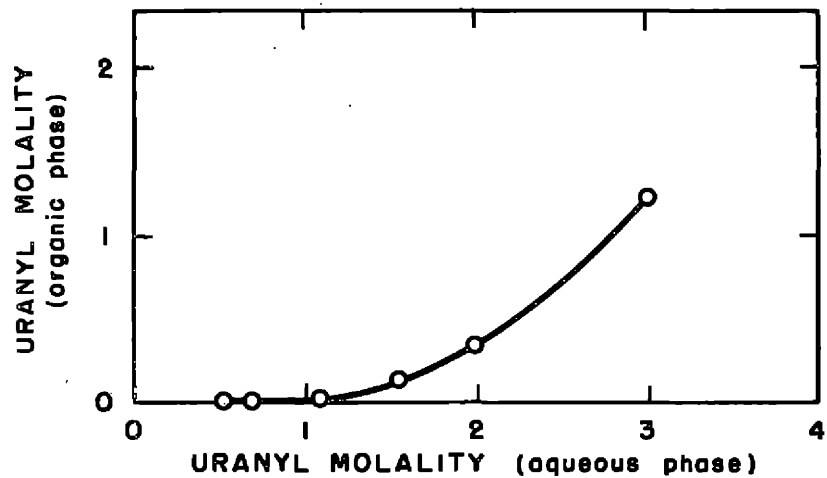


Figure 3-A. Partition of uranyl nitrate between water and isoamyl acetate. After E. Glueckauf, H. McKay, and A. Mathieson, reference 185. Temperature, 25°C.

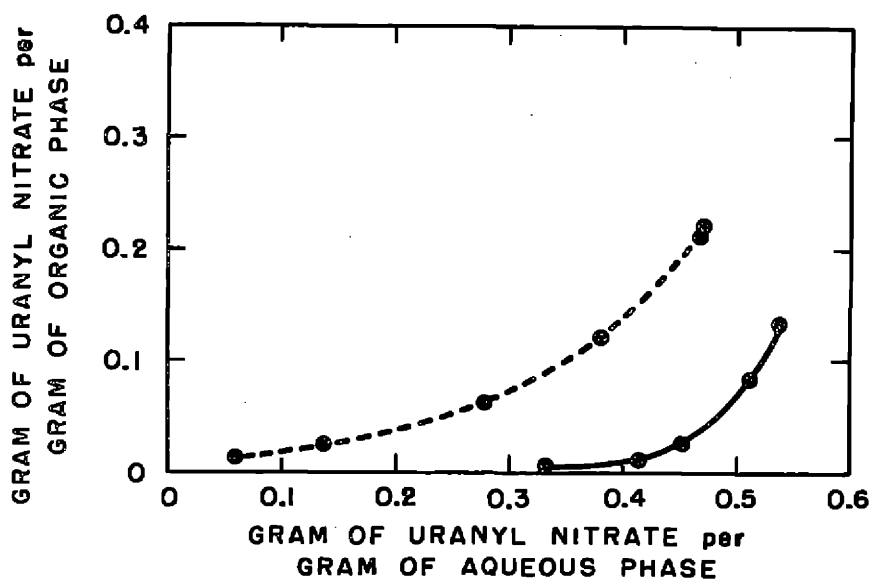
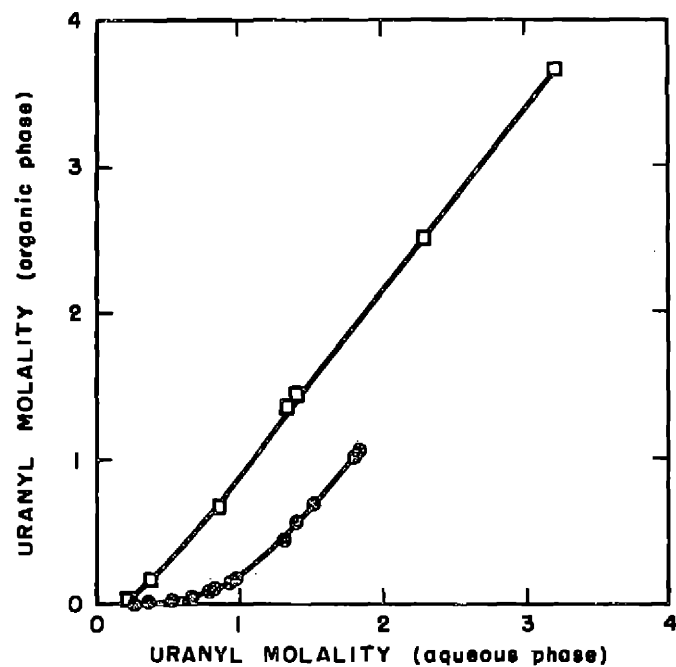
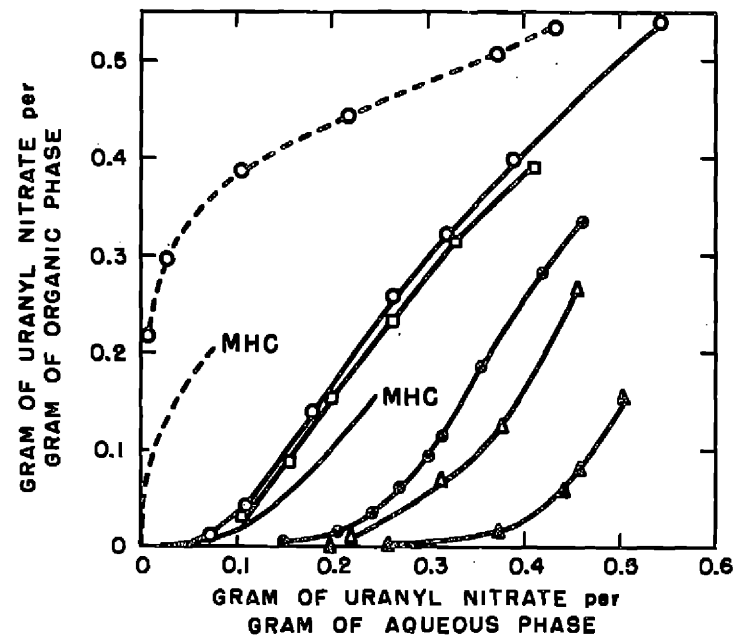


Figure 3-B. Partition of uranyl nitrate between water and nitromethane and saturated ammonium nitrate (dashed curve) and nitromethane. After R. K. Warner, reference 322. Temperature, 20°C.



(A)

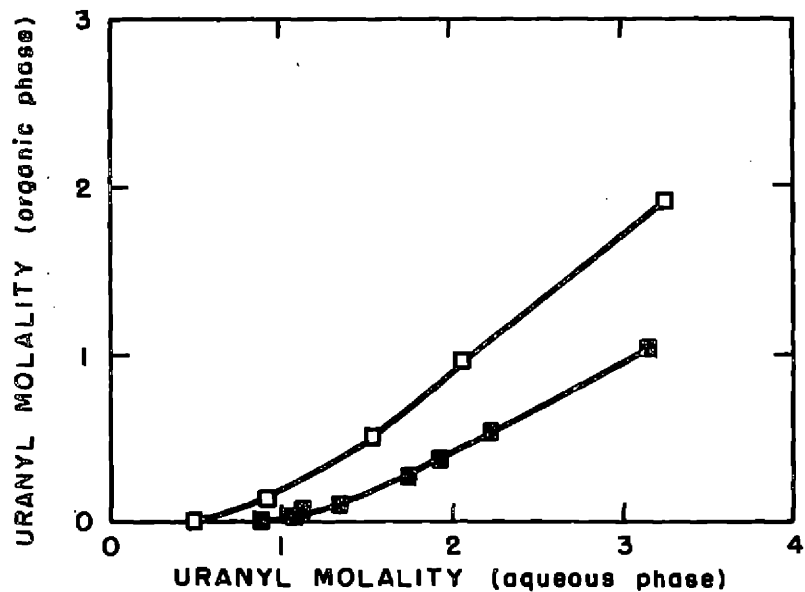


(B)

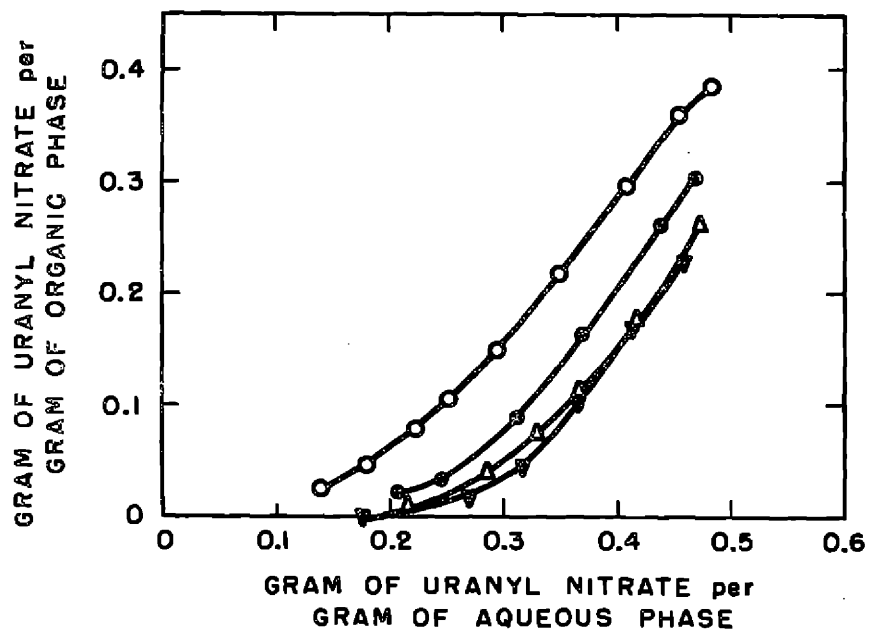
Figure 4. Partition of uranyl nitrate between water and ketones. O, methyl ethyl ketone. ⊙, methyl isobutyl ketone. △, methyl n-amyl ketone. ▲, di-isobutyl ketone. □, cyclohexanone. MHC, methyl cyclohexanone.

4-A. After E. Glueckauf, H. McKay, and A. Mathieson, reference 185. Temperature, 25°C.

4-B. MHC curves, after V. Veselý, H. Beranová, J. Malý, reference 323. Dashed curve--aqueous solution, 6M NH_4NO_3 . Remaining curves, after R. K. Warner, reference 321. Temperature, 20°C. Dashed curve--aqueous solution, saturated ammonium nitrate.



(A)



(B)

Figure 5. Partition of uranyl nitrate between water and alcohols. O, n-butanol. ●, n-pentanol. △, n-hexanol. ▼, methyl isobutyl carbinol. □, isoamyl alcohol. ■, sec-octyl alcohol

5-A. After E. Glueckauf, H. McKay, and A. Mathieson, reference 185. Temperature, 25°C.

5-B. After R. K. Warner, reference 321. Temperature, 20°C.

that it be replaced in continuous or multicontact extraction process. Large concentrations of nitric acid are generally not desirable. The formation of $\text{HNO}_3 \cdot \text{S}_n$ complexes reduces the amount of free solvent, the extraction of other elements is enhanced, and the danger of an explosive reaction between solvent and acid is increased. The formation of $\text{HM}(\text{NO}_3)_{x+1}$ species, which may be more easily extracted than $\text{M}(\text{NO}_3)_x$, is promoted by the addition of nitric acid. For uranium, however, the formation of the trinitrate-uranyl complex is far from complete, even in 16 M HNO_3 .⁷⁸

Effect of nitrate salts. The nitrate ion concentration may be increased by the addition of metal nitrates of significant solubility to the aqueous phase. This not only promotes the extraction of uranium but also the extraction of other elements whose nitrates are soluble in the organic solvent. In some cases, nitrates which serve as salting-out agents, eg. thorium may also be extracted in significant amounts. The extraction of other salting-out agents, eg., cesium, may be enhanced by the formation of uranyl trinitrate complexes, $\text{MUO}_2(\text{NO}_3)_3$.³²⁵ The ability of various nitrates to salt-out uranium has been related to the hydration of the cation,³²⁶ the activity coefficient of the pure nitrate salt,³²⁷ and the radius and charge of the cation.³²⁸ A salting-out agent which is highly hydrated facilitates extraction of uranium by reducing the water activity. In figure 6, the partition coefficient of uranium is plotted as a function of aluminum nitrate for several solvents. The partition of uranium between saturated ammonium nitrate solutions and diethyl ether, nitromethane, and methyl ethyl ketone is shown by the dashed curves in figures 1B, 3B, and 4B, respectively.^{321,322} Ammonium nitrate is widely used as a salting agent in spite of its relatively poor salting-out ability. The ease with which it is removed from solution or from heated samples

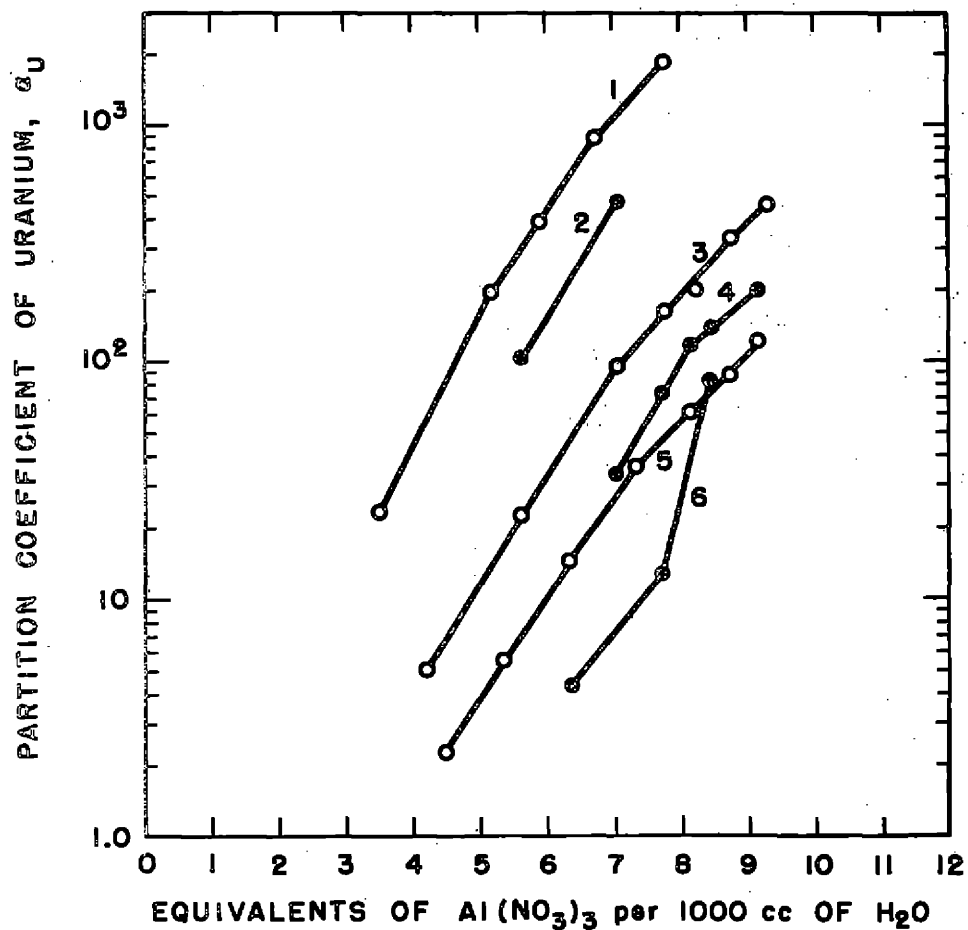


Figure 6. Effect of aluminum nitrate as salting-out agent on the extraction of uranium by various solvents. Curve 1, dibutoxytetraethylene glycol (pentaether); Curve 2, dibutoxytriethylene glycol; Curve 3, dibutoxydiethylene glycol (dibutyl carbitol); Curve 4, methyl isobutyl ketone (hexone); Curve 5, diethyl ether; and Curve 6, dibutylmonoethylene glycol (dibutyl cellosolve).

Data adapted from E. Evers and C. Kraus, reference 332.

Conditions: Aqueous phase - 2.0 to 6.0 grams of U per 100 cc of solution containing aluminum nitrate. Organic phase - solvent represented by curve. Equal phase volumes* equilibrated at 27°C.

*Equal or approximately equal phase volumes were employed in distribution experiments with dibutyl carbitol (C. A. Kraus, A-2322(1945)) and with hexone (C. A. Kraus, A-2324(1945)). It is assumed that the same volume ratio was used for other experiments.

makes its use advantageous.

The presence of nitrate salts which are sufficiently soluble in the organic solvent facilitates the extraction of uranium by formation of the trinitratouranyl complex, $\text{RUO}_2(\text{NO}_3)_3$.⁷⁸ This is discussed further under "Hexone."

Effect of other salts. Anions that complex uranium in the aqueous phase may seriously interfere with the extraction of the latter. Chloride, fluoride, sulfate, phosphate, and several organic anions have been studied for their interference. The adverse effects of these ions may be minimized by removing them from solution prior to uranium extraction, by complexing the anions with cations of salting-out agents, or by using an excess of an efficient salting-out agent to over-ride the anion interference. The inorganic anions may be precipitated as silver chloride, lanthanum fluoride, barium sulfate, zirconium phosphate, or ammonium phosphomolybdate. Fluoride ion is complexed by aluminum and calcium. Sulfate ion is complexed by ferric ion. Large amounts of sulfate ion are also precipitated by calcium nitrate. Phosphate ion is complexed by ferric and aluminum ions. Calcium nitrate has been used to counteract the effect of acetate and oxalate. The effect of chloride on the partition of uranium may be reduced in the presence of a strong salting-out agent. Chloride ion is more objectionable from the fact that it promotes the extraction of other elements, notably iron. In the presence of large amounts of interfering ions, particularly sulfate and phosphate, it is advisable to separate the uranium from solution prior to extraction. This may be done by precipitation with carbonate-free ammonium hydroxide. The precipitate is dissolved in nitric acid and the extraction is initiated. Ferric hydroxide may be used to carry trace amounts of uranium.

Uranium may be extracted from aqueous media other than

nitrate. Thiocyanate solutions have been found satisfactory. The extraction, however, is less selective from thiocyanate than from nitrate solutions.

Solvent action. The partition of uranyl nitrate is dependent upon the free solvent concentration. This is reflected in the coefficients of extraction of micro and macro amounts of uranium from highly salted aqueous solutions. The partition coefficient of trace amounts is larger than that of large amounts as a result of more available solvent. As mentioned previously, macro amounts of uranium extract more readily from water solutions or less highly salted aqueous solutions than do micro amounts. This effect may be attributed to the salting-out ability of uranyl nitrate itself.

The extraction of other elements is affected also by the uranium concentration. High loading of the solvent by uranium reduces the extraction of less preferred complex species. High uranium loading may be achieved by diluting the solvent with a secondary solvent in which uranyl nitrate is insoluble or significantly less soluble than in the primary extractant. Solvent dilution, in general, causes a decrease in the partition coefficient.^{319,321,329,330} Wohlhuter and Sauteron³³¹ have listed a number of aromatic and chlorine-substituted diluents in order of increasing harmfulness to uranium extraction: benzene, toluene, xylene, carbon tetrachloride and dichloroethylene, chloroform. Solvent dilution may be used also to improve upon the physical properties of the organic phase, eg., density, viscosity, etc.

The suitability of mixtures of oxygen-containing solvents as extractants for uranyl nitrate has been studied.^{318-320,329,330} Stover and co-workers³²⁰ reported that none of the mixtures they investigated were better than the pure solvent. Recently, however, Fomin and Morgunov³²⁹

and Vdovenko and Krivokhatskii³³⁰ have reported enhanced uranyl nitrate partition coefficients from solvent mixtures. Vdovenko and Krivokhatskii report that over ten such mixtures have been found. Among them are: di-isopropyl ether and β, β' -dichlorodiethyl ether, dibutyl ether and β, β' -dichlorodiethyl ether, diethyl ether and acetophenone, isoamyl alcohol and methyl isobutyl ketone. The enhanced extraction by solvent mixtures has been attributed to the formation of mixed solvates of uranyl nitrate.^{329,330}

Effect of temperature. The extraction of uranyl nitrate is decreased by a temperature increase. The partition coefficient of uranium is plotted as a function of temperature for several solvents in figure 7.³³²

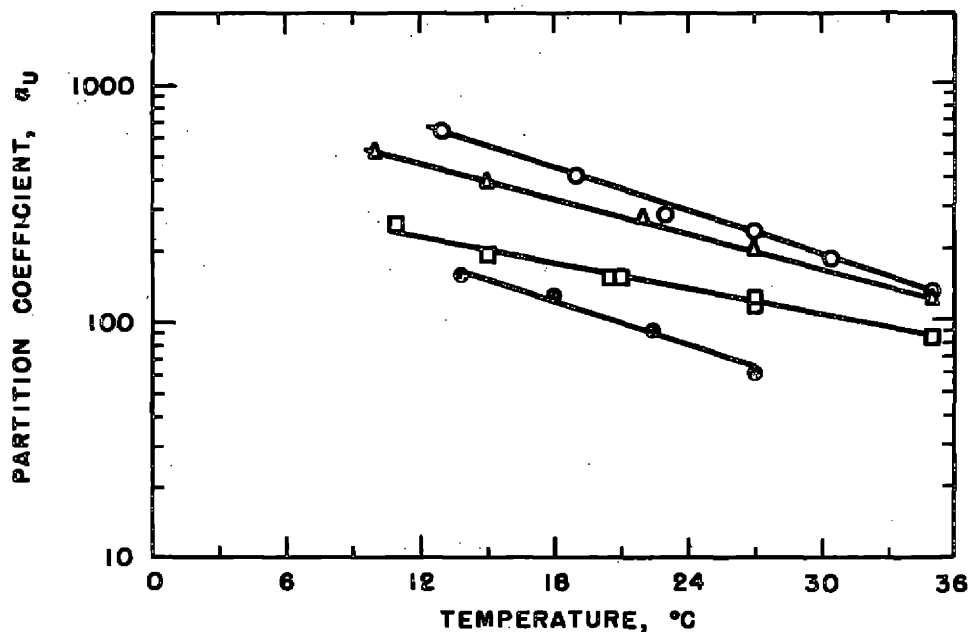


Figure 7. The effect of temperature on the extraction of uranium by organic solvents. O, dibutyl carbitol. □, hexone. ⊙, diethyl ether. Δ, pentaether.

After E. C. Evers and C. A. Kraus, reference 332.

The triangles represent aqueous solutions salted with 36.6 grams of $\text{Al}(\text{NO}_3)_3$ per 100 cc of water. All other symbols represent solutions salted with 58 grams of $\text{Al}(\text{NO}_3)_3$ per 100 cc of water. 2 to 6 grams of uranium per 100 cc of solution were extracted.

Re-extraction. Uranium is re-extracted from the organic solvents considered in this section by contact with water. Several water contacts may be required if large amounts of uranium or nitric acid have been extracted. Water-soluble salts whose anions complex uranyl ion, eg., ammonium sulfate, facilitate re-extraction.

Extraction of other elements. A number of elements other than uranium are extracted by oxygen-containing solvents. Those commonly found with irradiated uranium are hexavalent Np and Pu^{*}; pentavalent Pa; tetravalent Th, Np, Pu, Zr(+Nb), and Ce; and ruthenium complexes. Neptunium, plutonium, and cerium are made less extractable by reduction to lower oxidation states. Favorable separation of uranium from the other elements may be achieved by control of the nitric acid and salting-out agent concentrations. Free halogens are extracted. These elements may be eliminated from solution prior to uranium extraction. The halogens also combine chemically with a number of solvents; eg., iodine and hexone. The combined halogen is not re-extracted by water contacts.

General survey. The extraction of uranyl nitrate by polar oxygen-containing solvents has been investigated under a variety of conditions. The results of three surveys,³¹⁸⁻³²⁰ in which the experimental conditions were all different, are given in Table VIII.

* Am(VI) forms an extractable nitrate. Strong oxidizing conditions are necessary, however, for americium to be present in the (VI)-state. It is generally found in solution as Am(III).

Table VIII. Distribution of Uranyl Nitrate Between Various Oxygen-containing Solvents and Aqueous Nitrate Solutions.

Solvent	Survey 1 ^a			% Total volume in org. layer	Survey 2 ^b			Survey 3 ^c a _u
	Percent extracted				Percent extracted			
	U	Th	HNO ₃		U	Th	Pa	
<u>Ethers</u>								
Diethyl ether	96	60	-	62.5	35	0.0	0.19 [†]	0.09
n-Propyl ether	78	1.8	44	55	2.2	0.0	0.10	
Isopropyl ether	88	20	66.5	58				0.00
Dibutyl ether	48	< 1	45	53.5	1.6	0.0	0.05	0.00
Ethyl n-butyl ether	85	< 1	49	55				
Benzyl methyl ether	85	8	41	55				
β-β'-Dichloroethyl ether	45	< 1	37	53.5				
Diethyl cellosolve					55	24		0.29
Ethyl butyl cellosolve	95	60	57	61.5	59	7		
Dibutyl cellosolve	90	17	80	56.5				0.01
Phenyl cellosolve								0.00
Benzyl cellosolve						21		
Dibutyl carbitol	95	84	100	63.3	59	2.7	0.15 [†]	0.09
Dibutoxytetra-ethylene glycol (pentasther)	>95	99	100	66.6				0.54
Dimethyldioxane					34	3.8		
<u>Esters</u>								
Ethylacetate	88	77	-	65	43	0.6	0.26 [†]	
n-Propyl acetate	95	46	92	60	50	1.3	9.1	
Isopropyl acetate					54	1.2	11.9	
n-Butyl acetate	89	26	-	58	39	0.06	10.6	
sec-Butyl acetate					39		22.2	
Isobutyl acetate					21	0.02	0.18	
Amyl acetate	86(?)	16	-	56.5	31	0.3	31.9 [†]	
2-Ethylbutyl acetate					9	0.02		
2-Ethyl hexylacetate	72	< 1	95	54				
Butyl carbitol acetate	98	85	-	67				
Glycol diacetate								0.25
Ethylacetoacetate								0.11
Methyl proprionate					10-20	0.2	0.06	
Ethyl proprionate					22	0.03	0	
Ethyl butyrate	88	6	75	57.5	10	0		
Ethyl α-bromobutyrate	35	< 1	42	54				
Methyl benzoate					7.5	0		
Ethyl benzoate	80	1-3	54	56.5				
Ethyl caproate	80	3	60	55				
Diethyl maleate								0.07
Diethyl malonate					22			
<u>Ketones</u>								
Methyl ethyl ketone + 15% xylene					47	22		
Methyl n-propyl ketone					53	11		
Methyl isobutyl ketone (hexone)					57	20	20.9 [†]	0.004
Methyl n-amyl ketone					47	0.9	3.0 [†]	0.02
Methyl n-hexyl ketone								0.01
Diethyl ketone								0.08
Diisopropyl ketone					56	0.3	48.7 [†]	

Table VIII. - Continued

Solvent	Survey 1 ^a			% Total volume in org. layer	Survey 2 ^b			Survey 3 ^c
	Percent extracted				Percent extracted			
	U	Th	HNO ₃		U	Th	Pa	C ₁₄
Acetophenone								0.05
Cyclopentanone								0.80
Cyclohexanone						11		0.53
Methyl cyclohexanone						large		
m-Methyl cyclohexanone								0.24
p-Methyl cyclohexanone								0.30
Mesityl oxide	98	77.6	-	63		12.7		0.24
Isophorone						27		
<u>Alcohols</u>								
n-Butanol					44	5.2		
2-Ethyl butanol								0.01
sec-Butyl carbinol					30	0.6		0.01 ₃ **
tert-Amyl alcohol	88	63	-	61				
Diethyl carbinol					24	0.7		
Diisopropyl carbinol					11.6	0.0		
Tri-n-butyl carbinol	48	5.6	51	54				
2-Ethyl hexanol	75	10	77.5	56.5		0.0	72.1	
Heptanol					24	0.65	71.5 [†]	
Methyl t-amyl carbinol					36-50	4.7		
n-Octanol						0.5	74.8	
Heptadecanol						0.0	31.7	
2-Ethylhexanediol-1,3								0.09
<u>Miscellaneous</u>								
Nitromethane	60	32	62	54				
1-Nitropropane								0.00
2-Nitropropane					3.5	Trace	0.25	
Nitrobenzene					2.1	Trace	0.10	
o-Nitroanisole					3.5	Trace	0	
Trichloroethane					0	0	0	
Trichloroethylene					1.5	0.2		
Chloroform					0	0		
Dimethylsulfone								1.68

^a O. A. Johnson and A. S. Newton, reference 318. Equal volumes of organic solvent and aqueous solution (3M HNO₃, 3M Ca(NO₃)₂, 0.635M Th(NO₃)₄, 0.045 gm/ml UO₂(NO₃)₂) mixed 10 minutes. No extraction was detected with phenyl ether, p-bromoanisole, p-bromophenetol, anisole, isoamyl nitrite, chlorobenzene, xylene, ethyl iodide, butyl bromide, and trichloroethylene.

* indicates mutual solubility of aqueous solution and organic solvent.

^b E. K. Hyde and M. J. Wolf, reference 319. Equal volumes of organic solvent and aqueous solution (1N HNO₃, 3N NH₄NO₃, 2N Th(NO₃)₄, 5 x 10³ - 10⁴ c/min U²³³ at 52% counting yield) mixed 5 minutes.

[†] Aqueous phase: 1N HNO₃, 2N NH₄NO₃, 2N Th(NO₃)₄, 10⁴ - 3 x 10⁵ c/m Pa²³³ at 10% counting yield.

[‡] Aqueous phase: 1N HNO₃, 2N Al(NO₃)₃, 3N Th(NO₃)₄, 10⁴ - 3 x 10⁵ c/m Pa²³³ at 10% counting yield.

^c C. N. Stover, Jr., H. W. Crandall, D. C. Stewart, and P. C. Mayer, reference 320. Equal volumes of organic solvent and aqueous solution (0.34M HNO₃, 0.50M UO₂(NO₃)₂ · 6H₂O) mixed 2 hours at 25°C.

** Solvent identified as methyl isobutyl carbinol; formula given as CH₃CH₂CH(CH₃)CH₂OH (sec-butyl carbinol).

Diethyl ether

Aqueous nitrate systems. The extraction of uranyl nitrate by diethyl ether is widely used in radiochemical separations because of the selectivity of the extraction. Disadvantages of the method are the high volatility and low flash point of the solvent and the relatively low distribution of uranium into the solvent.

The partition of uranyl nitrate between diethyl ether and water is illustrated in figure 1.^{185,321} With both phases saturated with uranyl nitrate at 25-26°C, the distribution coefficient is about 0.68.³²⁶ The effect of nitric acid upon the partition of uranium^{333,334} and nitric acid^{326,333-335} itself is represented in figure 8. Furman, Mundy, and Morrison³²⁶ report that the pH of the aqueous phase should be 4 or less for complete extraction of uranium to occur. The influence of ammonium nitrate and calcium nitrate upon the extraction of nitric acid is also shown in figure 8.³²⁶

Figures 9 and 10 demonstrate the influence of various salting-out agents on the distribution of uranyl nitrate.^{326,328} The nitrate concentration plotted in figure 9 includes that of the salting-out agent plus that of uranyl nitrate left after extraction by an equal volume of ether. The latter contributes only a few percent to the total nitrate concentration in most cases. A notable exception is the iron (III) point at 1.18M nitrate concentration. In this instance, 0.82M nitrate ion is attributable to ferric nitrate and the remainder to uranyl nitrate. The nitrate concentration of the salting-out agent is plotted in figure 10. Uranyl nitrate contributes little to the nitrate concentration since only one gram of uranyl nitrate was used per 100 grams of initial aqueous solution. Furman, Mundy, and Morrison³²⁶

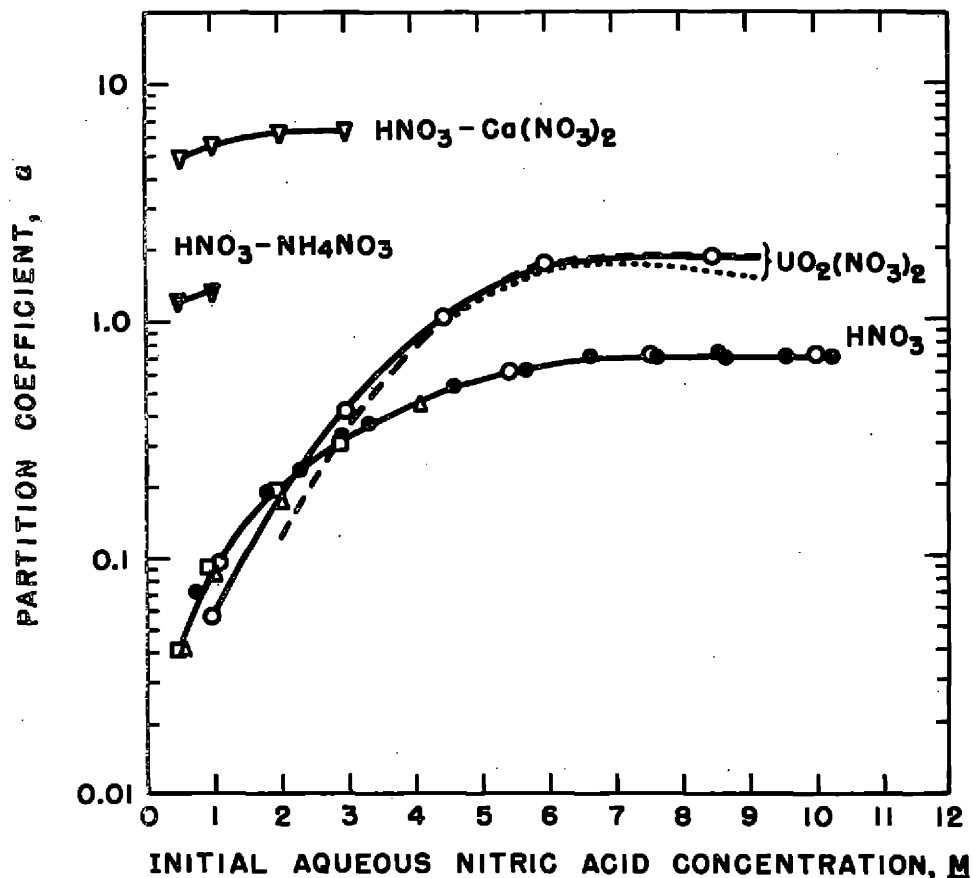


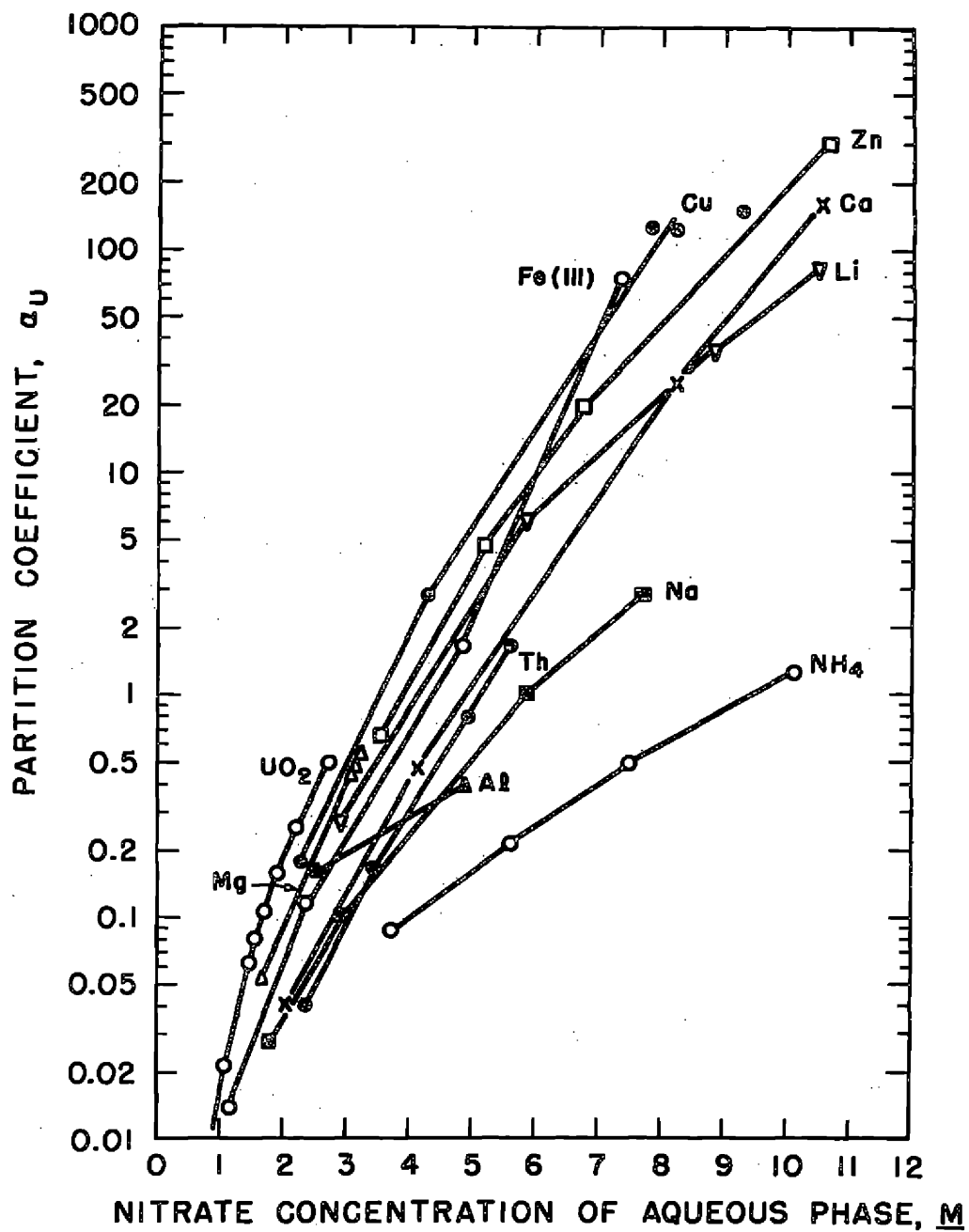
Figure 8. The extraction of uranyl nitrate and nitric acid by diethyl ether.

Uranyl nitrate extraction:

O, after R. Bock and E. Bock, reference 333. Initial U concentration, 0.1M; Temperature, $20^{\circ} \pm 1^{\circ}\text{C}$; $V_0/V_a = 1$. Dashed and dotted curves, after J. Kooli, reference 334. Temperature, 25°C ; $V_0/V_a = 1$; Initial U concentration, -----50, 150, 450 mg uranyl nitrate (hexahydrate) per 15 ml;1350 mg uranyl nitrate (hexahydrate) per 15 ml.

Nitric acid extraction:

O, after R. Bock and E. Bock, reference 333. Temperature, $19^{\circ} \pm 1^{\circ}\text{C}$; $V_0/V_a = 1$. ●, after J. Kooli, reference 334. Temperature, $25.0 \pm 0.1^{\circ}\text{C}$; $V_0/V_a = 1$. □, after A. Grinberg and G. Lozhkina, reference 335. Temperature, 20°C ; $V_0/V_a = 1$. Δ, ▽, ▼, after N. Furman, R. Mundy, and G. Morrison, reference $V_0/V_a = 1$; $\text{HNO}_3 - \text{Ca}(\text{NO}_3)_2$, 100 g. of $\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$ per 100 ml of initial solution plus HNO_3 ; $\text{HNO}_3 - \text{NH}_4\text{NO}_3$, 80 g. of NH_4NO_3 per 100 ml of initial solution plus HNO_3 .



(See facing page for legend.)

Figure 9.

Effect of various nitrates upon the partition coefficient of uranium with diethyl ether.

After N. Furman, R. Mundy, and G. Morrison, reference 326.

The nitrate concentration plotted as the abscissa includes that attributable to the equilibrium concentrations of uranyl nitrate in addition to that of the initial concentration of salting-out agent.

Conditions:

Salting-out Agent	[U], g/l	T. °C	V_o/V_a
NH_4NO_3	25	25-6	1
$NaNO_3$	(25)	24-6	1
$LiNO_3$	(25)		1
$Ca(NO_3)_2$	(25-100)	25-29	1
$Mg(NO_3)_2$	(25-100)	27-28	1
$Zn(NO_3)_2$	(50)	28-31	1
$Cu(NO_3)_2$	(~8->600)		
$Fe(NO_3)_3$	(~8->200)		1
$Al(NO_3)_3$	(25-100)	29	1
$Th(NO_3)_4$	(50)		1

The uranium concentrations in parentheses have been estimated by roughly adding the equilibrium uranium concentrations of both aqueous and solvent phase.

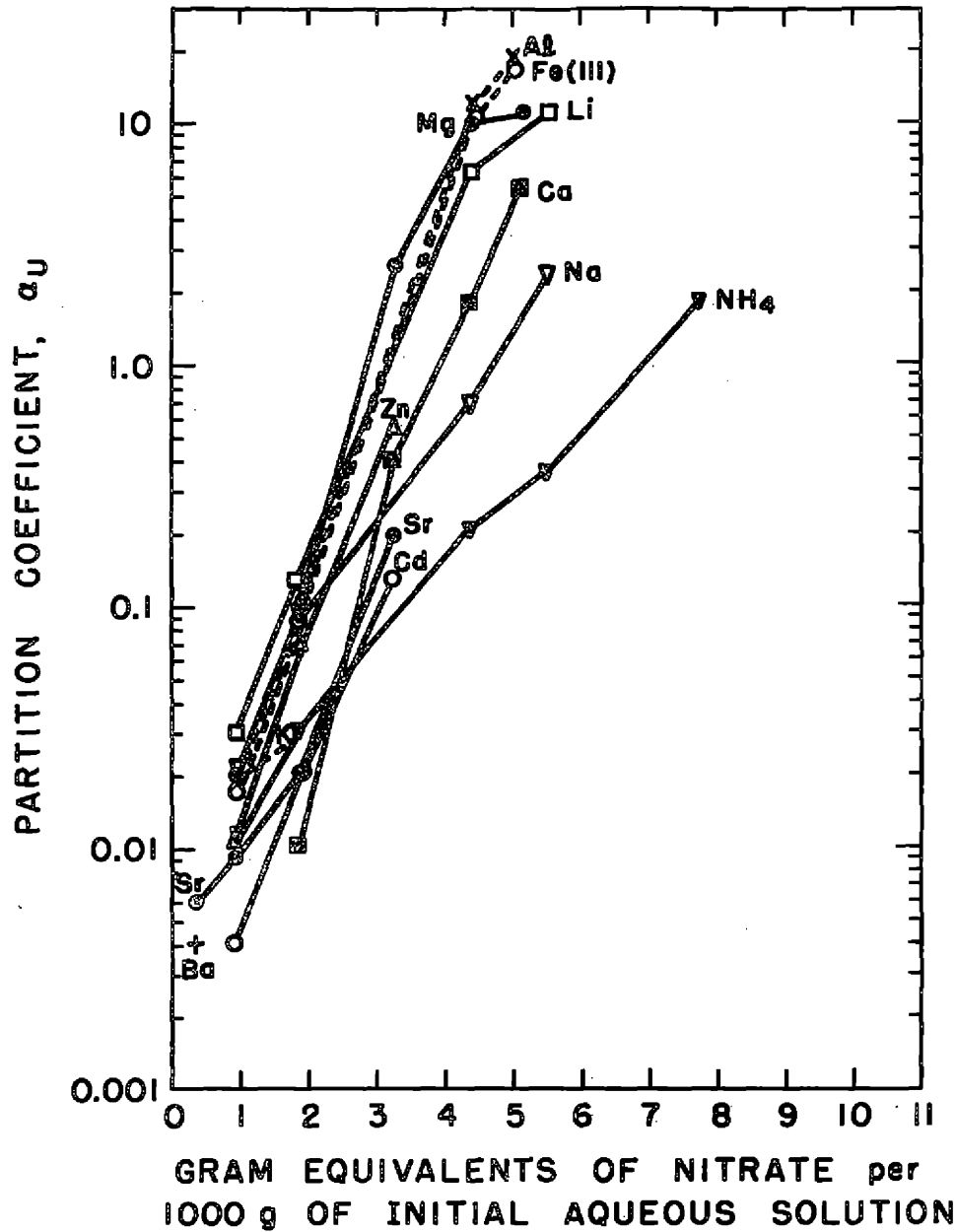


Figure 10. Effect of various nitrates upon the partition coefficient of uranium with diethyl ether.

After V. Vdovenko and T. V. Kovaleva, reference 328.

Conditions: 1 g. $\text{UO}_2(\text{NO}_3)_2$ in 100 g. of initial aqueous solutions; Temperature, 25°C ; and $V_o/V_a = 1$.

observed that the salting action of a mixture of nitrates could be reasonably predicted by the following method:

The logarithm of α_1 for each salt at a given total nitrate molarity is divided by the total nitrate molarity. These individual quotients are then multiplied by the nitrate molarity of the respective salts. The sum of the resulting products is then equal to the logarithm of the predicted partition coefficient.

Hellman and Wolf³³⁶ have studied the salting action of various nitrates in the presence of nitric acid and thorium nitrate. Some of their results are listed in Table IX. From the data it may be observed that (1) thorium nitrate is generally a less effective salting-out agent on a normality basis than other metal nitrates and (2) the extraction of uranium becomes less efficient as the amount of extracted thorium becomes appreciable.

The effect of several foreign anions on the extraction of uranyl nitrate by diethyl ether is given in figure 11³³⁷ and Table X.³²⁶ Arsenate, molybdate, and vanadate ions also interfere with the extraction of uranium. The effect of these ions may be offset by the addition of ferric nitrate to the solution.³³⁸

The partition of a large number of elements between various aqueous nitrate systems and diethyl ether is given in Table XI and illustrated in figures 12-14. The increased distribution of heavy elements and fission product elements with increased nitric acid concentration should be noted. For a selective uranium extraction the nitric acid concentration should be minimal. Hyde³³⁹ has recommended an aqueous phase 0.5-1M in nitric acid and 2.5M in magnesium nitrate for the quantitative extraction of uranium by diethyl ether. More selective extraction of uranyl nitrate may be made from a saturated ammonium nitrate-solution, 0.05-0.1M in nitric acid. The extraction can be made quantitative by repeated contacts with ether.

Table IX. Distribution of Uranium and Thorium between Diethyl Ether and Aqueous Solutions Containing Various Amounts of Metal Nitrates^a

Salting agent	Total nitrate normality	Composition of initial aqueous solution		
		0.5N HNO ₃ + salting agent		0.5N HNO ₃ + 1N Th(NO ₃) ₄ + salting agent
		U extracted, %	U extracted, %	Th extracted, %
HNO ₃	3	23	10	0.1
	5	52	40	3.3
	7	62	46	21
LiNO ₃	2	10	22	0
	3.5	36		
	5.5	74	31	2
	7.5	81	25	19
NH ₄ NO ₃	3		8	0.0
	5		30	0.1
	5.5	25		
	7		47	0.5
	9		57	0.8
Ca(NO ₃) ₂	10	59		
	12		59	1.2
	2.37	13		
	2.75		11	0.0
	3.62	32		
	4.0		25	0.0
Mg(NO ₃) ₂	4.87	63		
	5.25		59	0.1
	7.50	99	96	1.2
	1.50	3		
	2.17		9	0.0
	2.58	18		
Mn(NO ₃) ₂	3.15		18.5	0.0
	3.60	38		
	4.14		44	0.00
	5.5	99	99	0.31
	2.5	18		
	3.5		21.0	0.0
	5.0	74		
Cu(NO ₃) ₂	6.0		65	5.5
	7.0	87		
	7.5		70	
	8.5		72	43
	2.5	12		
	3.0		12	0.0
La(NO ₃) ₃	5.0	69	50	0.5
	7.5	91		
	8.0		62	19
	10.5	87		
	11.0		38	52
	2.5	14		
Al(NO ₃) ₃	4.5	61	5	0.0
	4.5		18	
	6.0	93	18	0.2
	6.5		39	2.1
	2.0	18		
Th(NO ₃) ₄	2.5		13	0.0
	4.0	54		
	5.5		63	2.8
	6.0	96		
	7.5		57	17.6
Th(NO ₃) ₄	3.0	8		0.0
	5.0	32		0.14
	9.0	58		8.0
	11.0	52		11.4

^a After N. N. Hillman and M. J. Wolf, reference 336.

Five ml of ether were shaken 10 minutes with 5 ml of the aqueous phase of appropriate composition. 5,000 to 8,000 c/m of U²³³ tracer (52% counting yield) were supplied to the aqueous phase. Variation of the tracer from 100 to 100,000 c/m in 5 ml did not change the % extracted.

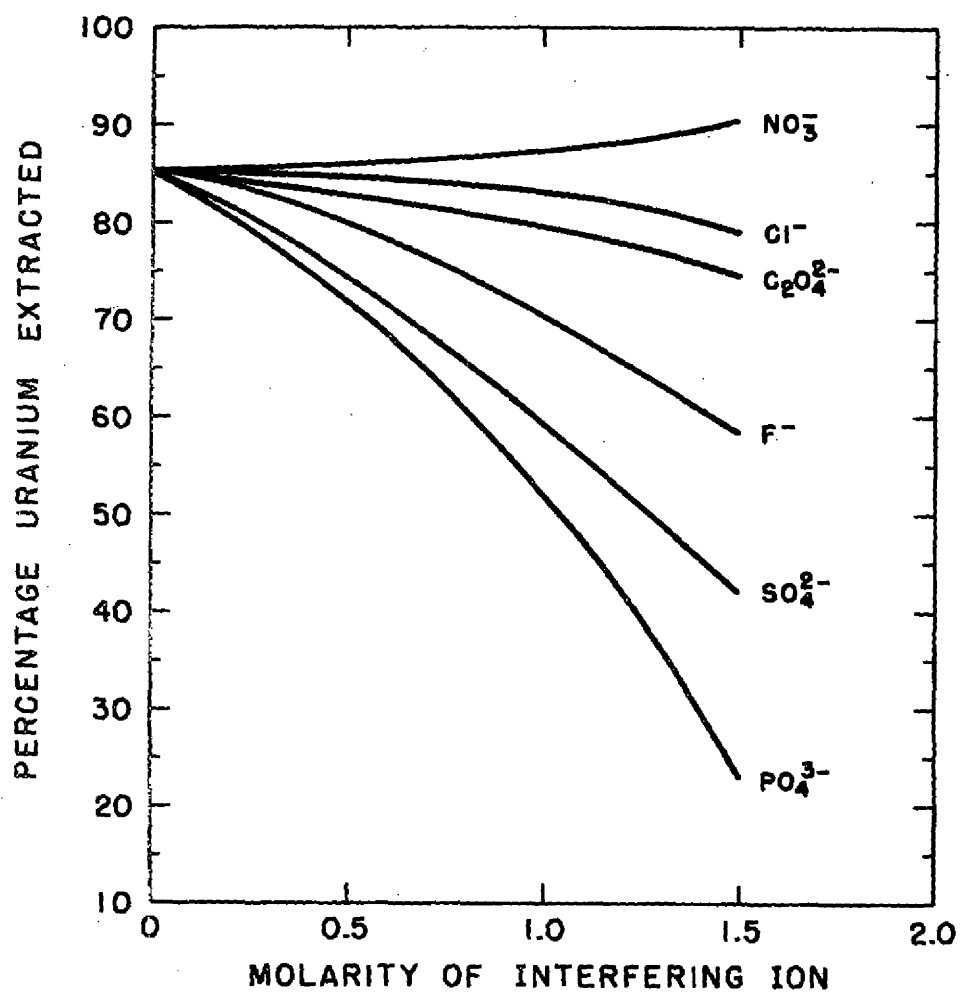


Figure 11. The effect of various anions on the extraction of uranyl nitrate by diethyl ether.

After T. R. Scott, reference 337.

Conditions: Aqueous phase - varying amount of anion, 3N HNO₃, 1M Fe(NO₃)₃; aqueous phase and organic phase shaken 1 minute at room temperature.

Table X. The Effects of Various Acids and Anions upon the Distribution Coefficient of Uranyl Nitrate to Diethyl Ether.^a

Acid or salt present	Composition of aqueous phase		
	4.23M $\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$ (1000 g/l)	8M NH_4NO_3 (640 g/l)	6.86M NH_4NO_3 (549 g/l)
None	19.95	0.645	0.364
HCl, 1N	10.24	0.336	
HCl, 2N	6.34	0.182	
HNO_3 , 1N	43.56	1.162	
HNO_3 , 2N	71.2	1.95	
$\text{CH}_3\text{CO}_2\text{H}$, 1N	15.7	0.662, 0.616	
$\text{CH}_3\text{CO}_2\text{H}$, 2N	10.52	0.720, 0.762	
H_2SO_4 , 0.0039N		0.613	
H_2SO_4 , 1N	29.6	0.024	
H_2SO_4 , 2N	23.5	0.019	
H_3PO_4 , 0.0058N		0.609	
H_3PO_4 , 1N	0.01	0.01	
H_3PO_4 , 2N	0.01	0.01	
H_3PO_4 , 1N	3.98	0.067	
HNO_3 , 2N			
$(\text{NH}_4)_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$, 0.7 g/100 ml			0.0847
$\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$, 0.7 g/100 ml			0.0800

^a After Furman, Mundy and Morrison, reference 326.

The initial volumes of ether and aqueous solution were equal. Room temperature.

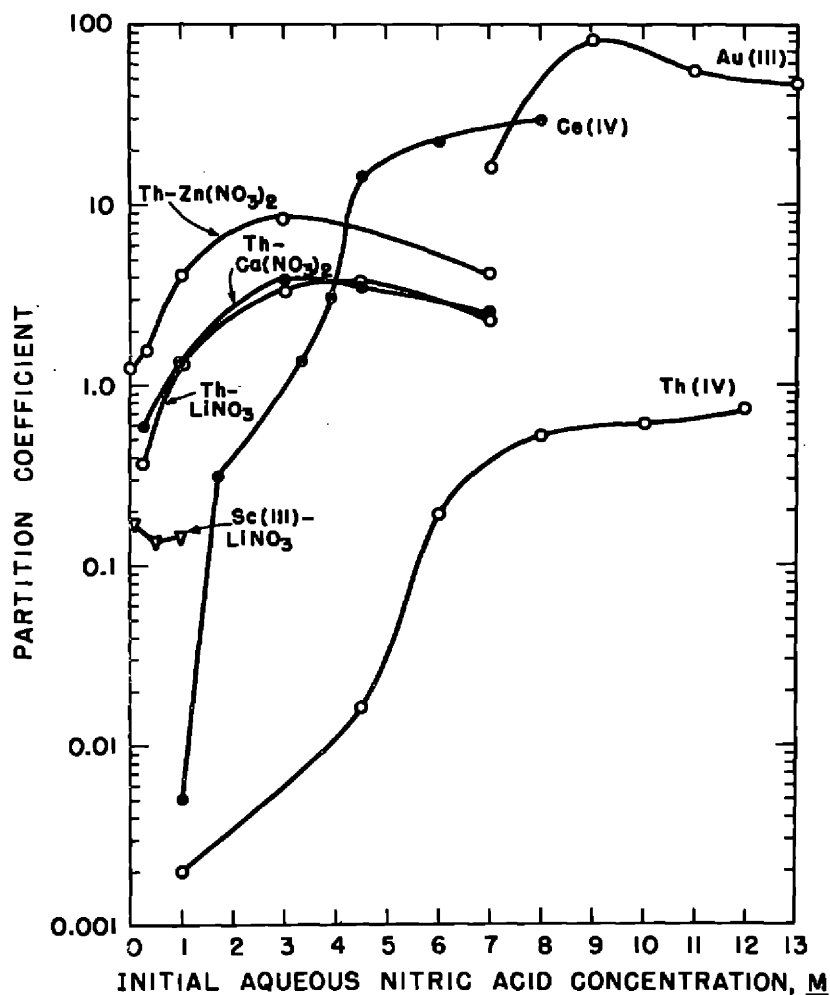


Figure 12. The extraction of various metal nitrates by diethyl ether.

After R. Bock and E. Bock, reference 333. Conditions: equal phase volumes.

Metal nitrate	Initial aqueous concentration	Temperature
Th	0.1M	20 + 1°C
Saturated LiNO ₃ , Ca(NO ₃) ₂ , or Zn(NO ₃) ₂	solutions	20 + 1°C
Ce(IV)	0.1M	room
Au(III)	0.1M	20°C
Sc(III)	0.1M	20°C
Saturated LiNO ₃	solution	

Table XI. Distribution Coefficients of Elements between Diethyl Ether and Various Nitrate

Element	Solutions.							
	H ₂ O ^a	0.8M HNO ₃ ^a 10M NH ₄ NO ₃	0.8M HNO ₃ ^a 4.2M Ca(NO ₃) ₂	1M HNO ₃ ^b satd NH ₄ NO ₃	1M HNO ₃ ^b satd LiNO ₃	8M HNO ₃ ^c	8M HNO ₃ ^c	1.4M HNO ₃ ^d 8M NH ₄ NO ₃
Al	<0.001	<0.001	<0.001			0.0006		<0.0001
Am(VI)				0.6 ^{bb}				
Sb(V)						<0.01		
As		0.0015*0.007*	0.048*			0.168		
Ba	<0.0005	<0.0005	<0.0005					<0.0001
Be						0.014		0.002
Bi	0.0001	0.0003	0.007			0.073		0.00031 0.021
B		0.01*	0.033-0.39*					
Cd	<0.00001	<0.00001	0.00001			0.003		<0.0001 0.0002
Ca	0.0003	0.0005	0.0005					<0.0001
Ce(III)								<0.0009
Ce(IV)			27*	0.095	0.314	30.3(29.4)		
Cr(III)	<0.0001	<0.0001	<0.0001					<0.0001
Cr(VI)						< 17		
Co	<0.0001	<0.0001	<0.001			0.002		<0.0001
Cu	0.0004	0.0002	0.0004			0.005		<0.0001 0.00082
Dy	0.0004	0.0003	0.002					
Gd	0.0001	0.00001	0.00076					
Ga						< 0.002		
Ge						0.022		
Au(III)						32.3(36.5-50)		
H						0.72(0.71)		
In	<0.0004	<0.0004	0.0003			0.001		
Fe	0.0005	0.0005	0.001			0.0013		<0.0001 0.0017
La								<0.0001
Pb	0.003?	<0.0002	<0.0002			0.005		
Li	<0.00002	0.0001	0.0002					0.0003
Mg	0.0002	0.0001	0.0002					<0.0001
Mn(II)	<0.0001	<0.0001	<0.0001			0.002		<0.0001
Mn(VII)						<0.005		
Hg	0.003	<0.0001	<0.0001			0.049		<0.0001 0.0003
Mo		<0.001****				0.0065		
Nd								<0.0009
Ni	<0.0001	<0.0001	<0.0001			0.0006		<0.0001
Np(IV)							0.2	
Np(V)							1.1	
Np(VI)							4.4	
P						0.256		
Pu(III)								<0.001
Pu(IV)							10.3	
Pu(VI)							2.4	1.5
K	0.0005	0.0002	0.002?					<0.0001
Ra	<0.00025	<0.00025	<0.00025					
Rare earths	0.00076	<0.00005	<0.0031					
Re	<0.015**	<0.015**	<0.015**					
Rb								0.00014
Sm								<0.0009
Se				0.013	0.146	0.001		
Ag						0.025		<0.0001 0.001

Table XI. - Continued

Element	H ₂ O ^a	0.8M HNO ₃ ^a 10M NH ₄ NO ₃	0.8M HNO ₃ ^a 4.2M Ca(NO ₃) ₂	1M HNO ₃ ^b satd NH ₄ NO ₃	1M HNO ₃ ^b satd LiNO ₃	8M HNO ₃ ^b	8M HNO ₃ ^c	1.4M HNO ₃ ^d 8M NH ₄ NO ₃	1.3M HNO ₃ ^d 6M Li(NO ₃) ₃
Na	<0.0001	<0.0001	<0.0001					<0.0001	
Sr	<0.0008	<0.0008	<0.0008					<0.0001	<0.0001
Tl(I)	<0.0005 ^{***}	<0.0005 ^{***}	<0.0005 ^{***}			<0.005			
Tl(III)						0.083			
Th		0.001*		0.0036	1.30	0.528(0.531)	0.5	0.0034	0.32
Ti						<<0.005			
U(VI)	0.68*	1.31*	165*	2.09	73.8	1.86(1.78)	1.85	2.3	208
V(IV)		0.0006-0.001*	0.004-0.008*						
V(V)	<0.0005*	0.019*	0.040*			0.02			
Y						0.001	<0.0009		
Zn	<0.0005	<0.0005	<0.0005			0.001	<0.0001		
Zr		0.001*				~0.087	0.0011	0.011	

^a After Furman, Mundy, and Morrison, reference 326.

Conditions were arranged, in general, so that approximately 5g. of U₃O₈ were recovered from the ether extract prior to spectrographic examination. In general survey studies, 0.1 g. of each of the elements: Ag, Al, B, Ba, Bi, Ca, Cd, Co, Cr, Cu, Fe, Hg, K, Li, Mg, Mn, Na, Ni, Pb, Sr, Tl, Zn, and 0.05 g. each of In & Re was present along with uranyl nitrate and a salting agent, 80 g. per 100 ml of ammonium nitrate and 100 g. per 100 ml of calcium nitrate. (The latter is equivalent to 4.23 M if the salt is considered to be Ca(NO₃)₂ · 4H₂O, 4.58 M if it is considered to be Ca(NO₃)₂). The description of the experiment given in reference 326 gives the nitric acid concentration as 5 ml of concentrated acid per 100 ml. A table of the partition coefficients gives the acid concentration as 5 g. of HNO₃ per 100 ml. If the concentrated acid concentration is considered to be approximately 16M, both of the above values give a nitric acid concentration of the aqueous solution as 0.8M. The aqueous solutions were shaken to equilibrium with an equal volume of ether at 20°C.

The uranium partition coefficients are the maximum values determined (see, also, figure 9).

Rare earths were present in amounts such that only limiting values could be given for individual elements by spectrographic means except for Dy and O.I.

* Large amounts of elements were used: V 1-2g, As or Cr 16-20g, Na or Ca nitrate 70-100g, Fe 1g, Mo 0.3g per 100 ml. Analysis of the ether phase made by means other than spectrographic.

** Limit set by lack of sensitivity of spectrographic test; a probably much smaller.

*** Valence state not designated.

**** a is much greater for Mo present as a heteropoly acid.

^b After Bock and Bock, reference 333.

Values of a for 8M HNO₃ solutions have been calculated from P-values given by Bock and Bock. Those values in parentheses are a-values given directly by Bock and Bock or interpolated from their data.

Concentration of element concerned in initial solution: 0.1M with the exception of Al 0.94M, Sb, Bi 0.2M, Ge 0.03M, H 8M HNO₃, Fe(III) 0.4M, La 0.27M, Pb 0.075M, Sc 0.2M for 8M HNO₃, Zn 0.4M.

Elements were present as the nitrate with the following exceptions: Ge as GeO₂, P as (NH₄)₂HPO₄, As as Na₂HAsO₄, V as NH₄VO₃, Cr as K₂Cr₂O₇, Mo as NH₄-molybdate.

The Sb(V) solution hydrolyzed strongly. The largest part of the antimony was precipitated as hydroxide and only traces were detected in the ether phase.

Cr(VI) extracted well but an exact value could not be determined because of reduction of the chromium.

Mn(VII) was reduced with the separation of MnO₂ from the ether so that practically no Mn was extracted.

Temperature: 20° for the saturated NH₄NO₃ and LiNO₃ solutions. Room temperature for the 8M HNO₃ solutions.

^{bb} Asprey, Stephanon and Penneman, reference 343.

Am(VI) was prepared by peroxydisulfate oxidation.

^c After Kooi, reference 334.

The values given in this column are values taken from curves given in this reference (see, also, figures 8 and 13).

Concentrations of the various elements in the initial aqueous solution: Th, 22 mg per ml of Th(NO₃)₄ · 4H₂O; U, 3.3, 10, 30 and 90 mg per ml of UO₂(NO₃)₂ · 6H₂O; Np and Pu, tracer quantities. No detectable difference was observed in the partition data for the first three quantities of U listed; a lower α (1.7) was observed for the latter concentration.

Shaking times for the Np and Pu experiments were kept short (15 sec to 3 min) to minimize disproportionation to other oxidation states. The shaking time for U solutions was never less than 5 minutes. An increase to two hours did not make a difference in the results.

^d After Vásovenko, reference 344.

The initial aqueous layer contained, in addition to the acid and salting agent, about 0.1M of the nitrate examined. Equal volumes of the aqueous solution and diethyl ether were shaken for 5 minutes. After 12 hours equilibration a sample of the organic solution was taken for analysis.

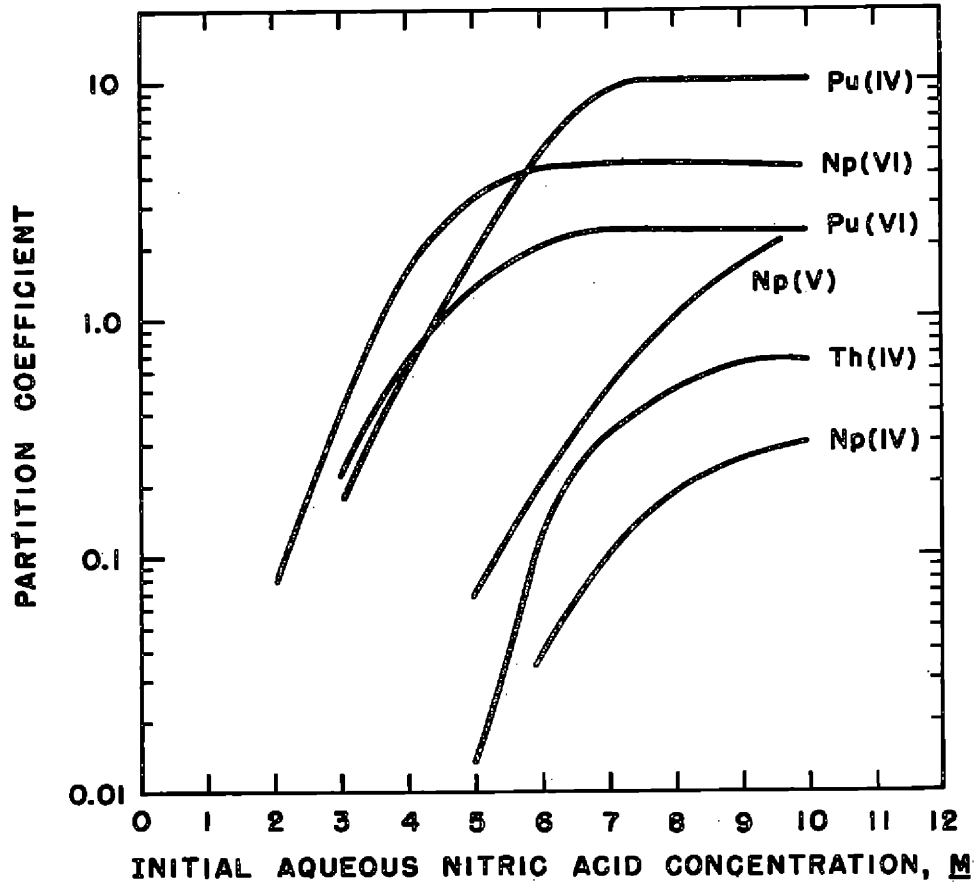


Figure 13.

The extraction of actinide nitrates by diethyl ether.

After J. Kooi, reference 334.

Conditions:

Tracer amounts of Np^{239} and Pu^{239} in aqueous solutions were equilibrated with an equal volume of diethyl ether at 25°C and room temperature, respectively. 330 mg of $\text{Th}(\text{NO}_3)_4 \cdot 4\text{H}_2\text{O}$ per 15 ml of initial aqueous solution was equilibrated with an equal volume of ether.

The oxidation states of neptunium have been the subject of some question.³⁴²

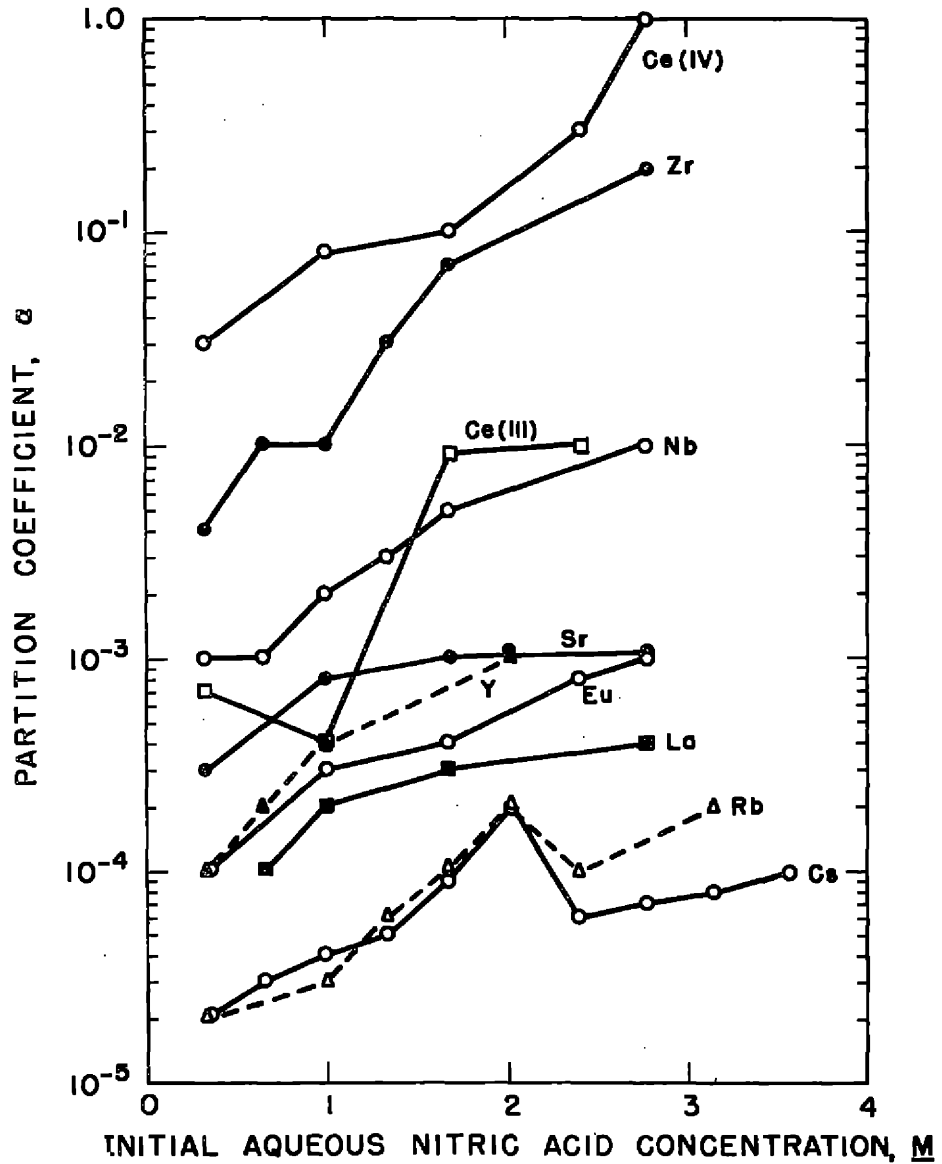


Figure 14. Partition coefficients of fission products between diethyl ether and aqueous solution containing $\text{Ca}(\text{NO}_3)_2$ and different initial acidities. After V. Vdovenko, reference 344. Conditions: Aqueous solution--required amounts of nitric acid and radioactive material added to 3.5M solution of $\text{Ca}(\text{NO}_3)_2$; V_0/V_a , 1.

Aqueous thiocyanate systems. Uranium may be extracted from aqueous thiocyanate solutions by diethyl ether.^{340,341}

Table XII lists the partition coefficients of several elements from aqueous solutions of various thiocyanate concentrations.³⁴¹ A number of substances not listed in the table give negligible distributions or distributions of only a few percent under the conditions tested: NH₄, Sb(III), As(III), As(V), Bi, Cd, Cu(I), Cr(III), Ge(IV), Li, Hg(II),

Table XII. - Partition Coefficients of Various Elements between Methyl Ether and Aqueous Thiocyanate Solutions.^a
Composition of the initial aqueous solution

	HCl M	t °C	NH ₄ SCN concentration			
			1M	3M	5M	7M
0.1M Al ^b	0.5	20		0.011	0.099	0.275 ^c
0.2M BeCl ₂	0.5	RT	0.039	0.987	5.29	11.9
0.1M CoCl ₂	0.5	RT	0.037	1.39	2.98	3.04
0.1M GaCl ₃	0.5	RT	1.89	9.56		152
HCl	0.5	20	2.67	12.7	31.1	59.4
0.1M InCl ₃	0.5	RT	1.06	3.05	2.15	0.908
0.1M FeCl ₃	0.5	RT	8.00	5.13	3.08	1.14
0.1M MoOCl ₃	~0.5	RT	140	34.9		36.4 ^d
0.1M ScCl ₃	0.5	RT	0.145	3.95		8.06
0.1M(NH ₄) ₂ SnCl ₆	0.5	RT	144	950	> 1000	> 1000
0.1M TiCl ₃	0.5	RT	1.43	5.25	3.94	3.22
0.1M TiCl ₄	0.5	RT				~0.15
0.1M UO ₂ Cl ₂	0.5	RT	0.821	0.417	0.160	0.072
0.1M VOCl ₂	0.5	RT	0.176	0.095		0.022
0.1M ZnCl ₂	0.5	RT	23.7	37.8	18.3	12.9

^a After R. Bock, reference 341.

^b as AlCl₃?

^c 6.2M

^d 6.6M

Equal phase volumes equilibrated at room temperature.

Ni, Pd(II). Figure 15 represents the change in α_u with thiocyanate concentration for aqueous solutions of different acidities.³⁴¹ Analysis of the ammonium, uranium, thiocyanate concentration of the ether phase indicates that uranium is extracted as $UO_2(SCN)_2$.³⁴¹

Aqueous fluoride systems. Uranium (VI) is poorly extracted by diethyl ether from aqueous fluoride solutions.³⁴² Table XIII lists the partition coefficients of a number of elements from aqueous solutions of various hydrofluoric acid concentrations.

Dibutyl ether

Aqueous nitrate systems. The applicability of dibutyl ether as an extractant for uranium has been investigated extensively by workers in the Soviet Union.^{344,346-349} Dibutyl ether offers several advantages over diethyl ether. It is less soluble in water, less volatile, and has a higher flash point. The distribution coefficient of uranyl nitrate is, however, less for dibutyl ether than for diethyl ether. The partition of uranium between water and dibutyl ether is represented in figure 1.^{185,321} The distribution of uranyl nitrate and nitric acid is plotted as a function of aqueous nitric acid concentration in figure 16.³⁴⁶ Karpacheva, Khorkhovina, and Agashkina³⁴⁸ have studied the effect of various salting-out agents on the distribution of uranyl nitrate. The salting-out action was found to increase with increasing valence of the cation. The partition coefficient of uranium from an aqueous solution initially 0.5M $UO_2(NO_3)_2$, 4.5M $Ca(NO_3)_2$, and 0.5M HNO_3 into an organic phase 85% (by volume) dibutyl ether and 15% carbon tetrachloride is 0.70; $\alpha_{Pu(VI)}$ is 0.42.³⁴⁷ Zirconium, niobium, and ruthenium are the main fission product elements extracted.^{344,348} Heyn and Banerjee³⁵⁰ have studied the extraction of bismuth nitrate by dibutyl ether and several other solvents.

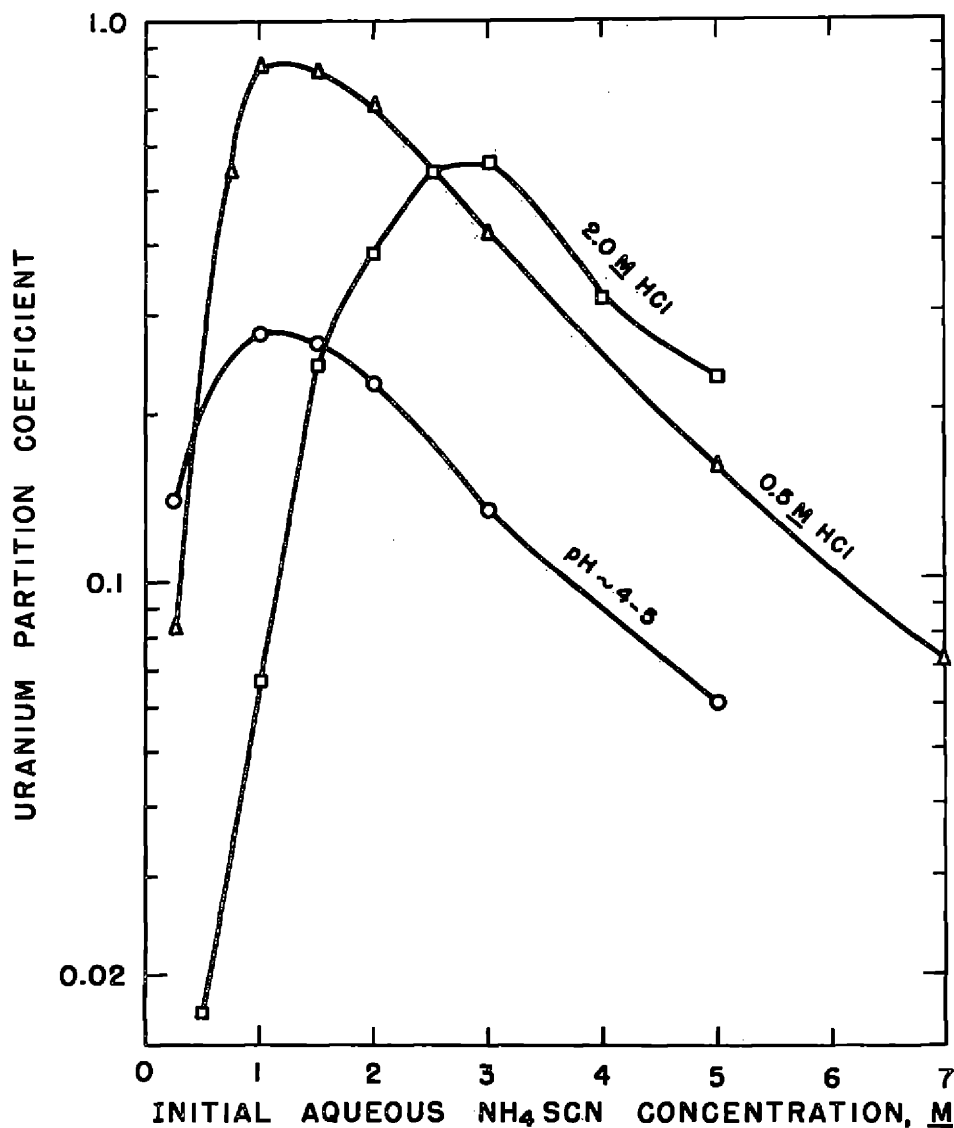


Figure 15.

The extraction of uranyl thiocyanate by diethyl ether at various initial NH_4SCN and HCl concentrations.

After R. Bock, reference 341.

Conditions:

Aqueous phase--initially 0.1M, UO_2Cl_2 and NH_4CNS and HCl concentration indicated.
Equal phase volumes equilibrated at room temperature.

Table XIII.- Partition Coefficients of Various Elements between Aqueous HF Solutions and Diethyl Ether^a

Element ^b	HF concentration of the starting aqueous solution				
	1.0M ^c HF	5.0M ^c HF	10.0M ^c HF	15.0M ^c HF	20.0M ^c HF
Sb(III)	<0.0005	0.003	0.019	0.040	0.067
As(III)	0.111	0.227	0.432	0.530	0.605
As(V)	<0.001	0.017	0.048	0.121	0.157
Be	<0.0005	<0.0005	0.005	0.019	0.042
Cd	<0.0005	0.002	0.006	0.009	0.014
Co	<0.002	<0.002	<0.002	0.005	0.017
Cu(II)	<0.001	<0.001	<0.001	0.010	0.013
Ge(IV)	<0.002	<0.002	0.005	0.028	0.072
Mn(II)	<0.0005	<0.0005	0.002	0.005	0.013
Hg(II)	<0.0005	<0.0005	<0.0005	0.009	0.028
Mo(VI)	0.007	0.018	0.031	0.062	0.103
Ni	<0.0005	<0.0005	<0.0005	0.005	0.007
Nb(V)	0.006	0.044	0.480	1.08	1.92
P(V)	<0.001	0.011	0.032	0.110	0.173
Re(VII)	0.0005	0.121	1.58	1.78	1.62
Se(IV)	0.0006	0.022	0.080	0.131	0.148
Ta	0.012	0.774	3.80	3.82	3.84
Te(IV)	0.0001	0.020	0.071	0.237	0.298
Sn(II)		0.020	0.029		0.052
Sn(IV)	0.006 ^c	0.006 ^d	0.053 ^e		0.055
U(VI)	<0.002	<0.002	<0.002	0.005	0.011
V(III)	<0.0005	0.003	0.03	0.10	0.13
V(V)	<0.001	0.004	0.017	0.056	0.093
Zn	<0.001	<0.001	<0.001	0.002	0.009
Zr	0.004	0.005	0.005	0.012	0.030

^a After Bock and Herrmann, reference 342.

Equal phase volumes equilibrated at $20.0 \pm 0.5^\circ\text{C}$.

^b The concentration in the initial aqueous solution of the ionic species of the element listed in the table was 0.1M in each case with the exception of Re(VII) which was 0.05M.

Fluoride stock solutions were prepared in the following manner: carbonates (Cd, Co, Cu, Mn(II), Ni, Zn), oxides (Sb(III), Ge(IV), Hg(II), Nb(V), Se(IV), Ta(V), Te(IV), V(III)) or hydroxides or hydrated oxides (Sn(IV), Zr) were dissolved in an excess of HF; As_2O_3 was dissolved in a known volume of 0.1N NaOH and the calculated amount of HF added to the solution; BeF_2 and SnF_2 were dissolved; K_2HAsO_4 , Na_2HPO_4 , KReO_4 , ammonium vanadate, ammonium molybdate, and sodium uranate were dissolved in HF.

^c 1.2 M HF

^d 5.4 M HF

^e 10.4 M HF

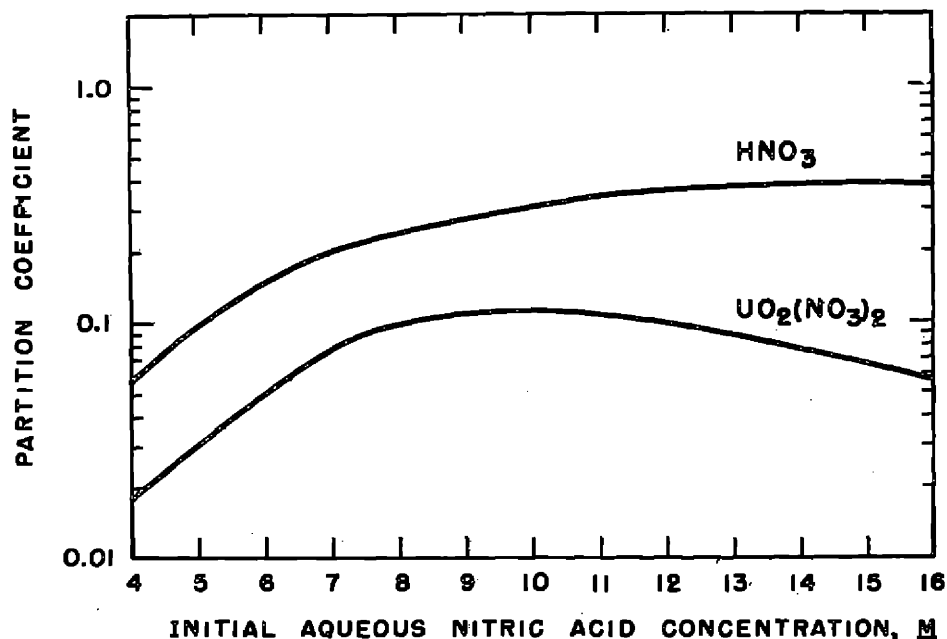


Figure 16.

Partition of uranyl nitrate and nitric acid between dibutyl ether and aqueous solution.

After V. Vdovenko, A. Lipovskii, M. Kuzina, reference 346.

Conditions:

Equal phase volumes equilibrated at room temperature for both $\text{UO}_2(\text{NO}_3)_2$ and HNO_3 extractions. For nitric acid, points corresponding to an acid content in the aqueous solution of greater than 13.4M were obtained by the extraction of previously acidified dibutyl ether with concentrated nitric acid. For uranyl nitrate, points greater than 12.0M HNO_3 were similarly obtained. The uranium concentration was 78 mg/ml.

Dibutyl "Cellosolve" (Dibutoxymonoethyleneglycol)

Aqueous nitrate systems. A number of cellosolve derivatives have been investigated for the extraction of uranium (Table VIII). Diethyl cellosolve is an excellent extractant.³³² Unfortunately its solubility in water is large (21% by weight at 20°C). Dibutyl cellosolve is less soluble in water (0.2% by weight at 20°C). However, it does not extract uranium as well as diethyl ether, either from water

solution (figure 2^{185,321}) or from aqueous aluminum nitrate solution (figure 6³³²). The partition coefficient of uranium into dibutyl cellosolve from nearly saturated solutions of ammonium, calcium, or ferric nitrate is 1, 50, and 20, respectively.³³²

Dibutyl "Carbitol" (Dibutoxydiethyleneglycol)

Aqueous nitrate systems. Dibutyl carbitol (Butex) is used in the recovery of irradiated fuel material.³⁵¹ As a solvent, it has been subject to considerable study.^{332,352} The partition of uranium between water solution and solvent is given in figure 2.^{176,185,321} The partition of uranium between nitric acid solution and dibutyl carbitol is illustrated in figure 17.³⁵³ For aqueous solutions in this range of acid concentration, the partition coefficient is observed to increase with increased uranyl nitrate concentration. The partition coefficient of nitric acid is plotted as a function of acid concentration in figure 18.^{334,353,354} In figures 19 and 20, the partition coefficients of uranium and several other heavy elements are plotted against nitric acid concentration. The initial acid concentration of the aqueous phase is plotted in figure 19.³³⁴ The equilibrium acid concentration of the aqueous phase is given in figure 20.³⁴⁵ Best, et al³⁴⁵ have observed that the steepness of the extraction curves (figure 20) is compatible with the formation of the species $\text{HMO}_2(\text{NO}_3)_3$ and $\text{H}_2\text{M}(\text{NO}_3)_6$ in the organic phase rather than just $\text{MO}_2(\text{NO}_3)_2$ and $\text{M}(\text{NO}_3)_4$. The curves given in figures 19 and 20 are in general agreement considering the difference in acid concentration plotted. There is a large discrepancy between Np(IV) data. The ability to maintain neptunium in the pentavalent state during extraction may be subject to question. The partition of some fission product elements is given in figure 21 for various

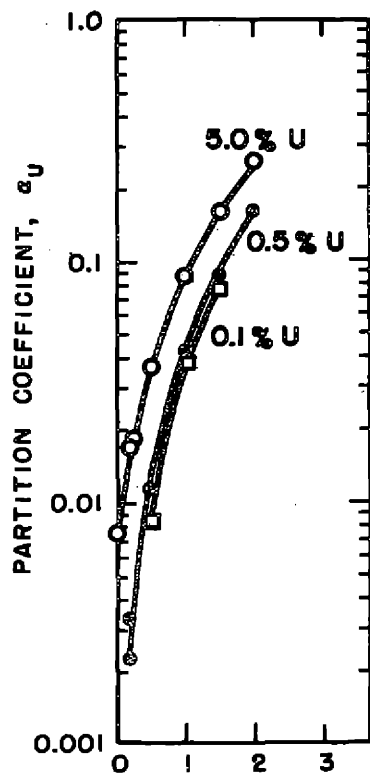


Figure 17. Distribution of uranyl nitrate between dibutyl carbitol and nitric acid solution. Adapted from C. A. Kraus, reference 353. Conditions: Approximately equal volumes of organic and aqueous phase, initially at the uranium and nitric acid concentrations indicated, equilibrated at about 27°C.

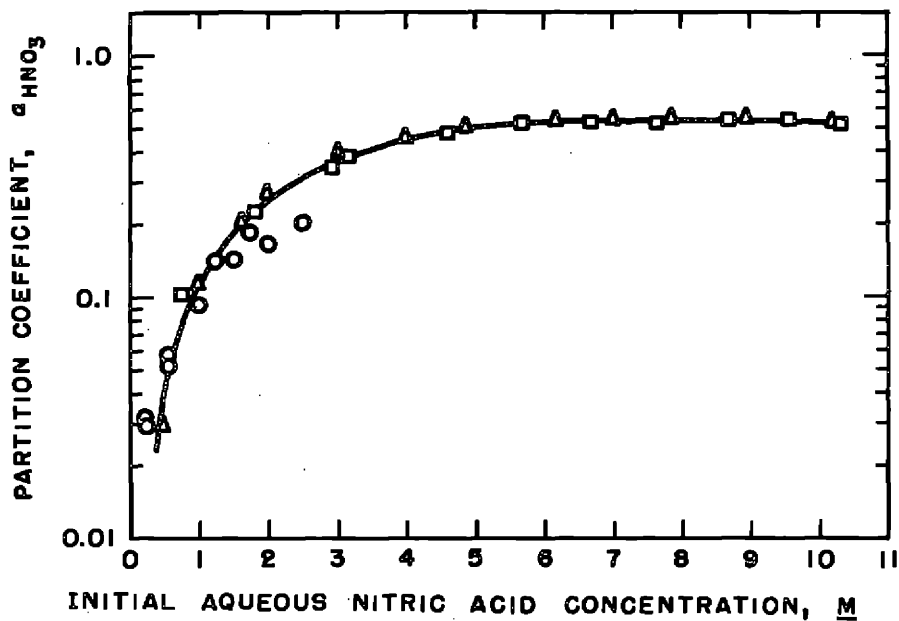


Figure 18. Distribution of nitric acid between dibutyl carbitol and aqueous solution. O, After C. A. Kraus, reference 353. □, After J. Kooi, reference 334. Δ, After D. G. Tuck, reference 354. Conditions: Equal^{334,354} or approximately equal³⁵³ volume portions of solvent and aqueous solution equilibrated at ~27°C,³⁵³ 25°C,³³⁴ and ~21°C.³⁵⁴

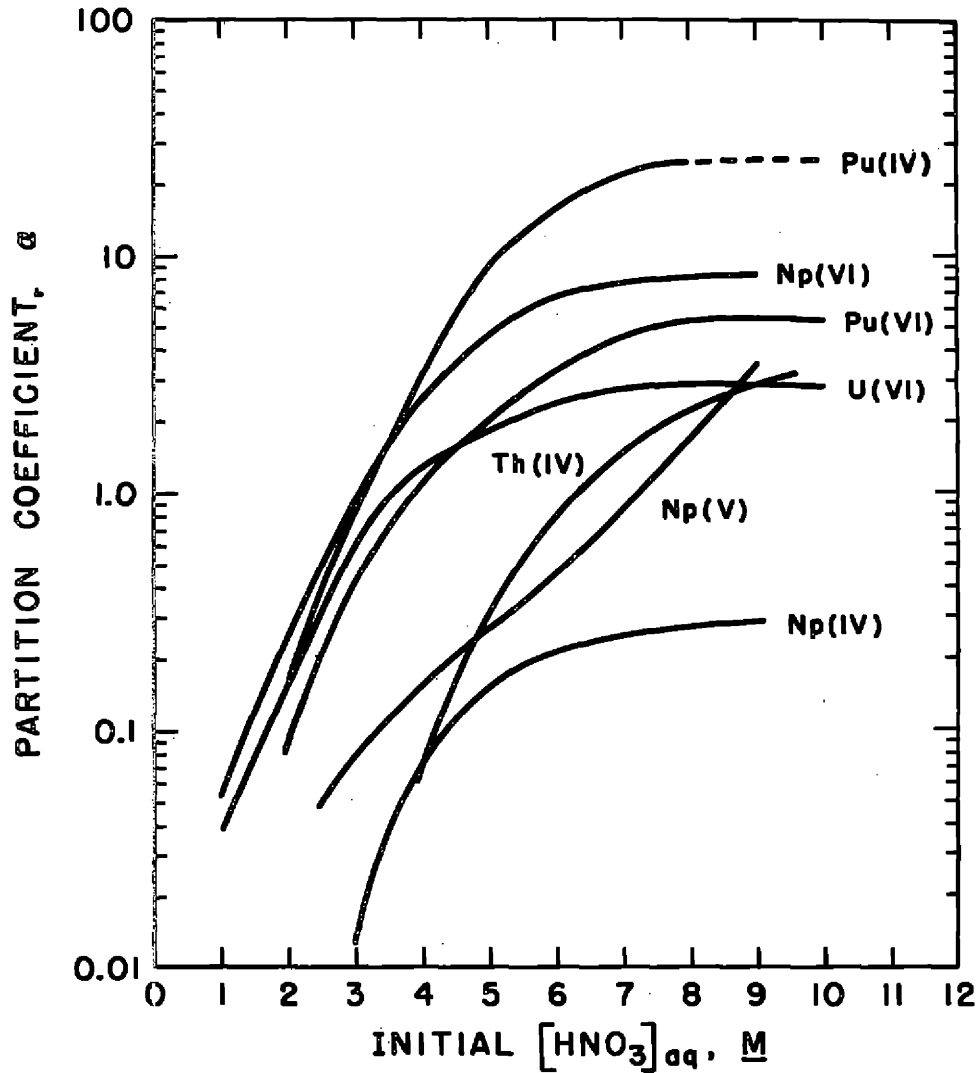


Figure 19.

The extraction of actinide nitrates by dibutyl carbitol.

After J. Kooi, reference 334.

Conditions:

Tracer amounts of Np^{239} or Pu^{239} , 330 mg of $Th(NO_3)_4 \cdot 4H_2O$ per 15 ml, or 300 mg of $UO_2(NO_3)_2$ (hexahydrate) per 15 ml in aqueous nitric acid solution equilibrated with an equal volume of dibutyl carbitol at $25^\circ C$ or room temperature.

The oxidation states of neptunium have been the subject of some question.³⁴⁵

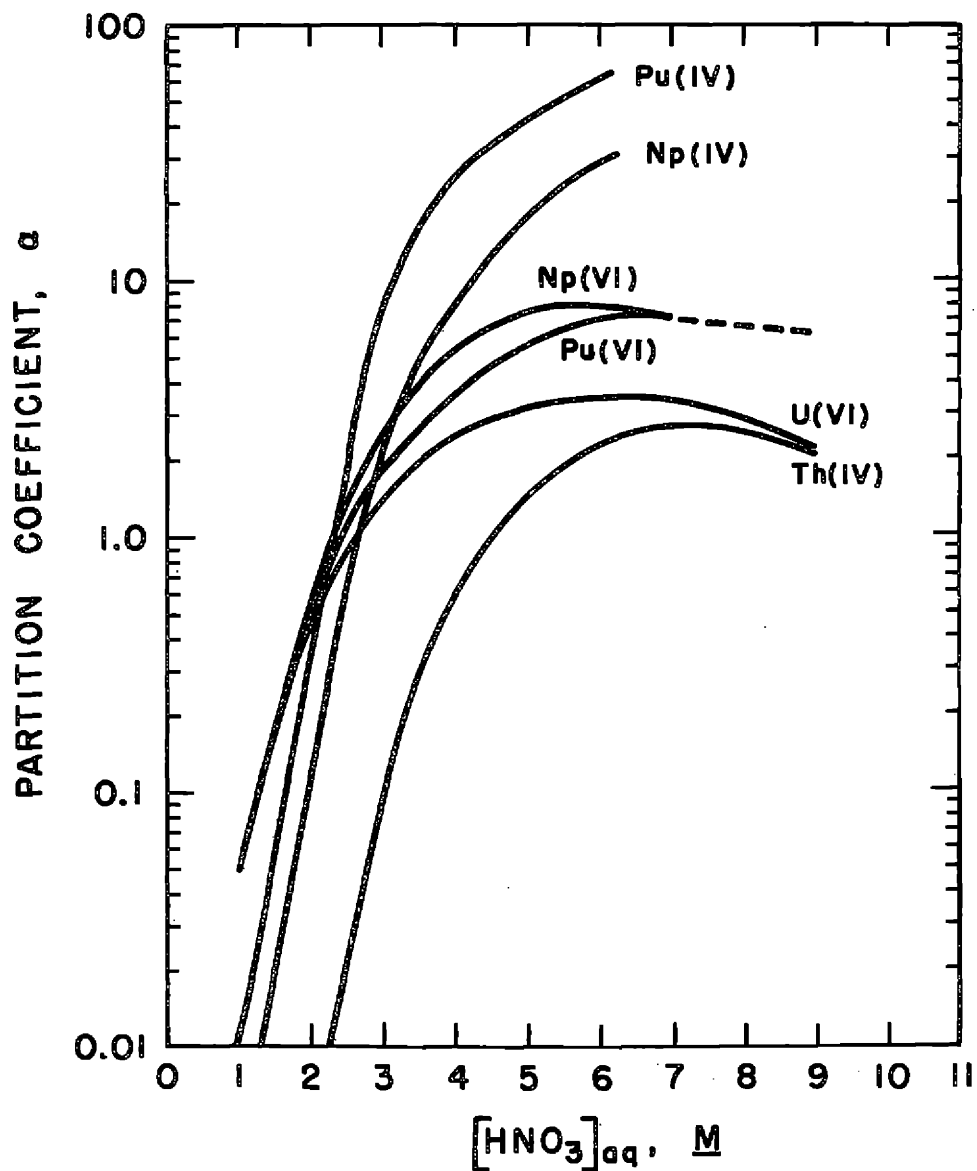


Figure 20.

The distribution of actinide elements between dibutyl carbitol and aqueous solution as a function of equilibrium aqueous nitric acid concentration.

After G. Best, E. Hesford, and H. McKay, reference 345.

Conditions:

Tracer concentrations ($\sim 10^{-3}\text{M}$) of actinide. Temperature, 25°C.

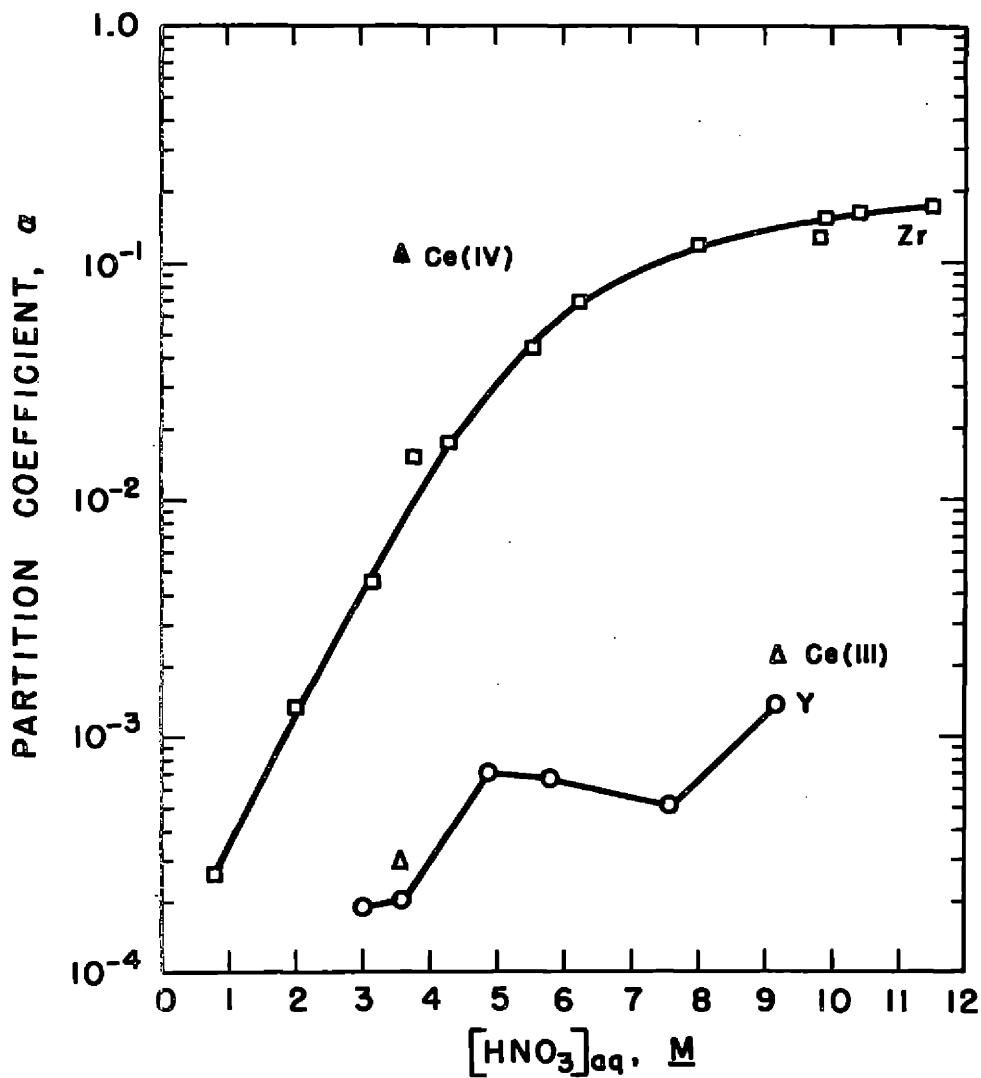


Figure 21.

The partition of tracer amounts of yttrium, cerium, and zirconium between dibutyl carbitol and aqueous solution as a function of aqueous nitric acid concentration.

After H. McKay, K. Alcock, and D. Scargill, reference 355.

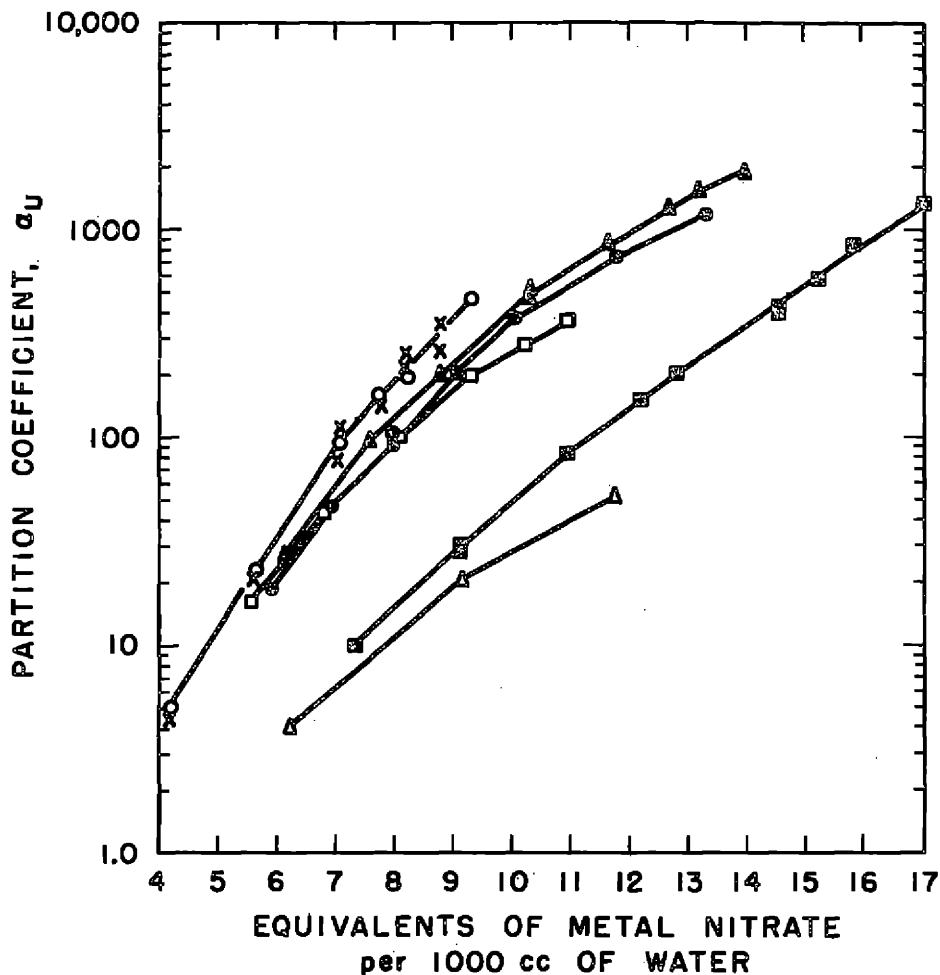


Figure 22. The effect of salting-out agents on the extraction of uranium by dibutyl carbitol. ● $\text{Cu(NO}_3)_2$, ■ $\text{Ca(NO}_3)_2$, ▲ $\text{Zn(NO}_3)_2$, ○ $\text{Al(NO}_3)_3$, □ $\text{Fe(NO}_3)_3$, △ $\text{La(NO}_3)_3$. After E. Evers and C. Kraus, reference 332. Conditions: Uranium concentration, 2-6 g/100 cc of phase. Temperature, 27°C; $V_o/V_a = 1$. X $\text{Al(NO}_3)_3$. After D. Lee, R. Woodward, G. Clewett, reference 358. Conditions: Trace amounts of uranium. Temperature, 27°C; V_o/V_a , varied.

aqueous nitric acid concentrations.³⁵⁵ The distribution of iron into dibutyl carbitol is increased by an increase in acidity.³⁵² Chloride ion promotes the extraction of iron. Boron is extracted by butex, especially in the presence of copper nitrate as salting-out agent.³⁵² Vanadium and molybdenum are extracted to several per cent.³⁵² The extraction of cadmium, chromium, nickel and titanium is small.³⁵²

The effect of salting-out agents on the distribution of uranium into dibutyl carbitol has been studied.^{332,352,356-358} Some of the results are presented in figure 22.

Aqueous chloride systems. Uranium (IV) and (VI) and thorium are poorly extracted by dibutyl carbitol from aqueous solutions 2-6M in hydrochloric acid.³⁵⁹ The extraction of protactinium is increased as the acid concentration is increased. From 6M HCl, α_{pa} is 10. The extraction of hydrochloric acid is negligible from aqueous solutions less than 6M in hydrochloric acid. A third phase is formed upon equilibration with 7.5M HCl. The third phase contains a large amount of the acid. One phase results upon equilibration with 8.5M HCl.

Pentaether (Dibutoxytetraethyleneglycol)

Aqueous nitrate systems. References 332 and 360 summarize much of the data pertinent to the extraction of uranium by pentaether. The distribution of uranyl nitrate between solvent and water is given in figure 2A.^{176,185} The partition coefficient of uranyl nitrate from various nitrate media is plotted in figure 23.³⁶⁰ The distribution of nitric acid as a function of aqueous acid concentration is also shown in figure 23.³⁶¹ The effect of salting-out agents on the partition of uranium is illustrated in figure 24.³³² Table XIV lists the partition coefficients of a number of elements other than uranium between pentaether and various aqueous media.³⁶² Uranium is extracted by pentaether from aqueous solutions containing ammonium nitrate and/or nitric acid in the presence of sulfate, phosphate, or silicate ions.³⁶² Phosphate ion, in large quantity, and soluble silicate ions are extracted by the solvent.³⁶² Fluoride ion, in significant quantity, interferes with uranyl nitrate extraction. This effect may be overcome by complexing the fluoride ion with calcium or aluminum nitrate.³⁶²

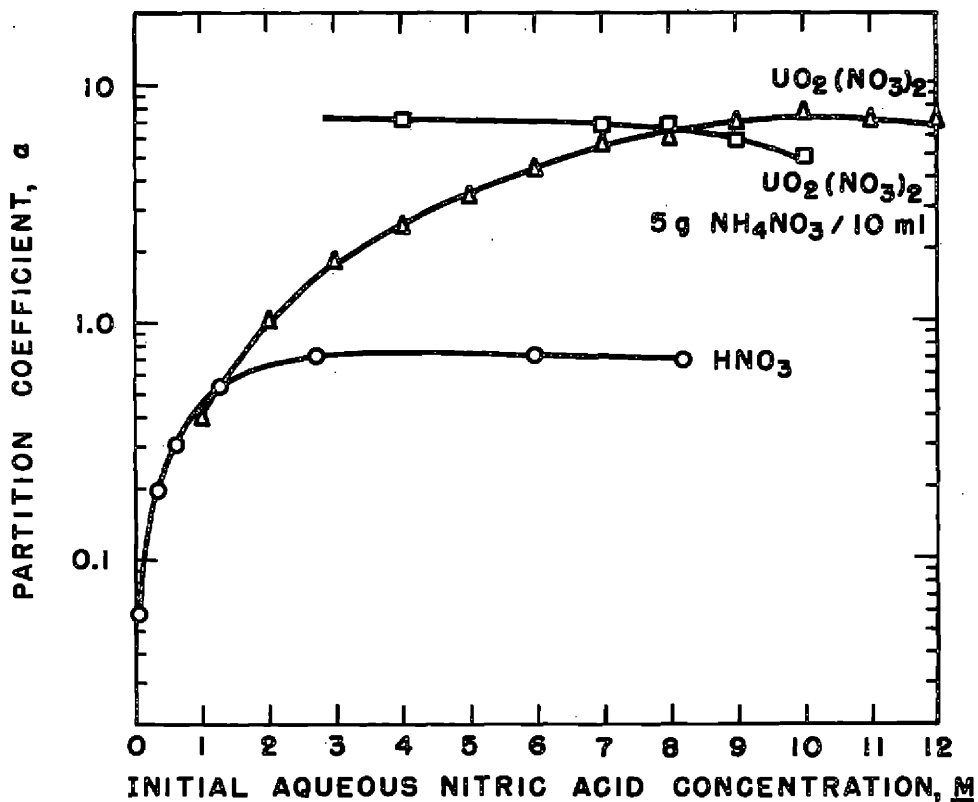


Figure 23. The partition coefficients of uranyl nitrate and nitric acid between pentaether and aqueous solution. Δ $\text{UO}_2(\text{NO}_3)_2$: 1.0 g. U_3O_8 dissolved in HNO_3 , diluted to 50 ml with acid of desired strength, and shaken 1 minute with an equal volume of pentaether. \square $\text{UO}_2(\text{NO}_3)_2 + \text{NH}_4\text{NO}_3$: 1.0 g. U_3O_8 dissolved in 10 ml of HNO_3 of desired strength after addition of 5 g. NH_4NO_3 ; shaken with an equal volume of pentaether for 1 minute at room temperature. Adapted from D. Musser, D. Krause, and R. Smellie, Jr., reference 360. \circ HNO_3 : equal volumes of nitric acid solution and pentaether equilibrated for 1 hour at 25°C. After C. Stover, Jr. and H. Crandall, reference 361.

Cyclic ethers.

A number of cyclic ethers have been investigated as extractants for uranium.^{363,364} Those solvents that contain the furane nucleus have been found to give good extractions of uranyl nitrate from aqueous solutions. Solvents of the hydrocarbon substituted tetrahydrofurane type have been found to be especially good.³⁶³ The extraction of uranium and thorium by four cyclic ethers is illustrated in figure 25 as a function of acid concentration in the aqueous phase.³⁶⁴

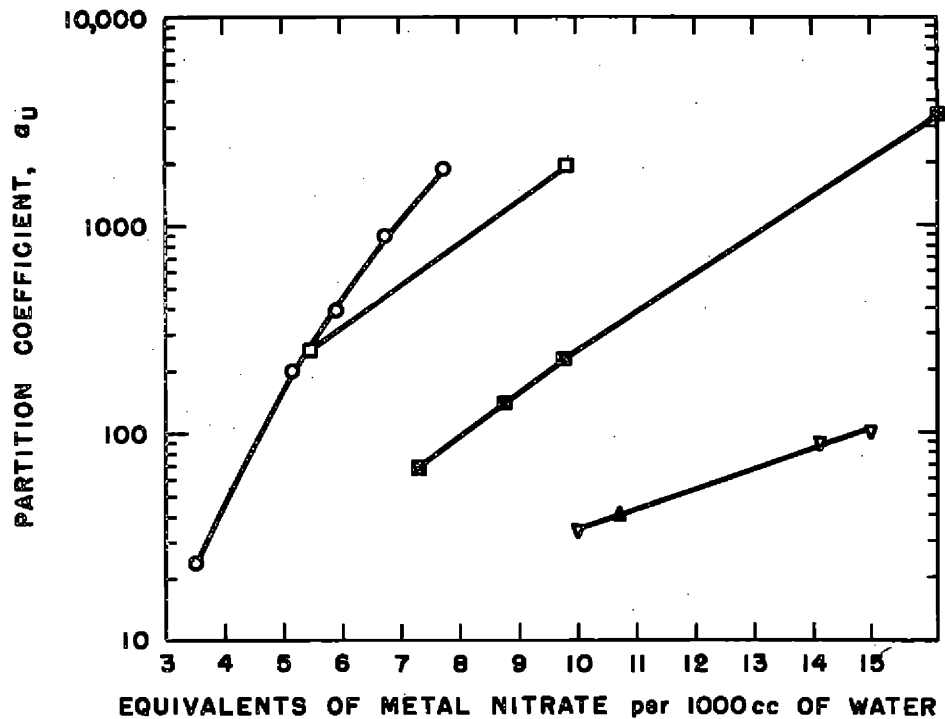


Figure 24. Effect of salting-out agents on the extraction of uranium by pentaether. \blacktriangle NaNO_3 , ∇ NH_4NO_3 , \blacksquare $\text{Ca}(\text{NO}_3)_2$, \circ $\text{Al}(\text{NO}_3)_3$, \square $\text{Fe}(\text{NO}_3)_3$. After E. Evers and C. Kraus, reference 332. Conditions: Uranium concentration, 2-6 g/100 cc of phase. Temperature, 27°C or room temperature.

From figure 25-A, it can be seen that uranyl nitrate is extracted more efficiently by the various solvents than is uranyl perchlorate. Better separation of uranium and thorium is also achieved from nitrate solution rather than perchlorate.

ESTERS

Information is less complete or less readily available for the extraction of uranium by esters than by ethers. The distribution of uranyl nitrate between iso-amyl acetate and water is represented in figure 3-A.¹⁸⁵ Karpacheva, et al.³⁴⁸ have found the extraction capacity of butyl acetate to be intermediate between diethyl ether and dibutyl ether. Hyde and Wolf,³¹⁶ in addition to their

Table XIV. Partition Coefficients of Elements between Penta-ether and

Element	Concentration in aqueous phase before extraction (mg/25 ml)	Various Aqueous Media ^a			
		Aqueous solution ^b			
		Nitrate	Nitrate + Chloride	Sulfate	Sulfate + Chloride
Al	500	0.003	0.003	0.000	0.003
Ba	362	0.02			
"	7	0.35			
Cd	501	0.024		0.003	
"	500		0.026		
"	10	0.053	0.01		
"	10.5			0.004	
Ca	521	0.011		-- <u>c</u>	
"	500		0.0001		-- <u>c</u>
"	10.4	0.020			0.00
Cl	496	0.026			
"	494			0.02	
"	10.3	0.025			
"	9.9			0.03	
Cr(III)	500	0.003	0.004	0.0001	
"	10	0.0023	0.013	0.0004	
Co	555	0.007			
"	500		0.002	0.00013	
"	11	0.012			
"	10		0.009	0.0065	
Cu(II)	500	0.026	0.024	0.000	
"	10	0.017	0.018	0.000	
Fe(III)	515	0.031		0.003	
"	500		1.2		
"	10.3	0.035		0.002	
"	10.0		0.046		
Pb	500	0.017		-- <u>a</u>	
"	10	0.007		0.005	
Mn(II)	500	0.0011		0.00006	
"	10	0.0014		0.00075	
Hg(II)	497		0.21		0.015
"	127	0.19		0.036	
"	10	0.41	0.176	0.23	0.03
Mo(VI)	500	0.028	0.10	0.001	0.001
"	10	0.10		0.015	

Table XIV. - Continued

Element	Concentration in aqueous phase before extraction (mg/25 ml)	Aqueous solution ^b			
		Nitrate	Nitrate + Chloride	Sulfate	Sulfate + Chloride
Ni	516		0.0001		
"	500	0.0018			
"	10.3		0.0008		
"	10	0.0032		0.00064	
P(PO_4^{3-})	500	0.00005		0.00001	
"	10	0.00024		0.00	
Ag	500	0.09		0.005	
"	10	0.32		0.005	
SO_4^{2-} (as $(\text{NH}_4)_2\text{SO}_4$)	500	0.00			
"	10	0.00			
Th	500	9.12	11.5	0.0001	0.0001
"	10	87.5		0.005	
Sn(IV)	500	0.0024	0.37	0.00015	0.000
"	10	0.019		0.0006	
Tl(IV)	11	0.003	0.00	0.035	
W	10	0.081		0.0025	
V(V)	140		0.22		
"	107	0.11			
"	14		0.12		
"	8.5	0.07		0.009	
Zn	497		0.15		
"	447	0.018			
"	10		0.14	-- ^c	
"	8.9	0.022		0.000	
Zr	10	0.040		0.013	

^a Adapted from A. G. Jones, C-4.360.3(1945).

Equal volume portions of aqueous solution and pentaether.

^b Nitrate : 3/4 saturated ammonium nitrate solution.

Sulfate : saturated ammonium sulfate solution.

Chloride: chloride added as ammonium chloride equivalent to the metal present.

^c Precipitates of insoluble sulfate obtained in ammonium sulfate layer.

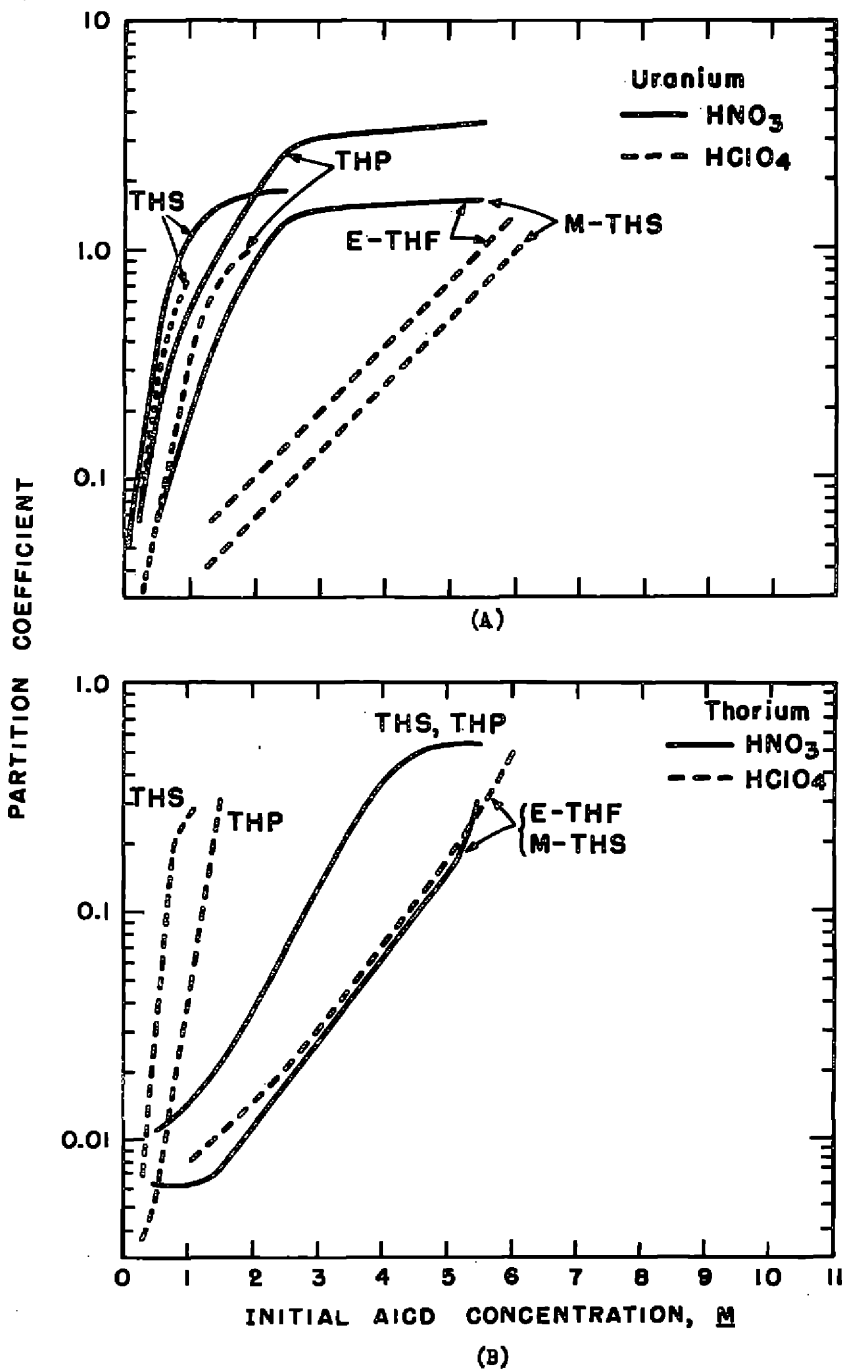


Figure 25. The effect of initial nitric and perchloric acid concentration on the extraction of uranyl salts (Fig. 25-A) and thorium salts (Fig. 25-B) by tetrahydrofuran (THF), tetrahydropyran (THP), 2-ethyltetrahydrofuran (ETHF), and 2,5-dimethyltetrahydrofuran (MTHF). After M. Branica and E. Bona, reference 364. Conditions: Uranium concentration, $2 \times 10^{-3}M$. Thorium concentration, tracer UX_1 . Temperature, $25 \pm 0.2^\circ C$. V_o/V_a , 1.

general survey work (Table VIII), have studied the extraction of thorium and uranium by ethyl acetate, n-propyl acetate, and iso-propyl acetate as a function of the nitrate concentration of the aqueous phase. It was the observation of the latter group³¹⁶ that the extraction of uranium tends to decrease with increasing molecular weight of the ester. Therefore, only acetates and propionates need to be considered seriously. Increased protactinium extraction was observed with increasing length of the alcohol portion of the ester.³¹⁶ It was further observed that hydrolysis of the ester tends to increase the extraction of both thorium and uranium.³¹⁶ It was not determined whether the addition of alcohol or organic acid causes the increased extraction.

Ethyl acetate

Aqueous nitrate systems. The distribution of uranyl nitrate between ethyl acetate and water has been studied by deKeyser, Cypres, and Hermann.¹⁶⁴ The partition coefficient was found to vary from 0.17 at 22% $\text{UO}_2(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ in the aqueous phase to 0.78 at 43% aqueous concentration. In laboratory practice, uranium is extracted by the solvent from aqueous nitrate media. The following conditions have been used by various groups to extract uranium:

Grimaldi and Levine³⁶⁵: 9.5 g. of $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ are added to 5 ml of solution approximately 2.4N in HNO_3 . 10 ml of ethyl acetate are added and shaken at least 30 seconds.

Rodden and Tregonning³⁶⁶: Uranium precipitated in the presence of aluminum (20 mg) with NH_4OH is dissolved in 1 ml of HNO_3 (1 to 1). 8 g. of $\text{Mg}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ is added and the volume adjusted to 10 ml with water. 5 ml of ethyl acetate are added and vigorously shaken for 2 minutes. (Used with 20-400 mg samples of U_3O_8 .)

Nietzel and DeSesa^{367,368}: Approximately 15 ml of saturated aluminum nitrate solution are added to 3 ml or less of sample containing 0.30 to 15 g. of U_3O_8 per liter. 20 ml of ethyl acetate are added and shaken for 1 minute.

Guest and Zimmerman³⁶⁹: To 5 ml of sample containing 5% concentrated HNO_3 by volume, 6.5 ml of hot aluminum nitrate solution, having a boiling point of $130^\circ C.$, are added. The resulting solution is cooled, 20 ml of ethyl acetate are added, and the mixture shaken for 45 to 60 seconds.

Steele and Taverner¹⁹⁸: Approximately 5 ml of aqueous solution are saturated with aluminum nitrate. The resulting solution is shaken with 10 ml of ethyl acetate for 1-2 minutes.

In the procedure of Rodden and Tregonning,³⁶⁶ aluminum nitrate is used instead of magnesium nitrate if extraction is to be made in the presence of phosphate. DeSesa and Nietzel^{367,368} found that 1 molar concentrations of phosphate, sulfate, or carbonate ion could be tolerated with no ill effect on uranium extraction. Small amounts of sodium phosphate have been used to suppress the extraction of thorium without affecting the extraction of uranium.¹⁹⁸ Steele and Taverner¹⁹⁸ report the extraction of appreciable amounts of thorium and zirconium and small amounts of vanadium, molybdenum, and platinum by ethyl acetate. Grimaldi and Levine,³⁶⁵ Guest and Zimmerman,³⁶⁹ and Nietzel and DeSesa^{367,368} have investigated the effect of a number of elements on the recovery and/or determination of uranium according to their respective procedures. Nietzel and DeSesa^{367,368} found vanadium, present in 100 mg amounts, was precipitated and uranium was occluded in the precipitate. Titanium was observed to partially extract. This was prevented by precipitation of titanium with p-hydroxyphenylarsonic acid before extraction.

Aqueous thiocyanate systems. Vanossi³⁷⁰ has described a procedure in which uranium is extracted by ethyl acetate from an aqueous phase containing an excess of ammonium thiocyanate. Dizdar and Obrenovic³⁷¹ have also investigated the extraction of the uranyl-thiocyanate complex by ethyl acetate.

KETONES

Methyl ethyl ketone.

Aqueous nitrate systems. The distribution of uranyl nitrate between methyl ethyl ketone and water and between methyl ethyl ketone and saturated ammonium nitrate solution is given in figure 4-B.³²¹ Palei¹⁹⁶ reports a uranium partition coefficient of approximately 25 between methyl ethyl ketone and an aqueous solution of 60% NH_4NO_3 and 1N HNO_3 . Methyl ethyl ketone is not as selective as diethyl ether.³²¹ Homogeneous solutions are formed between the ketone and an equal volume of saturated ferric or cupric nitrate at 20°C.³²¹

Miscellaneous aqueous systems. The partition coefficient of uranium between methyl ethyl ketone and an aqueous 60% NH_4NO_3 , 3% NH_4SCN solution is about 2000.¹⁹⁶ Iron is extracted.

Milner and Wood³⁷² report the separation of tantalum and niobium from uranium by extracting the fluorides of the former elements with methyl ethyl ketone.

Hexone (Methyl iso-butyl ketone).

Aqueous nitrate systems. The partition of uranyl nitrate between hexone and water is represented in figure 4.^{185,321} The partition coefficients of uranium, nitric acid, and several other actinide elements are plotted as a function of aqueous nitric acid concentration in figure 26.³³⁴ The effect of several salting-out agents on the partition coeffi-

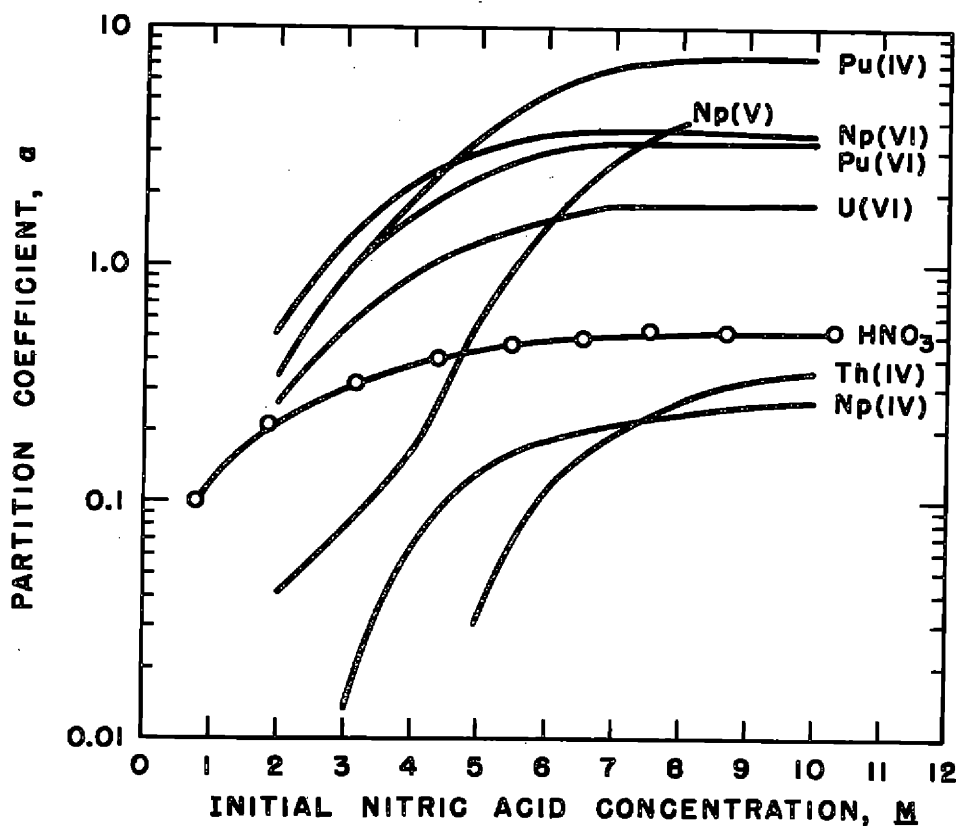


Figure 26.

The extraction of nitric acid and actinide nitrates, Th, U, Np, and Pu, by methyl isobutyl ketone (hexone).

After J. Kooi, reference 334.

Conditions:

Tracer amounts of Np²³⁹ or Pu²³⁹, 330 mg of Th(NO₃)₄ · 4H₂O per 15 ml, or 300 mg of UO₂(NO₃)₂ (hexahydrate) per 15 ml in nitric acid solution or nitric acid alone equilibrated with an equal volume of hexone at 25°C or room temperature.

The oxidation states of neptunium have been the subject of some question.³⁴⁵

cient of uranium is given in figure 27.³¹⁹ Vdovenko and co-workers³²⁵ have observed an increase in the partition coefficients of cesium, calcium, strontium, and lanthanum when the uranyl nitrate concentration in the initial aqueous solution is increased. This has been related to the extraction of the elements as metal uranyl trinitrate salts. The partition coefficient of uranium from a highly salted aqueous solution is decreased by an increase in uranium concentration. Kraus³⁷³ observed α_u to decrease from 153 to 78.3 as the initial uranium concentration was increased from 5 to 100 grams in an aqueous solution containing 580 grams of aluminum nitrate. Jenkins and McKay³⁷⁴ found α_u to decrease from 1.58 to 1.28 as the initial uranium concentration was increased from 144 to 348 grams per liter in an aqueous solution 8M in NH_4NO_3 and 0.3N in HNO_3 . In the latter case, commercial hexone adjusted to 0.15N HNO_3 was used as the extractant. Figure 28 represents the extraction of uranium by hexone from aqueous solutions containing various amounts of nitric acid and calcium or sodium nitrate.³⁷⁵ The distribution of U(VI), Pu(VI), Pu(IV), Th, La, Ca, Na, and HNO_3 by hexone from aqueous solutions containing nitric acid and calcium nitrate has been investigated by Rydberg and Bernström.³⁷⁶ Hyde and co-workers have studied the extraction of uranium³¹⁶ and thorium^{316,377} by hexone as a function of the total nitrate concentration of the aqueous phase. Distribution curves (α or P versus nitric acid or total nitrate concentration of the aqueous phase) are presented for the various elements in the different papers. The effect of aluminum nitrate concentration on the extraction of fission product gamma-activity in general and zirconium-niobium, cerium, and ruthenium in particular is shown in figure 29.³⁷⁸ Increased extraction is effected by an increase in salting-out agent. An increase in nitric acid concentra-

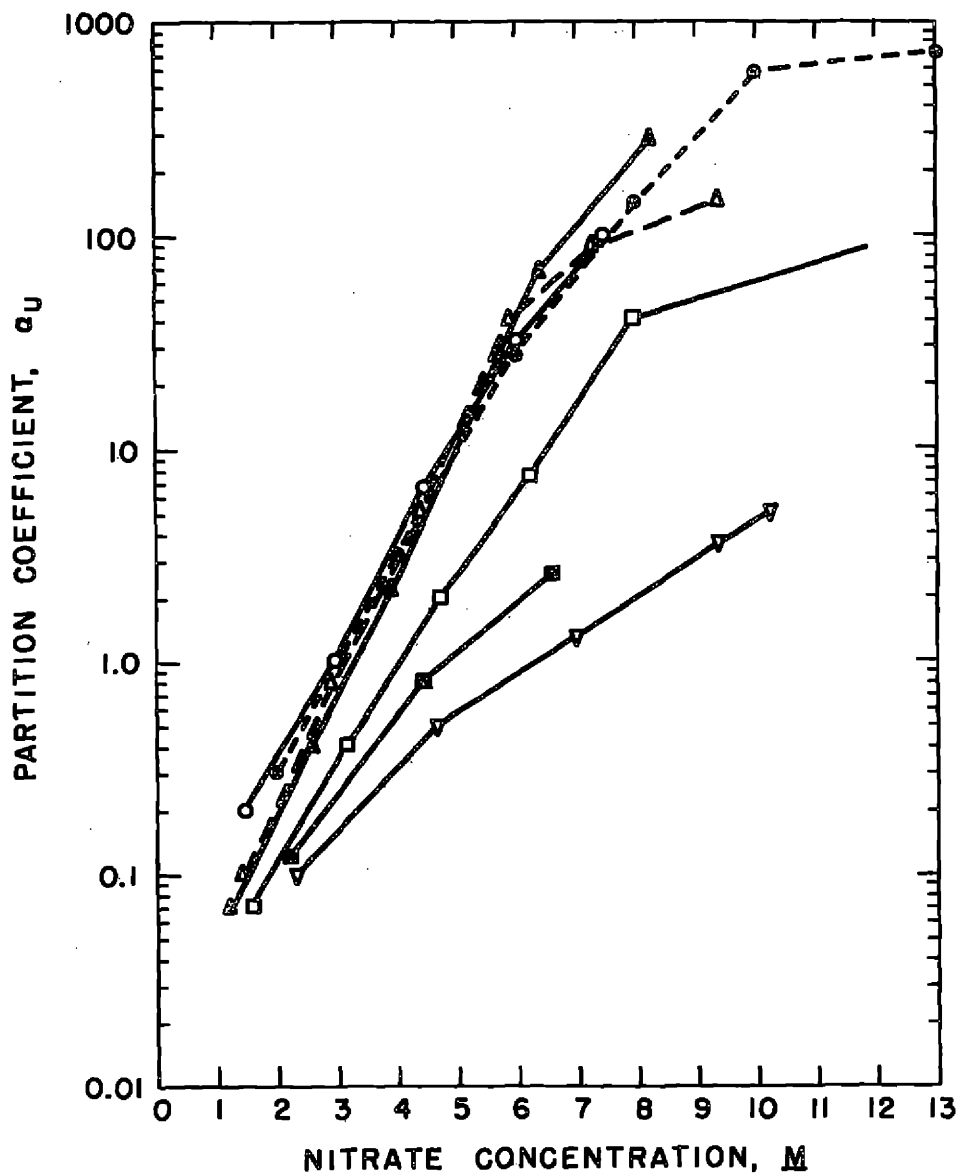


Figure 27. The effect of various salting-out agents on the extraction of uranium by hexone. ∇ NH_4NO_3 , \blacksquare NaNO_3 , \square $\text{Ca}(\text{NO}_3)_2$, \triangle $\text{Co}(\text{NO}_3)_2$, \blacktriangle $\text{Mg}(\text{NO}_3)_2$, \bullet $\text{Be}(\text{NO}_3)_2$, \circ $\text{Al}(\text{NO}_3)_3$. After W. H. Baldwin, reference 319. Conditions: Equal volumes of pure hexone used to extract aqueous containing 30 g U/liter.

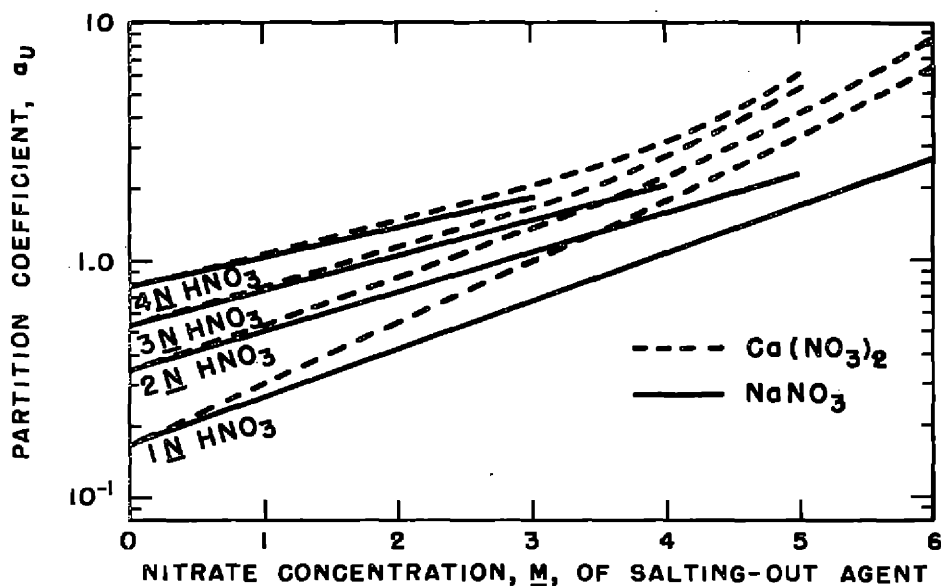


Figure 28.

The partition coefficient of uranium as a function of the nitrate concentration of the salting-out agents, $\text{Ca}(\text{NO}_3)_2$ and NaNO_3 , for an initial concentration in the aqueous phase of 100 g/l of uranium and 1, 2, 3, or 4M HNO_3 .

After A. Cacciari, R. DeLeone, C. Fizzotti, and M. Gabaglio, reference 375.

tion also causes an increased extraction of fission products (figure 29).³⁷⁸

The extraction of uranium by hexone is facilitated by the presence of substituted ammonium nitrates which are sufficiently soluble in the organic solvent. A number of these salts and their effect on the extraction of uranium are listed in Table XV.⁷⁸ Tri-n-butylamine, 2-hexyl pyridine, and dibenzoyl methane increase the extraction of fission products.³⁷⁸ Maeck, et al.³⁷⁹ have investigated the extraction of uranium by hexone from an aqueous solution containing aluminum nitrate and tetrapropylammonium nitrate. The extraction conditions adapted as a result of the

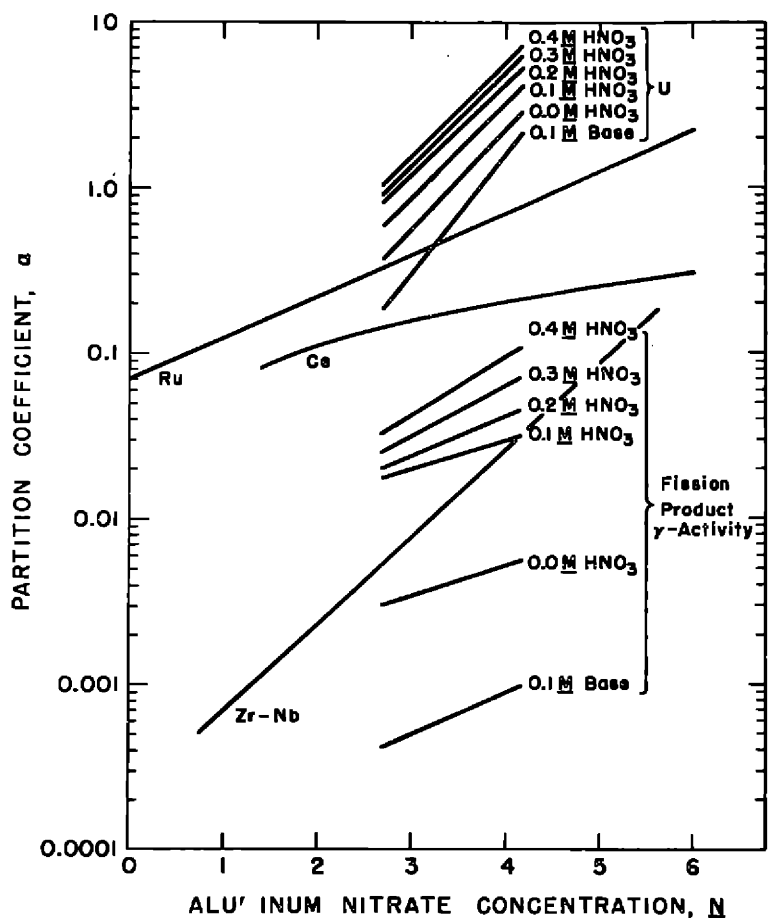


Figure 29. The effect of salting-out agent, $\text{Al}(\text{NO}_3)_3$, on the extraction of uranium and fission products by hexone from aqueous solution at various nitric acid concentrations. After F. R. Bruge, reference 378.

Conditions:

The results on uranium and gross fission product activity were obtained using $\text{Al}(\text{NO}_3)_3$ as salting-out agent and pretreated hexone as solvent. An irradiated uranium slug, cooled 144 days and dissolved in HNO_3 , was used as activity source. Extractions were made at 30°C from an aqueous phase oxidized 1 hour with 0.1M $\text{Na}_2\text{Cr}_2\text{O}_7$ at this temperature. The nitric acid is the sum of that in the aqueous and organic phases, expressed as moles of nitric acid per liter of aqueous phase.

Ruthenium extraction: aqueous phase-- 0.1M $\text{K}_2\text{Cr}_2\text{O}_7$, 0.2M HNO_3 , $\text{Al}(\text{NO}_3)_3$.

Cerium extraction: aqueous phase-- 0.025M $\text{Na}_2\text{Cr}_2\text{O}_7$, 0.5M HNO_3 , $\text{Al}(\text{NO}_3)_3$.

Zirconium-niobium extraction: aqueous phase-- 0.1M $\text{K}_2\text{Cr}_2\text{O}_7$, 0.25M HNO_3 , 8 g U per liter.

Table XV. Effect of Substituted Ammonium Nitrates (RNO_3) on the Extraction of Uranyl Nitrate by Hexone.^a

Cation, R	Total RNO_3 concentration (mol/l $\times 10^3$)	α_u
None	0	2.62
$(\text{C}_4\text{H}_9)_3\text{NH}$	2.1	6.2
	4.2	10.7
	11.0	32.4
	21.0	68
$(\text{C}_8\text{H}_{17})_2\text{NH}_2$ ^b	2.5	6.0
	5.0	9.4
	10.0	16.5
	20.0	33
$\text{C}_{11}\text{H}_{15}\text{NH}$ ^c	2.0	5.4
	5.0	13.4
	10.0	26.7
	20.0	57
$\text{C}_{12}\text{H}_{24}\text{N}_2\text{H}$ ^d	2.1	5.5
	10.5	25.6
$(\text{C}_4\text{H}_9)_4\text{N}$	10.0	97
	10.0	84
$\text{C}_6\text{H}_7\text{NH}$ ^e	10.0	2.8
$(\text{C}_2\text{H}_5)_3\text{NH}$	10.0	4.0
$(\text{CH}_2\text{CH}_2\text{OH})_4\text{N}$	10.0	2.67

^a After Kaplan, Hildebrandt, and Ader, reference 78.
Conditions:

equal volumes of hexone and of an aqueous solution. 8M in NH_4NO_3 , 0.4M in HNO_3 , and about 0.02M in uranyl nitrate.

^b di-2-ethylhexylammonium

^c 2-n-hexylpyridinium

^d methyl isobutyl ketazinium

^e 2-methylpyridinium

investigation were 4.0 ml of 2.8M aluminum nitrate, 1N acid-deficient, containing 0.1%(weight/volume) tetrapropylammonium nitrate; 2.0 ml hexone; and a sample size of 0.5 ml (~2 mg of uranium). These conditions provide a good separation of uranium from many ions. The separation from zirconium-niobium is particularly good. The recovery

of uranium is excellent even in the presence of foreign anions (10 to 1 mole ratio of anion to uranium). Of those anions tested, tungstate ion interferes most seriously (only 64.28% uranium extracted). Chloride, sulfate, phosphate, acetate, oxalate, etc., in the amounts tested, exhibit no appreciable interference in the extraction of uranium. Chloride does promote the extraction of those ions which form anionic chloride complexes, eg. gold (III). Certain other anions enhance the extraction of fission products, eg. dichromate and thiosulfate increase cerium extraction.

Aqueous thiocyanate systems. Uranium (VI) is extracted by hexone from aqueous thiocyanate solution.^{371,380} Reas³⁸⁰ has investigated the separation of uranium and thorium by this means. Some of his results are given in Table XVI. The effect of sulfate ion (experimental conditions B) is to hinder the extraction of both thorium and uranium. The effect, however, is greater for thorium than for uranium. Consequently, greater separation of thorium and uranium can be made in the presence of the complexing sulfate ion. The extraction of protactinium from an aqueous solution 1.2M in NH_4NO_3 , 0.20M in HNO_3 , about 0.01M in $\text{Th}(\text{NO}_3)_2$, 0.00987M in Na_2SO_4 , and 0.501M in KSCN by an equal volume of hexone was $\leq 4.4\%$. Decontamination from fission products is not too good. Equilibration of equal volumes of hexone and an aqueous solution approximately 0.04M in $\text{UO}_2(\text{NO}_3)_2$, 0.504M in $\text{Th}(\text{NO}_3)_4$, 0.485M in Na_2SO_4 , and 1M in HNO_3 resulted in a beta decontamination factor of about 6.6 and a soft gamma decontamination factor of about 1.5. Zirconium was found to be the principal fission product extracted.

Methylcyclohexanone

Aqueous nitrate systems. This solvent has been studied

Table XVI. Separation of U(VI) from Th(IV) by Thiocyanate Systems.^a

Experimental conditions	KSM M	U extracted %	Th extracted %
A	0.27	64.5	1.03-1.6
A	0.54	82	1.5-1.8
A	0.97	89.5	2.1-3.1
A	1.62	95	5.2-6.2
A	0.32 + 0.11M antipyrine	79	3.3
B	0.501	63	0.14*
B	0.25	35	0.015*

^a After W. H. Reas, reference 380.

Experimental conditions:

- A: 0.16M $UO_2(NO_3)_2$, 0.81M $Th(NO_3)_4$, volume of aqueous phase = 6.2 ml, volume of hexone = 10 ml.
 B: 0.0974M $UO_2(NO_3)_2$, 0.252M $Th(NO_3)_4$, 0.2M HNO_3 , 0.224M Na_2SO_4 , volume of aqueous phase = 10 ml, volume of hexone = 10 ml.

* The thorium extraction was performed under slightly different conditions in that NH_4NO_3 was substituted for $UO_2(NO_3)_2$. An ionium (Th^{230}) tracer was added to the solution and the distribution was measured by the determination of ionium in each phase.

by workers in Czechoslovakia as a means of separating uranium from thorium³⁸¹ and fission products.^{323,382} The extractability of uranium by methylcyclohexanone from sodium nitrate solution (6-8M) is considerably better than that of thorium.³⁸¹ From nitric acid solution (8M), the extractability of uranium is only two- to three-fold greater than that of thorium.³⁸¹ Ammonium nitrate is comparable to sodium nitrate as a salting-out agent for uranium.³²³ Aluminum nitrate is more effective than either. However, the order of salting-out agents in causing increased fission product extractability is $Al > Na > NH_4$. The best separation of uranium from fission products is achieved with ammonium nitrate as the salting-out agent.³²³ Methylcyclohexanol present in commercial methylcyclohexanone suppresses the extraction of uranium and fission

products. The separation factor between the two activities, however, is increased since the partition coefficient of fission products is decreased more than that of uranium.³²³ The partition of uranium between methylcyclohexanone and water and methylcyclohexanone and 6M ammonium nitrate solution is given in figure 4-B.³²³

Other ketonic solvents

Hyde and Wolf³¹⁶ have studied the extraction of uranium and thorium by methyl n-amyl ketone and diisopropyl ketone as a function of total nitrate concentration in the aqueous phase. In both cases, uranium was better extracted than thorium. The extraction of thorium did not become appreciable (<5%) until the aqueous nitrate concentration was greater than 5M. Diisopropyl ketone was found to be an excellent extractant of protactinium.³¹⁶

Vesely, Beranová, and Malý³⁸² have investigated the extraction of uranium and fission products by several methylalkyl ketones: methylhexyl, methylamyl, methylbutyl, and methylpropyl in addition to methyl isobutyl and methylcyclohexanone. The partition coefficients of both uranium and fission product activity were measured as a function of acid concentration in the range of 0.4 to 3M nitric acid. In this acidity range, fission product extraction was found to be maximum in the 0-1M nitric acid region. The partition coefficient, α_{FP} , in this region was greatest with methylpropyl ketone (22×10^{-3} at 0.61M) and least with methylhexyl ketone (2.4×10^{-3} at 0.03M). In the acid-deficient region, α_{FP} increased as the acid-deficiency was decreased (the solution was made more acidic). After the maximum α_{FP} was reached in the 0-1M acid region, the partition coefficient was decreased and then increased as the aqueous solution was made more acidic up to 3-4M. The partition

coefficient of uranium, α_u , increased as the nitric acid concentration was increased over the entire range. At 3M nitric acid, α_u varied from about 0.7 for methylhexyl ketone to about 2 for methylpropyl ketone and 3 for methylcyclohexanone. The greatest separation, β , of uranium from fission products was found in the 0.1M acid deficient region (-0.1M). For methylhexyl ketone, β was found to be >1500; for methylpropyl ketone, β was about 400. The extraction coefficients of uranium were < 0.2 for methylhexyl ketone and about 0.5 for methylpropyl ketone at this acid concentration.

Allen³⁸³ has tested diisobutyl ketone, diisopropyl ketone, and methylhexyl ketone as solvents for the purification of uranium from iron, copper, chromium, and nickel. Diisobutyl ketone was found most satisfactory under the conditions tested. Diisopropyl extracted some iron and chromium. Methylhexyl ketone extracted iron, chromium, and copper.

Uranium and thorium may be extracted quantitatively from a nitrate medium by mesityl oxide.³⁸⁴ Under the conditions tested* zirconium is extracted to a large extent; vanadium and yttrium to a lesser extent; cerium only slightly.

ALCOHOLS

Hyde and Wolf³¹⁶ found alcohols to be only fair extractants of uranium and the extraction capacity to decrease rapidly with the length of the carbon chain. This is borne out by the work of Poston, et al.³⁸⁵ who measured the extraction coefficients of uranium and ruthenium as a function of aluminum nitrate concentration of the aqueous phase for

* Experimental conditions: A salt of the elements tested was dissolved in 10 ml of HNO_3 (15 + 85). Nineteen grams of aluminum nitrate crystals were added and dissolved. The solution was shaken for 15 seconds with 20 ml of mesityl oxide. The extract was washed once with 20 ml of aluminum nitrate solution and analyzed.

hexone and several tertiary alcohols: tertiary amyl alcohol, 2-methyl 2-pentanol, 2-methyl 2-hexanol, 2-methyl 2-heptanol. Only tertiary amyl alcohol extracted uranium better than hexone (0.5 - 1.5M $\text{Al}(\text{NO}_3)_3$, 0.2M HNO_3) and all four alcohols extracted ruthenium better than hexone. Ruthenium was extracted as well or nearly as well as uranium by the alcohols.

Diisobutylcarbinol extracts ruthenium nearly as well as uranium.³⁸⁶ Thorium and zirconium-niobium are poorly extracted. Protactinium is extracted much more efficiently than uranium.

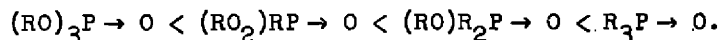
MISCELLANEOUS SOLVENTS

Nitromethane has been recommended by Warner³²² as an extractant for uranium. It is resistant to oxidation, stable to high concentrations of nitric acid, and highly selective. The distribution of uranyl nitrate between nitromethane and water and nitromethane and saturated ammonium nitrate solution is given in figure 3-B. The extraction of thorium nitrate by nitromethane from aqueous solution is much less than that of uranium. Color tests indicate that neither copper, cobalt, iron (III), nor chromium nitrate is extracted by the solvent. With diethyl ether, considerable amounts of copper nitrate and trace amounts of ferric nitrate are extracted. Nitric acid enhances the extraction of uranyl nitrate by nitromethane. However, above a critical acid concentration (~5N initial acid concentration with equal phase volumes at 20°C) only one liquid phase is formed.

ORGANOPHOSPHORUS COMPOUNDS. Within recent years, a large number of organophosphorous compounds have been developed and investigated as extractants for uranium. These compounds have been subdivided in the present paper into neutral and acidic organophosphorus compounds.

NEUTRAL ORGANOPHOSPHORUS COMPOUNDS

Solvents included in this category are trialkylphosphates, $(RO)_3P \rightarrow O$; dialkyl alkylphosphonates, $(RO)_2RP \rightarrow O$; alkyl dialkylphosphinates, $(RO)R_2P \rightarrow O$; and trialkylphosphine oxides $R_3P \rightarrow O$. The ability of the solvents to extract uranium is in the order

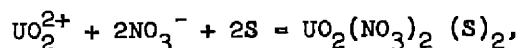


This is also the order of increasing base strengths of the phosphoryl oxygen.³⁰² In Table XVII, the four types of compounds are compared as extractants of uranium (VI), plutonium (IV), thorium, fission products, and acids.³⁸⁷

It should be noted that although uranium is extracted almost quantitatively by tributylphosphine oxide (Table XVII), other elements are also highly extracted. In fact, in spite of lower extraction coefficients, tributyl phosphate affords a better separation of uranium from thorium, plutonium (IV), and fission products under the conditions listed in Table XVII than does tributylphosphine oxide.

Tables XVIII^{388,389} and XIX³⁹⁰ list the distribution coefficients of uranium and some associated elements for a number of neutral organophosphorus extractants. Similar information on other solvents may be found in papers by Burger,³⁹¹ Healy and Kennedy,¹⁸⁸ and in numerous ORNL reports. The latter have been summarized by Blake, et al.³⁰² and Brown, et al.^{303,305}

The mechanism of extraction by neutral organophosphorus reagents appears to be similar to that of tributyl phosphate.^{188, 302,387} From nitrate systems, the extraction of uranium by tributyl phosphate and trioctylphosphine oxide is described fairly well by the equilibrium reaction



where S represents the solvent molecule.³⁰² Extraction may be made by tributyl phosphate from chloride solution. Stronger

Table XVII. Comparison of the Extractive Capacities of Various Types of Organophosphorus Compounds^a

	Extraction of $UO_2(NO_3)_2$					Extraction of U, Pu, and Fission Products ^b								Extraction of Th ^c
	Nitric acid concentration					No added HNO_3				0.6M HNO_3				
	0M	0.6M		3M										
	U ext'd (%)	U ext'd (%)	Acid ext'd (%)	U ext'd (%)	Acid ext'd (%)	U (VI) (%)	Pu (IV) (%)	Gross β (%)	Gross γ (%)	U (VI) (%)	Pu (IV) (%)	Gross β (%)	Gross γ (%)	Th ext'd (%)
Tributyl phosphate	11	56	4	96.5	8	17.4	0.7	0.01	0.01	58	6.6	0.07	0.08	3.5
Dibutyl butylphosphonate	55	97	6	99.4	11	64	1.13	0.05	0.13	97	54	0.72	1.0	18
Butyl dibutylphosphinate	98.5	99.9	15	99.9	14	94	20	1.9	7.9	99.9	98	23	38	74
Tributyl phosphine oxide	99.7	99.9	39	99.9	17	99.9	97.3	37	64	99.9	99.8	72	77	98.7

^a After Higgins, Baldwin, and Ruth, reference 387.

Experimental conditions: equal phase volumes equilibrated 30 minutes at $25 \pm 0.2^\circ C$; organic phase - 0.75 M phosphorus compound dissolved in CCl_4 ; aqueous phase - 0.1M uranyl salt with or without salting agent.

^b Aqueous phase - 0.1M $UO_2(NO_3)_2$ from dissolving irradiated U slugs in HNO_3 - 6 months cooling.

^c Aqueous phase - 0.1M $Th(NO_3)_4$.

	Extraction of $\text{UO}_2\text{SO}_4^{\text{d}}$		Extraction of UO_2Cl_2		Extraction of acids ^f						
	H_2SO_4 concentration		HCl concentration		% extracted						
	<u>0M</u>	<u>2M</u>	<u>0M</u>	<u>1M</u>	Acetic	Citric	Tartaric	HNO_3	H_2SO_4	H_3PO_4	HCl
	U ext'd (%)	U ext'd (%)	U ext'd (%)	U ext'd (%)							
Tributyl phosphate	0.001	0.001	0.3	0.8	25	0	0	7	0	0	0
Dibutyl butylphosphonate	0.1	0.03	1	26							
Butyl dibutylphosphinate	16	48	40	92	49	20	20	27	2	3	0
Tributyl phosphine oxide	95	96.8 ^e	90 some ppt. formed	all precipitated	58	39	23	39	0	10	7

^d Aqueous phase - 0.1M UO_2SO_4 ; acid as indicated.

^e Three phases were present, two of which were largely aqueous.

^f Aqueous phase - approximately 1N acid.

Table XVIII. Extraction of Actinides and Zirconium by Phosphates and Phosphonates at 30°C^a

Solvent	Extraction Coefficient						Zr
	Th	Np(IV) ^b	Pu(IV) ^c	U(VI)	Np(VI) ^d	Pu(VI) ^d	
Trialkyl phosphate^e							
n-butyl	2.9	3.2	16.1	26	15.6	3.5	0.22
isobutyl	2.4	2.7	11.8	22	15.9	3.4	
n-amyl	2.9	4.2	15.6	32	19.3	4.1	
iso-amyl	4.2	4.7	17.8	34	18.9	4.4	0.12
n-hexyl	3.0	3.6	15.6	38	20.0	4.5	0.14
n-octyl	2.4	3.4	15.3	33	15.7	3.9	0.14
2-ethylhexyl	2.5	4.3	25	58	23	5.7	0.14
2-butyl	0.45	4.9	28	42	20	4.6	
3-amyl	0.22	3.5	18.1	49	22	5.0	
3-methyl-2-butyl	0.18	3.0	24	47	25	5.4	
4-methyl-2-amyl	0.047	3.5	22	38	24	4.9	
cyclohexyl	3.5			106			0.64
Dialkyl alkylphosphonate^f							
di-n-butyl n-butyl	24			92			0.17
di-n-butyl cyclohexyl	17			125			0.14
di-n-amyl n-amyl	33			133			0.092
di-n-hexyl n-hexyl	26			89			0.070 E
di-2-ethylhexyl 2-ethyl-hexyl	10.6			176			0.12

^a After T. H. Siddall III, references 388 and 389.

^b Aqueous phase contained 0.01M ferrous sulfamate.

^c Aqueous phase contained 0.01M NaNO₂.

^d Aqueous phase contained 0.01M ceric ammonium sulfate.

^e 1.09M trialkyl phosphate in n-dodecane; extractants washed with 1M NaOH, water, and nitric acid before use; aqueous phase 3.0M HNO₃ at equilibrium; tracer concentration of element.

^f 1.09M phosphonate in n-dodecane; aqueous phase 0.8M HNO₃ at equilibrium; tracer concentration of element.

^E Extrapolated value.

extractants may extract uranium from sulfate and phosphate solutions, especially if a small amount of nitrate is added to the solution.

Tributyl phosphate (TBP)

Aqueous nitrate systems. Investigations on the extraction of metal nitrates by TBP indicate the cations are extracted as single, well-defined species: $M(NO_3)_3(TBP)_3$, $M(NO_3)_4(TBP)_2$, and $MO_2(NO_3)_2(TBP)_2$.^{168,187,311,392,393} This differs from the extraction of ethers, esters, and other oxygen-containing solvents, considered previously, in which a whole series of complexes containing varying numbers of nitrate, solvate, and water molecules is extracted. The stability of the TBP-solvate molecules increases in the order³¹¹ $H_2O(TBP) < Pu(NO_3)_3(TBP)_3 < Pa(NO_3)_5(TBP)_3 < HNO_3(TBP)$, $HNO_3(TBP)(H_2O) < Th(NO_3)_4(TBP)_2 < PuO_2(NO_3)_2(TBP)_2 < Pu(NO_3)_4(TBP)_2 < UO_2(NO_3)_2(TBP)_2$.

The unexpectedly large extraction of nitric acid, thorium, and zirconium at very high acid concentrations indicates higher complexes may be extracted.³¹¹

The distribution of uranyl nitrate between TBP and water is represented in figure 30.³⁹⁴ The extraction coefficient of uranium is plotted as a function of acid concentration for various concentrations of TBP in figure 31.³⁹⁵ The observed effects of (1) decreased α_u with increased acid concentration and (2) increased α_u with increased TBP concentration, may be explained on the basis of free solvent concentration. First, as the acid concentration is increased, more nitric acid is extracted resulting in less free solvent. Second, more free solvent is obviously available as a result of increasing the solvent concentration. Since the partition coefficient of uranium depends upon the second power of the free solvent concentration, α_u will decrease in the first case

Table XIX. Extractive Power of Tributyl Phosphate, Alkyl Phosphonates, Diphosphonates, and Phosphine Oxides².

Solvent	Initial concentration of HNO ₃ , N			0.5			1			2		
	0.5	1	2	0.5	1	2	0.5	1	2	0.5	1	2
Distribution coefficients												
Solvent	U			Pu(IV)			Zr-Nb			Nb		
(C ₄ H ₉ O) ₃ PO	0.68	1.42	4.65	0.19	0.67	2.73	.64	.86	1.05	.011	.014	.010
(C ₄ H ₉ O) ₂ P-CH ₃	4.50	10.35	21.50	5.46	11.15	14.90	.0058	.0040	.015	.0001	.0020	.017
(1-C ₅ H ₁₁ O) ₂ P-CH ₃	6.95	21.10	70.60	1.53	7.43	16.90	.15	.22	.33	.52	.34	.20
(C ₆ H ₁₃ O) ₂ P-CH ₃	5.72	18.45	45.70	3.51	11.65	18.95	.10	.053	.13	.0050	.012	.021
(C ₇ H ₁₅ O) ₂ P-CH ₃	21.30	107.1	103.1	3.25	10.95	17.15	.0078	.040	.12	.00040	.0015	.00080
(C ₈ H ₁₇ O) ₂ P-CH ₃	22.10	50.70	178.0	5.71	13.65	19.45	.037	.14	.28	.0061	.019	.044
(C ₉ H ₁₉ O) ₂ P-CH ₃	27.60	32.60	61.40	23.25	24.15	21.70	.97	.67	.55	.013	.014	.017
(C ₁₀ H ₂₁ O) ₂ P-CH ₃	6.73	14.65	34.30	16.60	21.65	21.50	.38	.40	.54	.0043	.0043	.0060
(o-C ₆ H ₁₁ O) ₂ P-CH ₃	17.45	44.10	101.0	4.47	17.35	25.15	.020	.061	.30	.015	.013	.0057
(C ₆ H ₅ O) ₂ P-CH ₃	0.033	0.053	0.15	1.84	1.34	1.71	1.31	1.42	.86	.40	.15	.086
(C ₄ H ₉ O) ₂ P-CH ₃	13.43	28.80	69.90	3.80	11.60	19.60	.24	.16	.24	.0056	.0053	.0062
(C ₆ H ₁₃ O) ₂ P-CH ₃	12.20	36.3	65.8	7.89	17.85	24.50	.80	.47	.47	.0038	.0041	.0051
(n-C ₄ H ₉ O) ₂ P-C ₂ H ₅	4.97	13.05	44.20	2.58	7.58	14.85	.031	.091	.33	-	-	-
(1-C ₄ H ₉ O) ₂ P-C ₂ H ₅	7.58	22.50	57.30	2.53	7.67	14.85	.0085	.018	.11	.0021	.0089	.039
(C ₄ H ₉ O) ₂ P-C ₃ H ₇ -n	10.45	22.20	74.70	2.58	8.65	16.75	.0039	.020	.10	<.00016	.0028	.0063

$(1-C_5H_{11}O)_2P - C_3H_7-n$	9.40	22.30	73.80	2.02	7.46	14.15	-	.013	.099	.0022	.0027	.0022
$(C_4H_9O)_2P - C_4H_9$	8.98	24.00	65.30	2.82	9.46	17.50	.0034	.024	.097	<.001	.0028	.025
$(C_4H_9O)_2P - C_5H_{11}$	8.26	25.75	103.5	2.45	8.91	17.65	.0032	.024	.11	.0035	.010	.032
$(1-C_5H_{11}O)_2P - C_5H_{11}$	13.65	28.40	58.40	2.50	9.00	15.75	.012	.036	.13	.011	.022	.074
$(1-C_5H_{11}O)_2P - C_5H_{11}-1$	7.68	21.10	30.95	2.18	7.69	13.05	.011	.031	.13	.0017	.0024	.0031
$(1-C_5H_{11}O)_2P - C_8H_{17}$	9.60	11.50	85.4	2.65	8.92	17.00	.016	.038	.13	.0012	.011	.029
$(C_4H_9O)_2P - CH_2C_6H_5$	2.45	6.34	21.10	0.47	1.91	6.00	.0034	.0084	.022	.0092	.0011	.0012
$(C_4H_9O)_2P - CH_2OCH_3$	1.13	3.09	7.79	0.28	1.16	4.24	.029	.075	.12	.0029	.0046	.0061
$(C_4H_9O)_2P - CH_2OC_2H_5$	1.31	3.11	8.86	0.34	1.50	5.03	.43	.42	.42	.0046	.0057	.0067
$(C_4H_9O)_2P - CH_2CH_2OC_4H_9$	3.15	8.68	42.20	0.80	3.26	10.95	.031	.077	.062	.0050	.0060	.012
$(C_4H_9O)_2P - CH(CH_3)CH_2C(=O)C_4H_9$	2.30	6.35	34.50	0.15	2.35	6.39	.021	.025	.043	.0031	.0041	.0058
$(C_4H_9O)_2P - CH_2CH_2COOC_4H_9$	2.14	5.57	16.46	0.79	2.38	6.24	.074	.12	.057	.0058	.0080	.0066
$(C_4H_9O)_2P - CH_2 - P(C_4H_9O)_2$	1.73	12.55	17.26	2.71	6.72	19.24	.23	.28	.29	.011	.011	.011
$(1-C_5H_{11}O)_2P - CH_2 - P(OC_5H_{11}-1)_2$	1.32	2.98	9.60	3.33	8.45	24.70	.41	.41	.42	.032	.031	.031
$(C_4H_9)_3PO$	1880	1360	352	365	299	83.2	1.00	1.84	2.45	.64	.60	.47
$(1-C_4H_9)_3PO$	5.33	7.26	7.19	16.45	21.85	22.55	.61	.55	.55	3.95	4.40	2.15

^a After Petrov, et al., reference 390.

Organic phase - 0.5M phosphorus compound in CCl_4 . Initial aqueous phase - 50 g/l U, 1 g/l Pu(IV), 1 c/l Zr^{95} + Nb^{95} , 1 c/l Nb^{95} , nitric acid as indicated. Volume ratio (organic/aqueous), 2. Time of shaking and of settling, 30 minutes. Temperature, $20^\circ \pm 1^\circ C$.
1 = iso; c = cyclo.

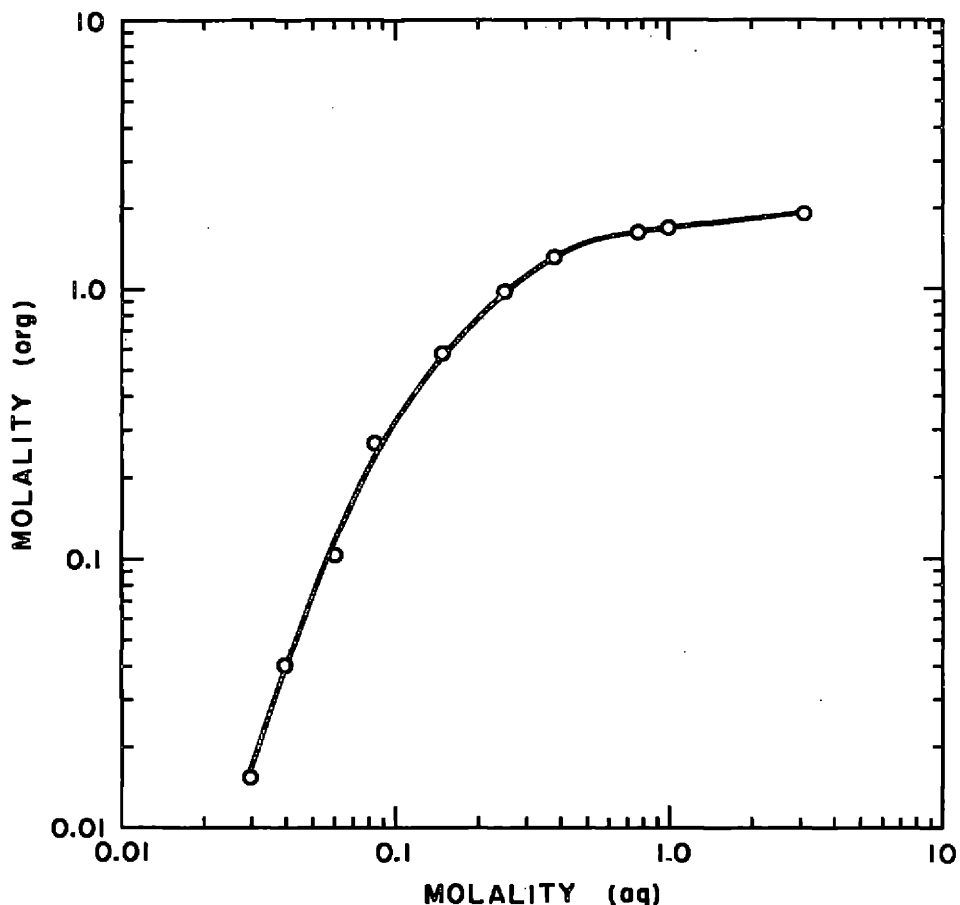


Figure 30.

The partition of uranyl nitrate between 100% TBP and water at 25°C.

After T. Healy, J. Kennedy, G. Waind, reference 394.

and increase in the second. The effect of uranium concentration on α_u is given also as a function of nitric acid concentration in figure 32.³⁹⁶ The decreased extraction with increased uranium concentration may again be interpreted in terms of the solvent available. The partition coefficients of other metal nitrates are also decreased, in general, by increased uranium concentration. More efficient separation may therefore be achieved by increased uranium loading of the solvent. For small amounts of uranium, a high uranium concentration may be attained in an organic phase suitable for handling by

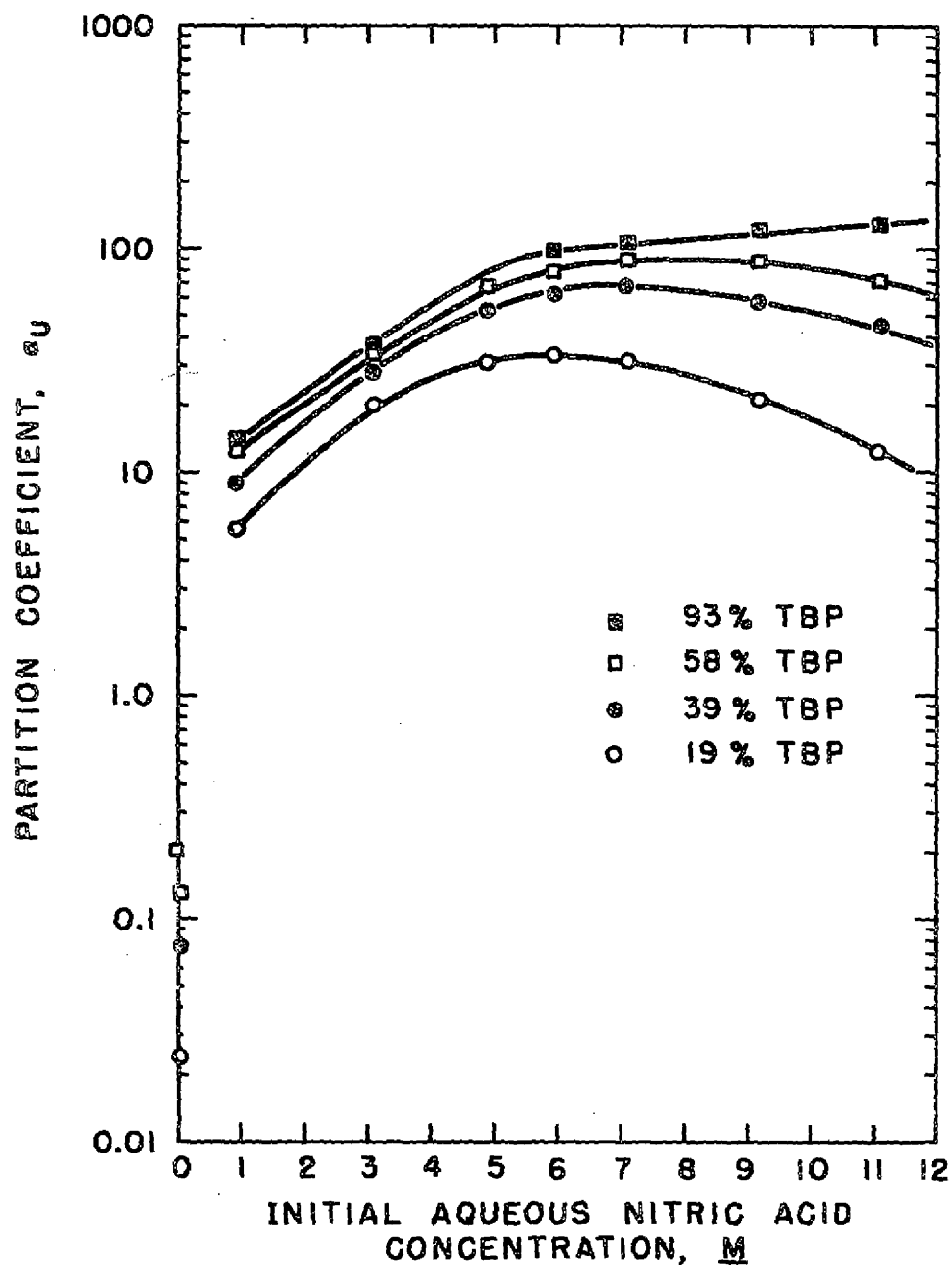


Figure 31. The extraction of uranyl nitrate by various concentrations of TBP in kerosene as a function of initial aqueous acid concentration. After T. Sato, reference 395. Conditions: Organic phase - volume % TBP in kerosene as indicated. Aqueous phase - 5 g. uranyl nitrate per liter, nitric acid concentration indicated. Temperature, 20°C; V_o/V_a , 1.

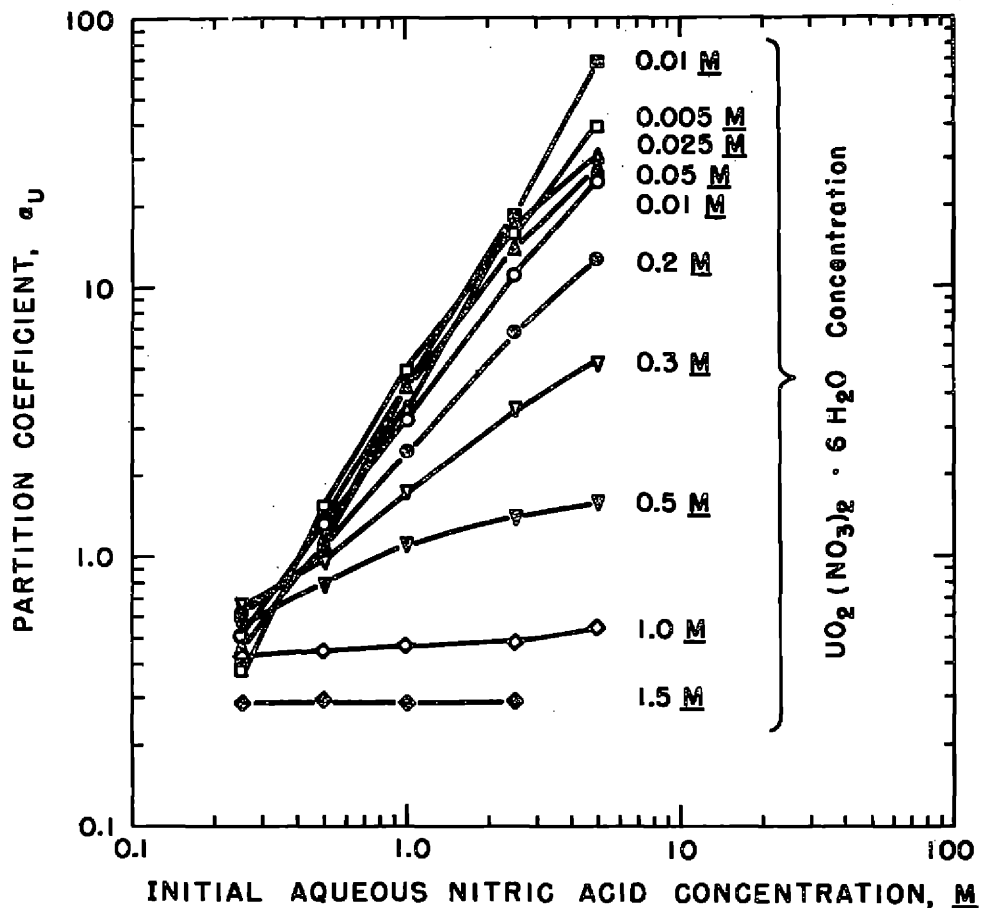


Figure 32.

The effect of initial uranium concentration on the extraction of uranyl nitrate by 20 volume per cent TBP in CCl_4 as a function of initial aqueous nitric acid concentration.

After R. L. Moore, reference 396.

Conditions:

Equal volumes of phases shaken in a water bath at $25^\circ C$.

dilution of the solvent. Duncan and Holburt³⁹⁷ have measured the distribution of uranium, initially present in 1.2 to 1200 micrograms per liter, between 20% TBP in kerosene and nitric acid solution. Although the results were somewhat erratic, it was generally shown that the partition coefficient is nearly constant over this range of uranium concentrations.

The extractive capacity of TBP is affected considerably by the choice of diluent. Taube³⁹⁸ extracted U(VI), Np(IV), Np(VI), and Pu(VI) from 5M HNO₃ aqueous solutions with 0.15M TBP dissolved in a number of solvents, including benzene and chloroform. Larger extraction coefficients were obtained for all the elements tested with benzene rather than chloroform as diluent. In the case of uranium, the difference in α_u was greater than ten-fold. Little difference in extractive capacity was observed with TBP diluted by benzene or carbon tetrachloride. Similar results were obtained by Dizdar, et al.³⁹⁹ Uranyl nitrate (0.0038M) was extracted from 2M nitric acid solutions by various concentrations of TBP diluted with carbon tetrachloride, xylene, kerosene, hexane, dibutyl ether, diethyl ether, and isopropyl ether. The partition coefficient was found to increase with increasing TBP concentration to a maximum for pure TBP. For carbon tetrachloride and xylene the maximum value was already attained at 40 mole per cent TBP. The other diluents are listed above in the approximate order in which they inhibit the extraction of uranium by TBP. Differences in α_u , for various diluents, were found to become smaller with increased uranium concentration. Bruce³⁷⁸ has found that the extraction of fission products is also affected by the choice of diluent.

The extraction of uranium by TBP is considerably enhanced by the presence of salting-out agents in the aqueous phase.⁴⁰⁰⁻⁴⁰⁶ The results of Sato⁴⁰⁶ are given in Table XX and figure 33.

The extraction of uranium by TBP decreases with increased temperature.⁴⁰⁷⁻⁴¹¹

Phosphate, sulfate, and fluoride ions reduce the extraction of uranium by TBP from nitrate media.⁴⁰⁹ Uranium is extracted from chloride solution but less efficiently than from nitrate solution. Silica causes poor phase separation

Table XX. Extraction of Uranyl Nitrate by TBP Using Various Nitrate Salting-out Agents.^a

Salting-out agent	Percentage extracted			
	0M HNO ₃	1M HNO ₃	3M HNO ₃	6M HNO ₃
(HNO ₃)	2.96	82.10	95.52	97.10
NH ₄ NO ₃	70.00	92.30	96.50	97.40
LiNO ₃	73.05	94.60	97.50	97.95
NaNO ₃	72.50	93.00	97.10	97.80
KNO ₃	65.00	90.50	96.00	97.10
Cu(NO ₃) ₂	86.02	97.50	97.70	99.00
Mg(NO ₃) ₂	84.35	97.20	97.50	99.40
Ca(NO ₃) ₂	82.48	96.60	96.60	97.50
Zn(NO ₃) ₂	79.75	98.05	98.10	99.40
Al(NO ₃) ₃	99.90	99.60	98.20	99.50
Fe(NO ₃) ₃	99.80	99.50	98.20	98.90

^a After T. Sato, reference 406.

Organic phase - 19% TBP in kerosene.
 Aqueous phase - 5 g/l uranyl nitrate, 1M salting-out agent, initial acid concentration indicated.
 Equal phase volumes shaken together for 30 minutes at 20°C.

and the formation of emulsions.

Uranium may be re-extracted from TBP by contact with sodium carbonate solution.⁴⁰⁹ Ammonium sulfate, sodium sulfate, and urea solutions have been used satisfactorily.⁴⁰⁹ Water or hydrogen peroxide is ineffective for TBP containing considerable nitric acid.⁴⁰⁹

The distribution of nitric acid between aqueous solution and 100% TBP is demonstrated in figure 34.⁴¹² The distribution of various metal nitrates between TBP and nitrate

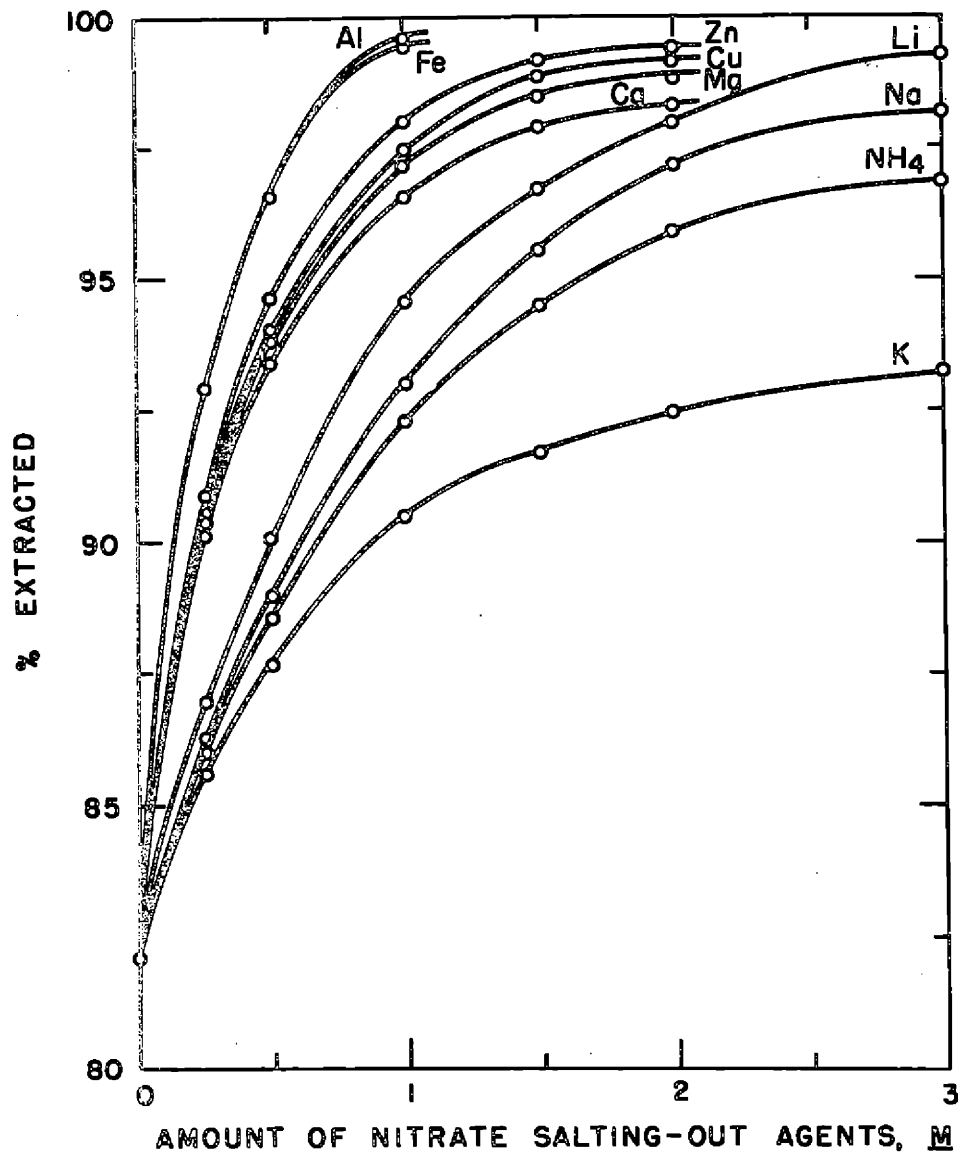


Figure 33.

Effect of nitrate salting-out agents upon the extraction of uranyl nitrate at 1M initial nitric acid concentration.

After T. Sato, reference 406.

Conditions:

Organic phase - 19% TBP in kerosene.
 Aqueous phase - 5 g/l uranyl nitrate, 1M HNO₃, salting-out agent concentration indicated.
 Equal phase volumes shaken together for 30 minutes at 20°C.

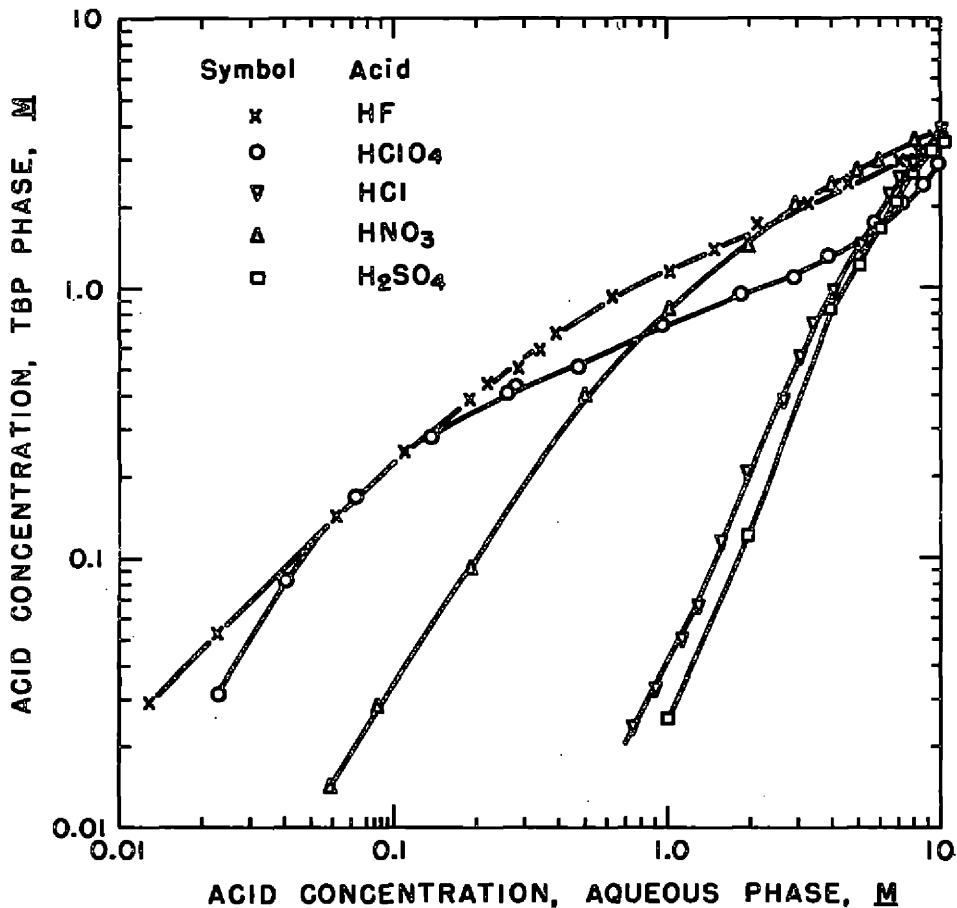


Figure 34.

The distribution of mineral acids between 100% TBP and aqueous solution at 25°C.

After E. Hesford and H. A. C. McKay, reference 412.

solutions has been extensively investigated. The extraction coefficients of some actinide elements are plotted against aqueous nitric acid concentration in figures 35 and 36.⁴¹³⁻⁴¹⁶ Ishimori and Nakamura⁴¹⁷ have also measured the partition coefficients of Hf, Th, Pa, U(VI), Np(IV)(V)(VI), and Pu(IV)(VI) at various aqueous nitric acid concentrations. Figure 37 represents the partition coefficient of several fission products as a function of the nitric acid concentration.⁴¹⁸⁻⁴²⁰

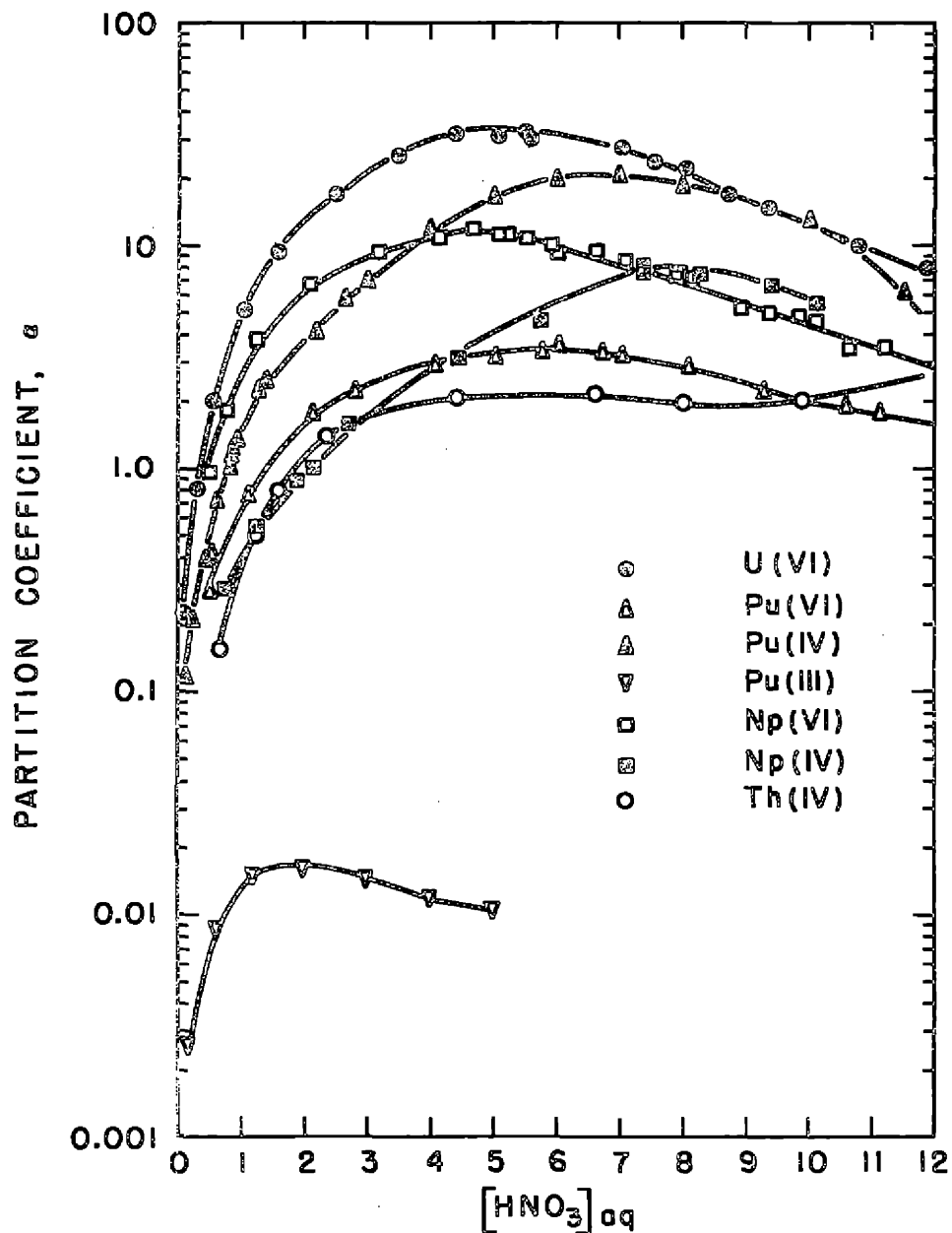


Figure 35. The partition coefficient of actinide nitrates between 19% TBP in kerosene and aqueous solution as a function of equilibrium nitric acid concentration. \odot U (VI), \square Np (VI), \blacksquare Np (IV), \triangle Pu (VI) at 20°-23°C., after K. Alcock, G. F. Best, E. Hesford, H. A. C. McKay, reference 413. \triangle Pu (IV), ∇ Pu (III), at 25°C. or 20-23°C., after G. F. Best, H. A. C. McKay, P. R. Woodgate, reference 414. \circ Th (IV) at 25°C., after E. Hesford, H. A. C. McKay, D. Scargill, reference 415.

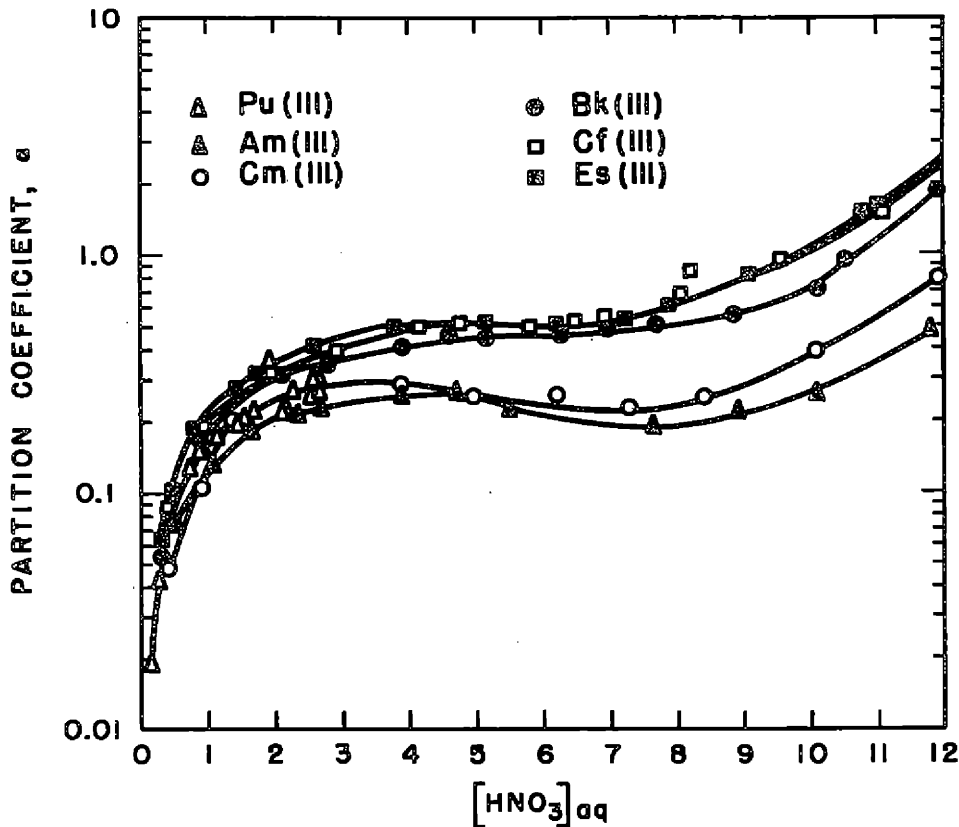


Figure 36.

The distribution of trivalent actinides between 100% TBP and aqueous solution as a function of equilibrium nitric acid concentration at 25°C.

After G. F. Best, E. Hesford, and H. A. C. McKay, reference 416.

The extraction of rare earths, Y, Zr, Sc, Th, and Am by TBP from aqueous nitric acid solution has been investigated by Peppard and co-workers.^{421,422} Iodine is extracted. It forms addition compounds with carbon-unsaturated compounds in the solvent. The extraction of iodine is minimized by keeping it in a reduced state and by careful selection of TBP diluents.³⁷⁸ Ruthenium is also extracted by TBP. Its extraction may be reduced by increased solvent saturation with uranium, by digestion in a nitrate solution of very high ionic

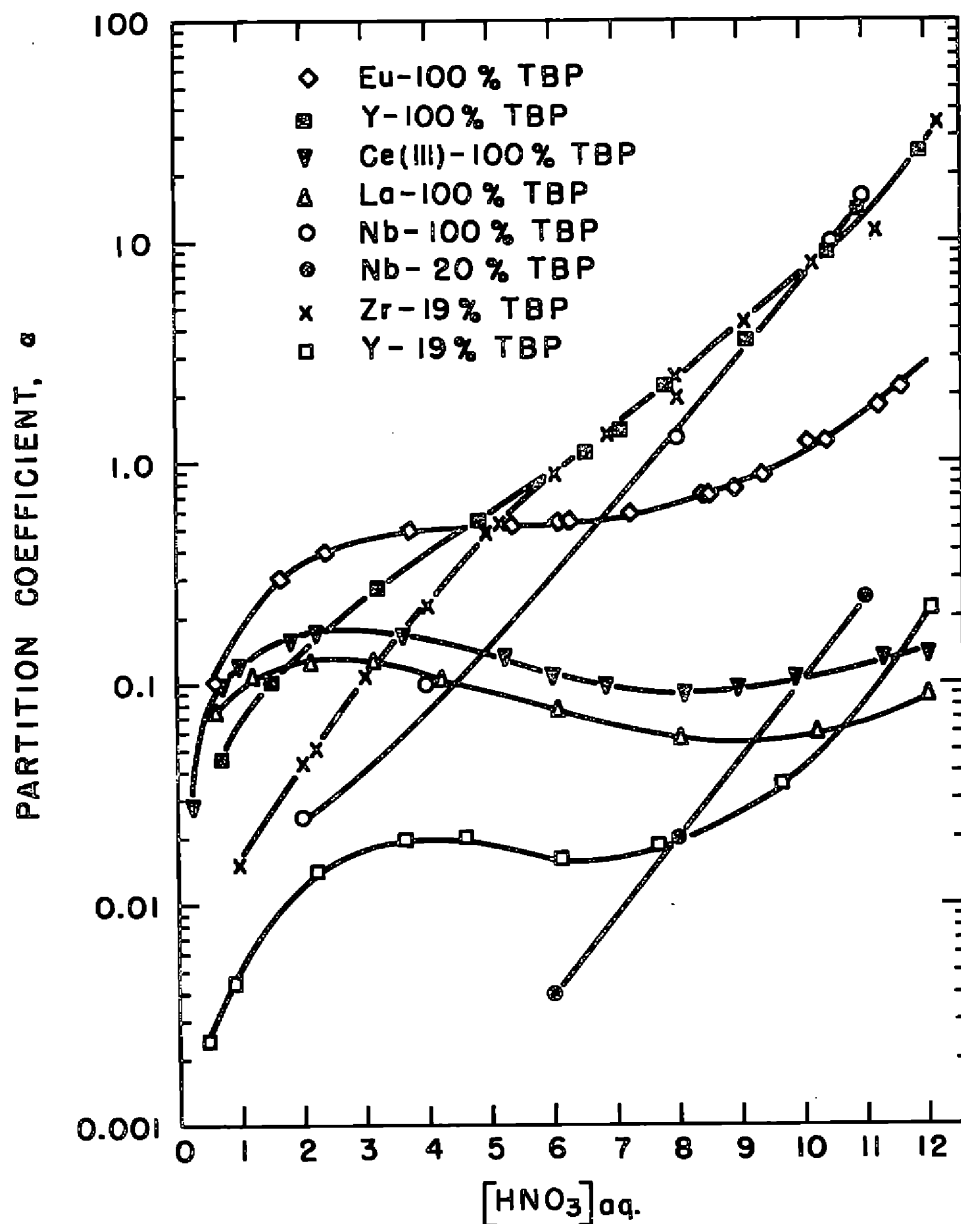


Figure 37. The distribution of fission product elements between TBP of several concentrations and aqueous solution as a function of equilibrium nitric acid concentration. Zr, after K. Alcock, F. C. Bedford, W. H. Hardwick, and H. A. C. McKay, reference 418. Y, La, Ce, Eu, after D. Scargill, K. Alcock, J. M. Fletcher, E. Hesford, and H. A. C. McKay, reference 419. Nb, after C. J. Hardy and D. Scargill, reference 420. Conditions: Tracer, carrier-free or with less than 1 g/l of carrier, used in all cases, Zr equilibrations made at 20-23°C.; Nb, ~20°C.; all others at 25°C. TBP diluted with kerosene.

strength, or by treatment with a reducing agent.³⁷⁸ Susic and Jelio⁴²³ and Sato⁴⁰⁶ have studied the TBP extraction of metal nitrates that may be used as salting-out agents. The order of extraction of 0.1 mg per ml concentrations of metal from 2N HNO₃ solutions by 20% TBP/kerosene with no uranium present is Bi > Co > Cu > Fe > Zn > Cd > Pb.⁴²³ The partition coefficient of bismuth under such conditions, with equal phase volumes, is about 0.1. The results of Sato⁴⁰⁶ are listed in Table XXI.

The partition of uranium and other metal nitrates between tributyl phosphate and aqueous solution is affected greatly by the presence of hydrolysis products in the organic

Table XXI. Extraction of Metal Nitrates by TBP.^a

Metal nitrate	Percentage extracted	
	0M HNO ₃	1M HNO ₃
LiNO ₃	-	-
NaNO ₃	-	-
KNO ₃	-	-
Cu(NO ₃) ₂	0.050	0.025
Mg(NO ₃) ₂	-	-
Ca(NO ₃) ₂	0.118	0.064
Zn(NO ₃) ₂	0.005	-
Al(NO ₃) ₃	0.004	0.003
Fe(NO ₃) ₃	0.010	0.008

^a After T. Sato, reference 406.

Aqueous phase - 5 g/l uranyl nitrate and 1M metal nitrate at initial nitric acid concentration indicated.
 Organic phase - 19% TBP diluted in kerosene.
 Equal phase volumes shaken together for 30 minutes at 20°C.

phase, eg., mono- and di-butyl phosphates. These products may be eliminated by washing or boiling the solvent with an alkaline solution. Two procedures for the removal of TBP impurities are given.

Procedure 1⁴²⁴: TBP is purified by boiling with a dilute caustic soda solution. Add 500 ml of 0.4% NaOH solution to 100 ml of impure TBP. Distill at atmospheric pressure until 200 ml of distillate have been collected. The remaining TBP is washed repeatedly with water. It may be dried by warming under vacuum.

Procedure 2⁴²¹: TBP is stirred with an equal volume of 6M HCl at 60°C for 12 hours. The separated TBP is cooled to room temperature and scrubbed with two equal-volume portions of water, three equal volume portions of 5% aqueous sodium carbonate solution, and three equal volume portions of water. The resultant TBP is dried by heating to 30°C under reduced pressure.

Aqueous chloride systems. Uranium is extracted from chloride solution as $UO_2Cl_2 \cdot 2TBP$ although higher uranyl chloride complexes may also be extracted.⁴²⁶⁻⁴²⁸ The partition of uranium between TBP and aqueous hydrochloric acid solution is shown in figures 38^{422, 429} and 39.⁴³⁰ The effect of uranium concentration on the distribution is given in figure 40;⁴²⁸ the effect of TBP concentration, in figure 41.^{426, 427} In Table XXII, the influence of salting-out agents on the extraction of uranyl chloride by 30% TBP in dibutyl ether is recorded.⁴³¹ The distribution of hydrochloric acid between TBP and aqueous solution is shown in figure 34.⁴¹² In figure 37, the partition coefficients of Pa, Th, Zr, and Sc are plotted as functions of aqueous HCl concentration.^{422, 429} In figure 39, the partition coefficients of Ni, Mn, Cu, Co, Zn, In, and Fe (III) are similarly plotted.⁴³⁰ Ishimori and Nakamura⁴¹⁷ have measured the partition coefficients of Hf, Th, Pa, U (VI), and Np (IV)(V)(VI) as functions of aqueous acid concentration. Gal and Ruvarec⁴³¹ have similarly

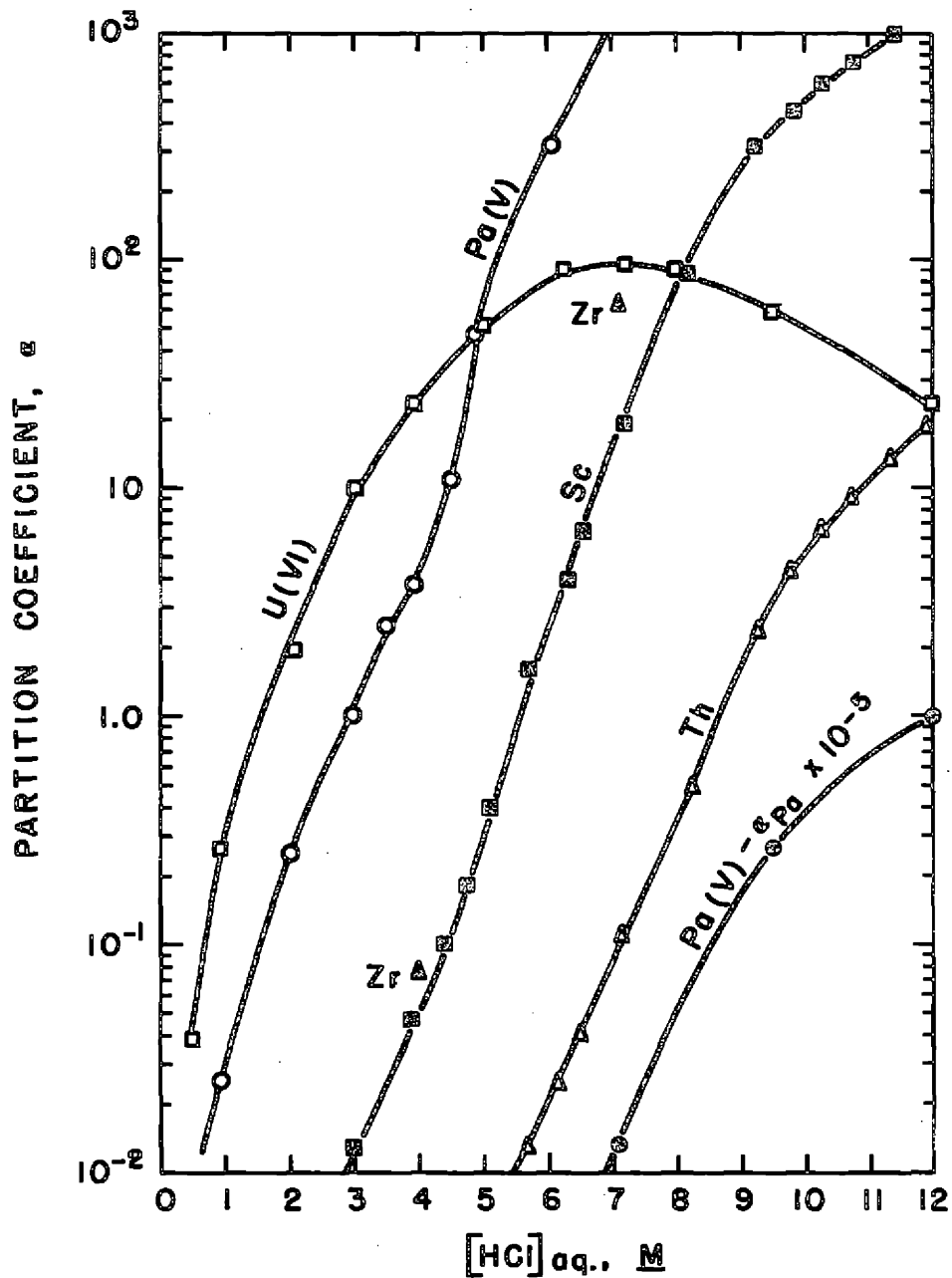


Figure 38. The extraction coefficient of U, Pa, Th, Zr, and Sc between pre-equilibrated 100% TBP and aqueous hydrochloric acid at $22^{\circ} \pm 2^{\circ}\text{C}$. After D. F. Peppard, G. W. Mason, and M. V. Gergel, reference 429, and D. F. Peppard, G. W. Mason, and J. L. Maier, reference 422.

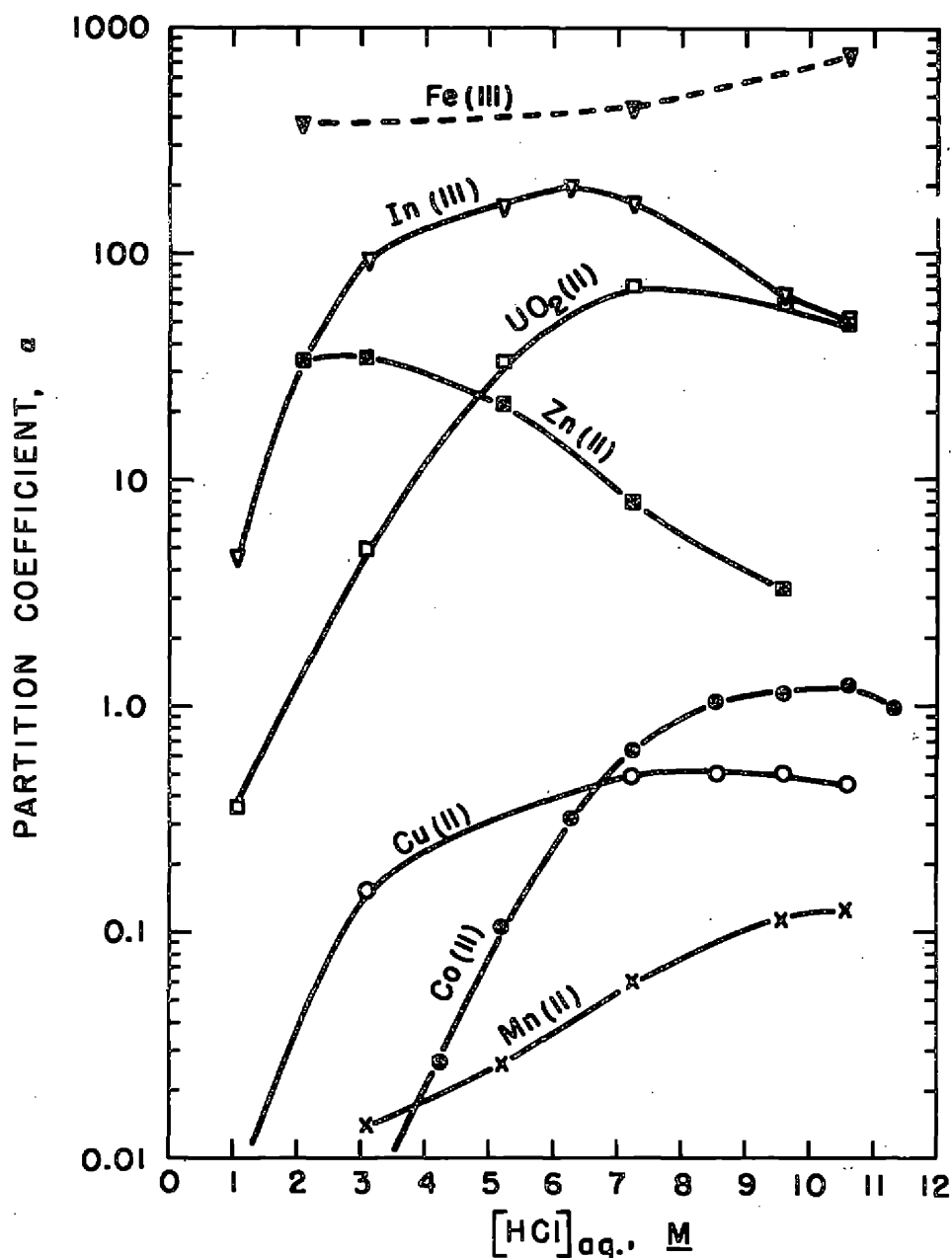


Figure 39. The extraction coefficients of U, In, Zn, Cu, Co, Fe, and Mn between pre-equilibrated 100% TBP and aqueous HCl solution at $21 \pm 0.1^\circ\text{C}$. After H. Irving and D. N. Edgington, reference 430. Conditions: TBP and HCl pre-equilibrated by stirring equal volumes together for about 10 minutes. Tracer concentrations or about 0.02M U and Cu used. Equal volumes of pre-equilibrated phases stirred together about 5 minutes.

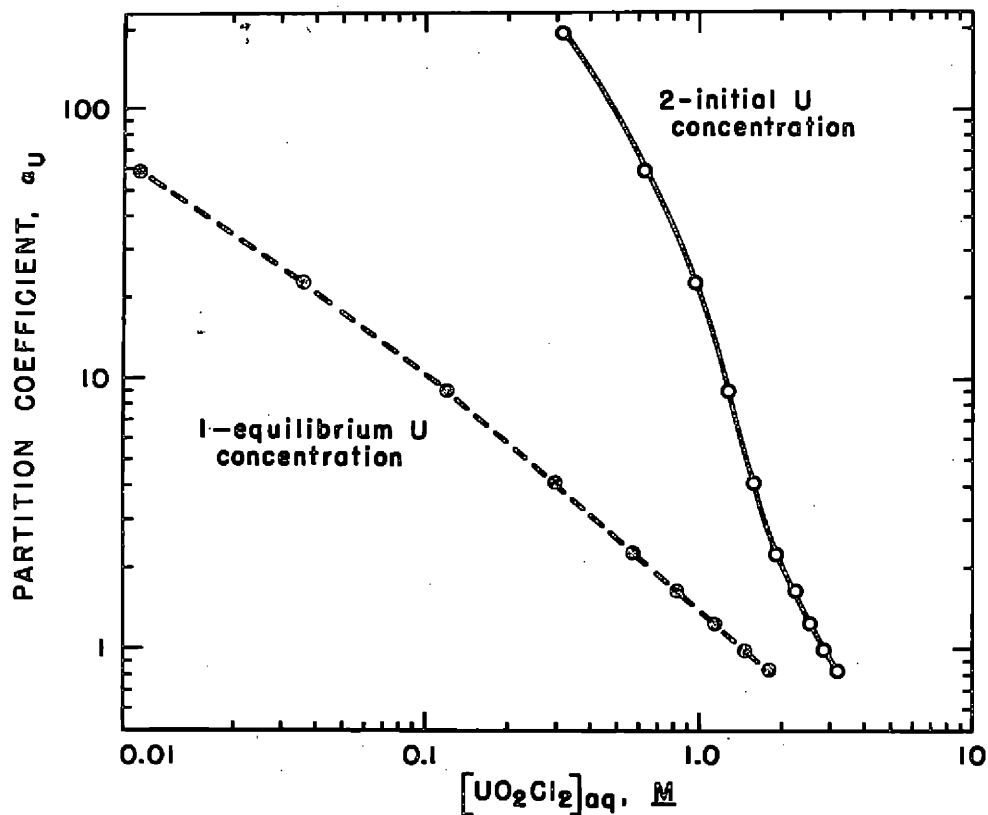


Figure 40.

The distribution of UO_2Cl_2 between 100% TBP and aqueous HCl solution as a function of uranium concentration of the aqueous phase. Curve 1 represents the partition with the equilibrium aqueous uranium concentration plotted as abscissa; curve 2, the partition with initial aqueous uranium concentration as abscissa.

After A. S. Kertes, and M. Halpern, reference 428.

Conditions:

Constant HCl acid concentration of 8.83M; equal phase volumes equilibrated for 15 minutes at room temperature, 18° - 22°C.

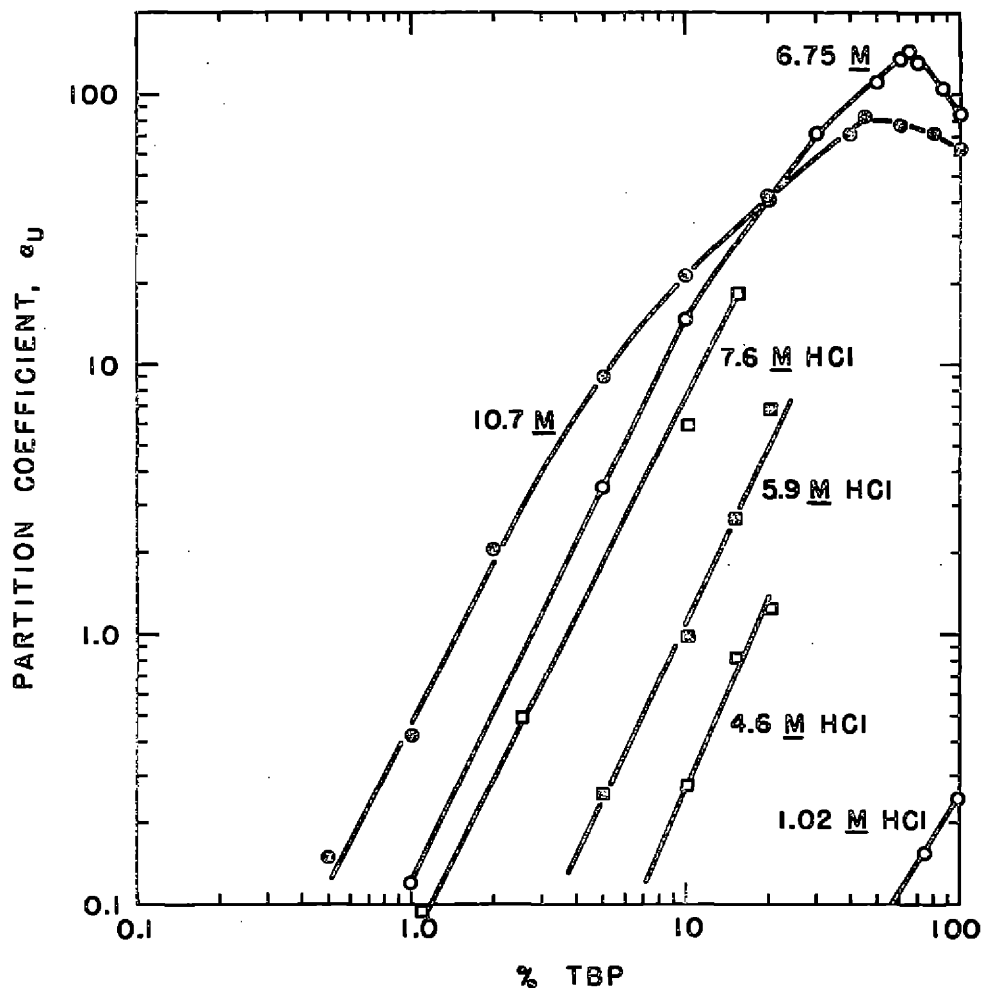


Figure 41. Partition coefficient of uranium as a function of TBP concentration for various initial aqueous HCl concentrations. 10.7M, 6.75M, and 1.02M HCl curves, after V. M. Vdovenko, A. A. Lipovskii, and S. A. Nikitina, reference 426. Conditions: TBP dissolved in benzene. Extraction made at room temperature using U²³³. 4.6M, 5.9M and 7.6M HCl curves, after V. B. Shevchenko, I. G. Slepchenko, V. S. Schmidt, and E. A. Nenarokomov, reference 427. Conditions: TBP dissolved in CCl₄. Equal phase volumes (10 ml) mixed together for 30 minutes and allowed to stand for 12-15 hours.

examined the partition coefficients of Fe (II)(III), U (VI), Cd (II), Sr (II), Zr (IV), Ce (III), Ru (IV), and V (V).

Aqueous perchlorate systems. The distribution of uranyl perchlorate between TBP and water at 25°C is shown in figure 42.¹⁶⁹ Figure 43 gives the partition coefficient of uranium as a function of the aqueous perchloric acid concentration.¹⁶⁹

Shevchenko *et al.*⁴³² have studied the extraction of uranyl perchlorate in the presence of HClO_4 , LiClO_4 , and NaClO_4 . The salting-out capacity of these salts increases in the order listed. The choice of TBP diluent also affects the extraction of uranyl perchlorate. From an aqueous solution of 0.065M HClO_4 and 1M NaClO_4 , the extraction of uranium by 2.20M TBP was found to decrease in the following order of diluents:⁴³²

isoamyl acetate > n-butyl acetate > isoamyl alcohol > toluene > xylene > benzene > carbon tetrachloride.

The distribution of perchloric acid between TBP and aqueous solution is given in figure 34.⁴¹² The partition coefficients of Th, Zn, Pm, Y, and Ce are plotted against aqueous perchloric acid concentration in figure 44.⁴³³ Ishimori and

Table XXIII. Effect of Salting-out Agents on the Extraction of Uranyl Chloride by TBP.^a

Salting-out agent	α_U
-	0.03
NaCl, sat'd	2.85
KCl, sat'd	0.38
NH_4Cl , 5M	0.71
LiCl , 5M	0.90
HCl , 5M	17.6
CaCl_2 , 2.5M	5.06
MgCl_2 , 2.5M	11.7
AlCl_3 , 1.67M	23.8

^a After Gal and Ruvarac, reference 431.

Initial composition of organic phase - 30% (v/v) TBP in dibutyl ether, sat'd with 1.225M HCl .
 Initial composition of aqueous phase - 1.225M HCl , 0.1M UO_2Cl_2 , salting-out agent at concentration indicated.

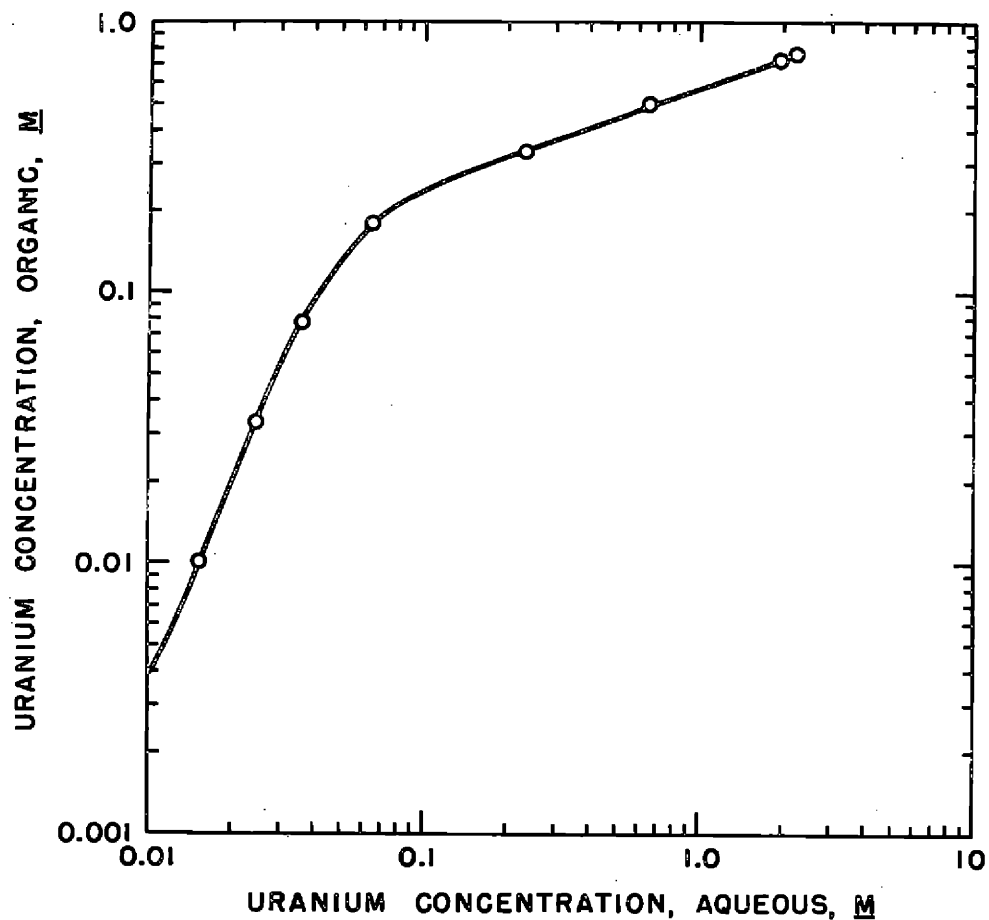


Figure 42.

The partition of uranyl perchlorate between 100% TBP and water at 25°C.

After E. Hesford and H. A. C. McKay, reference 169.

Nakamura⁴¹⁷ have studied the extraction of Pa, and Np (IV) (V) (VI) by TBP as a function of perchlorate concentration.

Aqueous sulfate systems. Sulfate ion is normally considered an interfering ion in the extraction of uranium from aqueous solution by TBP. Veereswararao,⁴³⁴ however, found that significant amounts of uranium may be extracted from sulfuric acid solution and that the extraction is increased as the acid concentration is increased (figure 45).⁴³⁴

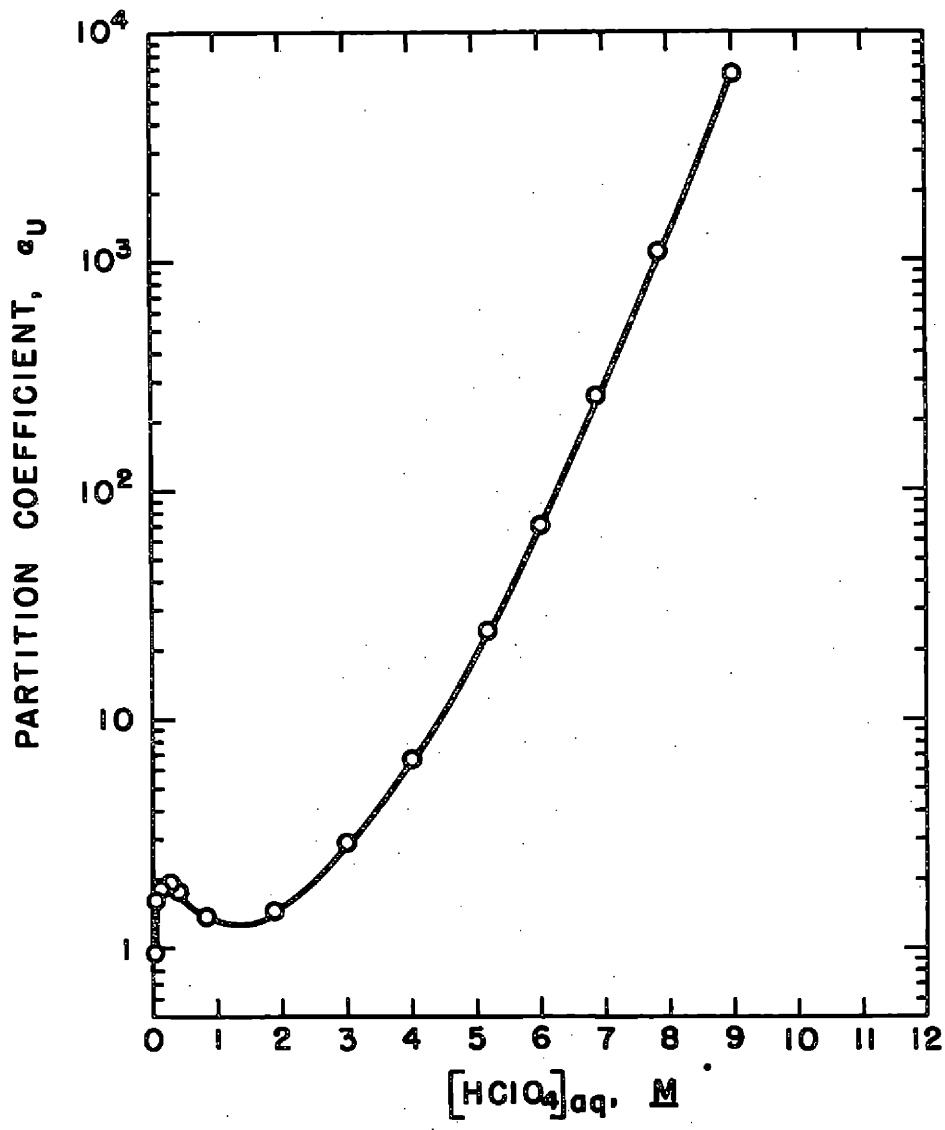


Figure 43.

The extraction coefficient of trace amounts of uranium between 100% TBP and aqueous solution as a function of equilibrium aqueous HClO₄ concentration. 25°C.

After E. Hesford and H. A. C. McKay, reference 169.

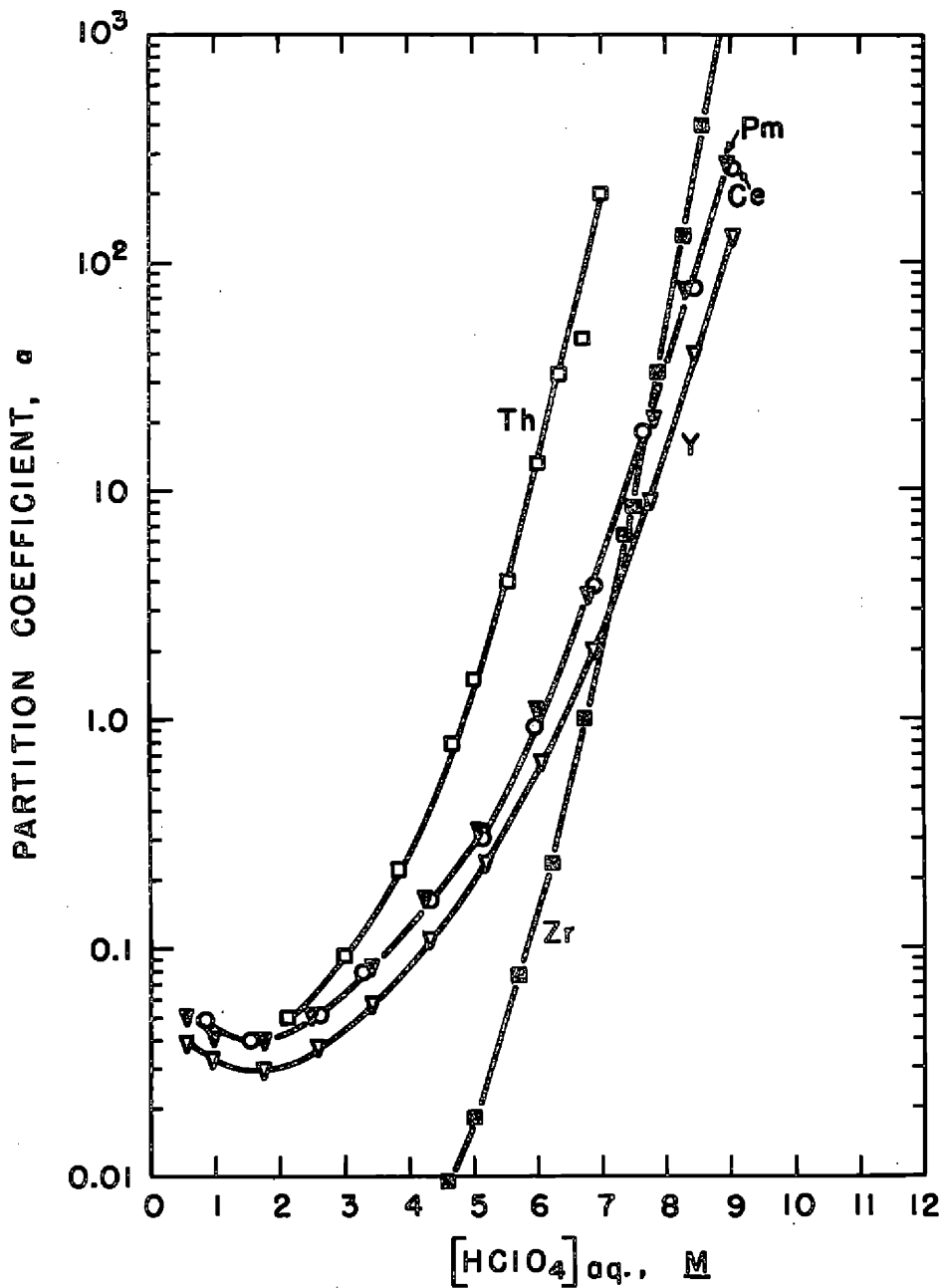


Figure 44. The distribution of Th, Pm, Ce, Zr, and Y between 100% TBP and aqueous solution as a function of initial equilibrium aqueous HClO_4 concentration. After S. Siekierski, reference 433. Conditions: Equal phase volumes (15 ml) shaken together for about 20 minutes at $21^\circ\text{-}25^\circ\text{C}$.

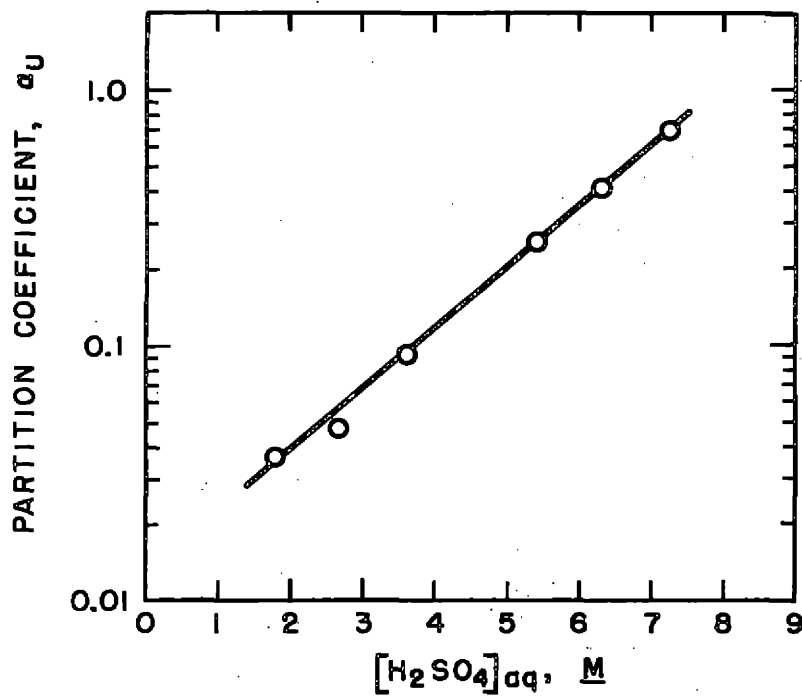


Figure 45.

The partition coefficient of uranium between 30% TBP (v/v) in kerosene and aqueous solution as a function of equilibrium aqueous H₂SO₄ concentration.

After U. Veereswararao, reference 434.

Conditions:

Equal phase volumes (10 ml) equilibrated by shaking for 5 minutes at 18° ± 1°C.

The presence of sodium chloride in sulfuric acid solution augments the extraction of uranium by TBP.⁴³⁴ Molybdenum and iron (III) are well extracted from such solutions. Vanadium and iron (II) are poorly extracted compared to uranium.⁴³⁴ The distribution of sulfuric acid between TBP and aqueous solution is represented in figure 34.⁴¹²

Aqueous thiocyanate solution. Petrow and Marenburg⁴³⁵ have studied the effect of thiocyanate ion on the extraction of uranium from sulfate liquors by TBP. From an aqueous solution containing 1.5 g/l of U₃O₈ as uranyl sulfate and

2.1 g/l total sulfate concentration at pH 1.5, the partition coefficient of uranium varied from 3.5 to 100 as the thiocyanate to uranyl molar ratio was increased from 2 to 6. Twenty per cent TBP dissolved in kerosene was used as extractant. The partition coefficient, α_u , increases with increased TBP concentration, increases with increased pH, and decreases with increased sulfate concentration. Vanadium and iron (III) are appreciably extracted by TBP from thiocyanate solutions. Copper, titanium, cobalt are weakly extracted. Iron (II), cadmium, molybdenum, magnesium and aluminum are essentially not extracted. Phosphate ion may cause the precipitation of uranium or complex formation when present in large amounts. Okada, et al.⁴³⁶ report the extraction of uranium by TBP, mesityl oxide, and methyl ethyl ketone from phosphoric acid solutions having 20 times as much ammonium thiocyanate as uranium.

Tri-n-octylphosphine oxide (TOPO)

Much of the work on this solvent has been reported by White and co-workers.⁴³⁷⁻⁴⁴⁴ Uranium is extracted by TOPO from nitrate and chloride solutions and to a smaller extent from sulfate and perchlorate solutions. It is essentially not extracted from phosphate solutions.⁴⁴¹ The extraction of U, Th, Bi, Mo, Zn, and Cr by 0.1M TOPO from aqueous solutions is given as a function of nitric acid concentration in figure 46; as a function of hydrochloric acid concentration in figure 47.^{440, 441} Iron and titanium extraction curves are also included in figure 47. The extraction of over 40 ions by 0.1M TOPO from hydrochloric, sulfuric, perchloric, and nitric acid solutions is qualitatively indicated in Table XXIII.⁴³⁸ The extraction of mineral acids by 0.1M TOPO as a function of acid concentration is given in figure 48.⁴⁴¹

Uranium may be stripped from TOPO solutions by contact

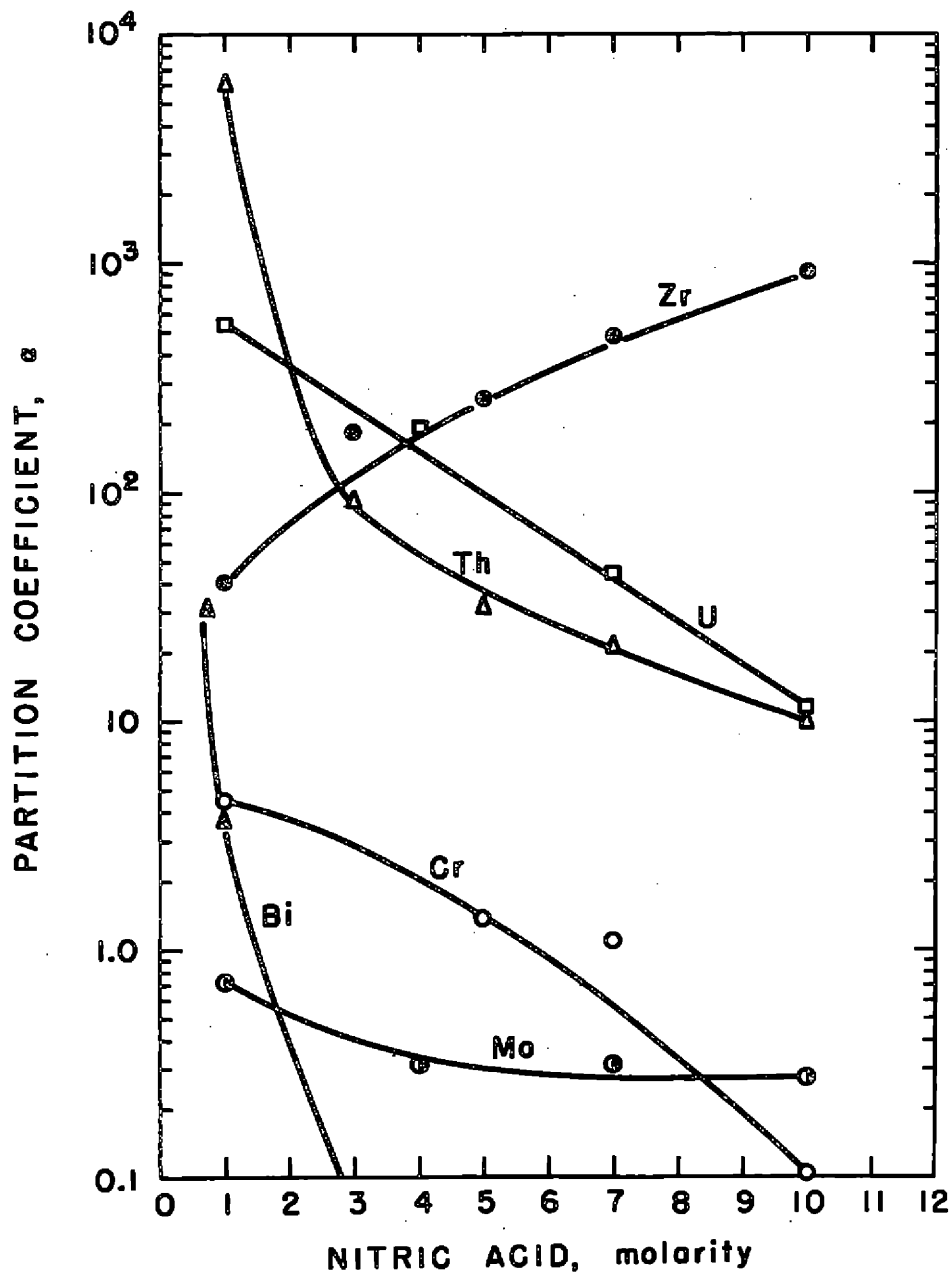


Figure 46.

Extraction of some metal ions by 0.1M tri-n-octylphosphine oxide from nitric acid solutions.

After J. C. White, references 440 and 441.

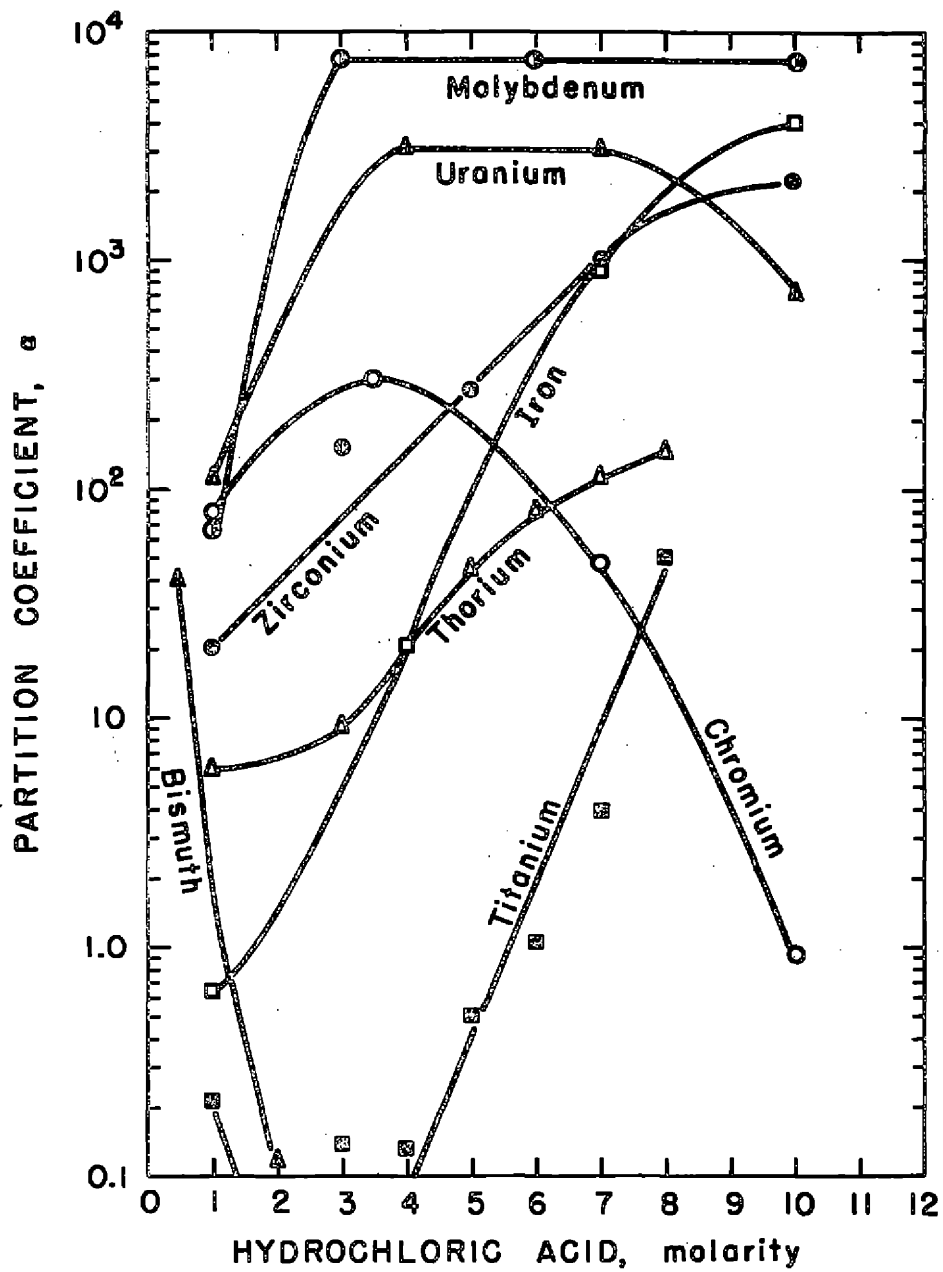


Figure 47.

Extraction of metal ions by 0.1M tri-n-octylphosphine oxide from hydrochloric acid solutions.

After J. C. White, references 440 and 441.

Table XXIII. Extraction of Ions from Acid Solutions with 0.1M TOPO in Cyclohexane.^a

Ion	HCl		H ₂ SO ₄		HClO ₄		HNO ₃	
	1M	7M	1M	11M	1M	7M	1M	7M
Al ⁺³	N	N	N	N	N	N	N	N
Sb ⁺³	P	E	E	E	P	E	P	P
As ⁺⁵	N	N	P	N	P	P	N	N
Ba ⁺²	N	N	-	-	N	N	N	N
Be ⁺²	N	N	N	N	P	N	N	N
Bi ⁺³	P	N	E	?	P	N	P	N
B ⁺³	N	N	-	-	N	N	N	N
Cd ⁺²	P	P	N	?	E	E	N	N
Ca ⁺²	N	N	-	-	N	N	N	N
Ce ⁺³	N	N	N	N	N	N	N	N
Cr ⁺⁶	E	E	E	E	E	E	E	E
Co ⁺²	N	N	N	N	N	N	N	N
Cu ⁺²	N	P	N	N	N	N	N	N
Dy ⁺³	N	N	N	N	N	N	N	N
Er ⁺³	N	N	N	N	N	N	N	N
Eu ⁺³	N	N	N	N	N	N	N	N
Gd ⁺³	N	N	N	N	N	N	N	N
Ga ⁺³	N	E	N	E	N	E	N	N
Ge ⁺⁴	?	P	?	P	?	P	?	?
Au ⁺	E	E	E	E	E	E	E	P
Hf ⁺⁴	E	E	E	?	E	E	P	P
Ho ⁺³	N	N	N	N	N	N	N	N
In ⁺³	P	P	E	E	P	N	N	N
Fe ⁺³	E	E	N	N	N	N	N	N
La ⁺³	N	N	N	N	N	N	N	N
Pb ⁺²	-	-	-	-	N	N	N	N
Mg ⁺²	N	N	-	-	N	N	N	N
Hg ⁺²	P	N	P	?	N	N	N	N
Mo ⁺⁶	E	E	E	E	E	E	N	P

Table XXIII.-Continued.

Ion	HCl		H ₂ SO ₄		HClO ₄		HNO ₃	
	1M	7M	1M	11M	1M	7M	1M	7M
Nd ⁺³	N	N	N	N	N	N	N	N
Ni ⁺²	N	N	N	N	N	N	N	N
Pd ⁺²	N	N	N	N	N	N	N	N
Pt ⁺²	P	P	P	N	N	N	N	N
Pr ⁺³	N	N	N	N	N	N	N	N
Ru ⁺²	N	N	N	?	N	?	N	N
Sm ⁺³	N	N	N	N	N	N	N	N
Ag ⁺	-	-	-	-	N	N	N	N
Sr ⁺²	N	N	-	-	N	N	N	N
Tb ⁺³	N	N	N	N	N	N	N	N
Th ⁺⁴	-	-	-	-	?	?	P	P
Tm ⁺³	N	N	N	N	N	N	N	N
Sn ⁺⁴	E	E	E	E	E	E	E	E
Ti ⁺⁴	N	E	P	E	P	E	P	P
U ⁺⁶	E	E	E	E	E	E	E	E
V ⁺⁴	N	E	N	N	N	N	N	N
Yb ⁺³	N	N	N	N	N	N	N	N
Y ⁺³	N	N	N	?	N	N	N	N
Zn ⁺²	P	P	N	N	N	N	N	N
Zr ⁺⁴	E	E	E	E	E	E	E	E

E = complete extraction P = partial extraction N = no extraction

^a After J. C. White, reference 438.

Equal phase volumes equilibrated 10 minutes.

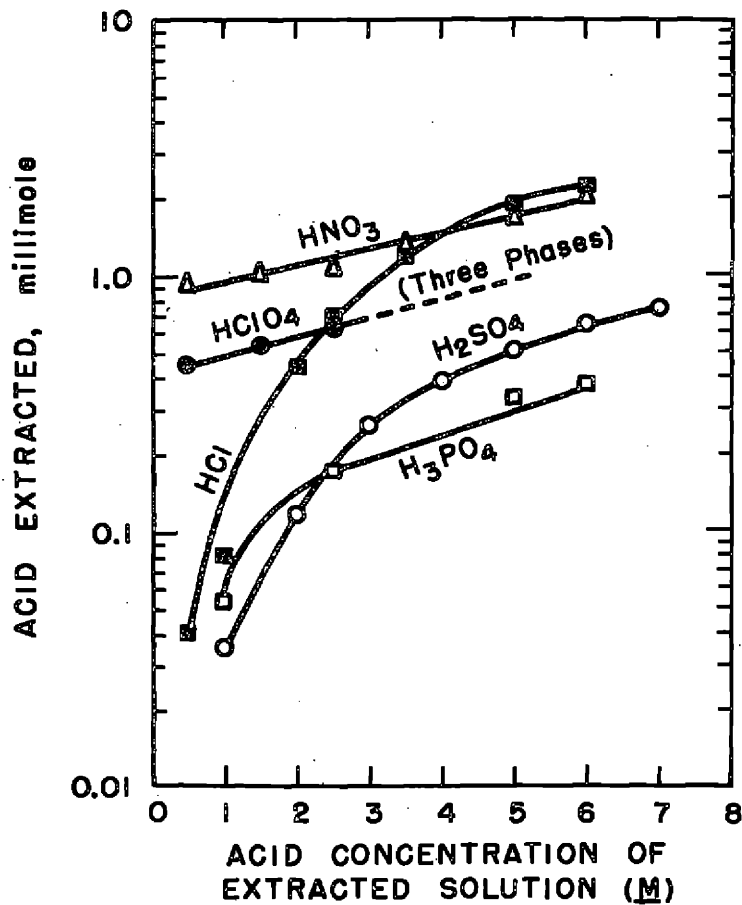


Figure 48.

Extraction of mineral acids by tri-n-octylphosphine oxide.

After J. C. White, reference 441.

Conditions:

Aqueous phases - acid solution of indicated molarity.
 Organic phase - 10 ml of 0.1M TOPO in cyclohexane.
 $V_o/V_a = 1$.

with acid (HF, H₃PO₄, or concentrated (NH₄)₂SO₄ solutions at pH 2), hydroxide (NaOH or NH₄OH), or carbonate [(NH₄)₂CO₃ or Na₂CO₃] solutions.⁴⁴⁵ Sodium carbonate is the most effective stripping agent.

Tetraphenylphosphonium chloride (TPPC)

A recent study has been reported in which uranium was extracted into chloroform as the tetraphenylphosphonium uranyl

tribenzoate complex.⁴⁴⁶ Uranyl ion was converted to an anionic form by benzoic acid. Tetraphenylphosphonium chloride was used as extractant. The extraction of uranium was found to depend upon pH, TPPC concentration, and uranium concentration. At pH 3-9, the extraction of uranium was nearly quantitative. The partition coefficient, α_u , was increased with increased TPPC concentration and was decreased with increased uranium concentration. The decrease in α_u with increased uranium concentration was observed with a constant uranium-to-TPPC molar ratio. At 25°C and pH 5.2, zinc, zirconium (niobium), and ruthenium were appreciably extracted (~10-20% compared to 100% for uranium). The extraction of zinc and zirconium may be depressed by the use of a complexing agent, EDTA, in solution.

ACIDIC ORGANOPHOSPHORUS COMPOUNDS

Uranium is efficiently extracted by acidic organophosphorus compounds which include di- and mono-alkylphosphoric acids, $(HO)(RO)_2P \rightarrow O$ and $(HO)_2(RO)P \rightarrow O$; dialkylphosphinic acids, $(HO)R_2P \rightarrow O$; alkylphosphonic acids, $(HO)_2RP \rightarrow O$; and dialkylpyrophosphoric acids, $H_2R_2P_2O_7$. The latter acids are discussed separately.

Table XXIV compares the extractive capacities of several dialkylphosphoric, dialkylphosphinic, and monoalkylphosphoric acids for uranium.³⁰² The ability to extract uranium, within a given class, appears to decrease with increased branching of the alkyl chain near the phosphate group. The acidity of the reagent decreases roughly in the same order.³⁰² Where comparisons can be made for the same alkyl group between classes of reagent, the extraction coefficient of uranium increases in the order

dialkylphosphoric acid < dialkylphosphinic acid < monoalkylphosphoric acid.

The choice of diluent affects the extraction of uranium. For

Table XXIV. Extraction of Uranium by Acidic Organophosphorus Reagents.^a

Reagent	Uranium extraction coefficient, α_u	
	Carbon tetrachloride	Kerosene
Dialkylphosphoric acids		
n-octyl	90	450
3,5,5-trimethylhexyl	40	260
2-ethylhexyl	17	135
2-ethyl-4-methylpentyl	-	90
2-propyl-4-methylpentyl	-	60
octyl-2	11	-
diisobutylmethyl	2	10
Dialkylphosphinic acids		
γ -phenylpropyl	300	-
phenyl-2-ethylhexyl	300	-
n-decyl	180	-
n-octyl	160	-
3,5,5-trimethylhexyl	120	-
2-ethylhexyl	30	-
Monoalkylphosphoric acids		
n-octyl	580	-
3,5,5-trimethylhexyl	>1000	(>1000)
2-ethylhexyl	>1000	(>1000)
diisobutylmethyl	450	-
2,6,8-trimethylnonyl-4	-	650
1-isobutyl-4-ethyloctyl	-	600
3,9-diethyltridecanol-6	-	550

^a After C. A. Blake, Jr., C. F. Baes, Jr., K. B. Brown, C. F. Coleman, J. C. White, reference 302.

Aqueous phase: 0.5M SO_4^{2-} , pH 1, 0.004M U(VI) initially.

Organic phase: 0.1M reagent in solvent indicated.

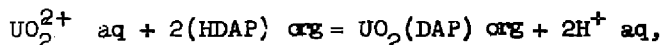
Temperature, 25°C; V_o/V_a , 1.

dialkylphosphoric and dialkylphosphinic acids, α_u generally increases as the dielectric constant of the solvent increases.³⁰²

For monoalkylphosphoric acids, a reverse trend is indicated.³⁰²

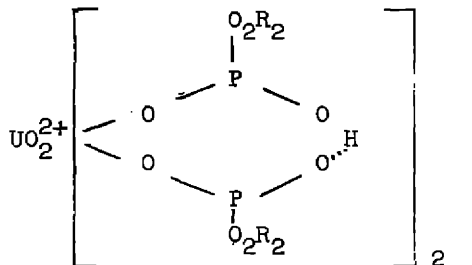
The mechanism of extraction of uranium by dialkylphosphoric acids has been studied by various groups.^{158,159,302,447}

At low uranium concentrations, the extraction mechanism appears to be consistent with the reaction



where HDAP represents a dialkylphosphoric acid. However, in

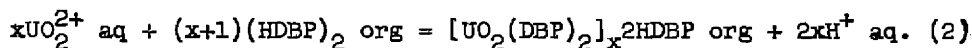
organic solvents, dialkylphosphoric acids are largely associated as dimers.^{158,159,448} On this basis, the reaction $UO_2^{2+} \text{ aq} + 2(\text{HDAP})_2 \text{ org} = UO_2(\text{DAP})_2(\text{HDAP})_2 \text{ org} + 2H^+$ (1) is indicated.^{158,159} The number of dialkylphosphate groups associated with the uranyl ion in equation (1) may be accounted for by a chelate structure ^{158,302}



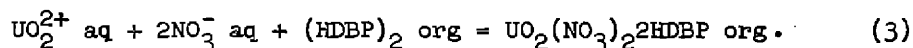
At higher uranium concentrations, isopiestic and viscosity measurements indicate that polymeric uranyl-dialkylphosphate chains are formed.¹⁵⁸

The extraction coefficient of uranium by dibutylphosphoric acid, HDBP, is given in figure 49 as a function of nitric acid concentration.¹⁸⁸ The shape of the curve has been explained by Healy and Kennedy in the following manner:¹⁸⁸

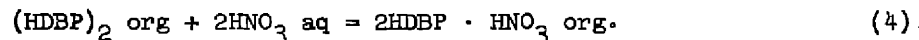
The initial decrease in α_u between 0.1M and 3M HNO_3 is expected on the basis of hydrogen ion replacement by UO_2^{2+} ion. However, for α_u greater than 10 not enough HDBP is present in the organic phase to give the monomeric species $UO_2(\text{DBP})_2(\text{HDBP})_2$ demanded by equation (1). In this region, the extraction mechanism is likely to be governed by the reaction



The shape of the extraction curve from 3M to 10M HNO_3 is similar to that obtained with TBP and indicates a change in extraction mechanism. The likely reaction is



The decrease in α_u above 7M HNO_3 is probably due to the competing reaction



It is likely that mechanisms (2) and/or (3) also occur to some extent at high acid concentrations.

The extraction mechanism of dialkylphosphinic acids is expected to be similar to that of dialkylphosphoric acids. The former are often found as dimers in organic solvents and the partition coefficient of uranium, α_u , exhibits a power dependence on extractant concentration at low uranium levels similar to that of dialkylphosphoric acids.³⁰² Monoalkylphosphoric and monoalkylphosphonic acids have been found in larger polymeric aggregates.^{302,448} Partition coefficients for these extractants exhibit first to second power dependencies on extractant concentration.³⁰²

Interference to uranium extraction by anions increases in the order $\text{ClO}_4^- < \text{Cl}^- < \text{SO}_4^{2-} < \text{PO}_4^{3-}$.

Stripping is essentially the inverse process of extraction. Uranium may be stripped from dialkylphosphoric acids by contact with hydrofluoric, sulfuric, phosphoric, or even hydrochloric acids. The stripping efficiency is generally increased with increased acid concentration.⁴⁴⁹ Ammonium or sodium carbonate stripping is efficient.⁴⁴⁹

Synergism.* In a search for reagents to modify kerosene as the diluent for dialkylphosphoric acids,** it was discovered that neutral organophosphorus compounds provided a synergistic enhancement of the uranium partition coefficient. The enhancement is increased in the following order of neutral reagent:

trialkylphosphate < alkyl dialkylphosphonate < dialkyl alkylphosphinate < trialkylphosphine oxide.

Table XXV lists α_u for several synergistic systems.³⁰² The reason for the enhanced partition coefficient, α_u , has been

* Co-operative action of discrete agencies such that the total effect is greater than the sum of the two effects taken independently.

**Kerosene is modified to prevent separation of a dialkylphosphate salt as a separate phase when alkaline stripping is used. Long chain alcohols have been used as chemical modifiers. These, however, depress the extraction coefficient of uranium and other metals.

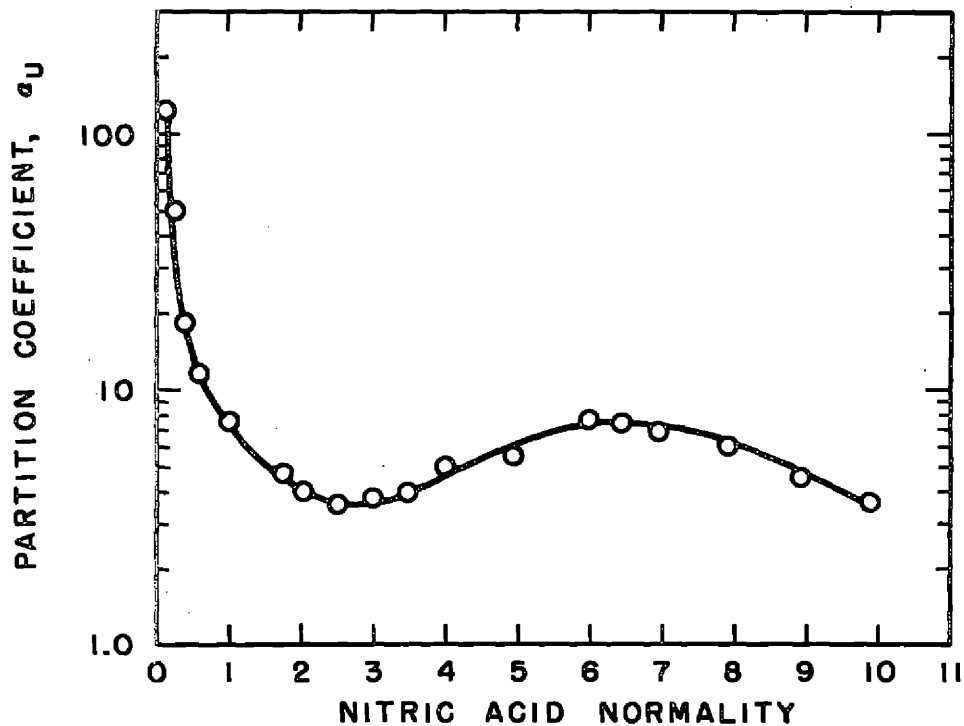


Figure 49.

Variation of α_U with nitric acid concentrations for 0.14M dibutylphosphoric acid in benzene using 20 ml organic phase, and 50 ml aqueous phase, and an initial uranium concentration of 0.018M.

After T. V. Healy and J. Kennedy, reference 188.

explained on the basis of (1) the addition of neutral reagent to the uranyl-dialkylphosphate complex through hydrogen bonding³⁰² or of (2) eliminating the need of monomerizing a mole of dimeric extractant⁴⁵⁰ in the extraction mechanism (see equation (1)). A recent study of the synergistic system, thenoyltrifluoroacetone-neutral organophosphorus compound, indicates that more investigation is necessary for a more precise explanation of synergistic effects.⁴⁵¹ Much of the work done on synergistic systems involving dialkylphosphoric acids is summarized in reference 452.

Table XXV. Synergistic Enhancement of Uranium Extraction Coefficient.^a

Organophosphorus reagent	Conc., M	α_u	
		Reagent Alone	In synergistic combination with 0.1M D2EHPA
Di(2-ethylhexyl) phosphoric acid (D2EHPA)	0.1	135	-
Phosphates			
tri-n-butyl	0.1	0.0002	470
tri-2-ethylhexyl	0.1	0.0002	270
Phosphonates			
di-n-butyl n-butyl	0.1	0.0002	1700
di-n-amyl n-amyl	0.1	0.0003	2000
di-n-hexyl n-hexyl	0.1	0.0004	2200
di-2-ethylhexyl 2-ethylhexyl	0.1	0.0002	870
Phosphinates			
n-butyl di-n-butyl	0.1	0.002	3500
n-butyl di-n-hexyl	0.1	0.002	3500
Phosphine oxides			
tri-n-butyl	0.05	0.0025	7000
tri-n-octyl	0.1	0.06	3500
tri-2-ethylhexyl	0.1	0.02	650

^a After G. A. Blake, Jr., C. F. Baes, Jr., K. B. Brown, C. F. Coleman, J. C. White, reference 302.

Aqueous phase: 0.5M SO_4^{2-} , pH 1, 0.004M U(VI) initially.
Organic phase: Reagents in kerosene diluent.
Temperature, 25°C.; V_o/V_a , 1.

Di(2-ethylhexyl) phosphoric acid (D2EHPA, HDEHP)

This reagent may also be known by a less descriptive name, dioctylphosphate (DOF). The extraction of uranium by D2EHPA is reviewed in reference 453. The effect of acid concentration on the extraction of uranium by D2EHPA is shown in figure 50.⁴⁵³ The uranium extraction curve for D2EHPA from nitric acid is similar in shape, for the few points given, to that for dibutylphosphoric acid given in figure 49. Figure 51 illustrates the effect of nitrate ion on the extraction of uranium by D2EHPA.⁴⁵³ The presence of a small amount of nitrate in an aqueous sulfate solution increases the extraction of uranium significantly. An increase in temperature causes a decrease in uranium extraction.⁴⁵³

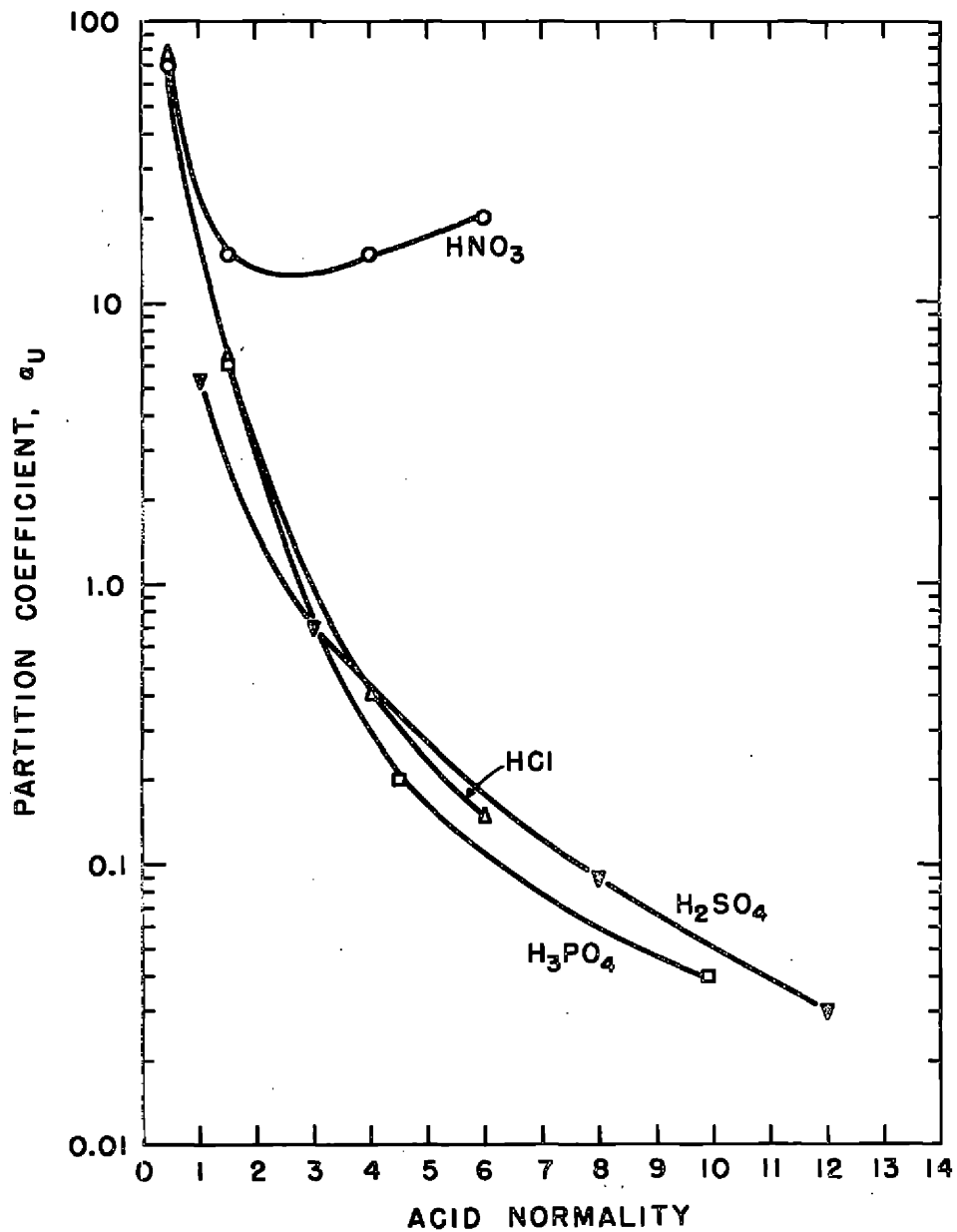


Figure 50. Extraction of uranium by di(2-ethylhexyl) phosphoric acid in kerosene from mineral acid solutions. After C. A. Blake, K. B. Brown, and C. F. Coleman, reference 453. Conditions: Organic phase - 0.1M D2EHPA in kerosene, 2% (w/v) 2-ethylhexanol. Aqueous phase - 1 gU/l for all acid solutions except H₃PO₄ in which case the U concentration was 100 ppm. Agitation time - 2 minutes. $V_O/V_A = 1$ for all acids but HNO₃ in which $V_O/V_A = 2$.

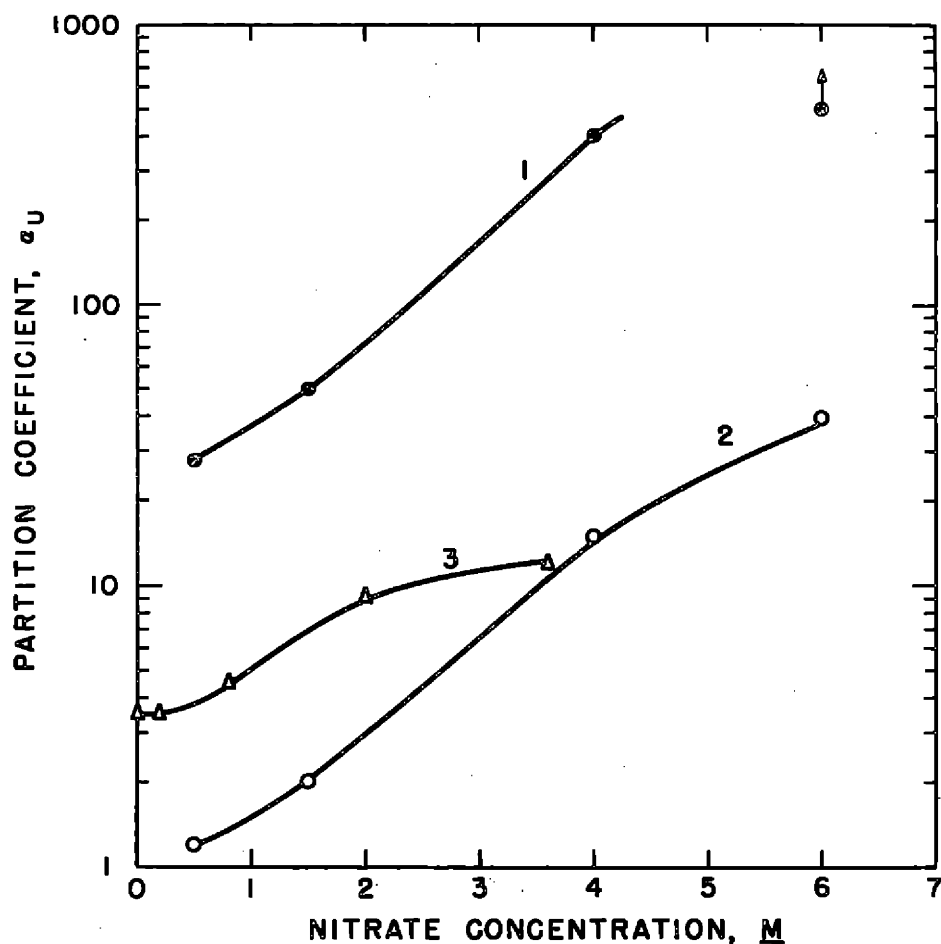


Figure 51. The effect of nitrate ion on the extraction of uranium by di(2-ethylhexyl) phosphoric acid. Curve 1 - initial pH = 1.5 - 1.85; Curve 2 - initial pH = 0.5 - 0.75. Conditions: 0.01M D2EHPA in kerosene (1.3% 2-ethylhexanol), 1 g U/l in aqueous phase, $V_o/V_a = 2$, 2 min. contact time. Curve 3 - 0.5M SO_4 , pH = 1.2. Conditions: 0.05M D2EHPA in CCl_4 , 1 g U/l in aqueous phase, $V_o/V_a = 1$, 20 min. contact time. After C. A. Blake, K. B. Brown, and C. F. Coleman, reference 453.

The effect of diluent on α_U is given in Table XXVI.⁴⁵³ The enhanced extraction of uranium by D2HFA in synergistic combination with neutral organophosphorus reagents has already been noted (Table XXV).³⁰² The extent to which other ions are extracted is indicated qualitatively in Tables XXVII and XXVIII.⁴⁵⁴

Table XXVI. Choice of Diluent with Di(2-ethylhexyl)
Phosphoric Acid.^a

<u>Diluent</u>	<u>α_u</u>
Kerosene	135
Hexane	110
Carbon tetrachloride	20
Isopropyl ether	17
Benzene	13
Chloroform	3
2-Ethylhexanol	0.1
Octanol-2 (capryl alcohol)	0.08

^a After C. A. Blake, K. B. Brown, and C. F. Coleman, reference 453.

0.1M D2EHPA, 0.004M U (VI), 0.5M SO_4^{2-} , pH = 1.1,
 $V_a \sqrt{V} = 1$, agitation time = 10 min. (wrist-action shaker).

Dialkylpyrophosphoric acids

Dialkylpyrophosphoric acids are used in the recovery of uranium from low-grade phosphate ores. Much of the work that has been reported in project literature has been summarized by Ellis,⁴⁵⁵ by Long, Ellis, and Bailes,⁴⁵⁶ and by Brown and Coleman.³⁰⁵ The acids are prepared just prior to use by adding alcohol to a slurry of P_2O_5 in kerosene with stirring and cooling. A concentration of about 0.1 g P_2O_5 per ml of kerosene is optimum.⁴⁵⁵ A 2:1 mole ratio of alcohol: P_2O_5 is used to form the dialkylpyrophosphoric acid. A 3:1 mole ratio should give about equal mole quantities of mono- and di-alkyl orthophosphoric acids.⁴⁵⁵ The reactions are complex and mixtures of various phosphoric acids are formed. With pyrophosphoric acids, uranium extraction increases with carbon chain length from butyl to octyl.⁴⁵⁵ Nonyl and

Table XXVII. Extraction of Metal Ions from Acidic Solutions with
0.1M D1(2-ethylhexyl) Phosphoric Acid in Cyclohexane.^a

Metal Ion	Sodium Chloride (1M)			Ammonium Sulfate (1M)	Sodium Nitrate (1M)
	pH 0	pH 0.5	pH 1.5	pH 1.5	pH 1.5
Al ⁺³	-	-	N	N	N
As ⁺⁵	-	-	N	N	N
Ba ⁺²	-	-	N	-	N
Be ⁺²	P	P	P	P	P
B ⁺³	-	-	N	-	N
Cd ⁺²	-	-	N	N	N
Ca ⁺²	-	-	N	-	N
Cr ⁺⁶	P	P	P	P	P
Co ⁺²	-	-	N	N	N
Cu ⁺²	-	-	N	N	N
Ga ⁺³	-	-	N	N	N
Au ⁺	-	-	N	N	N
Hf ⁺⁴	E	E	E	E	E
In ⁺³	N	N	P	P	P
Fe ⁺³	P	P	P	P	P
La ⁺³	N	N	N	N	N
Mg ⁺²	-	-	N	-	N
Hg ⁺²	-	-	N	N	N
Mo ⁺⁶	E	E	E	E	E
Ni ⁺²	-	-	N	N	N
Pd ⁺²	-	-	N	N	N
Pt ⁺²	-	-	N	N	N
Ru ⁺²	-	-	N	N	N
Sr ⁺²	-	-	N	-	N
Th ⁺⁴	E	E	E	E	E
Sn ⁺⁴	E	E	P	N	E
Ti ⁺⁴	E	E	E	E	E
U ⁺⁶	E	E	E	E	E
Y ⁺³	N	E	E	E	E
V ⁺⁴	N	N	P	E	P
Zn ⁺²	N	N	P	N	P
Zr ⁺⁴	E	E	E	E	E

E = complete extraction, P = partial extraction, N = no extraction,
- = no test was conducted.

^a After W. J. Ross and J. G. White, reference 454.

Aqueous phase: 1-2 mg of ion, salt at concentration indicated,
pH indicated.
Organic phase: 0.1M D2HPA in cyclohexane. 5 ml portions of each
phase shaken together for one hour.

Table XXVIII. Extraction of Rare Earths from Chloride Solutions with Di(2-ethylhexyl) Phosphoric Acid in Cyclohexane.^a

Ion	pH 1.9	pH 1.0	pH 1.0 ^b	pH 0.5	pH 0.5 ^b	pH 0
Y ⁺³	E	E	E	E	P	N
La ⁺³	N	N	N	N	N	N
Ce ⁺³	P	N	N	N	N	N
Pr ⁺³	P	N	N	N	N	N
Nd ⁺³	E	N	P	N	N	N
Sm ⁺³	E	N	P	N	N	N
Eu ⁺³	E	E	E	P	N	N
Gd ⁺³	E	E	E	P	N	N
Tb ⁺³	E	E	E	E	E	N
Dy ⁺³	E	E	E	E	E	N
Ho ⁺³	E	E	E	E	E	N
Er ⁺³	E	E	E	E	E	N
Tm ⁺³	E	E	E	E	E	P
Yb ⁺³	E	E	E	E	E	P

E = complete extraction, P = partial extraction, N = no extraction.

^a After W. J. Ross and J. C. White, reference 454.

Aqueous phase: standard solution of 2 mg/ml Ce⁺³; 1 mg/ml Pr⁺³, Nd⁺³, Sm⁺³, Dy⁺³, Ho⁺³; 0.5 mg/ml Tb⁺³; 0.2 mg/ml Y⁺³, La⁺³, Eu⁺³, Gd⁺³, Er⁺³; 0.1 mg/ml Yb⁺³;

1 ml of standard solution, 1 ml 5M NaCl, NaOH or HCl to give desired pH in 5 ml of solution.

Organic phase: 5 ml of 0.1M D2EHPA in cyclohexane extraction for 1 hour.

^b Without NaCl.

decyl give about the same extraction as octyl.⁴⁵⁵ No appreciable difference in extracting ability was observed between pyrophosphoric acids prepared with octanol-1 or octanol-2.⁴⁵⁵ Most of the studies have been made with octylpyrophosphoric acid (OPPA). Pyrophosphoric acids deteriorate fairly rapidly with time at room temperature. At elevated temperatures, the rate of deterioration is even greater. Contact with mineral acid causes pyrophosphoric acids to hydrolyze to orthophosphoric acids. The rate of hydrolysis is slower with basic solutions.⁴⁵⁵

Kerosene is a satisfactory diluent for OPPA.⁴⁵⁵ The acid is used in 1-10% concentration.

The partition coefficient of uranium, α_u , is considerably higher with OPPA than with the corresponding mixture of orthophosphoric acids, OPA. The partition coefficient is a function of the oxidation potential of the acid. With OPPA, satisfactory uranium recovery can be made if the e.m.f. is -0.250 volts or greater.⁴⁵⁵ Reduction of the acid increases the extraction of uranium considerably. At zero to +100 volts α_u is about twenty times that at -300 to -200 volts.⁴⁵⁶ The extraction of iron is decreased in reduced solution, i.e. $\alpha_{\text{Fe(II)}} < \alpha_{\text{Fe(III)}}$.

Uranium is stripped from the organic solvent by precipitation as uranous fluoride.

Several papers have recently appeared in open literature publications concerning the extraction of uranium by pyrophosphoric acids. Zangen⁴⁵⁷ has shown that OPPA prepared by the alcoholysis of P_2O_5 is a mixture of several components. OPPA prepared in this manner was found to be more effective in the extraction of uranium than pure dioctylphosphoric acid by two orders of magnitude. The pure acid was prepared by synthesis, starting from POCl_3 .

In an effort to determine the uranium species extracted by OPPA, Grdenic and Korpar⁴⁵⁸ have isolated the species $\text{U}(\text{Oct}_2\text{P}_2\text{O}_7)_2$. The species, however, was insoluble in ligroin, the OPPA diluent. It was soluble in ligroin containing OPPA in a ratio of one mole of U(IV)-salt and 2 moles of OPPA. This indicates $\text{U}(\text{Oct}_2\text{HP}_2\text{O}_7)_4$ is the extractable species. The same formula was obtained by determination of the uranium content in a saturated ligroin phase.

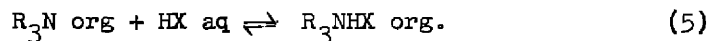
Habashi⁴⁵⁹ has investigated the extraction of uranium and other metals by OPPA from phosphoric acid solutions. Uranium (VI) was found to be more highly extracted than uranium

(IV). This is surprising in view of the increased extraction of uranium from reduced acid solution mentioned previously. Also, cerium (III) was found to extract more readily than cerium (IV). The partition coefficients of several metal ions are given for various phosphoric acid concentrations in Table XXIX. The partition coefficients decrease with H_3PO_4 concentration for all the metal ions tested except cerium. The extraction coefficients of both cerium (IV) and (III) pass through maxima in the region of $4M H_3PO_4$. The partition coefficient of uranium is decreased by increased initial uranium concentration. The addition of Na_3PO_4 to the solution causes α_u to increase greatly--apparently by decreasing the hydrogen ion concentration in the aqueous phase. Fluoride ion interferes most seriously with the extraction of uranium by OPPA.

Zangen⁴⁶⁰ has studied the extraction of uranium (IV) from phosphoric acid by di(2-butyloctyl) pyrophosphate, BOPPA.

AMINES AND QUATERNARY AMMONIUM SALTS. A large number of amines, quaternary ammonium salts, and other organonitrogen compounds have been investigated as possible extractants of uranium.⁴⁶¹⁻⁴⁶³ The physical chemistry of uranium extraction by amines has been studied by McDowell, Baes, and Allen⁴⁶⁴⁻⁴⁶⁸ and Boire^{469,470} Much of the above work has been summarized by Coleman, et al.³⁰⁴ More recently, Moore⁴⁷¹ has reviewed the extraction of a large number of elements, including uranium, by amines.

The reactions involved in the extraction of uranium by amines have been reviewed by Coleman, et al.³⁰⁴ Organic solutions of amines extract acids from aqueous solution to form alkylammonium salts



The amine salt in the organic phase can exchange its ion for another in the aqueous phase

Table XXIX. Partition Coefficients of Several Metal Ions
Between OPPA and H_3PO_4 .^a

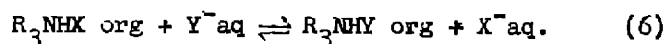
Ion ^b	Partition coefficient			
	H_3PO_4 concentration			
	2M	4M	6M	8M
U(VI)	190	46	23	20
U(IV)	18.6	14.2	13.5	
Th(IV)	24	23	18	13
Fe(III)	18	8	5	
Fe(II)	1	< 1	< 1	
V(IV) ^c	2	0.8	0.1	
Ce(IV)	5	8	3	1
Ce(III)	7	22	4	2

^a After F. Habashi, reference 459.

The coefficients, with the exception of $\alpha_{u(IV)}$, were determined from figures which appear in reference 459.

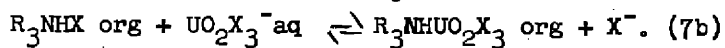
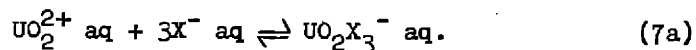
^b 0.4 mg metal ion per ml; 2% OPPA in n-hexane; $V_o/V_a = 0.1$.

^c $V_o/V_a = 1$.

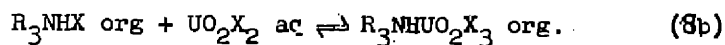
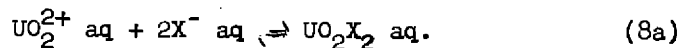


The order of preferences for anions in the organic amine solution is $ClO_4^- > NO_3^- > Cl^- > HSO_4^- > F^-$.³⁰⁴

In this anion exchange representation, metals are then extracted from aqueous solutions in which they are present as anions or anionic complexes. For example,



This mechanism, however, is indistinguishable from one in which a neutral complex is extracted.



The factors that influence uranium extraction have been studied most extensively for amine-sulfate systems.³⁰⁴ The effect of amine structure on the extraction of uranium and other metal ions is illustrated in Table XXX.³⁰⁴ Uranium (IV) is efficiently extracted by primary amines. The efficiency decreases with secondary and tertiary amines. With uranium (VI) there does not seem to be much correlation between α_u and amine class. With primary, secondary, and tertiary laurylamines, $\alpha_{u(VI)}$, under the conditions given in Table XXX, is < 0.1 , 80, and 140, respectively.³⁰⁴ With primary laurylamine an emulsion is formed.³⁰⁴ The extraction of uranium is also affected by carbon chain branching near the nitrogen atom in tertiary amines (Table XXX). Certain *n*-benzyl-branched-alkyl secondary amines have been found to extract uranium extremely well.³⁰⁴ The uranium (VI) partition coefficients of *N*-benzylheptadecylamine, *N*-benzyltetradecylamine, *N*-benzyl-dodecylamine, and *N*-(2-naphthylmethyl) dodecylamine, under the conditions outlined in Table XXX, are 2000, > 1000 , > 1000 , ~ 1000 , respectively.³⁰⁴ The partition coefficient depends upon the amine-diluent combination. The effect of diluent on $\alpha_{u(VI)}$ is indicated in Table XXX.³⁰⁴

The partition coefficient, $\alpha_{u(VI)}$, is influenced by uranium concentration in that it changes the amount of free amine sulfate concentration.³⁰⁴ In sulfate solution, bisulfate complexes the amine more strongly than does sulfate. The uranium partition coefficient, therefore, decreases with increased acidity.³⁰⁴ Excess aqueous sulfate causes a decrease in $\alpha_{u(VI)}$.³⁰⁴ The partition coefficient is also decreased by increased temperature.³⁰⁴ Extraction isotherms indicate that four to six amine molecules are associated with each uranium (VI) ion.³⁰⁴ The number depends upon the particular amine used. With vigorous shaking, the partition coefficient, α_u , varies approximately as the first power of the free amine

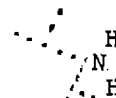
Table XXX. Effect of Amine Structure on the Extraction of Metal Sulfates.^a

Metal ion	Partition coefficient, α							
	Primary Amines		Secondary Amines		Tertiary Amines			
	Amine, 21F81 ^b	Primene JM-7 ^c	Dilauryl- amine	Di(tri- decyl) amine	Amine S-24 ^d	Methyl- di-n- decyl- amine	Tri-iso- octylamine	Tris(2- ethylhexyl) amine
Mg, Ca, Al V(IV), Cr(III), Mn(II), Fe(II), Co(II), Ni(II), Cu(II), Zn	<0.1		<0.01	<0.01	<0.01	<0.01	<0.01	
V(III)		4 ^{f,g}		0.1 ^{f,g}				
Fe(III)	40		<0.5	<0.5	<0.1		<0.01	
R.E.(III)	20			<0.1	<0.01		<0.01	
Ce(IV)	>50	15 ^{f,g}	<0.1		<0.01		<0.01	
Tl(IV)	10		5		0.2		<0.1	
Zr	>1000 ^f			350			200	<1
Th(0.5M SO ₄)	>5000 ^f	>5000 ^{f,g}		>500 ^f	2 ^f		<0.1 ^f	
U(IV)	3000	>5000 ^{f,g}	200		50 ^{f,g}		<1 ^f	
V(V)			<1	<1	<1		<1	
V(V)(pH2) ^e	-20		-20		-20	-20	-20	
Mo(VI) ^e	150		200		400	400	150	3
Mo(VI)(pH2) ^e	>1000		>1000		>1000	>1000	>1000	
U(VI)	40	10	80	120	20	50	90	0.2
U(VI)	25 ^g	3 ^g		80 ^h	110 ^g			
U(VI)	50 ^h	90 ^h		40 ^h	2 ^h			

^a After C. F. Coleman, K. B. Brown, J. G. Moore, K. A. Allen, reference 304.

^b 1M SO₄, pH 1, ~1 g metal ion per liter except as noted. V₀/V_a = 1, 0.1M amine in aromatic hydrocarbon diluent.

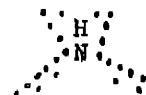
b 1-(3-ethylpentyl)-4-ethyloctylamine



c trialkylmethylamine, homologous mixture, 18-24 carbon atoms



d bis-(1-isobutyl-3,5-dimethylhexyl) amine



e Coefficients at loadings of ~5 g V or ~3 g Mo per liter of extractant. Extraction coefficients of these metals decrease as their concentration decreases.

f Extraction from 0.5M SO₄ solution.

g Diluent kerosene instead of aromatic hydrocarbon.

h Diluent chloroform instead of aromatic hydrocarbon.

sulfate concentration,

$$\alpha_{U(VI)} = k [M (\Sigma \text{ amine}) - nM (U (VI)) \text{ org}]^n,$$

where n has a value between 4 and 6, characteristic of the amine.³⁰⁴ With slow equilibration, in which the liquid-liquid interfacial area is strongly limited and interfacial turbulence is prevented, nearly theoretically ideal results have been obtained;⁴⁶⁸ i.e.,

$$c_{U(VI)} = k [M (\Sigma \text{ amine}) - nM (U (VI)) \text{ org}]^n.$$

Small amounts of foreign anions added to sulfate solutions hinder the extraction of uranium more than similar amounts of added sulfate. The order of increasing interference is $SO_4 < PO_4 < Cl < F < NO_3$.³⁰⁴

Effective separations between uranium and other metal ions may be made by choice of amine and/or diluent (Table XXX).³⁰⁴ Modification of the diluent with long-chain alcohols or other modifiers affects the extractive powers of the organic solvent phase. A possible synergistic enhancement of $\alpha_{U(VI)}$ has been found with 3,9-diethyltridecyl-6 amine and di(2-ethylhexyl) phosphoric acid.⁴⁵²

The amine extraction of uranium (VI) from aqueous phosphate or fluoride solutions is qualitatively similar to that from sulfate solution. Uranium is extracted from relatively low anion concentrations. As the latter concentration is increased, α_U is decreased.³⁰⁴ The opposite is true for chloride or nitrate solutions. Uranium extraction is increased as the concentration of either of the latter two anions is increased.³⁰⁴

Uranium may be stripped from the amine solvent phase by a number of methods. Uranium extracted as the amine-sulfate complex may be stripped by contact with a nitrate or chloride solution. Alkaline stripping with sodium carbonate results in an aqueous uranyl tricarbonate solution. Ammonium or sodium hydroxide forms precipitates that are difficult to handle. A slurry of magnesium oxide causes uranium to precipitate

as a magnesium polyuranate.³⁰⁵

Tri-n-octylamine (TnOA)

The partition coefficients obtained by Keder, *et al.*⁴⁷²,⁴⁷³ for the extraction of actinide metals from nitric acid solutions by 10 volume percent TnOA in xylene are given in figure 52. Carswell³⁵⁹ has studied the extraction of uranium and thorium by 0.2M TnOA in toluene, also, from nitric acid solution. Thorium appears to be more strongly extracted than uranium in the latter system. Uranium, however, is extracted practically to the same extent in both systems for acid concentrations up to 6M.

The extraction of uranium from hydrochloric acid solutions by TnOA in CCl₄ has been studied by Bizot and Tremillon.⁴⁷⁴ The extraction curve as a function of HCl concentration is similar in shape and magnitude to that for triisooctylamine plotted in figure 53.

Allen and co-workers⁴⁶⁵⁻⁴⁶⁸ have made fundamental studies on the extraction of uranium from sulfate solution by TnOA.

Extraction of uranium from acetic acid solution by TnOA in Amsco D-95 appears to be intermediate between extraction from sulfuric and phosphoric acids on one side and hydrochloric and nitric acids on the other.³⁰⁴

Triisooctylamine (TiOA)

The results of Moore⁴⁷⁵ for the extraction of uranium (VI), thorium, and fission products from hydrochloric acid solution by 5% TiOA in xylene are presented in figure 53. The extraction of strontium-85 is negligible from 2-11M HCl. Americium (III) and curium (III) are not extracted. Elements which are extracted include Fe(III), Co(II), Zn(II), Hf(IV), V(V), Pa(V), Cr(VI), Mo(VI), U(IV), Np(VI,V,IV), and Pu(VI,IV) in addition to those shown in figure 53. The extraction of iron, vanadium, and chromium may be suppressed by reduction to a lower oxidation state. Ruthenium remains in the organic

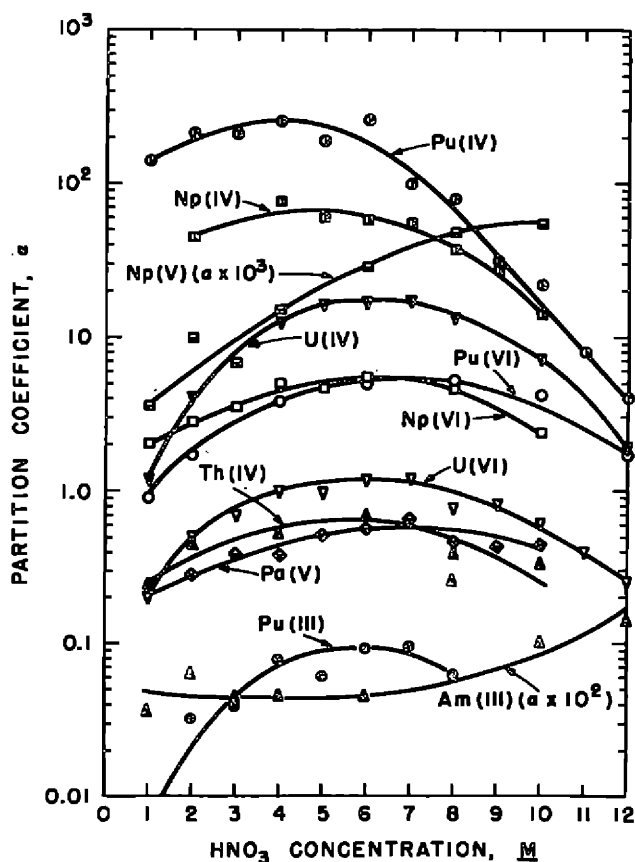


Figure 52. The extraction of actinide ions by ten volume percent tri-n-octylamine in xylene from aqueous nitric acid solution. After (1) W. E. Keder, J. C. Sheppard, and A. S. Wilson, reference 472 and (2) A. S. Wilson and W. E. Keder, reference 473.

Conditions:

- (1) Ten volume percent TnOA in xylene were stirred with an equal volume of nitric acid of the desired composition for 3-5 minutes at room temperature ($\sim 25^{\circ}\text{C}$). Phases were separated by centrifugation after contacting.
- (2) Uranium (IV) data only. Aqueous solutions were prepared at each nitric acid concentration by dilution of a stock solution which was $\sim 1\text{M}$ U(IV), 0.1M HClO_4 , and $\sim 1\text{M}$ Zn(II). Solutions for extraction experiments were 0.015M U(IV). TnOA was contacted by an equal volume of 12M HNO_3 followed by three contacts of one volume each of the nitric acid concentration used. Equal volumes of aqueous and amine solutions were contacted at room temperature for 5 minutes. Phases were separated by centrifugation.

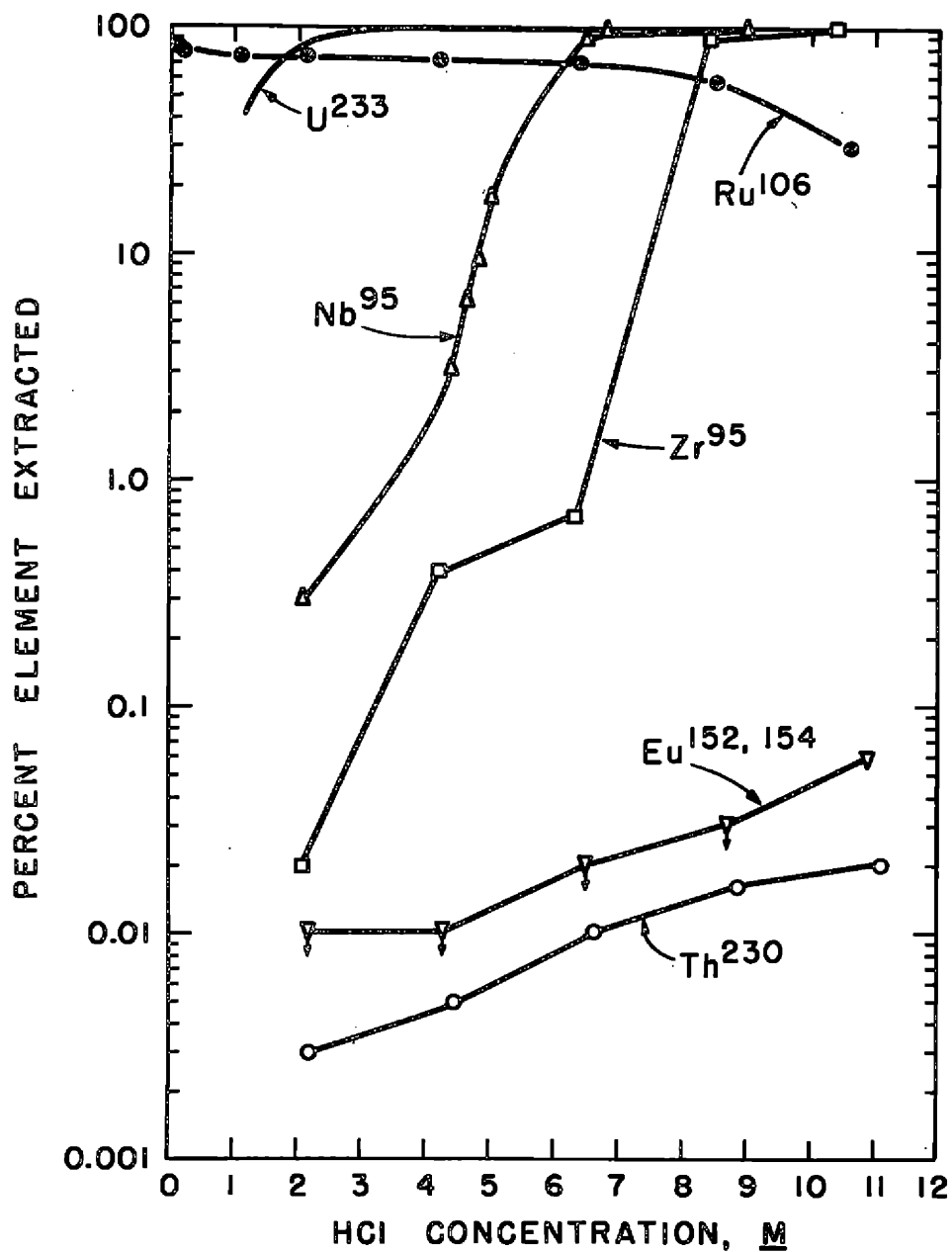


Figure 53. The extraction of U²³³, Th²³⁰, and fission products by 5% (w/v) triisooctylamine in xylene as a function of HCl concentration. After F. L. Moore, reference 475. Conditions: Equal phase volumes extracted for two minutes at room temperature (24°C.).

phase when washed with 0.1M HCl solution. Uranium is stripped into the aqueous phase. Excellent extraction (>90%) of macro amounts of uranium (60.4 mg U/ml initial aqueous concentration) can be obtained from 9M HCl with 20% T10A in hexone.

Moore⁴⁷⁶ has also investigated the extraction of uranium (VI) from acetic acid solution by T10A. Extractions were carried out in the same manner as those from hydrochloric acid solution (figure 53). Aqueous solutions of varying acetic acid concentration containing 2×10^4 alpha counts per minute per ml of U²³³ tracer were extracted with equal volumes of 5% (w/v) T10A in xylene. It was found that maximum uranium extraction (>90%) is obtained from 0.5M to 1M acetic acid solutions. The addition of 3%(v/v) butyl cellosolve to the T10A-xylene solution inhibits foaming during the extraction process. By increasing the T10A concentration, macro amounts of uranium are efficiently extracted. Greater than 95% stripping may be achieved by contacting the amine-xylene phase with an equal volume of 0.5M HNO₃, 3M H₂SO₄, 6M H₂SO₄, 1M NH₄HCO₃, concentrated NH₄OH, or 0.25M HF-0.25M HNO₃ solution. From 0.5M-1M acetic acid solution, ruthenium (11.5%), zirconium (27.9%), and niobium (11.1%) are extracted. Separation is made from strontium (alkaline earths), cesium and europium (rare earths), plutonium (III) (trivalent actinides), thorium, protactinium, hafnium, tantalum, iron, lead, nickel, cobalt, manganese, chromium (III), aluminum, copper, zinc, bismuth, tin, and antimony.^{471, 476} The selectivity may be improved if the uranium is first precipitated by hydroxide, dissolved with 1M acetic acid, and then extracted as previously described. Iron hydroxide is used to carry trace amounts of uranium in the precipitation step.⁴⁷⁶

Other amine extractants.

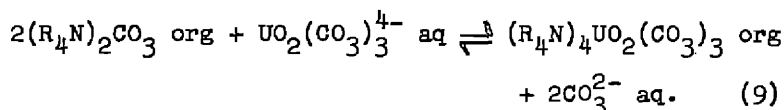
As stated at the beginning of this section, many organo-nitrogen compounds have been investigated as extractants of

uranium. A large number of these investigations are reported in ORNL reports (eg., ORNL-1922⁴⁷⁷, ORNL-2099⁴⁷⁸). For further information, one may refer to these reports, the summaries^{304, 461-463} previously mentioned, or the review by Moore.⁴⁷¹

Quaternary ammonium salts.

The enhanced extraction of uranium by hexone containing tetrabutylammonium nitrate⁷⁸ or tetrapropylammonium nitrate³⁷⁹ has already been noted (see Hexone). Haeffner, Nilsson, and Hultgren⁴⁷⁹ have also used tetrabutylammonium nitrate to extract uranyl nitrate with chloroform.

Quaternary ammonium salts, unlike amines, may be used to extract uranium from alkaline carbonate solutions.^{477, 478, 480} The Rohm and Haas compound Quaternary B-104* converted to the carbonate form has been used successfully to extract uranium from aqueous solutions having carbonate concentrations up to one molar.⁴⁸⁰ Amsco G alone or modified with a long-chain alcohol, tridecanol, and kerosene modified with tridecanol have been used as diluents. The alcohol modifier improves both the phase separation time and the extraction coefficient.⁴⁸⁰ The partition coefficient exhibits a negative two power dependence on carbonate concentration⁴⁸⁰ in accord with the reaction³⁰⁵



The extraction coefficient is virtually independent of the bicarbonate concentration with the carbonate-bicarbonate total concentration held constant.⁴⁸⁰ The coefficient is decreased by an increase in temperature.⁴⁸⁰ Uranium may be stripped from the organic phase by solutions of HCl, HCl-NH₄Cl, HNO₃, and HNO₃-NH₄NO₃.^{478, 480} Nitrate solutions

* An isopropanol solution of dimethyldiodecylammonium chloride.

are more effective than chloride.⁴⁷⁸ Sodium hydroxide (2M-3M) may also be used as a stripping agent.⁴⁸⁰

Clifford, *et al.*⁴⁸¹ report the extraction of uranium from aqueous carbonate solutions by (1) forming a singly charged anion, $UO_2X_3^-$, with a complexing agent, and (2) extracting this anion into an organic solvent with a singly charged cation. Extractions were obtained with benzoin 2-oxine, cupferron, hydroxylamine, peroxide, pyrogallol, and 8-quinolinol (oxine). The latter was used for further study. Arquad 2C, $R'_2N(CH_3)_2Cl$, where R' is about a 16-carbon chain, was found to be the most effective extractant tested. Hexone was found to be the most effective solvent tested. Kerosene gave no extraction. With oxine as complexing agent, the extracted species was identified as $R_4NUO_2(Ox)_3$. The extraction coefficient of uranium was found to increase with increased pH; to increase with increased oxine concentration and with increased R_4NCl concentration (to an optimum value); to decrease with increased carbonate concentration. An extraction coefficient, α_u , of 10.9 was obtained by extracting two volumes of an aqueous solution containing 0.01M $UO_2(NO_3)_2$, 0.92M Na_2CO_3 , 0.04M NaOH, and 0.02M Arquad 2C with one volume of hexone containing 0.100M oxine. Both uranium and oxine were removed from the organic phase by strong acids. Sodium bicarbonate was found the most efficient stripping agent on a counter-current basis.

CARBOXYLIC ACIDS.

Hök-Bernstrom^{138,482} has studied the extraction of uranium (VI), thorium, and lanthanum by several carboxylic acids: salicylic, methoxybenzoic, 3,5-dinitrobenzoic, and cinnamic. Table XXXI lists the pH at which 50 percent of the metal ions are extracted from perchlorate solutions by 0.1M solutions of the carboxylic acid in hexone. Chloroform was found to be a poor solvent for the extraction of

Table XXXI. pH for 50 Percent Extraction of U(VI), Th, La
by Carboxylic Acid.^a

Acid	pH ₅₀		
	UO ₂ ²⁺	Th ⁴⁺	La ³⁺
Salicylic	3.1 ^b	3.33 ^b	4.93 ^c
Methoxybenzoic	3.4 ^b	3.8 ^b	
3,5-Dinitrobenzoic	2.75 ^c	2.85 ^c	4.38 ^c
Cinnamic	3.60 ^c	3.07 ^c	6.13 ^c

^a After B. Hök-Bernstrom, references 138, 482.

Aqueous phase: metal concentration, 10⁻⁵M Th or La, 10⁻³M U;
ionic strength, 0.1M adjusted by the addition of NaClO₄;
pH adjusted with NaOH and HClO₄.

Organic phase: 0.1M carboxylic acid in hexone.

V_o/V_a, 1; temperature, 25°C.

^b Log α = 0, reference 138.

^c Calculated from data given in reference 482.

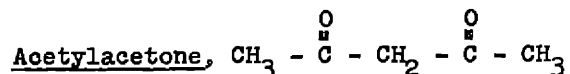
the metals by the carboxylic acids studied.

Cole and Brown⁴⁸³ have studied the extraction of U(VI), Th, Hf and Zr from aqueous nitrate solutions by salicylic acid in furfural. Satisfactory separations between uranium and thorium were obtained, depending largely upon the two metal concentrations.

Sudarikov, *et al.*⁴⁸⁴ have studied the extraction of U(VI), Th, Ce, La, Y, and Sc from aqueous solutions by salicylic acid in isoamyl alcohol. The uranium complex was observed to extract at pH 1.5 and to be completely extracted at pH 2.5 to 5.0. Up to pH 6.5, α_u was found to decrease from 100 to 0.3-0.4 and to remain unchanged at higher pH values.

Mills and Whetsel⁴⁸⁵ have extracted uranium (VI) with perfluorobutyric acid dissolved in diethyl ether.

CHELATING AGENTS. The chelating agents described below are listed in the same general order as they may be found in the book by Morrison and Freiser.²⁹⁰



The extraction of both uranium (VI) and (IV) from perchlorate solutions with acetylacetone as chelating agent has been investigated by Rydberg. The percentage extracted is given as a function of equilibrium pH in figures 54-A, -B, and -C for the three solvents, chloroform, benzene, and hexone, respectively. The extraction of other actinides, fission products, and hafnium is also included in the figures.^{14b,51, 486-488} Strontium and potassium are poorly extracted by acetylacetone into chloroform.⁴⁸⁷ Lanthanum and samarium are poorly extracted by the chelating agent into all three solvents.⁴⁸⁷

Krishen⁴⁸⁹ has investigated the extraction of uranium (VI) with acetylacetone used both as chelating agent and solvent. The results are given in figure 55-A together with the extraction curves of several other metals. The effect of masking agents, ethylenediaminetetraacetate, fluoride, and tartrate, on the extraction of these metals is given in figures 55-B, -C and -D, respectively.⁴⁸⁹

The extraction of uranium by acetylacetone-chloroform in the presence of sodium chloride and EDTA has been studied by Tabushi.⁴⁹⁰ Sodium chloride increases the extraction yield and broadens the favorable pH range. EDTA permits the separation of uranium from thorium and fission products by more effective masking of the latter. Uranium has also been extracted with acetylacetone using butylacetate as solvent.⁴⁹¹

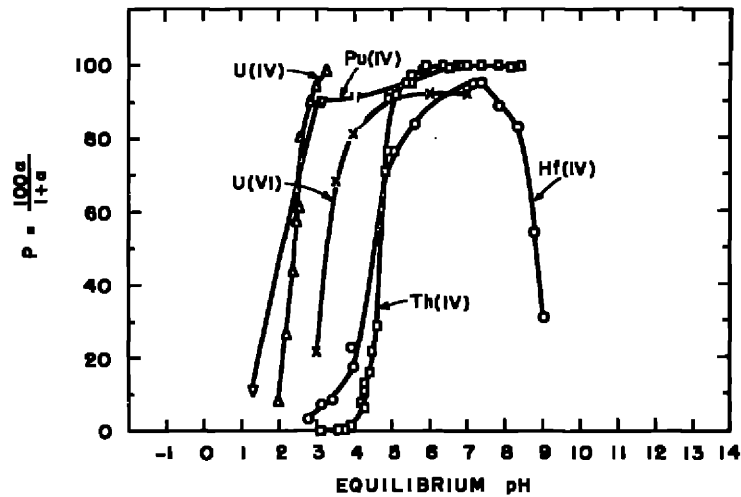


Figure 54-A. The extraction of various elements from 0.1M NaClO₄ solutions by an equal volume of acetylacetone-chloroform solution at 25°C. Acetylacetone concentrations used: U(VI), 0.0210M [HAA]_{aq}; U(IV), 0.50M [HAA]⁰_{org}; Pu(IV), 1.00M [HAA]_{init}; Th(IV), 0.0489M [HAA]_{init}; Hf(IV), 0.050M [HAA]⁰_{org}. After J. Rydberg, references 51, 486-488.

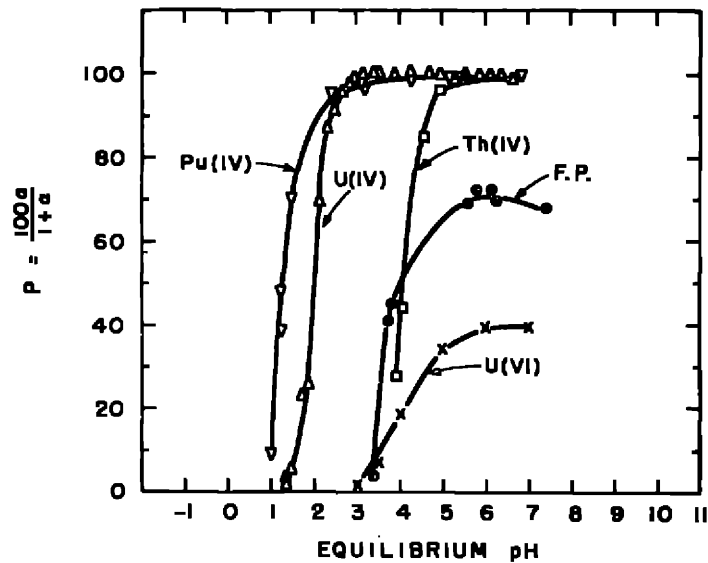


Figure 54-B. The extraction of various elements from 0.1M NaClO₄ solutions by an equal volume of acetylacetone-benzene solution at 25°C. Acetylacetone concentrations used: U(VI), 0.0210M [HAA]_{aq}; U(IV), 0.072M [HAA]_{aq}; Pu(IV), 1.00M [HAA]_{init}; Th(IV), 0.0673M [HAA]⁰_{org}; F.P., 0.70M [HAA]⁰_{org}. F.P. irradiation time = cooling time = 1 year. After J. Rydberg, references 51, 487, 488, 492 and J. Rydberg and B. Rydberg, reference 14b.

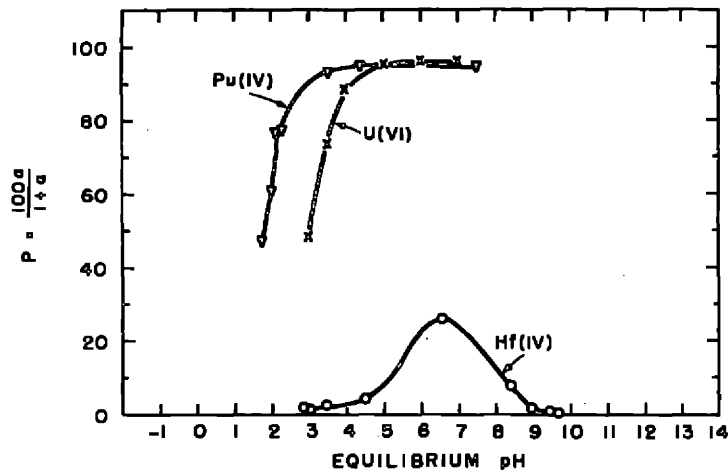
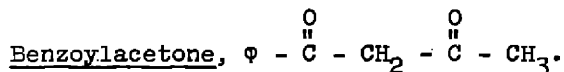
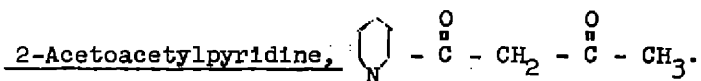


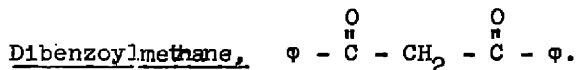
Figure 54-C. The extraction of various elements from 0.1M NaClO₄ solutions by an equal volume of acetylacetone-hexone solution at 25°C. Acetylacetone concentrations used: U(VI), 0.0210M [HAA]_{aq}; Pu(IV), 1.00M [HAA]_{init}; Hf, 0.050M [HAA]_{org}. After J. Rydberg, references 51, 487, 488.



Stary⁴⁹³ has determined the stability constants of uranyl acetate, oxalate, tartrate, and EDTA complexes. The effect of these ions was observed on the extraction of uranium (VI) from 0.1M NaClO₄ solutions by 0.1M benzoylacetone in benzene.



The extraction of uranium from a 0.2N NaOH, 0.2N acetic acid solution at pH 5.0 to 6.5 by 0.12% acetoacetylpyridine in butylacetate is reported by Hara.⁴⁹⁴



Uranium (VI) (0.05 - 0.5 mg) is extracted from aqueous solution by a 0.5% solution of dibenzoylmethane in ethyl acetate.⁴⁹⁵ In the presence of other cations, the extraction is made more selective by the addition of complexone III (EDTA sodium salt). Excess complexone is complexed by a 1% Ca(NO₃)₂

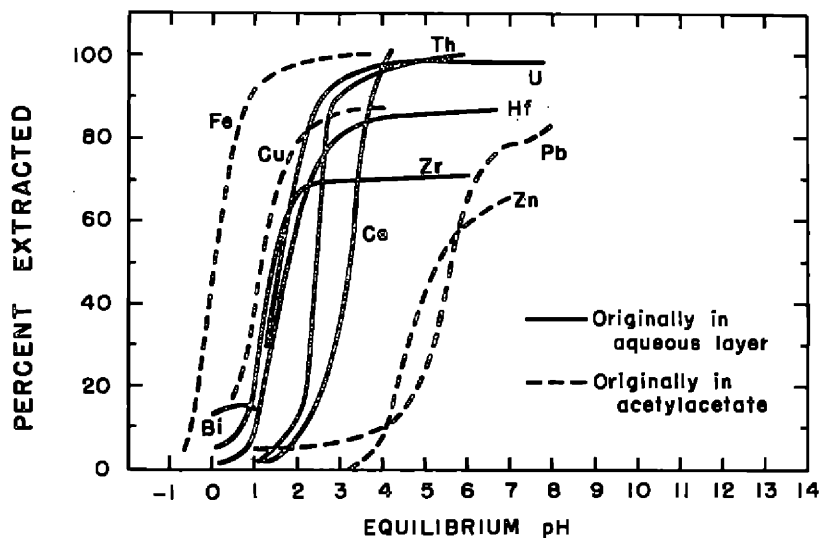


Figure 55-A. The extraction of various metals from aqueous solution by an equal volume of acetylacetone at 25°C. Solid lines indicate the metal was originally contained in the aqueous phase. The dashed lines indicate the metal was originally in the organic phase. The pH was adjusted to the desired value by sulfuric acid or sodium hydroxide. After A. Krishen, reference 489.

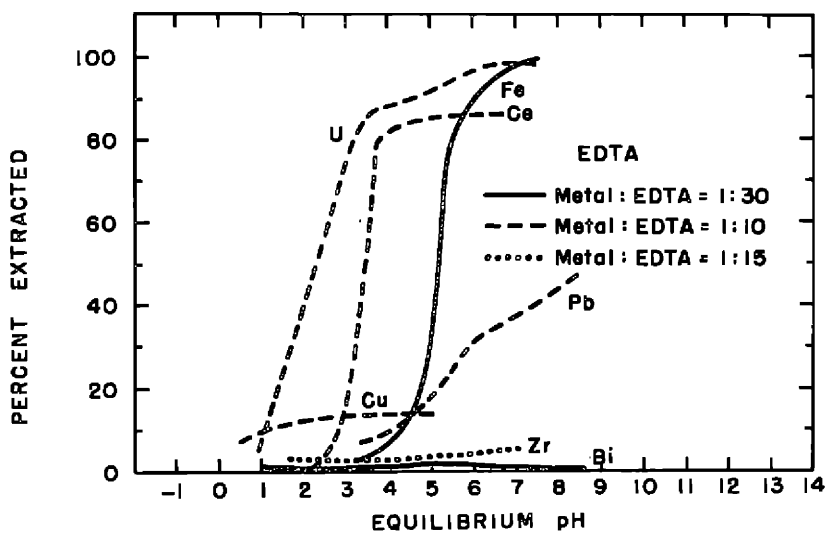


Figure 55-B. The effect of ethylenediaminetetraacetate (EDTA) on the extraction of various metals from aqueous solution by an equal volume of acetylacetone at 25°C. The mole ratio of metal to EDTA is shown by the line texture. After A. Krishen, reference 489.

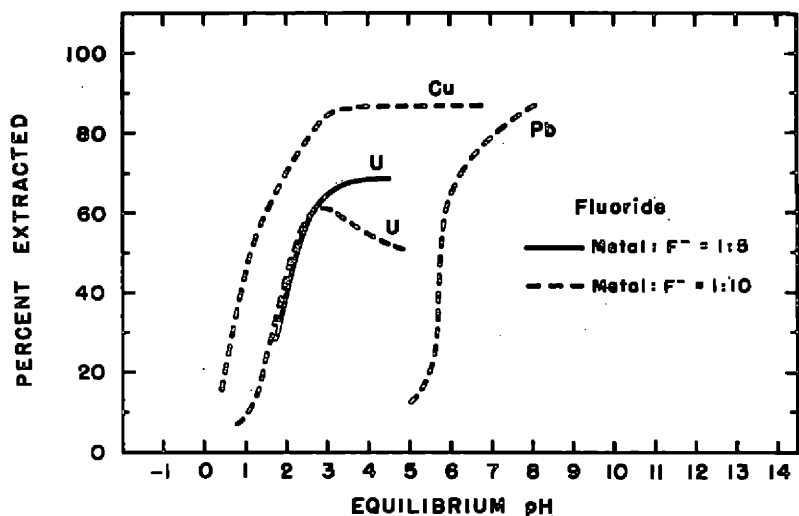


Figure 55-C. The effect of fluoride on the extraction of various metals from aqueous solution by an equal volume of acetylacetone at 25°C. The mole ratio of metal to fluoride is shown by the line texture. After A. Krishen, reference 489.

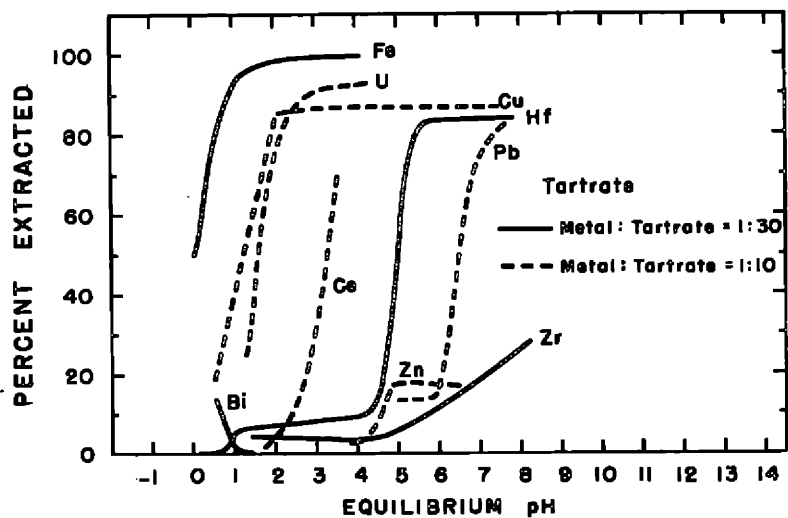
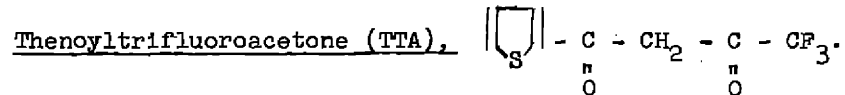


Figure 55-D. The effect of tartrate on the extraction of various metals from aqueous solution by an equal volume of acetylacetone at 25°C. The mole ratio of metal to tartrate is shown by the line texture. After A. Krishen, reference 489.

solution. The resulting solution is neutralized with ammonia to pH 7 and is then contacted several times with the extracting solution.

The dibenzoylmethane extraction of uranium with chloroform, benzene, and carbon tetrachloride has been investigated by Moucka and Stary.⁴⁹⁶



Considerable effort has been expended in the study of TTA as an extractant for uranium. King,⁴⁹⁷ Orr,⁴⁹⁸ Heisig and Crandall,⁴⁹⁹ Walton, et al.,⁵⁰⁰ and Peterson⁵⁰¹ have made fundamental studies on the extraction of uranium (VI) from aqueous perchlorate^{497, 498} and nitrate^{499, 500} media by TTA dissolved in benzene,^{497, 498, 500, 501} hexone,⁴⁹⁹ cyclohexanone,⁴⁹⁹ and pentaether.⁴⁹⁹ The partition coefficient, $\alpha_{u(VI)}$, is increased by increased TTA concentration in the organic phase; decreased by increased initial uranium concentration.⁵⁰⁰ The effect of pH and various salting agents on the extraction of uranium (VI) and thorium from nitrate solutions by 0.2M TTA in benzene is shown in figure 56.⁵⁰² Salting agents increase the extraction of uranium by TTA-benzene from low pH solutions. There is no apparent effect on the extraction of thorium with or without 1N Al(NO₃)₃. A 4M NH₄NO₃ concentration in the aqueous phase (not shown), in fact, depresses the extraction of thorium.⁵⁰² The effect of foreign anions on the extraction of U(VI)⁵⁷ and U(IV)⁵⁸ from aqueous perchlorate solutions by 0.5M TTA in benzene is shown in figures 57-A and 57-B, respectively. Poskanzer and Foreman⁵⁰³ have recently reviewed the extraction of elements throughout the periodic table by TTA. The pH values for 50 percent extraction into an equal volume of 0.2M TTA in benzene at room temperature or 25°C. listed by these authors

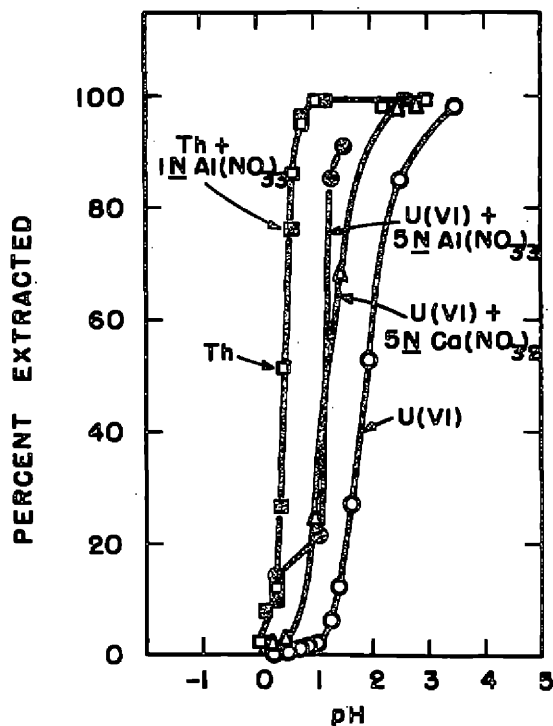


Figure 56. The effect of pH and salting-out agents on the extraction of uranium (VI) and thorium by TTA-benzene solutions. After E. K. Hyde and J. Tolmach, reference 502. Conditions: An equal volume of 0.2M TTA in benzene was stirred vigorously for 20 minutes with an aqueous solution containing 0.003M thorium or trace amounts of uranium-233 with or without the salting-out agent indicated at the pH given.

are: for U(VI) from dilute nitric acid,⁵⁰² $pH_{50} = 1.97$;
 for U(VI) from $HClO_4 + LiClO_4$,⁴⁹⁸ $\mu = 2$, $pH_{50} = 1.79$; for
 U(IV) from $HClO_4 + NaClO_4$,^{14a} $\mu = 2$, $pH_{50} = -0.58$; for U(IV)
 from HNO_3 ,⁵⁰⁴ $pH_{50} = -0.31$.

Irving and Edgington⁴⁵¹ have observed a synergistic enhancement of the uranium partition coefficient with tributylphosphate (TBP) - or tributylphosphine oxide (TBPO) - TTA mixtures. The results, α_U versus percent TBP or TBPO in the extractant mixture, are given in figure 58.

The analysis of metals with TTA has been reviewed by Moore,⁵⁰⁵ Sheperd and Meinke⁵⁰⁶ have published, with references, the extraction curves of a large number of elements with TTA.

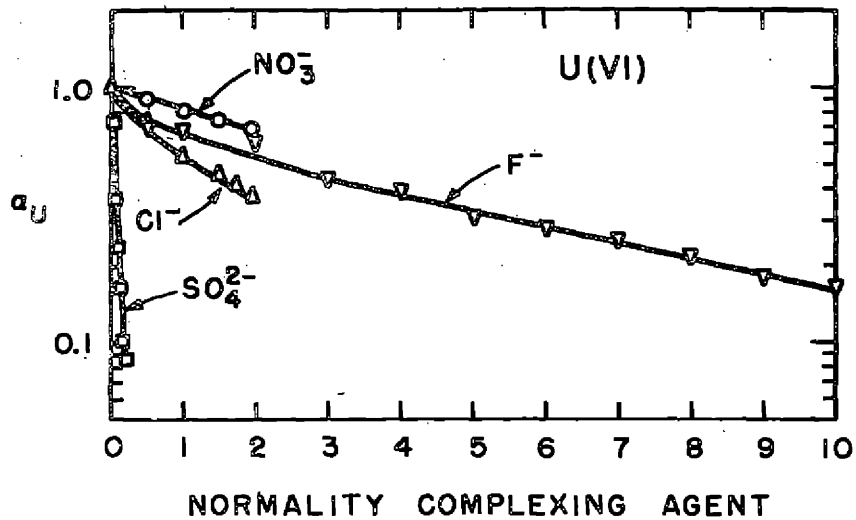


Figure 57-A. The effect of foreign anions on the extraction of uranium (VI) from aqueous perchlorate solution by TTA-benzene. After R. A. Day, Jr. and R. M. Powers, reference 77. Conditions: Organic phase - 0.50M TTA in benzene pre-treated by shaking with dilute perchloric acid overnight. Aqueous phase - $\sim 10^{-5}$ M U233, anion at concentration indicated, 0.05M HClO₄, plus sufficient NaClO₄ to maintain an ionic strength of 2.0. Equal phase volumes shaken together for 2 hours at 25°C.

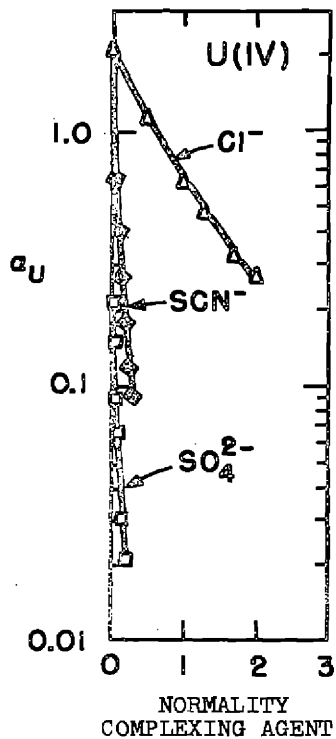


Figure 57-B. The effect of foreign anions on the extraction of uranium (IV) from aqueous perchlorate solution by TTA-benzene. After R. A. Day, Jr., R. N. Wilhite, F. D. Hamilton, reference 58. Conditions: Organic phase - 0.05M TTA in benzene pre-treated with dilute acid. Aqueous phase - 0.0016M - 0.0037M U(IV), anion at concentration indicated, 1.000M H⁺ (HClO₄ used for all experiments except chloride in which HCl was used), plus sufficient NaClO₄ to maintain an ionic strength of 2.0. Equal phase volumes shaken together for 30 minutes.

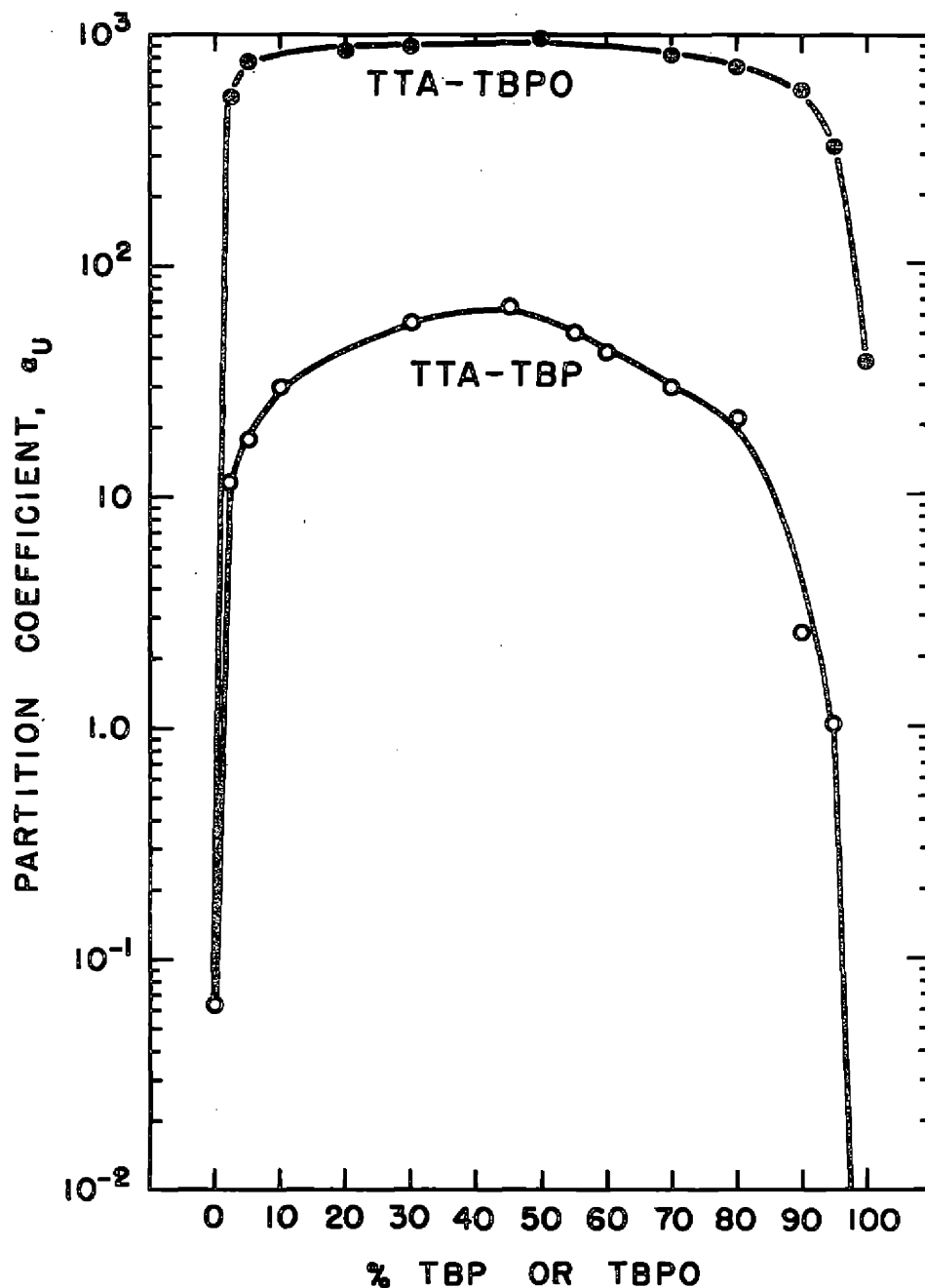
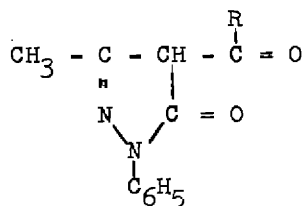


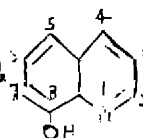
Figure 58. The synergistic enhancement of the uranium (VI) partition coefficient between aqueous nitrate solutions and mixtures of TTA and TBP or TBPO in cyclohexane. After H. Irving, D. N. Edgington, reference 451. Conditions: Organic phase - 0.02M mixture of TTA and TBP or TBPO in cyclohexane. Aqueous phase - $1.025 \times 10^{-4}M$ U^{233} , 0.01N HNO_3 . Equal phase volumes shaken together for 24 hours at room temperature ($21^\circ - 23^\circ C.$)

Substituted 1-phenyl-3-methyl-4-acyl-pyrazolones-5,



Skytte Jensen¹⁶⁰ has studied the possibility of using substituted 1-phenyl-3-methyl-4-acyl-pyrazolones-5 as extractants for a number of elements including uranium (VI), thorium, and lanthanum. The pH for 50% extraction of trace amounts of these elements by a 1M solution of chelating agent in chloroform is given in Table XXXII. The pH₅₀ for TTA is given for comparison.

8-Quinololinol (8-hydroxyquinoline, oxine)



Hok,⁵⁰⁷ and Dyrssen and Dahlberg¹⁴³ have studied the extraction of uranium (VI) from aqueous perchlorate solutions by oxine dissolved in chloroform or hexone. Results of the latter group,¹⁴³ percent extracted versus final aqueous pH, are shown in figure 59. These results are in agreement with those of Hok⁵⁰⁷ (0.1M oxine - CHCl₃, 10⁻³M U, aqueous perchlorate solution, $\mu = 0.1\text{M}$, 25°C.). No appreciable difference was observed with uranium concentrations of 10⁻⁴ and 10⁻³M (open and solid circles, respectively, in figure 59). Chloroform is shown to be a slightly better solvent for the uranyl-oxine complex than hexone. The extraction curves for Th,⁵⁰⁸ La and Sm⁵⁰⁹ are also shown in the figure. A tabulation of pH for 50% extraction of various metal oxinates by chloroform has been made by Dyrssen and Dahlberg¹⁴³ and is reproduced in Table XXXIII.

The extraction of uranium (VI) by solutions of 1% oxine in chloroform from buffered aqueous solutions is shown in figure 60 as a function of aqueous pH.⁵¹⁰

Table XXXII. pH for 50% Extraction of Tracer Amounts of Uranium (VI), Thorium, and Lanthanum by 1M Solutions of Substituted (R) 1-Phenyl-3-methyl-4-acyl-pyrazolones-5 in Chloroform.^a

<u>R</u>	pH ₅₀ of Metal Ion		
	UO ₂ ²⁺	Th ⁴⁺	La ³⁺
acetyl	-0.15	0.10	2.60
propionyl	0.05	0.05	2.65
butyryl	0.52	0.42	2.47
valleryl	0.24	0.24	2.84
capronyl	0.7	-0.25	3.15
ethoxycarbonyl	1.00	not meas.	2.50
chloroacetyl	0.65	0.05	2.28
trifluoroacetyl	0.8	not meas.	not meas.
benzoyl	1.0	0.4	2.45
p-bromobenzoyl	0.9	0.30	2.3
p-nitrobenzoyl	-	-	-
TTA	0.70	-0.30	3.75

^a Values for pH₅₀ were calculated from data presented by B. Skytte Jensen, reference 160.
Aqueous perchlorate media.

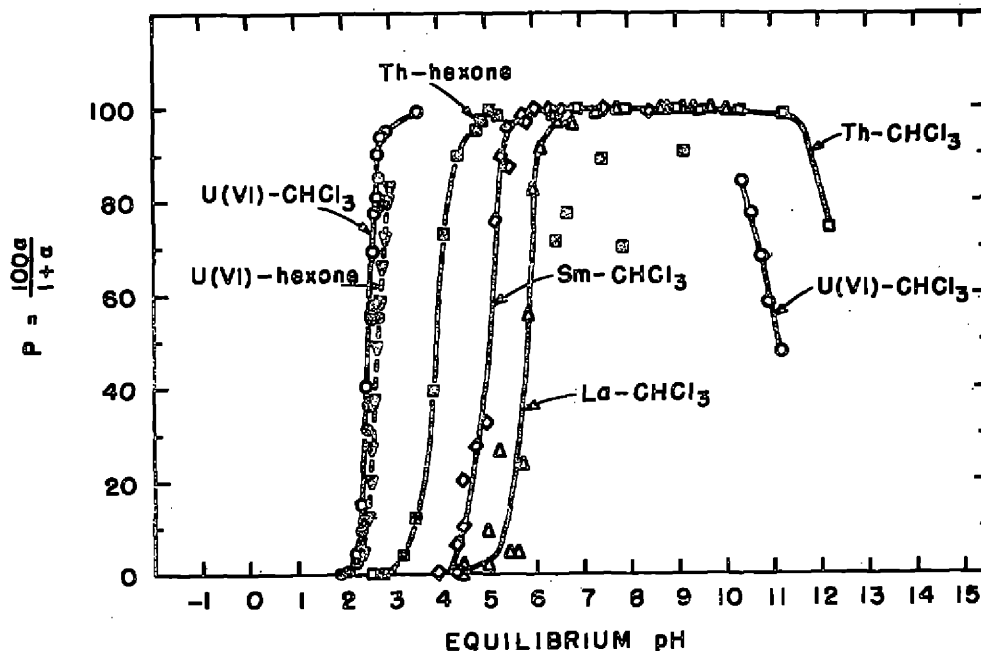


Figure 59. The extraction of tracer amounts of uranium (VI), thorium, samarium, and lanthanum from perchlorate solution by solutions of oxine-chloroform or oxine-hexone. After D. Dyrssen and V. Dahlberg, reference 143; D. Dyrssen, references 508 and 509. Conditions: Aqueous phase - ionic strength = 0.1M with NaOH, HClO₄, and NaClO₄; for uranium, open circles represent 0.0001M U concentrations, solid circles and triangles, 0.001M U. Organic phase - oxine concentrations: for U, 0.100M; for Th, 0.050M; for La and Sm, 0.5M; solvent indicated. Equal phase volumes equilibrated at 25°C.

Substituted quinolinols.

Rulfs, et al.⁵¹⁰ and Dyrssen, et al.⁵¹⁴ have studied the extraction of uranium by dihalogen derivatives of 8-quinolinol. The uranium extraction curves with 1% solutions of 5,7-dichloro- and 5,7-dibromo-8-quinolinol in chloroform are shown as functions of final aqueous pH in figure 60.⁵¹⁰ Use of the halogen-substituted oxines permits extraction of uranium from more acidic aqueous solutions. Similar curves for uranium, thorium, and lanthanum are given in figure 61 for extraction with 0.05M 5,7-dichloro-oxine in chloroform.⁵¹⁴

Hynek⁵¹⁵ has studied the extraction of various metals by 8-hydroxyquinaldine (2-methyl-8-quinolinol). The uranium complex was found to be extracted, but not quantitatively,

Table XXXIII. pH for 50% Extraction of Metal Oxinates with Chloroform.^a

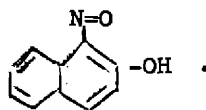
Metal ion	pH	Procedure	Reference
Ga ³⁺	1.0	V aq = V org, 0.1M total oxine.	511
In ³⁺	2.1	Anions in aqueous solution: chloride.	
Al ³⁺	3.4		
Fe ³⁺	1.6	Four successive extractions with	512
Cu ²⁺	2.0	0.01M solution of oxine in CHCl ₃ .	
In ³⁺	2.2	Anions in aqueous solution: sulfate,	
Bi ³⁺	3.0	acetate, nitrate, chloride.	
Al ³⁺	4.2		
Ni ²⁺	6.1		
Co ²⁺	6.5		
Sn ⁴⁺	0.0	V aq = 5V org, 0.07M oxine.	513
Mo	1.0	Anions in aqueous solution:	
Fe ³⁺	2.0	acetate, chloride, tartrate.	
Cu	2.1		
Ni	3.7		
Al	3.8		
Mn ²⁺	6.4		
Hf ⁴⁺	1.3 [*]	V aq = V org, 0.1M total oxine.	143
UO ₂ ²⁺	2.6	Anions in aqueous solution:	
Th ⁴⁺	3.1	perchlorate.	
Sm ³⁺	5.7		
La ³⁺	6.5		

^a After D. Dyrssen and V. Dahlberg, reference 143.

* pH = -log[H⁺] + 0.1.

by chloroform from an aqueous phase at pH 9.5 containing tartrate and acetate ions. Cyanide or H₂O₂ prevented extraction.

1-Nitroso-2-naphthol,



Alimarín and Zolotov⁵¹⁶ have investigated the extraction of uranium (VI) by organic solutions of 1-nitroso-2-naphthol. It was found that a mole ratio of naphthol to U₃O₈ of 125

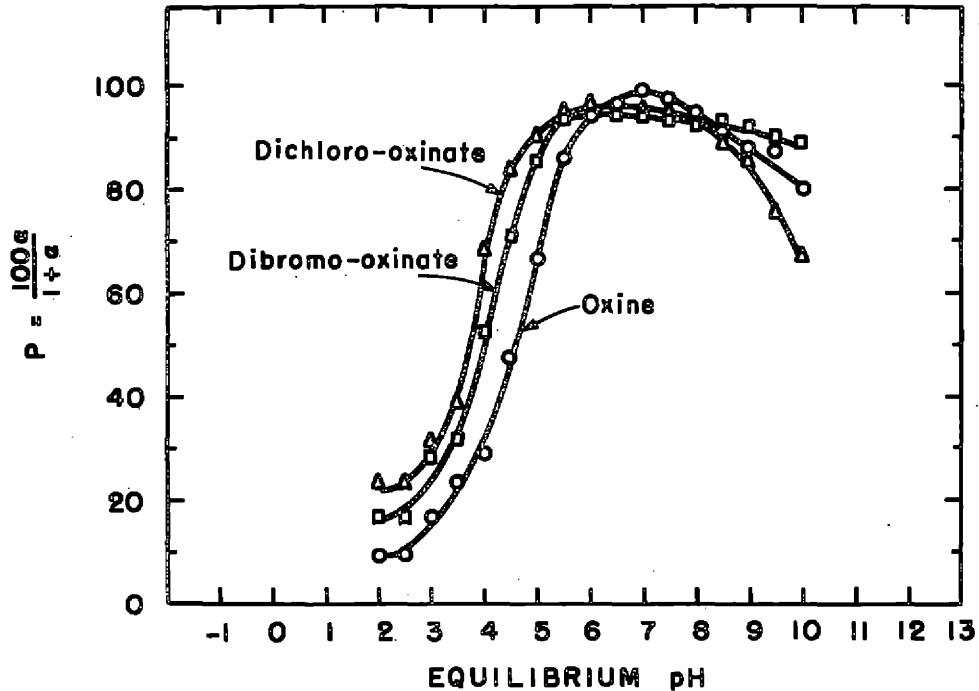


Figure 60.

The extraction of uranium (VI) by oxine and its 5,7-dichloro- and 5,7-dibromo-derivatives. The percent extracted, P , was calculated from the values of the distribution coefficient given in the paper by C. L. Rulfs, A. K. De, Jr., J. Lakritz, and P. J. Elving, reference 510.

Conditions:

2.1 mg of uranium in 10 ml and 25 ml of an approximately 1M buffer solution were shaken 6 to 8 minutes with 20 ml of a 1% oxine-chloroform solution. The aqueous phase was rinsed twice with 5 ml of chloroform. The pH of the final aqueous phase was measured.

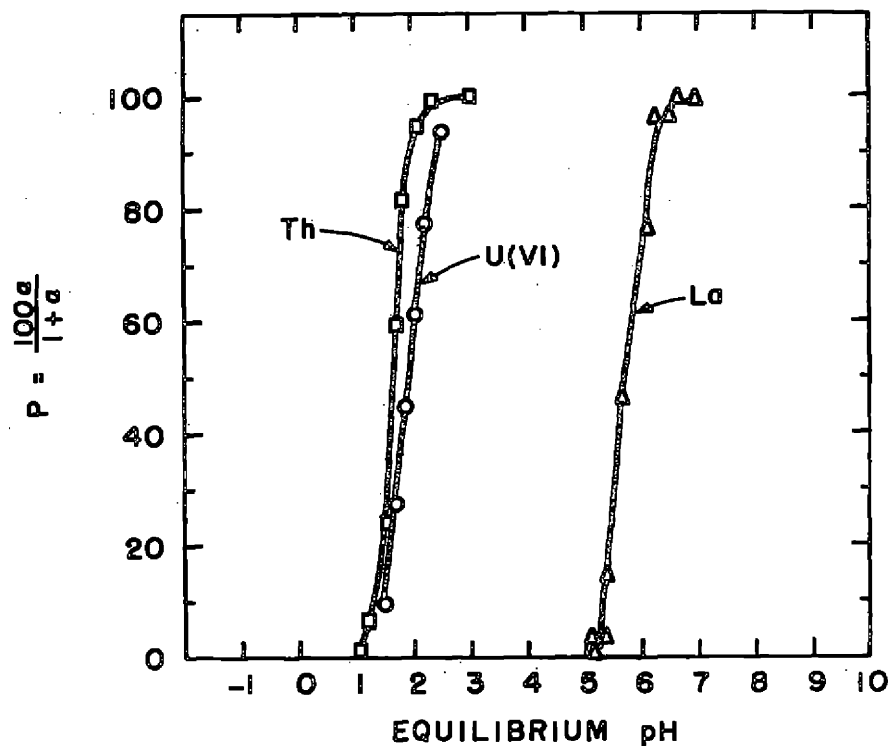


Figure 61.

The extraction of tracer amounts of U (VI), Th, and La by 0.05M 5,7-dichloro-oxine dissolved in chloroform from 0.1M HClO_4 - NaClO_4 solutions at 25°C.

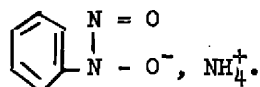
After D. Dyrssen, M. Dyrssen, and E. Johansson, reference 514.

and a volume phase ratio of organic to aqueous of 0.25 is more than adequate to give quantitative extraction of uranium into isoamyl alcohol at a pH of 5 to 6. Two minute shaking is sufficient for quantitative uranium extraction. Ethyl acetate, n-butanol, diethyl ether, amyl acetate, benzene, and chloroform also extract the uranium-naphtholate complex. Quantitative extraction is obtained with ethyl acetate and n-butanol at a pH of 3.0 to 8.5; with isoamyl alcohol at pH 4.5 to 7.5. Quantitative extraction can be achieved at temperatures of 0° to 100°C. Chloride or nitrate ions at con-

concentrations up to 0.2M do not seriously interfere with the extraction of uranium. Iron (III) is completely extracted; vanadium (IV) and (V) and thorium are partially extracted. The extraction of all four metal ions is considerably suppressed by complexing with complexone III (sodium salt of EDTA). The pH range for quantitative separation of uranium with isoamyl alcohol is increased in the presence of complexone III (~25 mg complexone per mg of metal) to 6.5 - 9. Aluminum and zinc are not extracted with 1-nitroso-2-naphthol.

Dyrssen, et al.⁵¹⁷ have studied the extraction of uranium and thorium from aqueous perchlorate solutions ($\mu = 0.1M$) by 0.1M 1-nitroso-2-naphthol in chloroform. Fifty percent of the uranium was extracted at pH 3.07 and fifty percent of the thorium at pH 1.66. Lanthanum and samarium were not extracted. Other metals that have been extracted as nitroso-naphtholates include Mn(II), Fe(II), Co, Ni, Cu(II), Pd(II), Ag, Cd, Hg(II), Pu(IV)⁵¹⁷ and Np(V).⁵¹⁸

Ammonium salt N-nitrosophenylhydroxylamine (cupferron),



Cupferron is an important reagent in the analytical separation of uranium. The reagent precipitates uranium (IV) from acidic (H_2SO_4 or HCl) solution but not uranium (VI). By converting uranium to its two oxidation states, separation can be made alternatively from elements not precipitated by cupferron and from those precipitated by the reagent. The uranium (IV) cupferrate complex, $\text{U}(\text{Cup})_4$, was found by Auger⁵¹⁹ to be soluble in chloroform and neutral organic solvents. Furman, et al.²⁵⁰ found milligram amounts of uranium (IV) to be incompletely extracted from aqueous acid solution by hydrogen cupferrate in chloroform but to be almost completely extracted by ethereal hydrogen cup-

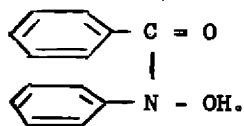
ferrate; i.e., cupferron extracted by ether from an acid solution. Ethereal hydrogen cupferrate was also found to extract quantitatively macroamounts of uranium (IV) from (1 + 19) sulfuric acid containing hydroxylamine hydrochloride and submilligram amounts from (1 + 19) sulfuric acid in the presence of saturated mercury-zinc amalgam.²⁵⁰ The partition coefficient, $\alpha_{u(IV)}$, is increased with increased cupferrate concentration and is decreased with increased acid concentration.²⁵⁰

A uranium (VI) cupferrate complex is precipitated by the reagent from neutral solutions. There appear to be two forms, one of which is soluble in ethyl ether.²⁵⁰ From (1 + 9) sulfuric acid, milligram amounts of uranium (VI) are extracted by an equal volume of chloroform with an excess of cupferron present.²⁵⁰

The extraction of uranium (VI) cupferrate from aqueous perchlorate solution by hexone and chloroform is given in figure 62 as a function of the pH of the final aqueous solution.¹⁴³ Chloroform is a poor solvent for the complex. Hexone is better, but quantitative extraction is not achieved by a single contact of the solvent with an equal volume of the aqueous solution. The extraction curves for Th,⁵⁰⁸ Sm, and La⁵⁰⁹ are also given in the figure.

The properties of other metal cupferrates have been reviewed by Furman, Mason, and Pekola.²⁵⁰

N-Benzoylphenylhydroxylamine,



Dyrssen⁵²⁰ has studied the extraction of uranium (VI) with N-benzoylphenylhydroxylamine in chloroform from aqueous perchlorate solutions. The results, P versus pH, are shown in figure 63 together with those for thorium and lanthanum.

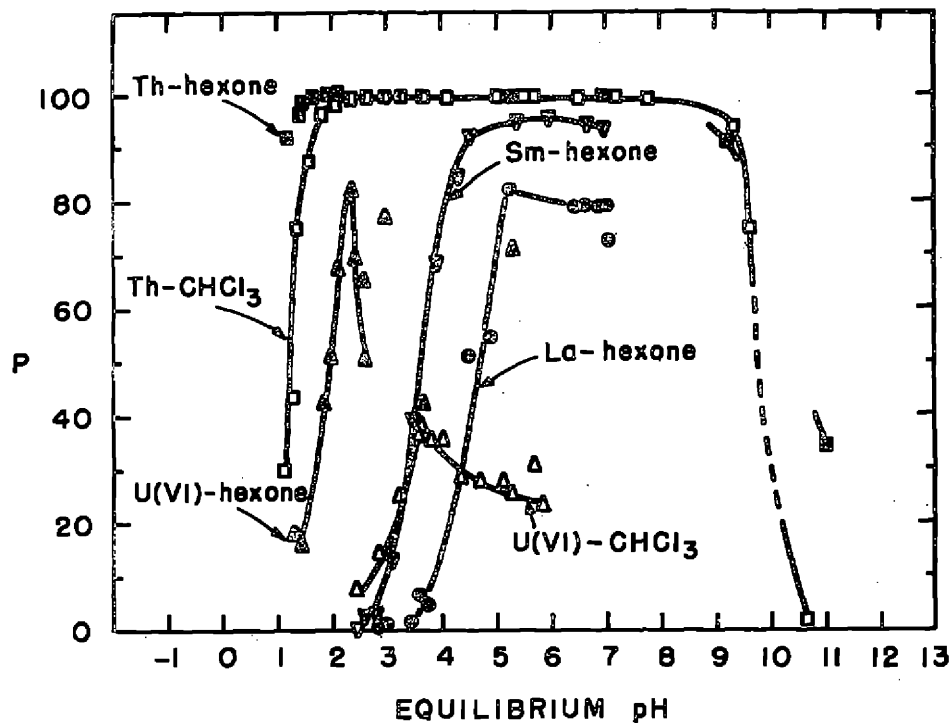


Figure 62.

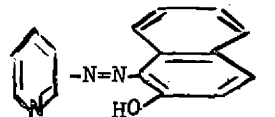
The extraction of tracer amounts of uranium (VI), thorium, samarium, and lanthanum cupferrates from perchlorate solutions by hexone or chloroform.

After D. Dyrssen and V. Dahlberg, reference 143; D. Dyrssen, references 508 and 509.

Conditions:

Aqueous phase - ionic strength = 0.1M with NaOH, HClO_4 , and NaClO_4 . Na cupferrate added to aqueous phase: for U, 0.01M, for Th, Sm, and La, 0.005M.
Equal volumes of aqueous and organic solvent indicated equilibrated at 25°C.

1-(2-Pyridylazo)-2-naphthol (PAN),



1-(2-Pyridylazo)-2-naphthol forms colored complexes (generally red) with a large number of polyvalent metal ions.⁵²¹ The uranyl-PAN complex is insoluble in alcohols, carbon tetrachloride, chloroform, and ethers.⁵²² Ortho-

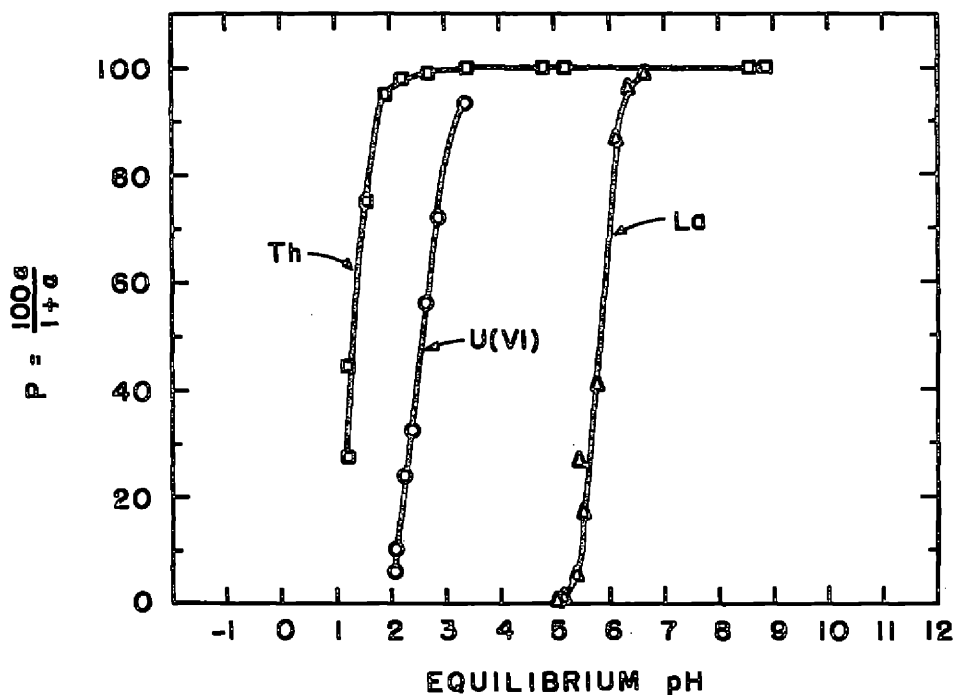


Figure 63.

The extraction of tracer amounts of uranium (VI), thorium, and lanthanum from perchlorate solutions by N-benzoyl-phenylhydroxylamine dissolved in chloroform.

After D. Dyrssen, reference 520.

Conditions:

Aqueous phase - ionic strength = 0.1M with HClO_4 , NaOH , and NaClO_4 . The aqueous phase was sometimes buffered with 1 ml of 0.1M anilinium perchlorate, sodium acetate, or hydrozinium perchlorate per 15 ml.

Organic phase - 0.1M N-benzoylphenylhydroxylamine in chloroform.

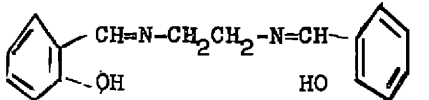
Temperature, 25°C.

or meta-dichlorobenzene and bromobenzene are excellent solvents for the complex. The maximum color of the uranyl-PAN complex is developed at pH 10. At pH less than 5 or greater than 12 little complex formation occurs.⁵²² Uranium may be selectively separated from a large number of elements by PAN-dichlorobenzene extraction in the presence of masking

agents (EDTA, trinitrilotriacetic acid, cyanide) and with proper pH control.

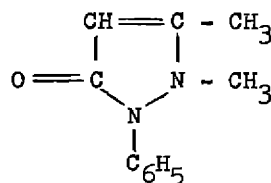
Sodium diethyldithiocarbamate (DDTC), $(C_2H_5)_2N - C \begin{matrix} =S \\ \backslash S^- \end{matrix} Na^+$.

Bode⁵²³ reports that the U(VI) - DDTC complex, unlike other heavy metals, is soluble in water. A precipitate is formed only with high concentrations of uranium and reagent. The uranyl-DDTC complex is practically inextractable by carbon tetrachloride but is readily extracted by isoamyl alcohol, diethyl ether, and amyl acetate.⁵²³ Others have used hexone,⁵²⁴ ethyl acetate,⁵²⁵ chloroform,⁵²⁶ and benzene⁵²⁷ to extract the complex. Employing the above solvents, the U(VI)-DDTC complex has been extracted from aqueous solutions having a wide range of pH, e.g., pH 1-5⁵²⁷ and pH 6.5-8.3.⁵²⁵ Sodium tartrate has been used to prevent hydrolysis at higher pH values.⁵²⁵ The U(VI)-DDTC complex is extracted in the presence of EDTA. Uranium may then be separated from elements, such as thorium, that form strong EDTA complexes.^{524,527,528} Uranium may be further separated from those elements extracted as DDTC complexes by stripping the former into an ammonium carbonate solution.^{196,525}

Disalicylethylenediimine, 

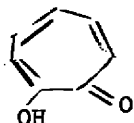
Dyrssen⁵²⁹ reports that uranium (VI) is somewhat extractable (60-90%) with solutions of disalicylethylenediimine in chloroform. Hafnium and thorium are extracted (90-99%) from weakly acidic solutions (pH 1.5) with a 0.1-0.5M chloroform solution of the reagent. Lanthanum and samarium are not extracted.

Antipyrine,



Rodden¹⁹¹ has mentioned that chloroform extracts uranyl complexes with antipyrine. Reas³⁸⁰ has reported that both uranium (VI) and uranium (IV) are almost completely extracted with antipyrine-chloroform solutions from perchlorate media. Uranium may be separated from thorium using the antipyrine-chloroform system. From an aqueous solution of 20.6 ml containing 5 mmoles of $\text{Th}(\text{NO}_3)_4$, 1 mmole of $\text{UO}_2(\text{NO}_3)_2$ and 48.6 mmoles of HCl, 93-94% of the uranium and only 5% of the thorium was found to extract with 36 mmoles of antipyrine in chloroform. The uranyl-antipyrine complex is soluble in nitrobenzene, but not very soluble in hexone.

Tropolone.



The extraction of U(VI) and Th from 0.1M perchlorate solutions by 0.05M tropolone in chloroform is given as a function of pH in figure 64.⁵³⁰ The pH of 50% extraction for U(VI), Th, and Y under the above conditions is approximately 0.9, 1.1, and 4.0, respectively. Less than 50% lanthanum is extracted at pH 6.5.⁵³⁰

Dyrssen⁵³¹ reports the extraction of a uranium (VI)-beta-isopropyl tropolone complex with chloroform and hexone.

Ion Exchange.

A number of articles are available in which the behavior of uranium toward ion exchange resins is reviewed and in which reference to much of the literature is given. Hyde,³³⁹ Katz and Seaborg,² Choppin,⁵³² Palei,¹⁹⁷ and Kuznetsov, et al.²⁰⁰ have reviewed the ion exchange of a number of the actinide elements including uranium. Steele and Taverner¹⁹⁸ have outlined several anion exchange separations of uranium. Clegg and Foley⁵³³ have described the use of ion exchange resins in the processing of uranium ores.

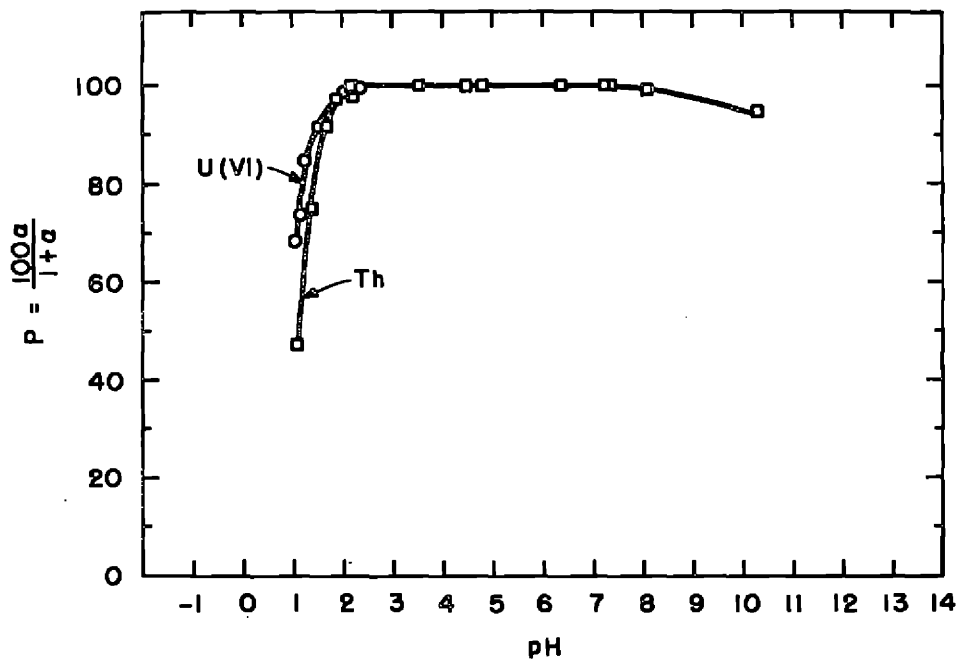


Figure 64.

The extraction of trace amounts of uranium (VI) and thorium from 0.1M perchlorate solution by 0.05M tropolone in chloroform at 25°C.

After D. Dyrssen, reference 530.

In the following paragraphs, the distribution of uranium (and of other elements) between an ion exchange resin and a particular solution is described in terms of the distribution coefficients, D and D_v . These are defined as

$$D = \frac{\text{amount } M^{+x} / \text{gram dry resin}}{\text{amount } M^{+x} / \text{ml solution}}$$

and

$$D_v = \frac{\text{amount } M^{+x} / \text{ml resin bed}}{\text{amount } M^{+x} / \text{ml solution}}$$

The two coefficients are related by the density of the resin bed, $D_v = \rho D$. The coefficient D is referred to as K_D by many authors.

Anion exchange. Anion exchange resins commonly used in the radiochemical laboratory are the strong base resins such as Dowex-1 and-2 and Amberlite IRA-410 and IRA-400. The capacity of these resins is approximately 2.5 milli-equivalents per gram of resin. Weak base resins are also available. However, their use is more limited. These resins have capacities ranging from about 6 to 10 milli-equivalents per gram of resin.

Spivey, et al.⁵³⁴ have investigated various factors such as resin capacity, resin phase volume anion adsorption, etc., that affect the sorption of uranium. Trivisonno⁵³⁵ has made a literature survey of factors that influence the adsorption and elution of uranium by and from strong base anion exchange resins. These are similar to the factors influencing solvent extraction and include, other than those already mentioned, uranium concentration, anion concentration, pH, the presence of other metallic ions and foreign anions, temperature, resin size, porosity, cross-linkage, etc.

The various systems from which uranium may be adsorbed by anion exchange resins are described below. The resin may be converted to a particular anionic form by washing with an appropriate solution.

Chloride systems.

Kraus and Nelson⁵³⁶ have measured the distribution coefficients for a number of elements between a strong base anion exchange resin (Dowex-1, 10% DVB, ~200 mesh) and hydrochloric acid solutions of varying molarity. Their results are shown in figure 65. The results of Marcus,⁵³⁷ obtained under conditions similar to those used by Kraus and Nelson,⁵³⁶ are given in figure 66. The concentrations of the various elements used in the study by Marcus were such that the oxidation states could be determined spectro-

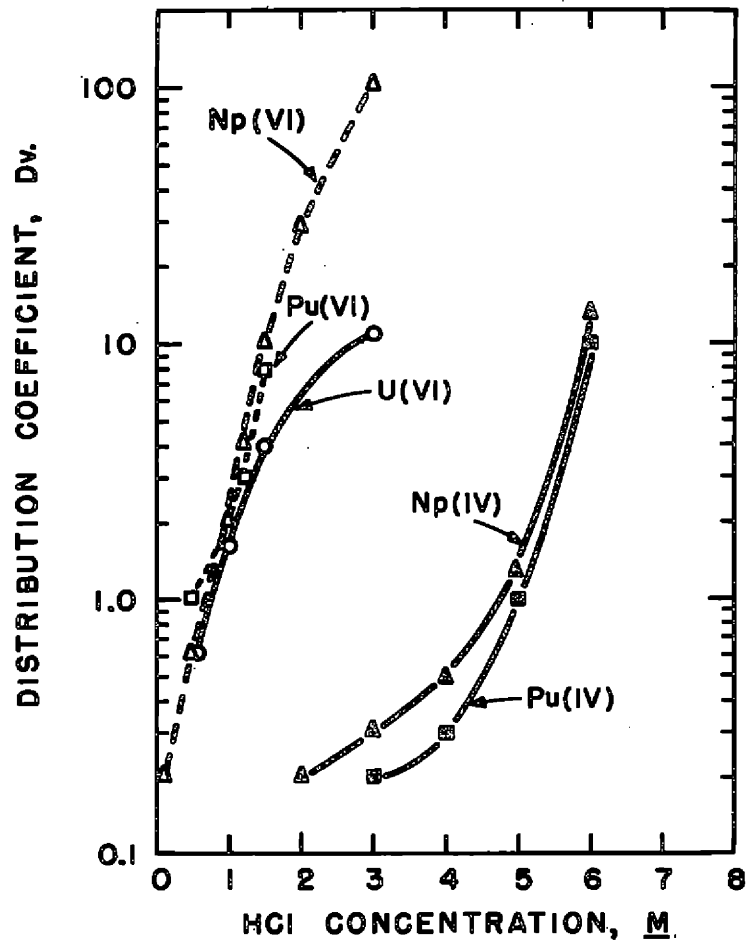


Figure 66.

Adsorption of elements from hydrochloric acid solution with Dowex-1 anion exchange resin (10% DVB).

After Y. Marcus, reference 537.

photometrically. No adsorption of Np(V), Pu(III), or Am(III) was found. The data presented by Wish⁵³⁸ for the adsorption of various elements by Dowex-2 (x8,200-400 mesh) is represented in figure 67. Ward and Welch⁵³⁹ have studied the distribution of neptunium in various oxidation states between Amberlite IRA-400 and hydrochloric acid solutions of varying concentration. Their results show that Np(VI) is strongly adsorbed at $>6M$ HCl, Np(V) is exponentially adsorbed from $3M$ to $6M$ HCl (D increases from 1 to 10), and Np(IV) is similarly adsorbed from about $6M$ to $10M$ HCl (D increases from about 2 to 400). Prevot, et al.⁵⁴⁰ have investigated the adsorption of U, Pu, Th, Fe, Ce, and Zr by anion exchange resin A300D from hydrochloric acid solutions ranging in molarity from 4 to 7. Quantities of 7 mg Pu, 6.9 mg U and 5.9 mg Fe per ml of solution and 2 grams of resin were used in the determination of D. Their results are considerably different than those shown in figures 65-67. The distribution coefficients of U(VI), and Fe(III) are lower roughly by an order of magnitude. The distribution coefficient of Pu(IV) is almost an order of magnitude higher. For Pu(III), D is about 0.1 at $4M$ HCl and about 1 at $7M$ HCl. Zirconium adsorption is similar to that shown in the figures. Thorium and cerium are poorly or not at all adsorbed.

Korkisch, et al.^{541,542} have found the distribution coefficient of uranium between Dowex-1 and hydrochloric acid solutions to increase with increased alcohol concentration of the solution. With 80% ethanol,* D is increased from about 40 to 6000 as the HCl concentration is increased from $0.2M$ to $2.4M$.⁵⁴¹ The distribution coefficient at $2.4M$ HCl without alcohol is about 40. Alcohol also increases the adsorption of thorium, titanium, and zirconium. The distribution co-

* 95% alcohol denatured with benzene is considered 100% alcohol.

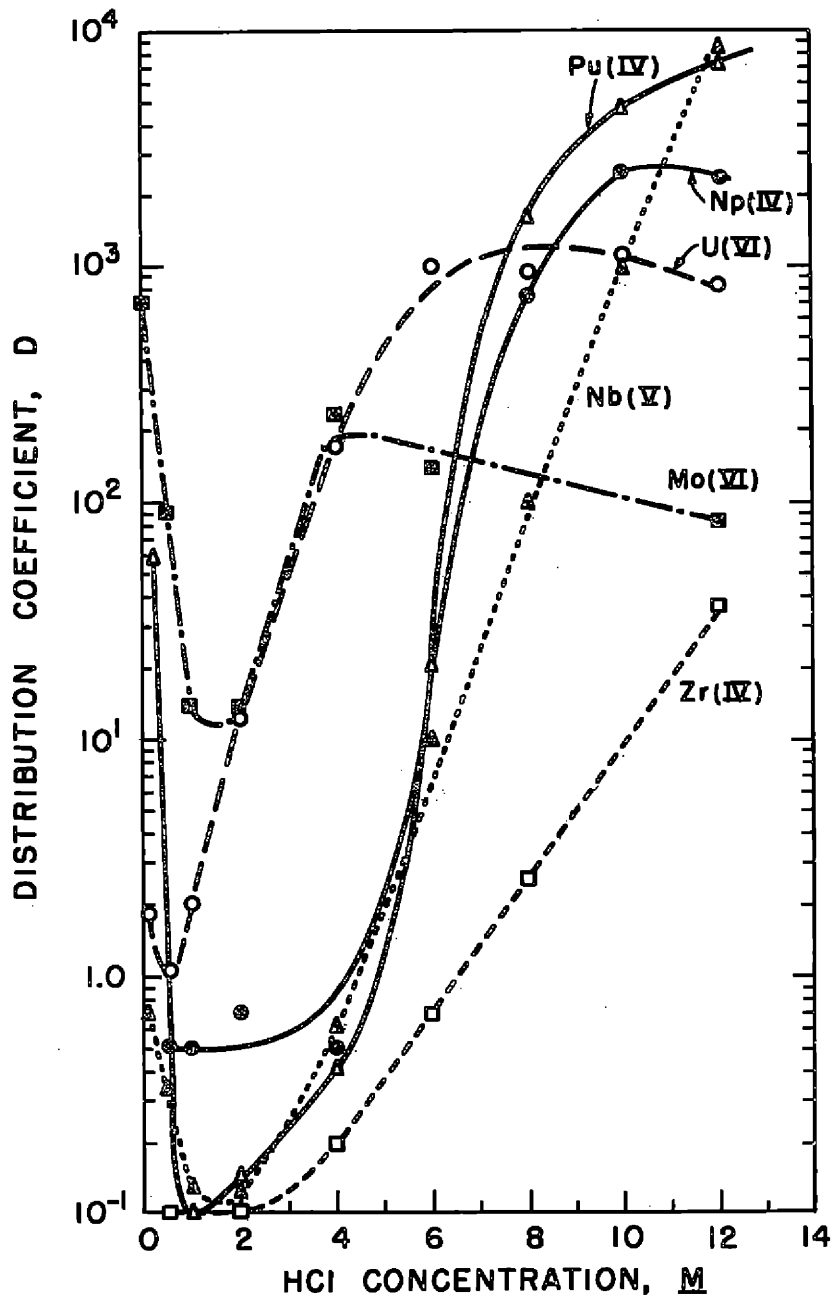


Figure 67.

Adsorption of elements from hydrochloric acid solution with Dowex-2 anion exchange resin (x8, 200-400 mesh).

After L. Wish, reference 538.

efficients of these elements vary roughly between 1 and 10 from 80% alcohol solutions containing HCl in the range of 0.2M to 2.4M.⁵⁴²

Numerous separations of uranium from other elements are possible using hydrochloric acid systems. The more obvious ones are those in which uranium is adsorbed and the other element is not. Consideration of figure 65, indicates that uranium can be separated from alkali metals, alkaline earths, aluminum, yttrium, rare earths, actinium, and thorium by adsorption as uranium (VI) on a strong base anion exchange resin from a concentrated hydrochloric acid solution. Trivalent actinide elements are not adsorbed from hydrochloric acid solutions. Plutonium is eluted as Pu (III) with 12M HCl containing hydroxylamine hydrochloride and NH_4I , NH_4I alone, or HI. Separations may be made by adsorption of the contaminating element and elution of uranium with dilute hydrochloric acid. For example, molybdenum is adsorbed from 0.1M HCl.⁵⁴³ Bismuth is also adsorbed from dilute (<1M) HCl.^{544,545} Other elements that show strong adsorption from dilute HCl include many of the transition metals, tin, tellurium, and polonium.⁵³⁶ Kraus and Moore⁵⁴⁶ have effected the separation of protactinium and uranium by adsorbing them from 8M HCl on a column of Dowex A-1 resin and developing the column with 3.8M HCl. Protactinium appeared first in the eluent, separated from uranium. The uranium fraction contained, however, a fair amount of protactinium 'tailing'.

Advantage may be taken of the different distribution coefficients exhibited by ions in various oxidation states to effect their separation from uranium. Iron reduced to ferrous ion by hydrogen iodide⁵⁴⁷ or ascorbic acid⁵⁴⁸ is separated by elution with 4M HCl. U(IV) may be separated from Pa(IV) and Th(IV). U(IV) is adsorbed by Amberlite IRA-401 (100 mesh) and Dowex-1 (100-200 mesh) from >8M HCl.

Neither Pa(IV) nor Th(IV) is adsorbed from $6M - 12.6M$ HCl.⁵⁴⁹ The elution of Pu(III) by $12M$ HCl from strong base anion exchange resin has already been mentioned. Wish and Rowell⁵⁵⁰ have effected the separation of Th, Pu, Zr, and Np from U by elution with hydrochloric acid in a sequence of concentrations. The elements are adsorbed on the resin (Dowex-2) from $12M$ HCl. Thorium does not adsorb. Plutonium is eluted in the trivalent state with $12M$ HCl saturated with hydroxylamine hydrochloride and ammonium iodide. Zirconium is eluted with $7.5M$ HCl; neptunium (IV) with a $6M$ HCl - 5% $NH_2OH \cdot HCl$ solution. Uranium is finally eluted with $0.1N$ HCl.

Korkisch, et al.⁵⁵¹ have separated uranium from tungsten by means of anion exchange. The uranium is adsorbed on Dowex-1 resin from a solution containing 20% $4M$ HCl and 80% ethanol (volume %). Ascorbic acid is used to reduce any iron present. The resin is washed with a similar solution and uranium is eluted with an ether-saturated $0.1M$ HCl solution. No tungsten is observed in the final eluate.

Fluoride systems.

Faris⁵⁵² has reported the adsorption of elements from hydrofluoric acid solutions with Dowex-1 anion exchange resin ($\times 10$, 200 mesh). His results are shown in figure 68. Uranium (VI) adsorption is strong from dilute HF solutions and decreases with increased acid concentration. Separation from elements exhibiting no or strong adsorption from HF solutions may be achieved by proper selection of the acid concentration. Elements such as Be, B, Sc, Ti, Zr, Mo, Sn, Te, Hf, Ta, W, Re, and Hg have adsorption curves similar in shape to that of uranium (VI). Separation from these elements using an HF system should prove difficult to almost impossible, depending upon the distribution coefficients involved.

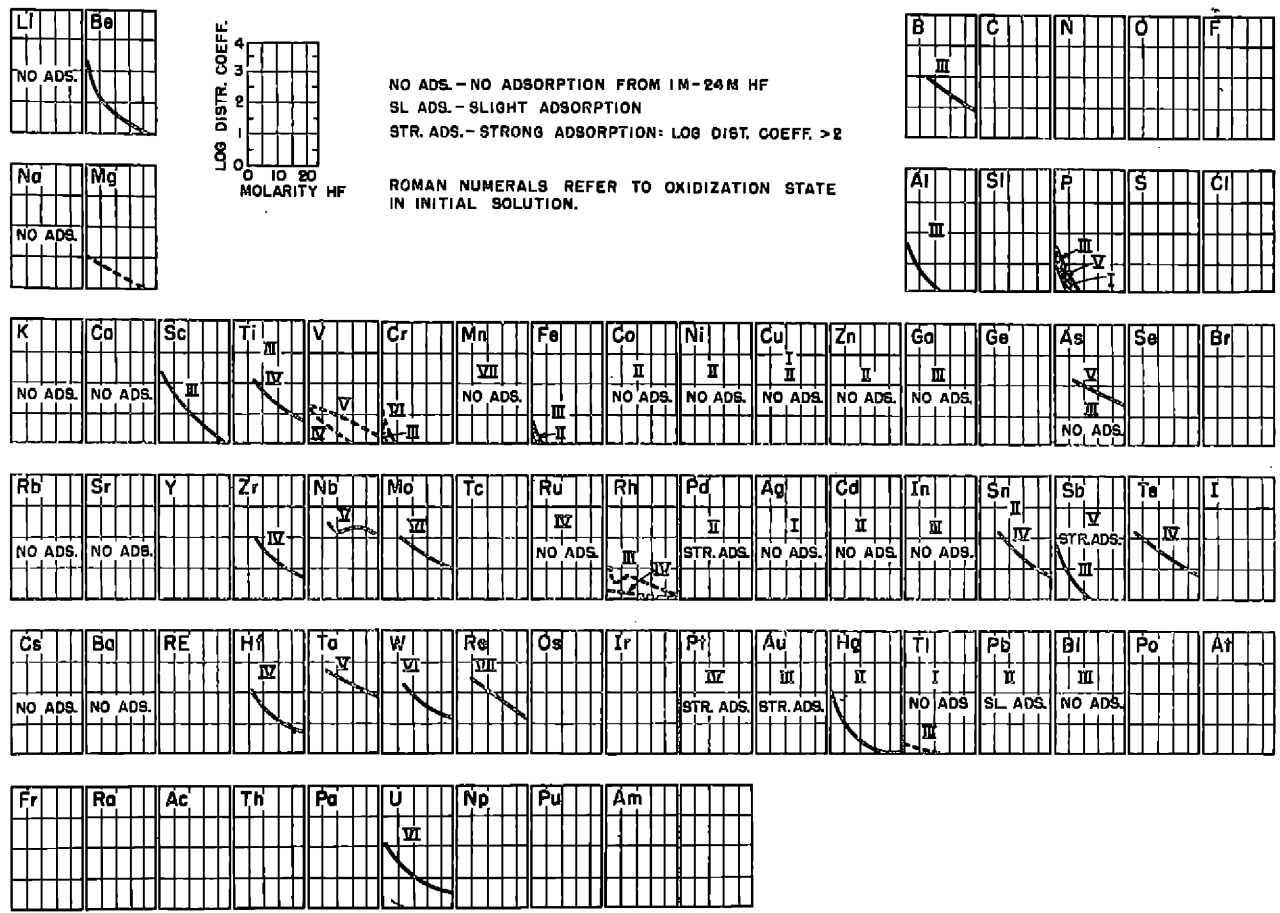


Figure 68. Adsorption of elements from hydrofluoric acid solution with Dowex-1 anion exchange resin (x10, 200 mesh). After J. P. Paris, reference 552.

Bhat and Gokhale⁵⁵³ have found evidence for the adsorption of the anionic species UO_2F_3^- with Amberlite IRA-400. HCl-HF systems.

Certain elements are efficiently separated from uranium by anion exchange when a combined HCl-HF eluting system is used. Such systems have been studied by a number of workers.^{538,554-559} The results of Nelson, Rush, and Kraus⁵⁵⁸ are shown in figure 69. Faris and Brody⁵⁵⁹ have examined the distribution coefficient D of uranium as a function of HCl concentration in the presence of 0, 1, and 8M HF and as a function of HF concentration in the presence of 0 and 0.2M HCl. The former three curves are similar in shape but decrease in magnitude as the HF concentration is increased. The presence of 0.2M HCl also causes a decrease in magnitude of the D vs.[HF] curve. However, the shapes of the 0 and 0.2M HCl curves for varying HF concentration are dissimilar for HF concentrations less than 4M.

Table XXXIV lists a number of separations of U from other elements using HCl-HF eluting solutions.

Nitrate systems.

The distribution of uranium between anion exchange resins and nitric acid solutions has been reported by a number of workers.^{536,540,543,550,557,560-562} The results of Buchanan and Faris⁵⁶² are given in figure 70. From the non- or only slight adsorption of most of the elements from nitric acid media, it appears that anion exchange affords an excellent means for purifying uranium. Uranium is adsorbed more strongly from nitrate salt solutions than from nitric acid solutions alone.^{536,557,562,563} With DeAcidite FF resin, the adsorption of uranium (VI) is greatest from $\text{Al}(\text{NO}_3)_3$ solutions and decreases in the order $\text{Ca}(\text{NO}_3)_2 > \text{LiNO}_3 > \text{NH}_4\text{NO}_3$.⁵⁶³ Ethanol increases the distribution of uranium to the resin phase.⁵⁴¹ With an 80% alcoholic solution, the distribution coefficient

Table XXXIV. Separation of Uranium from Various Elements by Anion Exchange Using HCl-HF Eluting Solutions.^a

Elemental mixture	Element eluted	Eluting solution	Reference
W, U	W	7M HCl-1M HF	555
	U	0.1M HCl	
U, W, Mo	U	0.5M HCl	555
	W	7M HCl-1M HF	
	Mo	1M HCl	
W, Nb, Ti, V, Zr, U, Ta	W, Ti, V, Zr	7M HCl-4M HF	559
	Nb	7M HCl-0.2M HF	
	U	1M HCl-4M HF	
	Ta	24M HF or 4M NH ₄ Cl- 1M NH ₄ F	
Fe(III), U	Fe(III)	1M HF-0.01M HCl	555
	U	1M HCl	
R.E. as Eu(III), U(IV), U(VI), Zn(II)	R.E.	8M HCl	556
	U(IV)	8M HCl-0.1M HF	
	U(VI)	0.5M HCl	
	Zn(II)	0.01M HCl	
Te(IV), U(VI)	U(VI)	3M HCl-1M (to 8M) HF	559
	Te(IV)	1M HCl	
Th(IV), Pa(V), U(VI)	Th(IV)	10M HCl	556
	Pa(V)	9M HCl-1M HF	
	U(VI)	0.1M HCl	
Pa, U		9M HCl	554
	Pa	7M HCl-0.11M HF	
	U	0.5M HCl	
Zr, Np, Nb, U, Mo, Tc	Zr	12M HCl-0.06M HF	538
	Np	6.5M HCl-0.004M HF	
	Nb	6.0M HCl-0.06M HF	
	U	0.1M HCl-0.06M HF	
	Mo, Tc	12M HNO ₃	

^a Dowex-1 or -2 anion exchange resin used.

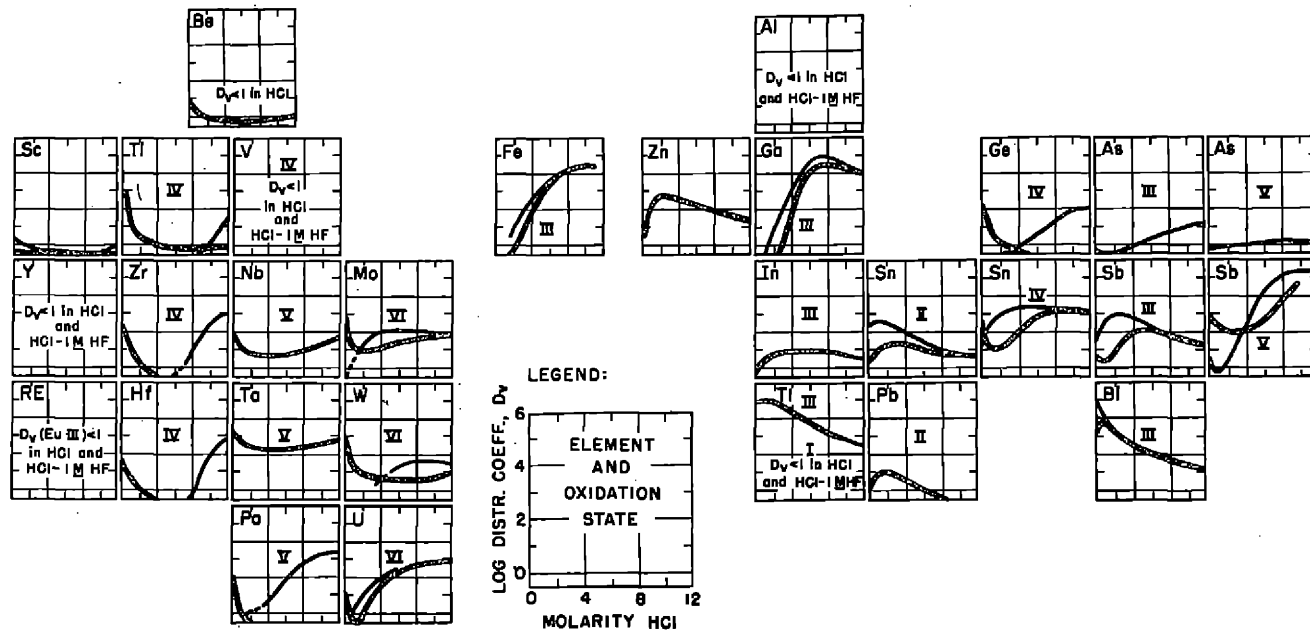


Figure 69. Adsorption of elements from HCl and HCl-HF solutions with anion exchange resin. — Distribution coefficients in absence of HF. - - - Distribution coefficients in HCl-HF mixtures (usually 1M HF except Zr(IV), Hf(IV), Nb(V), Ta(V), and Pa(V) where $M_{HF} = 0.5$). After F. Nelson, R. M. Rush, and K. A. Kraus, reference 558.

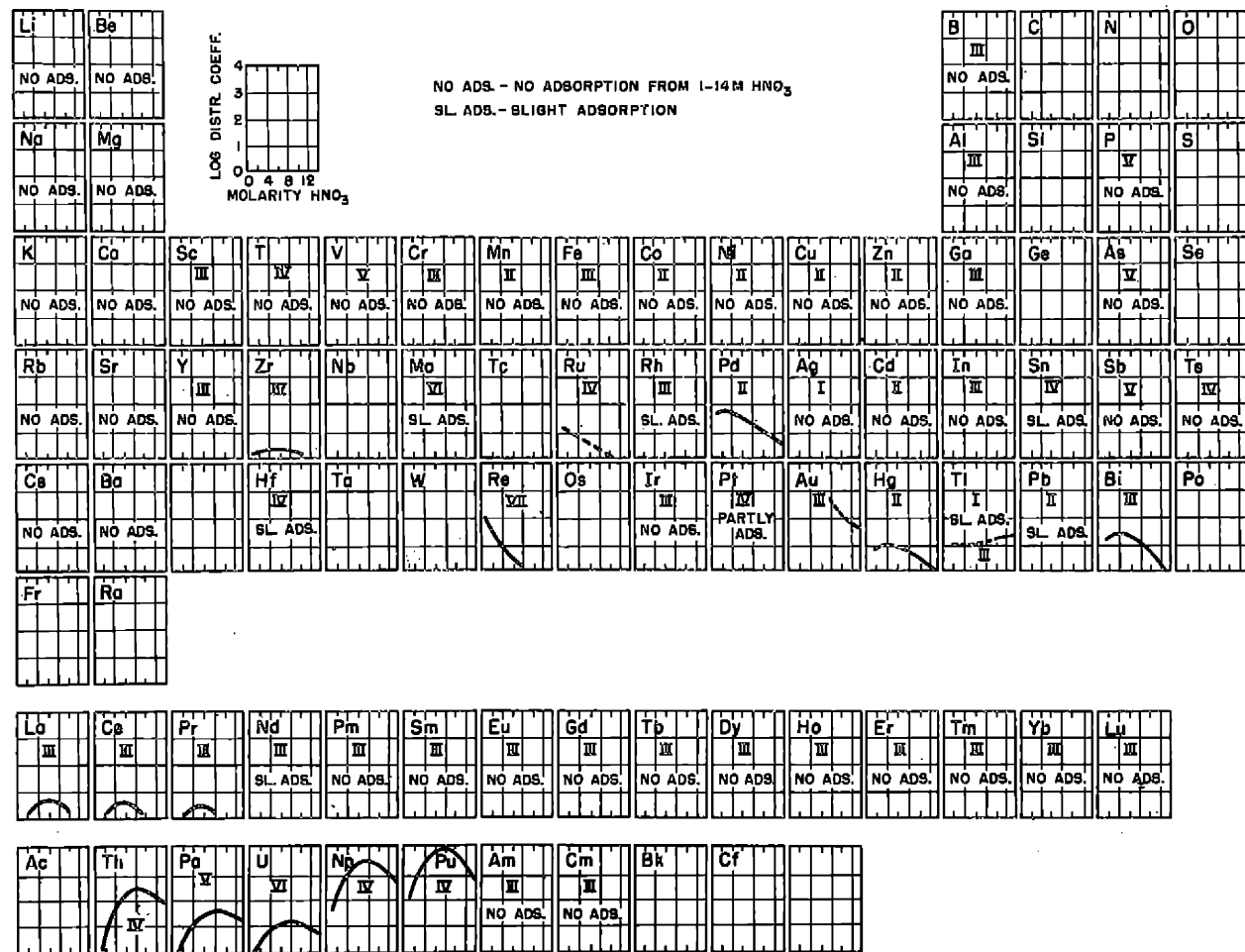


Figure 70. The adsorption of elements from nitric acid solutions with Dowex-1 anion exchange resin (x10, 200-400 mesh). After R. F. Buchanan and J. P. Paris, reference 562.

Table XXXV. Separation of Uranium from Various Elements by Anion Exchange Using Nitric Acid Solutions.

Elemental Mixture	Resin	Element Eluted	Eluting Solution	Reference
Eu(III), U(VI)	Dowex-1	Eu(III) U(VI)	8M HNO ₃ 0.2M HNO ₃	557
U, Ru	Dowex-2	U	1M HNO ₃	543
Zr, U, Th	Dowex-2	Zr U Th	Column developed with 8M HNO ₃ . Zr elutes first, followed by U and then Th.	543
U, Th	DeAcidite FF (77°C)	U Th	6M HNO ₃ 4M HNO ₃ H ₂ O	560
U, Th	Dowex-1	U Th	90% methanol-1M HNO ₃ 1M HNO ₃	564
U, Np	Dowex-1 or Dowex-21K	U Np	6M HNO ₃ -ferrous sulfamate - hydrazine or semicarbazide 0.35M HNO ₃	565
Th, R.E., trans-Pu, Pu, U, Zr, Np	Dowex-2	Th, R.E., trans-Pu Pu - U, Zr - Np	conc.HCl-trace HNO ₃ conc.HCl sat'd with NH ₂ OH·HCl and NH ₄ I conc.HCl(80%)-conc.HNO ₃ (20%) 12M HNO ₃ conc. HCl 4M HCl with 5% NH ₂ OH·HCl	550

for uranium is about 12 between Dowex-1 resin and 1.2M HNO₃ solution; for a 40% alcoholic solution, D is about 7.

Table XXXV lists a number of separations of U from other elements using nitrate media. The last procedure listed in the table may be revised to include Pa separation. Following the elution of U with 12M HNO₃, in which a small amount of Pa is eluted, the remaining Pa is eluted with 12M HNO₃-0.1M HF. 550

Sulfate systems.

The recovery of uranium from sulfate liquors by anion exchange methods is important industrially. Laboratory-wise, a number of procedures have been developed for the determination of uranium that make use of the adsorbability of the anionic uranyl sulfate complexes. The nature of these complexes has also received considerable study^{98,566,567} (see section on complex ions). The distribution of uranium between anion exchange resins and sulfate media has been reported by a number of investigators.^{536,543,550,557,561,568} The results of Bunney, *et al.*⁵⁴³ are given in figure 71. Strontium, yttrium, cerium, and americium do not show any significant adsorption by the resin (Dowex-2) at any acid concentration.⁵⁴³ The distribution coefficient of Pu(IV) is approximately twice that of U(VI) in the acid range of 0 to 10N H₂SO₄.⁵⁵⁰ The adsorption of uranium (and thorium) from solutions of (NH₄)₂SO₄ is similar to that from H₂SO₄. The decrease in adsorption is less rapid, however, with increased ammonium sulfate concentration than with sulfuric acid.^{536,557,568} The adsorption of uranium from sulfate solution exhibits a pH dependence (D increases as the pH is increased from 1 to 4) which decreases as the sulfate concentration is decreased.⁵³⁴ The distribution coefficient of uranium between Dowex-1 resin and 0 to 1.2N sulfuric acid solutions is one to three orders of magnitude greater from 80% ethanol solutions than from aqueous solutions.⁵⁴¹

A number of procedures have been developed for the separation of uranium from various elements by anion exchange in sulfate solution. These generally involve the adsorption of the uranyl complex from a sulfate solution at pH 1 to 2 from which the foreign element is not adsorbed. After thoroughly washing the resin bed to remove impurities, uranium is eluted with a dilute solution of hydrochloric, nitric, or perchloric acid. Mixtures of elements that have been or may be separated

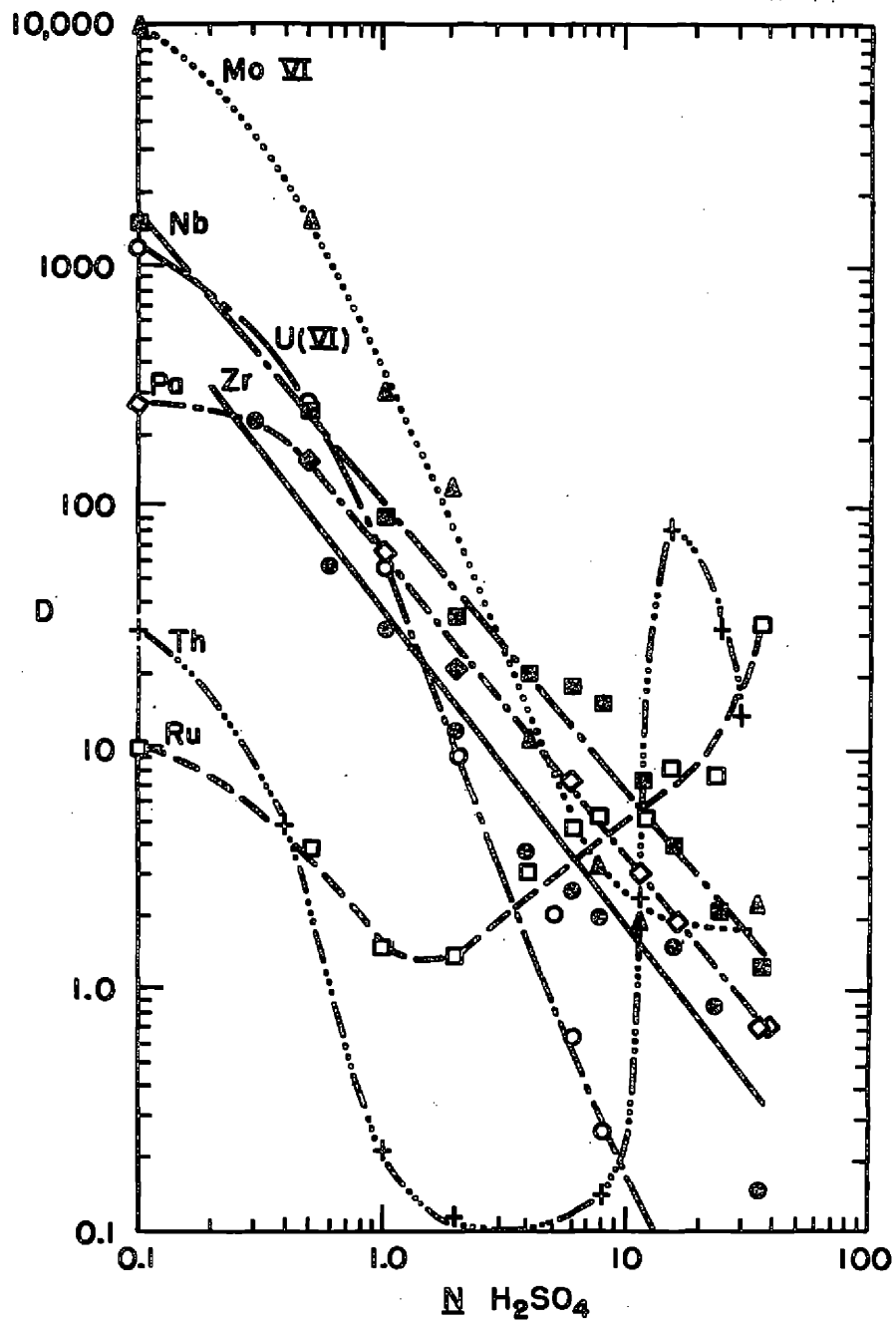


Figure 71.

The adsorption of elements from sulfuric acid solutions with Dowex-2 anion exchange resin (x8, 200-400 mesh).

After L. R. Bunney, N. E. Ballou, J. Pascual, S. Foti, reference 543.

from uranium by procedures similar to the one described include Am, Th;⁵⁴³ Fe, Al, Mg;⁵⁶⁹ Fe, V;⁵⁷⁰ Zn, Ni, Co, Cd, Mn, Cu, Fe, alkali metals;⁵⁷¹ Zr, Ce, Cs, Ag, Cd, V;⁵⁷² Bi;⁵⁷³ R.E.;^{543,571,574} and metals contained in ores.^{575,578} The literature references should be consulted for exact experimental conditions.

Carbonate systems.

Uranium is recovered from carbonate leach liquors by anion exchange in industrial operations. Its anion exchange behavior is similar in carbonate solution to that in sulfate solution. That is, the distribution coefficient is decreased with increased carbonate concentration. This is illustrated in figure 72 for ammonium carbonate solutions.⁵⁷⁵ A similar decrease in D is observed for increased sodium carbonate concentration.⁵⁶⁸ The distribution coefficient is also decreased by a decrease in pH of the solution. The increase in bicarbonate concentration at the lower pH interferes with uranium adsorption. Other anions such as sulfate, nitrate, and chloride may also interfere with uranium adsorption from carbonate solution. To prevent gassing with carbon dioxide, uranium is eluted with salt solutions rather than acids.

Vanadium⁵⁷⁶ and phosphate and molybdate⁵⁷² have been separated from uranium in carbonate solutions by anion exchange. The impurities are adsorbed on the resin together with uranium and eluted with a 10% Na_2CO_3 solution. Uranium is eluted with a 5% NaCl solution.

Phosphate systems.

The distribution of uranium and other elements between Dowex-2 anion exchange resin and phosphoric acid solutions is represented in figure 73.⁵⁷⁹ Marcus⁹¹ has studied the Dowex 1-uranyl phosphate system. His distribution coefficients are lower by factors of 2 to 50 at 0.1M H_3PO_4 and 3M H_3PO_4 , respectively,⁵⁷⁹ than those shown in figure 73.

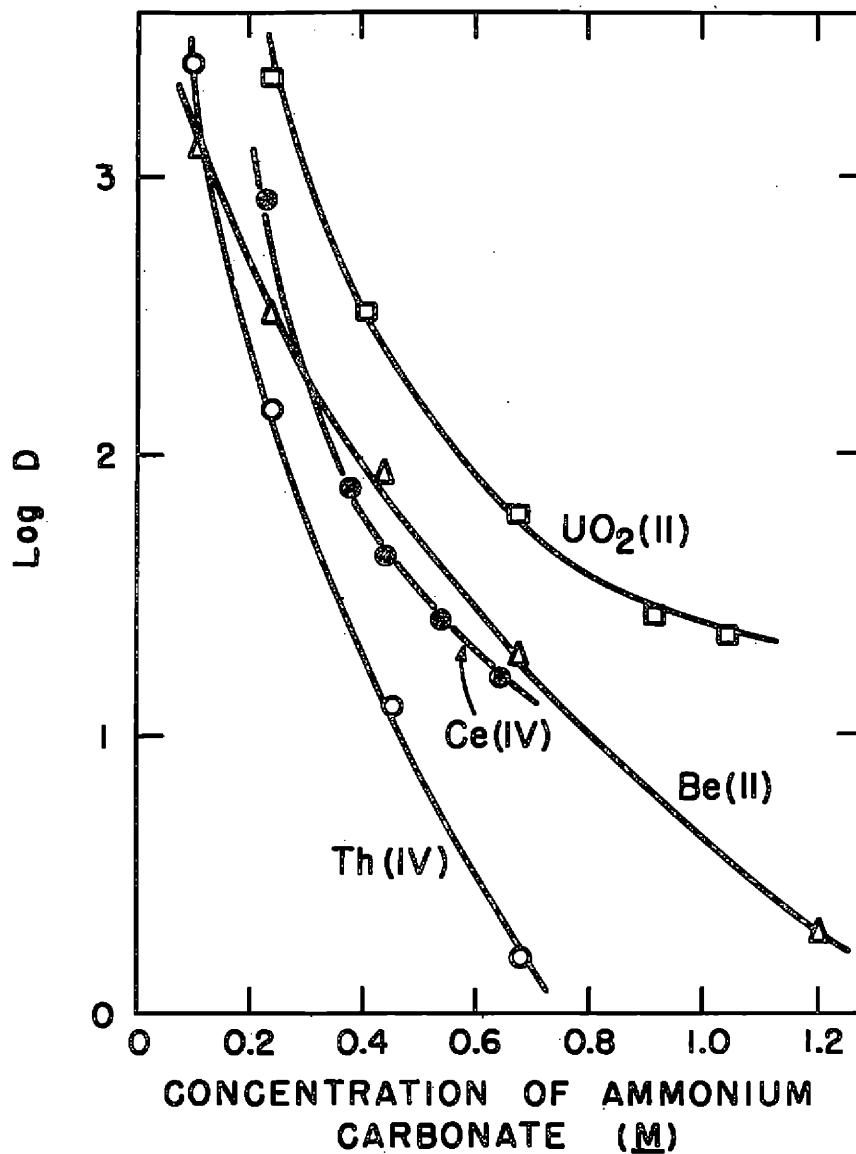


Figure 72. The adsorption of elements from ammonium carbonate solutions with Dowex-1 anion exchange resin (x8, 50-100 mesh). After S. Misumi, T. Taketatsu, reference 575. Conditions: Amounts taken, Be^{2+} , 11.7 mg (BeO); Ce^{4+} , 7.3 mg (CeO_2); Th^{4+} , 26.9 mg (ThO_2); and UO_2^{2+} , 61.0 mg (U_3O_8). 1 gram of resin and 200 ml of solution in contact 12-20 hours at 20°C.

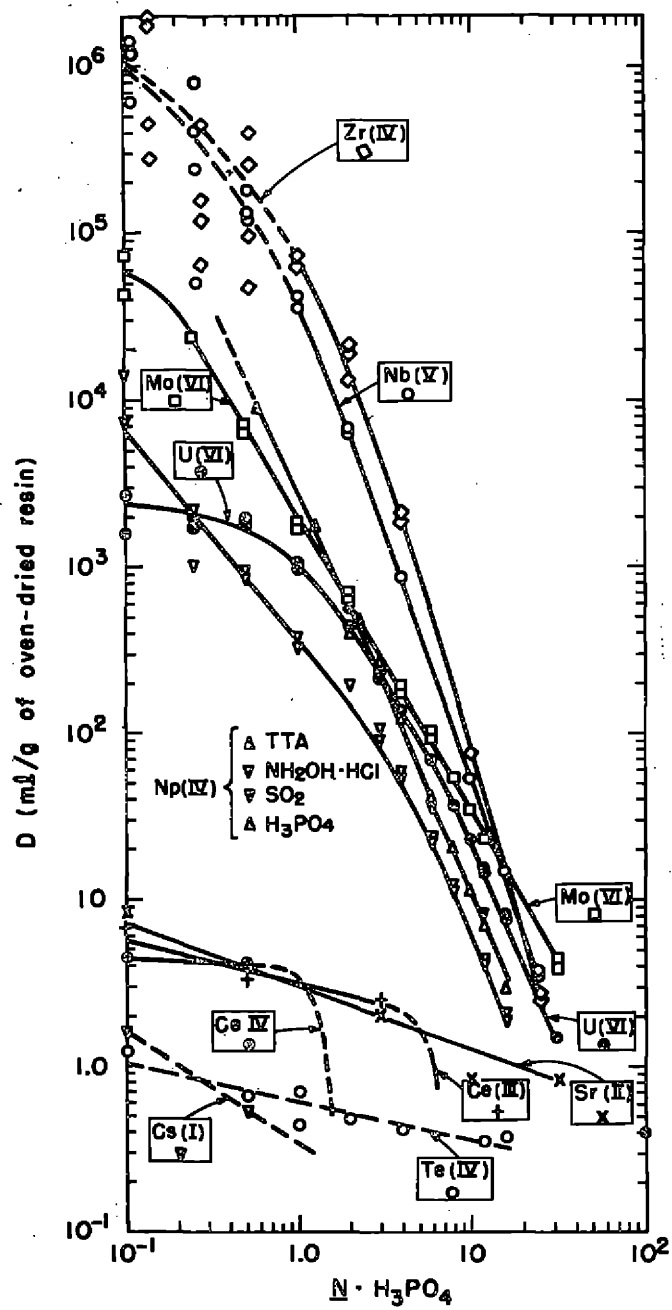


Figure 73.

The adsorption of elements from phosphoric acid solutions with Dowex-2 anion exchange resin (x8, 200 mesh).

After E. C. Freiling, J. Pascual, and A. A. Delucchi, reference 579.

Separations between uranium and the rare earths (Ce^{3+} , Ce^{4+}), alkaline earths (Sr^{2+}), alkali metals (Cs^+), and tellurium are possible using Dowex 2- H_3PO_4 systems. The latter separation has been used by Wish⁵⁸⁰ who loaded a column of Dowex-2 resin from a 0.1N H_3PO_4 solution. Tellurium passed through and uranium was adsorbed. The column was then converted to the chloride form with concentrated hydrochloric acid and uranium was eluted with 0.1M HCl - 0.06M HF. An alternative method involved loading the column from concentrated HCl. Tellurium was eluted with 1.0N H_3PO_4 after washing the column with an alcoholic phosphoric solution. The column was then washed with an alcoholic HCl gas solution. Uranium was eluted with a 0.1M HCl - 0.06M HF solution and molybdenum with 12M HNO_3 .

Miscellaneous systems.

Uranyl ion forms an anionic complex with acetate ions at pH 4.25 to 5.25. The complex has been adsorbed on Amberlite IRA-400 strong base resin in the determination of small amounts of uranium in stones and natural waters,^{581,582}

Uranium is also adsorbed on Amberlite IRA-400 resin as an ascorbate complex.^{583,584} Thorium, titanium, zirconium, tungsten, and molybdenum are also adsorbed.

Uranium complexed with sulfosalicylic acid has been separated from Zn, Cu, Ni, and Cd.⁵⁸⁵ The latter are complexed with EDTA. The pH of the solution is kept between 8 and 10. Separation has been made on Amberlite IRA-401 and Dowex-1 resins.

The uranyl cyanate complex formed by adding potassium cyanate solution to a uranyl salt is adsorbed by Dowex-1 anion exchange resin.⁵⁸⁶ Uranium is eluted by a dilute hydrochloric acid solution.

Cation exchange. Although a number of separations of uranium from various elements have been reported in the literature, the amount of quantitative data reported is rather meager. Hardy⁵⁸⁷ has summarized much of the data available on the dis-

tribution of uranium between cation exchange resins and nitric and hydrochloric acid solutions. His curves are reproduced in figure 74. Distribution coefficient curves for other actinide elements: Th, Pa, Np, Pu, are given for comparison. Prevot, et al.⁵⁴⁰ have published the distribution curve for uranium between cation exchange resin C.50 and nitric acid solution. Its shape and magnitude are similar to that shown for the Zeokarb-225-HNO₃ system in figure 74. Ishimori and Okuno⁵⁶⁸ have found that increasing amounts of methanol in nitric acid solution (0.18M) increase the distribution coefficient of uranium for Dowex-50 resin.

The elution peak positions of a number of ions including uranium (IV) and (VI) are given in figure 75 for various hydrochloric acid concentrations. The conditions under which the peak positions were determined are described in the figure caption. Ionescu, et al.⁵⁸⁸ have studied the effect of acetone on the distribution of several elements between cation exchange resins, KU-2 and R-21, and dilute hydrochloric acid solutions. For fixed hydrochloric acid concentrations of 1, 2, and 3%, maximum uranyl distribution coefficients were found between 60 and 80% acetone solutions. For 4 and 5% acid solutions, D was found to be considerably lower than for the more dilute acid solutions.

Ishimori and Okuno⁵⁶⁸ have investigated a number of cation exchange systems other than those already noted. Some of their results, D versus sodium sulfate, sodium acetate, and oxalic acid concentration, are illustrated in figure 76. The adsorption of uranium by Dowex-50 from solutions of hydroxylamine was found to be pH dependent. As the pH of the solution was increased, a sharp decrease in D was observed between pH 5 and 6. Uranyl ion was not adsorbed by Dowex-50 resin from carbonate solution.

Khopkar and De⁵⁸⁹ have investigated the behavior of

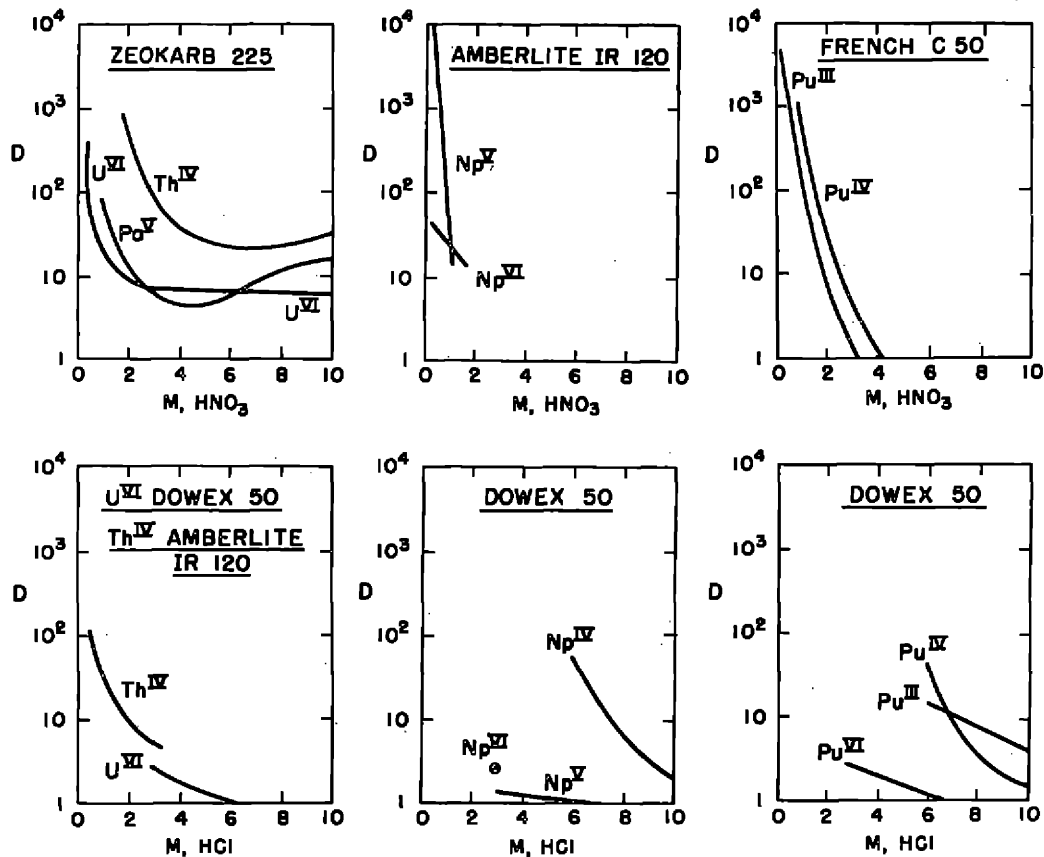


Figure 74.

The adsorption of uranium and other actinide elements by cation exchange resins from nitric and hydrochloric acid solutions.

After C. J. Hardy, reference 587.

The curves present the data of the following investigators:

C. J. Hardy, D. Scargill, J. M. Fletcher, *J. Inorg. Nuclear Chem.* **7**, 257 (1958): Th(IV), Pa(V), U(VI)-HNO₃.

K. F. Sculz, M. J. Herak, *Croat. Chem. Acta* **29**, 49 (1957): Th(IV)-HCl.

R. M. Diamond, K. Street, Jr., G. T. Seaborg, *J. Am. Chem. Soc.* **76**, 1461 (1954): U(VI), Np(IV), (V), (VI), Pu(III), (IV), (VI)-HCl.

M. Ward, G. A. Welch, Unpublished data, U.K.A.E.A., Windscale: Np(V), (VI)-HNO₃.

I. Prevot, P. Régnaut, *Progress in Nuclear Energy, Series III*, **2**, 377 (1958): Pu(III), (IV)-HNO₃.

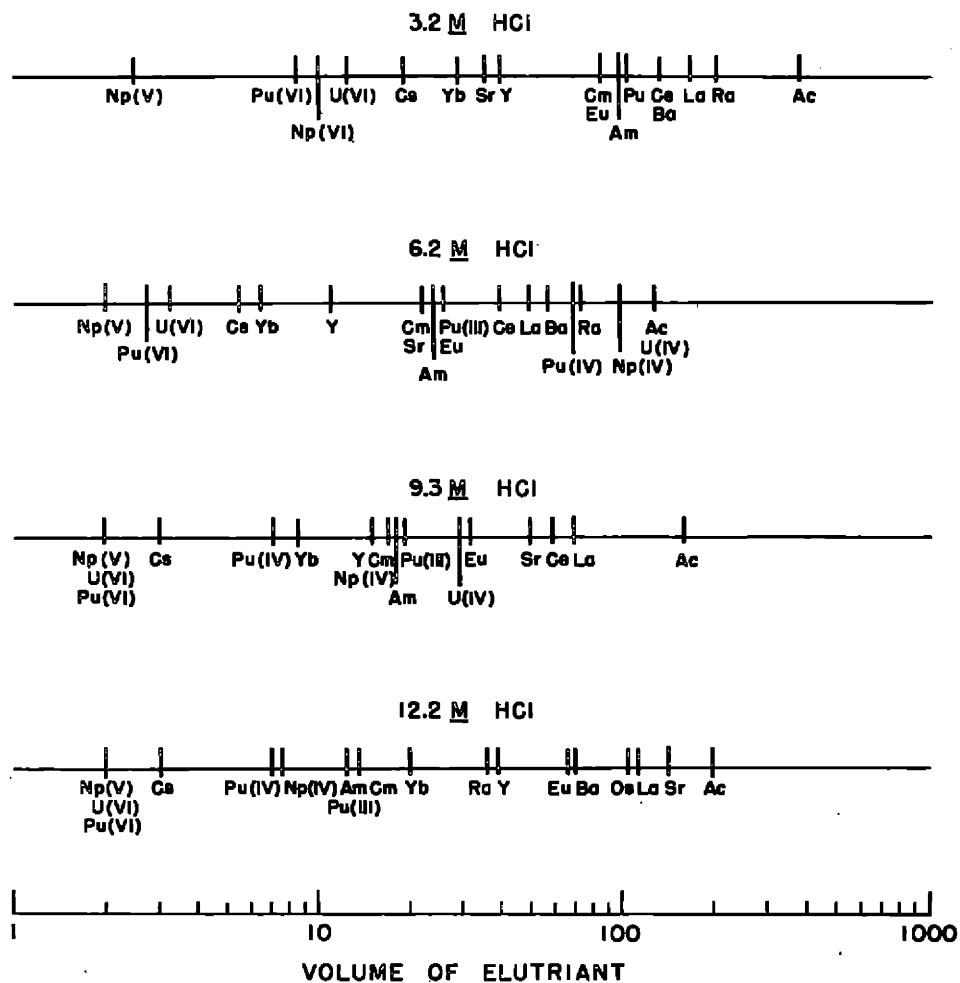


Figure 75.

Elution peak positions of various ions with 3.2, 6.2, 9.3, and 12.2M HCl from Dowex-50 cation exchange resin. Elutrient volume given in drops.

After R. M. Diamond, K. Street, Jr., and G. T. Seaborg, reference 591.

Conditions:

Dowex-50 resin, HR form, 250-500 mesh, settling rate approximately 0.5 cm/min.
 Column, 10 cm long x 1 mm diameter.
 Flow rate, approximately 0.1 cm/min.
 Room temperature.

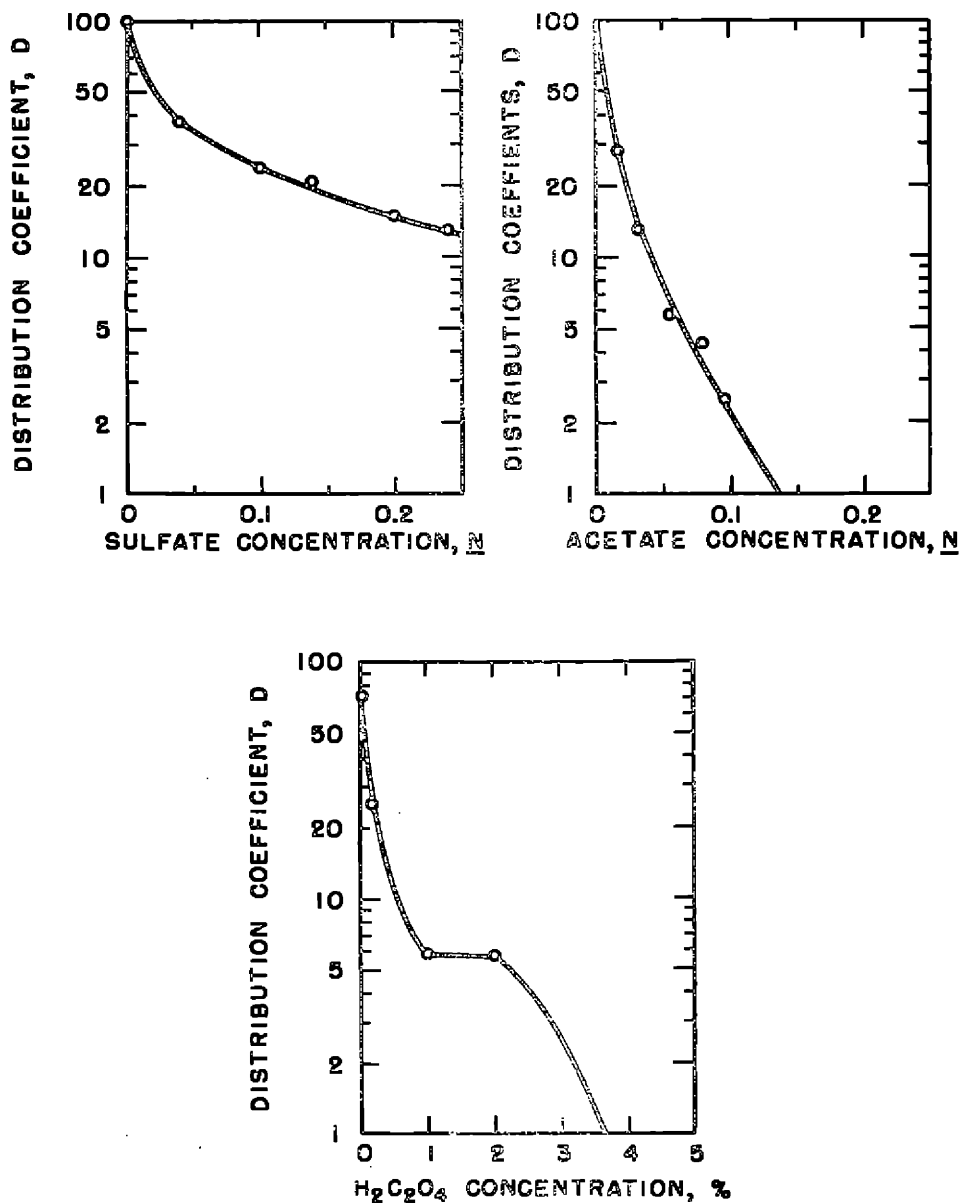


Figure 76. The adsorption of uranium by Dowex-50 cation exchange resin from solutions of sodium sulfate, sodium acetate, and oxalic acid. After T. Ishimori and H. Okuno, reference 568.

Conditions:

Sulfate - 0.5 g resin, NaR form; 5.4 mg U/25 ml; Na₂SO₄-NaNO₃ mixed solution, [Na] = 0.30N.
 Acetate - 0.5 g resin, NaR form; 5.4 mg U/25 ml; NaOAc-NaNO₃ mixed solution, [Na] = 0.16N.
 Oxalic acid - 0.5 g resin, HR form; 10.8 mg U/25 ml.

uranium (VI) on Amberlite IR-120 cation exchange resin with hydrochloric, nitric, sulfuric, acetic, citric, and perchloric acids. Uranium (1.7 mg) adsorbed on a resin bed (1.4 x 14.5 cm) was eluted with 200 ml of various eluants. Uranium was quantitatively recovered with 2-4M HCl, 2-4M HNO₃, and 1-2M H₂SO₄. Uranium was incompletely recovered with 1M HCl, 1M HNO₃, 2M HClO₄, 2M acetic acid, and 2-5% citric acid.

Sullivan, *et al.*⁵⁹⁰ have investigated the distribution of uranium between Dowex-50 ion exchange resin and perchloric acid media as a function of time and bisulfate ion concentration.

Table XXXVI lists a number of separations of uranium from various elements that have been achieved by cation exchange.

4. Chromatography. The subject of paper and cellulose chromatography for the separation of uranium has been reviewed by Rodden¹⁹² and by Steele and Taverner.¹⁹⁸ Work of Soviet scientists in the field has been reviewed by Palei¹⁹⁷ and by Senyavin.⁶¹² References to much of the literature may be found in the review article by Kuznetsov, Savvin, and Mikhailov.²⁰⁰ Books by Pollard⁶¹³ and by Blasius⁶¹⁴ include chromatographic separations of uranium.

One of the most successful separations of uranium by filter-paper chromatography makes use of the solvent, 2-methyltetrahydrofuran.⁶¹⁵ Of thirty-one metals tested, only ruthenium and rhodium, measured as Ru¹⁰⁶-Rh¹⁰⁶, and tungsten (W¹⁸⁵) were incompletely separated from uranium (U²³³). The results for tin (Sn¹¹³) and antimony (Sb¹²⁴) were inconclusive and the behavior of mercury (Hg²⁰³) was similar to that of uranium.

An example of the use of cellulose columns in combination with organic solvent for the separation of uranium is that given by Burstall and Wells.⁶¹⁶ An ethereal solution containing 5 per cent v/v of nitric acid is used to extract uranium from a cellulose column. The nitrates of the alkali metals, alkaline earths, rare earths, Cu, Ag, Zn, Cd, Al, In, Tl, Tl, Hf, Ge, Sn, Y,

Table XXXVI. Separation of Uranium from Various Elements by Cation Exchange.

Elemental mixture	Element eluted + eluting agent	Resin	Reference
U, Th	U-1.70M HCl; Th-1.1M (NH ₄) ₂ CO ₃	Amberlite IR-120	592
U, Th	U, Th-90% acetone, 5% HCl, H ₂ O (U eluted first)	KU-2	588
U, Th	U-1M HCl; Th-3M H ₂ SO ₄	Amberlite IR-120	589
U, Th	U-dil. HCl; Th-complexing agent (HSO ₄ ⁻)	phenol-formaldehyde type	593
U, Th	U-2M HCl; 0.5M H ₂ C ₂ O ₄	Wofatit KS Amberlite IR-100	594
U, Th	dilute HNO ₃	alginic acid Wofatit KS	595
U, Th	U-0.1 to 0.4M H ₂ SO ₄ ; Th-0.5M H ₂ C ₂ O ₄		596
U, Np(IV)	U-dil. HCl; Np(IV)-complexing agent	Amberlite IR-120	597
U, Np	adsorb from 1M HNO ₃ ; eluted Np ahead of U with 2M HNO ₃	Amberlite IR-120	597
U, F.P.(Cs,Sr,Y,Ce)	Sr-0.1M HCl; Ce,Y-1M HCl; U-6M HCl; Cs-5M NH ₄ Cl	zirconium phosphate	598
U, F.P., Pu	adsorb from uranyl nitrate solution at pH 1-3; U-0.2 to 0.3M H ₂ SO ₄ ; F.P.-phosphoric acid and 1M HNO ₃ ; Pu-0.8M H ₃ PO ₄ and 1M HNO ₃	sulfonated phenol formaldehyde type	599
U, F.P.	F.P.-3% Na ₂ EDTA; U-3% NaOAc, 0.25M Na ₂ CO ₃	sodium dialkyl phosphate	600
U, La	La-0.06% Na ₂ EDTA, pH 4.0; U-3% NaOAc, 0.25M Na ₂ CO ₃	sodium dialkyl phosphate	600
U, Eu	U-0.75M H ₂ SO ₄ ; Eu-6M HCl	Dowex-50	601,602
U, Ce, Eu, Y	U-1M H ₂ C ₂ O ₄ ; Ce, Eu, Y-5M HCl	Amberlite IR-120	603
U, R.E.	R.E.-Na ₂ EDTA	Amberlite IRC-50	604
U, R.E.	U-2.5M HF	Dowex-50	605
U, Fe, Cu, Cd, Ni, Co, Mn, R.E.	U, Fe,Cu-0.5N H ₂ C ₂ O ₄ ; Cd,Ni,Co,Mn-1M HCl; R.E.-5% ammonium citrate	Amberlite IR-120	606

U, Fe(III), Co(II), Cu(II)	Fe, Co, Cu-2% Na ₂ EDTA, pH 3.0; U-3% NaOAc, 0.25M Na ₂ CO ₃	sodium dialkyl phosphate	600
U, Fe(III)	U, Fe-0.8M HCl (U eluted first)	Lewatit S100	607
U, Co, Cu, Ca	adsorb from 0.1M HNO ₃ ; Co, Cu, Ca-0.2M HNO ₃ ; U-resin removed from column and washed with 0.25M Na ₂ CO ₃	dialkyl- phosphoric acid	600
U, Cd	Cd-0.5N HCl	Dowex-50	608
U, Th, Ac, Bi, Ra, Pb	U, Th, Ac, Bi-5% H ₂ C ₂ O ₄	KU-2	609
U, many ions	foreign ions-EDTA (Na salt); pH 7; U-H ₂ SO ₄	Amberlite IRC-50	610
U, many ions	foreign ions-EDTA (Na salt), (a)pH 1.7-1.9 or (b) pH 5.5-7.0 following Fe(OH) ₃ precipitation; U-3N H ₂ SO ₄	(a)KU-2 or (b) Amberlite IRC-50	611
U, Zr, Ce(III), Cu, Ni, Hg(II)	Zr as anionic oxalate complex, Ce(III), Cu, Ni as anionic EDTA complexes, Hg(II) as anionic iodide complex are not adsorbed; U-4M HCl	Amberlite IR-120	589
U, phosphate	phosphate (Na ₂ HPO ₄) not adsorbed; U-4M HCl	Amberlite IR-120	589

F.P. [=] fission products.

R.E. [=] rare earth elements.

EDTA [=] ethylenediaminetetraacetic acid.

Pb, Nb, Ta, Cr, W, Te, Mn, Fe, Co, and Ni remain stationary or move only slightly. Gold reduced with FeSO_4 is retained by the column. Mercury (II), selenium, arsenic, antimony, and bismuth move less rapidly through the column than uranium. Ceric nitrate is extracted as are thorium, zirconium, and scandium nitrates. Cerium in the III-state is not extracted. Thorium extraction is sensitive to the acid concentration. Zirconium extraction is inhibited by phosphate, sulfate, oxalate, and tartrate ions. Scandium extraction is also inhibited by tartrate ion. Tin is precipitated as meta-stannic acid. Large amounts of tin may be first removed by volatilization as the iodide. Vanadium is retained if peroxides are absent. Ferrous sulfate reduces vanadium to an immobile salt. Phosphoric acid is extracted. Ferric nitrate inhibits the extraction of this acid. The behavior of molybdenum is complex. Iridium and rhodium are not extracted. Traces of ruthenium and platinum may be found in the eluent. Palladium is extracted. Reduction of platinum and palladium with FeSO_4 results in retention of bulk amounts by the column. Small amounts of sulfate do not interfere with the extraction of uranium. Sulfuric acid is retained by the column under normal conditions. Halides increase the extraction of other elements, e.g., Au, Sn. Under normal conditions, HCl is retained in the column; HBr, HI, bromine and iodine move slowly down the column. Molybdenum and arsenic may be adsorbed by the use of activated alumina in conjunction with cellulose.⁶¹⁷

The use of silica gel columns combined with organic solvents, dibutyl carbitol and tributyl phosphate, and nitric acid have been used for the separation of uranium and plutonium.^{618,619}

A non-ionic phosphorylated resin, diethyl polystyrene-methylenephosphonate, may be used to separate uranium (VI) from iron (III), lanthanum, zirconium, niobium, thorium, and mixed fission products.⁶²⁰ Uranium is adsorbed by the resin from 2 per cent solutions of dibutyl phosphoric acid. The other elements

are not absorbed. Uranium is eluted with a dimethyl formamide-benzene solution.

The feasibility of using tributyl phosphate gels for the separation of uranium from iron (III) and thorium has recently been demonstrated.⁶²¹

5. Volatilization. Uranium may be separated from many elements by fractional distillation of the volatile compound, uranium hexafluoride. This method of separation has been applied to the recovery of uranium from irradiated fuel elements. Katz and Rabinowitz¹ have reviewed many of the early methods for the preparation of UF_6 : fluorination of various uranium compounds with elemental fluorine or cobalt trifluoride, disproportionation of UF_5 which results in both UF_4 and UF_6 , and the reaction between UF_4 and dry oxygen which results in UO_2F_2 and UF_6 . The latter two methods are not very practical from an analytical standpoint. Other fluorinating agents that form UF_6 include ceric fluoride, manganic fluoride, silver difluoride, halogen fluorides (e.g., BrF_3 and ClF_3) and fused metallic fluorides. Most elements form fluorides under the conditions that UF_6 is obtained. However, only a small number of these fluorides are volatile. Hyman, et al.³ have published a table of some 26 elements having fluorides with boiling or sublimation points of $550^\circ C$ or less. Included in this group are the fluorides of boron, silicon, phosphorus, vanadium, sulfur, tungsten, bismuth, plutonium, and the fission products, germanium, arsenic, selenium, niobium, molybdenum, ruthenium, antimony, tellurium, and iodine. The boiling point of UF_6 is $54.6^\circ C$. Non-volatile fluorides from which uranium is readily separated include those of the alkali metals, alkaline earths, rare earths, Fe, Co, Ni, Ag, Al, Be, Mn, Tl, Pb, Zn, Cu, Hg, Cd, and Zr.^{34, 622}

Uranium does not form a volatile compound by interaction with anhydrous hydrogen fluoride. Materials such as Nb, Ta, As,

Sb, Si, Te, Se, etc. do.³⁴ The oxides of titanium, tungsten, and molybdenum react slowly with HF. V_2O_5 and VN also react slowly. VO, V_2O_4 and V_2O_3 are not volatilized.³⁴ Rodden and Warf³⁴ describe a procedure that makes use of both anhydrous hydrogen fluoride and fluorine in the separation of uranium. Possible contaminants of the separated UF_6 include Cr, Ta, W, Mo, or V.

Uranium hexachloride and uranium (IV) borohydride are volatile compounds for which procedures might be developed for the separation of uranium.

Uranium may be separated from arsenic, antimony, bismuth, selenium, and tin by volatilization of the latter elements with a mixture of hydrobromic acid and bromine.¹⁹⁸

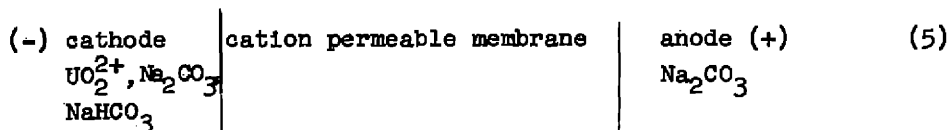
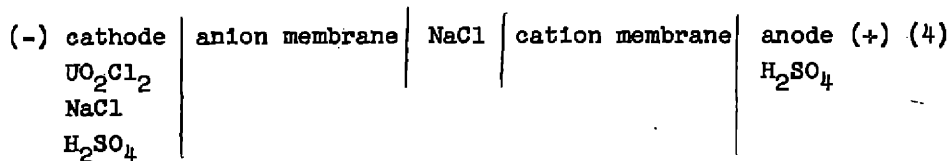
6. Electrochemical methods. The electrolysis of dilute sulfuric acid solutions with a mercury cathode results in the quantitative deposition of Cr, Fe, Co, Ni, Cu, Zn, Ga, Ge, Mo, Rh, Pd, Ag, Cd, In, Sn, Re, Ir, Pt, Au, Hg, and Tl in the cathode.²¹³ Arsenic, selenium, tellurium, osmium, and lead are quantitatively separated from the electrolyte, but are not quantitatively deposited in the cathode.²¹³ Manganese, ruthenium, and antimony are incompletely separated.²¹³ Uranium and the remaining actinide elements, rare earth elements, the alkali and alkaline earth metals, aluminum, vanadium, zirconium, niobium, etc. remain in solution.²¹³ Casto⁶²³ and Rodden and Warf³⁴ have reviewed the effects of many variables in the electrolytic separation of the above-named elements from uranium. According to Rodden and Warf,³⁴ optimum conditions for the purification of uranium in sulfuric acid solutions with a mercury cathode are: electrolyte volume, 50 ml; free sulfuric acid concentration, 1N; current density, as high as practicable with the given acid concentration (about 10 amp maximum); anode, flat platinum spiral or grid just making contact with the surface of the electrolyte; cathode area,

as large as practicable; stirring, surface of the mercury cathode is stirred rather rapidly; temperature of electrolyte, between 25° and 40°C; mercury for the cathode, pure; anions, chloride, nitrate, and phosphate ions should be absent, or present in only small amounts.

Uranium may be deposited electrolytically at the cathode of a cell from acetate,⁶²⁴⁻⁶³¹ carbonate,⁶³² oxalate,⁶³³⁻⁶³⁸ formate,⁶³⁹ phosphate,⁶⁴⁰ fluoride,^{632,641} and chloride⁶⁴² solutions. Many of the uranium electrodeposition procedures have been developed in an effort to prepare thin, uniform films for alpha and fission counting rather than to separate the element from any particular impurity. However, in the work of Smith and co-workers^{624,626} and Coomans⁶²⁷ uranium was separated from alkali and alkaline earth metals and zinc. Casto⁶²³ and Rodden and Warf³⁴ review much of the material pertinent to the electrodeposition of uranium.

Electrodialysis has achieved a certain amount of importance in the recovery of uranium from leach liquors. In a review article by Kunin,⁶⁴³ the following cells are presented for consideration:

(-) cathode $\text{UO}_2(\text{NO}_3)_2$ H_2SO_4 NH_4NO_3	anion permeable membrane	anode (+) NH_4NO_3	(1)
(-) cathode UO_2SO_4 H_2SO_4 H_3PO_4	anion permeable membrane	anode (+) H_2SO_4	(2)
(-) cathode UO_2Cl_2 NaCl H_2SO_4	anion permeable membrane	anode (+) NaCl	(3)



In acid systems (cells 1, 2, 3, and 4), the transport of sulfate, nitrate, and chloride ions to the anode results in a removal of acid and a subsequent increase in pH in the cathode compartment. The uranium, reduced during electrolysis, is precipitated as UO_2 or $\text{U}(\text{HPO}_4)_2$. In the alkaline system, the transport of sodium ions is also accompanied by a rise in pH in the cathode compartment and uranium is again precipitated as the dioxide or as a mixture of dioxide and sodium polyuranate.

The electrodialytic separation of uranium from metals in a complex mixture has been demonstrated by Willard and Finley.⁶⁴⁴ An ammonium bicarbonate solution containing U, Fe, Ni, Cu, Cr, Zn, Al, Mo, Mg, and Na salts, and traces of other elements was electrolyzed in a two-compartment cell having a cation permeable membrane and a mercury cathode. The solution was first made the catholyte (electrolyte in the cathode compartment) and electrolyzed. Iron (80%), nickel and copper (95%), tin, and zinc were removed from solution by deposition. The bicarbonate solution was then made the anolyte and electrolyzed at a platinum anode. All aluminum, molybdenum, ammonium, and silicon, and some sodium and magnesium were separated from the uranium by migration. Uranium was retained as the carbonate complex and was recovered as the oxide by evaporation of the anolyte.

Other features of the electrodialytic behavior of uranium that may be useful in its separation and purification are (1) the retention of uranium during electrodialysis from a perchloric

catholyte using an anion selective membrane, (2) the dissolution and separation of impure UF_4 , and (3) the feasibility of electro-dialysis in organic solutions. 644

7. Pyrometallurgical processes. Although pyrometallurgical or high temperature processes have been designed primarily for large scale recovery of fertile material from irradiated fuel elements, some of the methods may find application in the radiochemistry laboratory. Types of pyrometallurgical operations that have received considerable attention are (1) distillation, (2) salt extraction, (3) molten metal extraction, (4) oxidative slagging, (5) electro-refining, and (6) decomposition of uranium iodide. These methods have been reviewed by Lawroski. 211

1. Plutonium is concentrated by vacuum distillation from molten uranium at 1500-1800°C.
2. Plutonium is extracted from molten uranium by salts such as UF_4 or $MgCl_2$. Uranium remains in the metallic state. Plutonium is recovered as a halide salt.
3. Plutonium is extracted by molten metals, such as silver or magnesium, that are immiscible with molten uranium. Fission products are also extracted.
4. Oxidative slagging involves the preferential formation of the most stable oxides by a molten irradiated fuel element in a limited oxygen environment. These oxides (rare earths) float to the surface of the molten material and are skimmed off. Other oxides diffuse into the crucible and through the slag layer.
5. In electro-refining, uranium is dissolved anodically in a fused salt bath of alkali or alkaline earth halides that contain a uranium compound. Noble metals do not dissolve and are deposited as anode sludge. Uranium and chemically similar materials are deposited at the cathode. Alkali, alkaline earth, and rare earth fission products concentrate in the salt bath.
6. Uranium is recovered as the metal from the thermal decomposition of UI_4 . Zirconium and niobium are the principal contaminants.

It is not the purpose of this section to describe the techniques involved in pyrometallurgical processes. The interested reader may consult the many papers presented in "Progress in Nuclear Energy, Series III, Process Chemistry," volumes 1(1956) and 2(1958), and in the "Proceedings of the International Con-

ference on Peaceful Uses of Atomic Energy," volumes 9(1956) and 17(1958).

IV-E Determination of Uranium

The amount of uranium in a sample may be determined by standard methods of analysis: gravimetric, volumetric, colorimetric, spectrophotometric, etc.^{34, 191-195, 197, 198, 200, 645} Because of its natural radioactivity, uranium may also be determined by counting techniques. The applicability, in terms of mass range, of various methods for the determination of uranium is given in Table XXXVII.¹⁹⁹

1. Counting techniques. Principles of alpha, beta, and gamma counting are considered in review articles by Steinberg,⁶⁴⁶ Hanna,⁶⁴⁷ Deutsch and Kofold-Hansen,⁶⁴⁸ Crouthamel,⁶⁴⁹ and Jaffey.⁶⁵⁰ All three methods of counting are applicable to the radiometric determination of uranium since both alpha- and beta-emitting isotopes exist (Section III). Spontaneous fission half-lives have been determined for several uranium isotopes: U^{232} , U^{234} , U^{235} , U^{236} , U^{238} . These isotopes are too long-lived, however, to make fission counting a practical method for their determination.

Ionization chambers are most commonly used for the detection of alpha particles. In figures 77 and 78 are shown the alpha spectra of U^{235} and U^{233} , respectively. The spectra were obtained with a parallel plate, Frisch grid ionization chamber using P-10 (90% argon, 10% methane) gas. A multi-channel analyzer was used in conjunction with the ionization chamber. Both U^{235} and U^{233} samples were prepared by volatilization. Figure 79 represents the U^{233} alpha spectrum obtained with a surface barrier silicon solid state detector. Data for this figure was taken from the same sample as that for figure 78. It is readily apparent from the two figures that the solid state detector gives much better resolution of the alpha groups than does the ionization chamber.

Table XXXVII. Range of Application of Various Methods for the Determination of Uranium.^a

Method	Range of application (micrograms)	Range of error (per cent)
Neutron activation	10^{-4} - 10^4	±2 to 5
Fluoroscopy	10^{-4} - 1	±5 to 50
Emission spectroscopy	5×10^{-2} - 50	±1 to 10
Visual chromatography on paper	10^{-1} - 10^2	
Volumetric (including microvolumetric) methods	1 - 5×10^5	±0.5 to 5
Autoradiography (α emission) counting of tracks	1 - 10^4	±1 to 10
Colorimetry dibenzoylmethane	10 - 10^4	±1 to 3
thiocyanate	50 - 5×10^4	±1 to 3
H ₂ O ₂ - HClO ₄	10^3 - 10^5	±1 to 5
Alpha counting ^b	50 - 5×10^3	±1 to 10
Polarography	10^2 - 10^4	±2 to 5
Potentiometry	2×10^2 - 10^4	±1 to 5
Gravimetric methods	5×10^4 -	±0.1 to 2

^a Adapted from a table given by A. Simenauer, reference 199.

^b 50 μg of U²³⁸ gives about 20 cpm at 52% geometry. Uranium-238 may be detected in samples having much lower counting rates than this, depending upon the physical condition of the sample and the presence of extraneous alpha activity. For a thin source with low alpha background, 1 alpha cpm of uranium should be readily detected. For U²³⁸ this is the equivalent of about 2 μg; for other uranium isotopes, the mass is even less.

The data for figure 79, however, was taken in 1000 minutes. The data for figure 78 was taken in 10 minutes.

Alpha particles may be counted also by gaseous, liquid, plastic, and crystalline scintillation detectors. The resolution of these detectors is, in general, less than ionization chambers and their application more limited. Nuclear emulsions are used to record alpha activity. Such devices as cloud chambers are generally not used in the radiochemistry laboratory.

Geiger-Müller counters, proportional counters, and liquid, plastic, and crystalline scintillation detectors are suitable for the counting of β⁻-emitting isotopes, U²³⁷, U²³⁹, and U²⁴⁰.

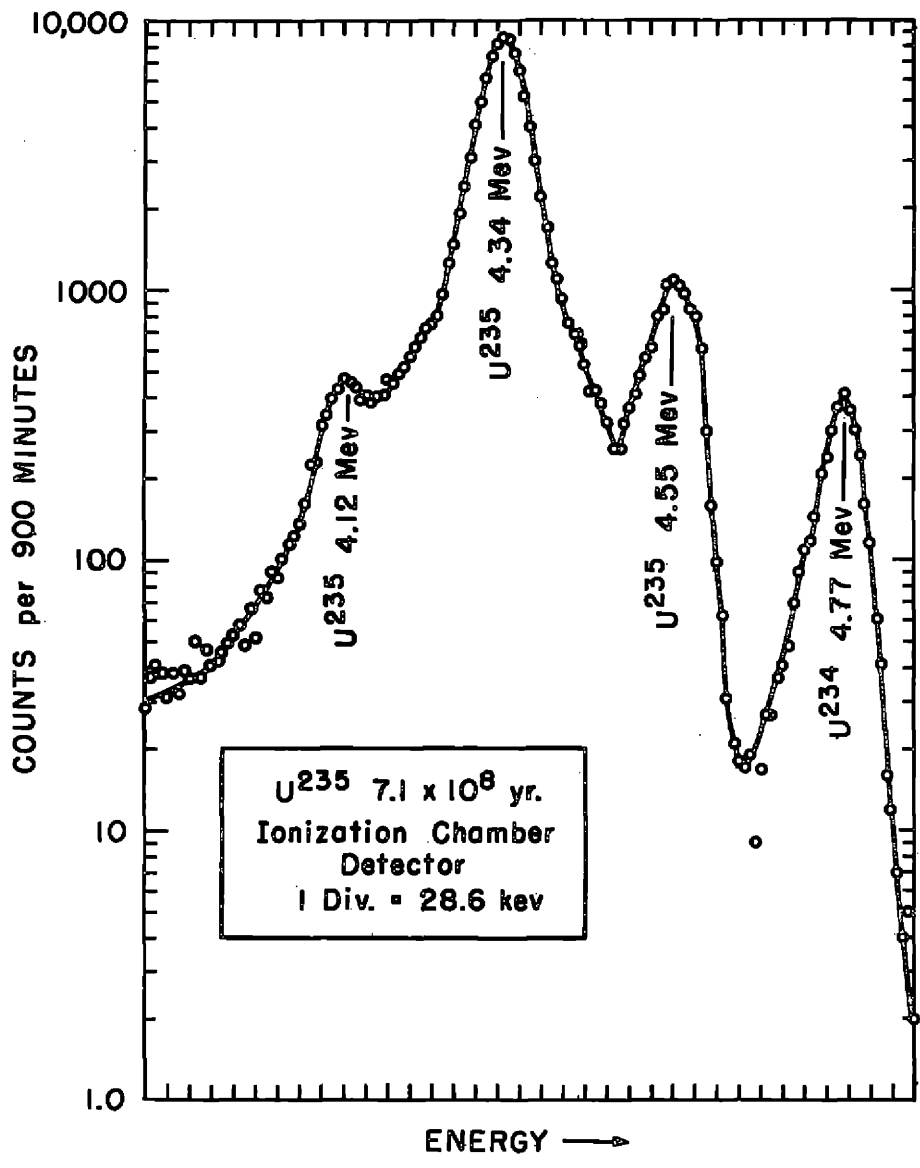


Figure 77.

Alpha spectrum of a volatilized source of U²³⁵ obtained with a parallel plate, Frisch grid ionization chamber using 90% argon - 10% methane gas.

D. J. Henderson, Argonne National Laboratory, Unpublished data.

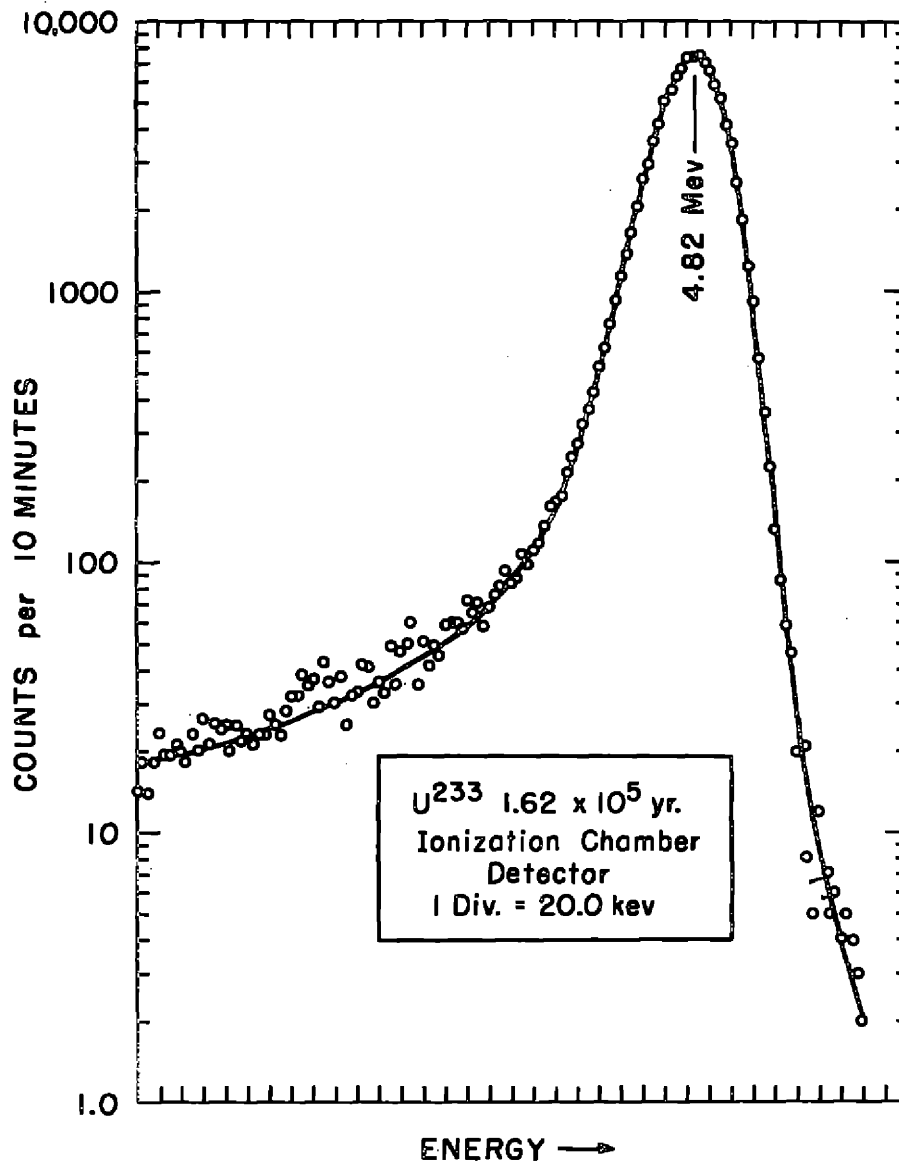


Figure 78.

Alpha spectrum of a volatilized source of U^{233} obtained with a parallel plate, Frisch grid ionization chamber using 90% argon - 10% methane gas.

D. J. Henderson, Argonne National Laboratory, Unpublished data.

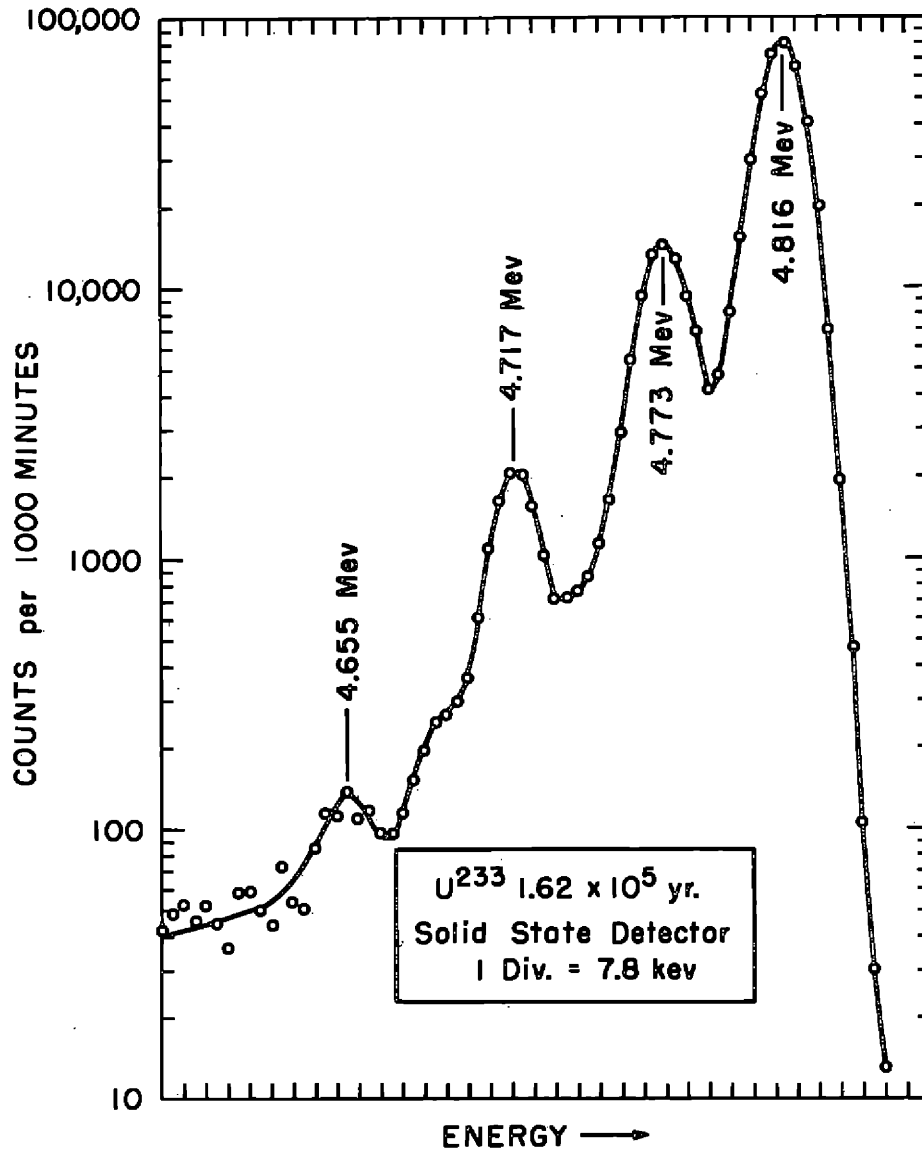


Figure 79.

Alpha spectrum of a volatilized source of U^{233} obtained with a surface barrier silicon detector.

D. W. Engelkemeir, Argonne National Laboratory, Unpublished data.

Sodium iodide-thallium activated crystals have gained widespread acceptance as detectors of gamma radiation. The gamma-ray spectra of U^{239} , U^{237} , U^{235} , U^{233} , and a uranium ore from the Belgian Congo are illustrated in figures 80-84. The spectra were measured by Crouthamel, Gatrouris, and Goslovich⁶⁴⁹ with a 4-inch diameter x 4-inch thick cylindrical NaI(Tl) crystal.

Not only can the amount of uranium in a sample be determined directly by measuring the disintegration rates of the various isotopes, it can also be measured indirectly by determining the activity of daughter products. For such a measurement to be meaningful, however, the equilibrium condition between the uranium isotope and its daughter must be known.

Information on the radioactive decay of the uranium isotopes is given in Section III. Further information on these isotopes and their daughter products may be obtained by consulting the "Table of Isotopes" compiled by Stromminger, Hollander, and Seaborg⁶⁵¹ and the references given therein. Volumes 8 (1956), 3 (1958), and 28 (1958) of the "Proceedings of the International Conference on the Peaceful Uses of Atomic Energy" contain a number of articles on the radiometric determination of uranium. Reference to many more articles is made in the review papers by Meinke.⁶⁵²

2. Sample preparation. One of the most important problems to overcome in the detection of alpha particles and in direct fission counting is the preparation of thin foils or sample deposits. This subject has received considerable attention and has been reviewed by several authors.^{632,647,650,653,654} Several techniques are available. The simplest and most quantitative is the direct evaporation of an aliquot of a sample. The distribution of such deposits are generally not very uniform. This may be improved upon by the addition of a spreading agent such as tetraethyleneglycol, TEG. Painting techniques may be used to build up fairly uniform deposits of several milligrams

per square centimeter. ^{655,656} Uranyl nitrate is dissolved in alcohol and added to a dilute solution of Zapon in Zapon thinner or cellulose in amyl acetate. This solution is painted over a metal backing, allowed to dry, and then baked or ignited at a suitable temperature. For aluminum backing, temperatures of 550° to 600°C are satisfactory. For platinum, higher temperatures (800°C) are preferred. The thickness of the uranium deposit is increased by repeatedly painting and baking the foil.

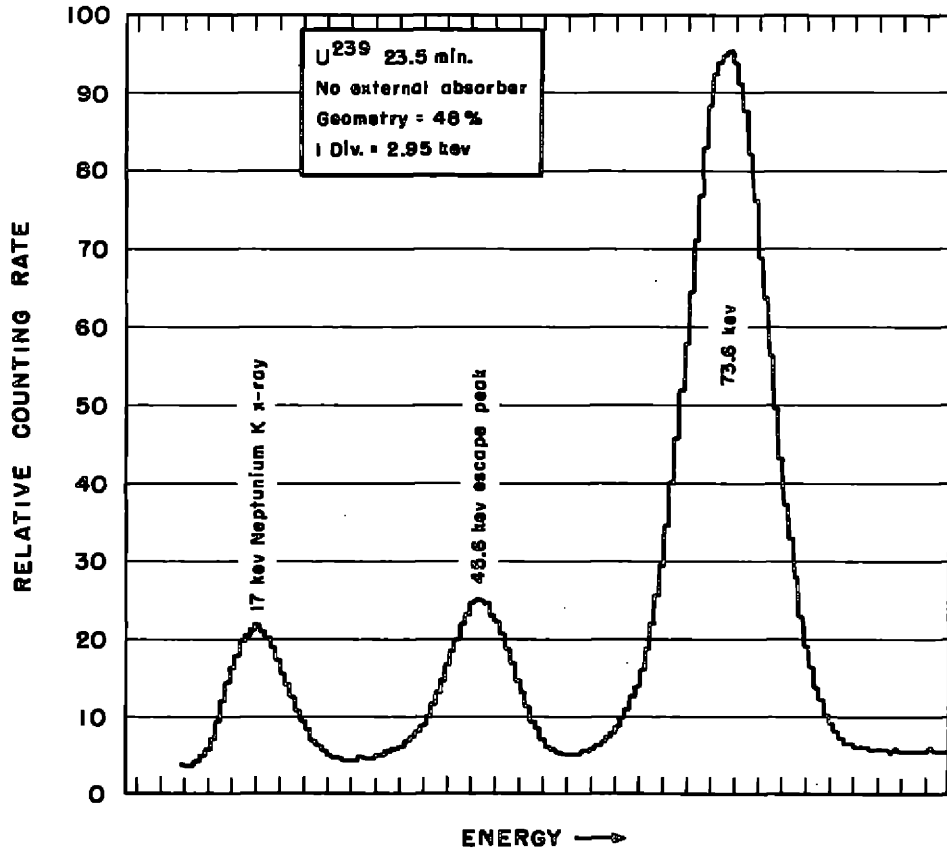


Figure 80.

Gamma-ray spectrum of U^{239} obtained with a 4-inch diameter x 4-inch thick cylindrical NaI(Tl) crystal.

After C. E. Crouthamel, C. Gavouris, S. J. Goslovich, reference 649.

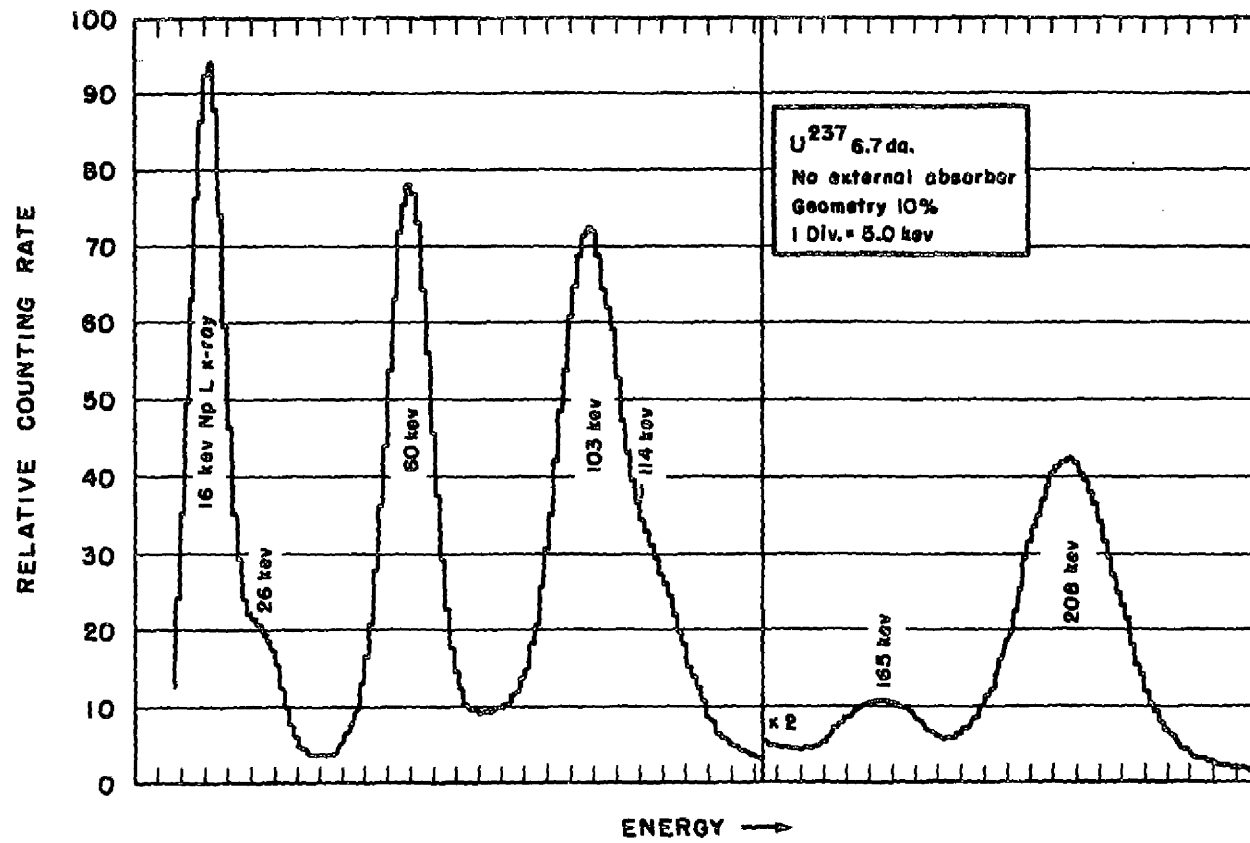


Figure 81. Gamma-ray spectrum of U^{237} obtained with a 4-inch diameter x 4-inch thick cylindrical NaI(Tl) crystal. After C. E. Crouthamel, C. Gatrousis, S. J. Goslovich, reference 649.

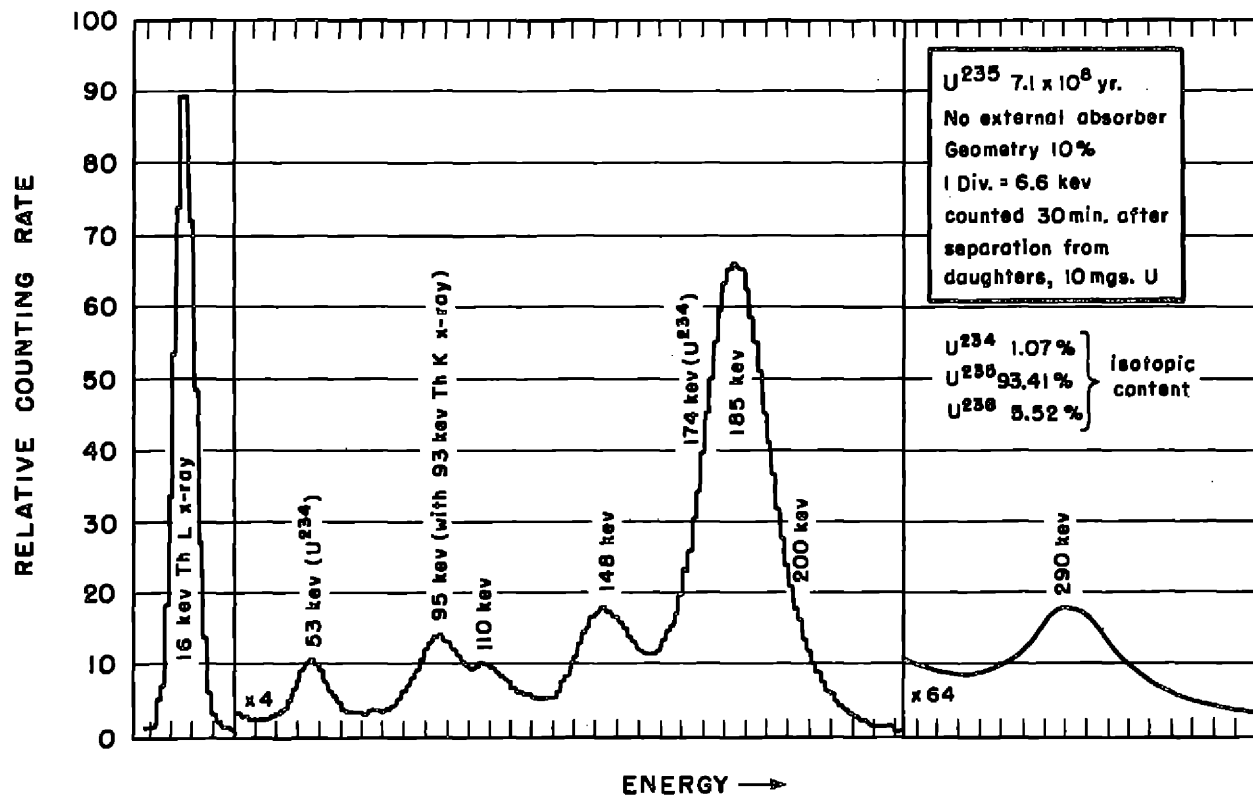


Figure 82. Gamma-ray spectrum of U^{235} obtained with a 4-inch diameter x 4-inch thick cylindrical NaI(Tl) crystal. After C. E. Crouthamel, C. Gatrousis, S. J. Goslovich, reference 649.

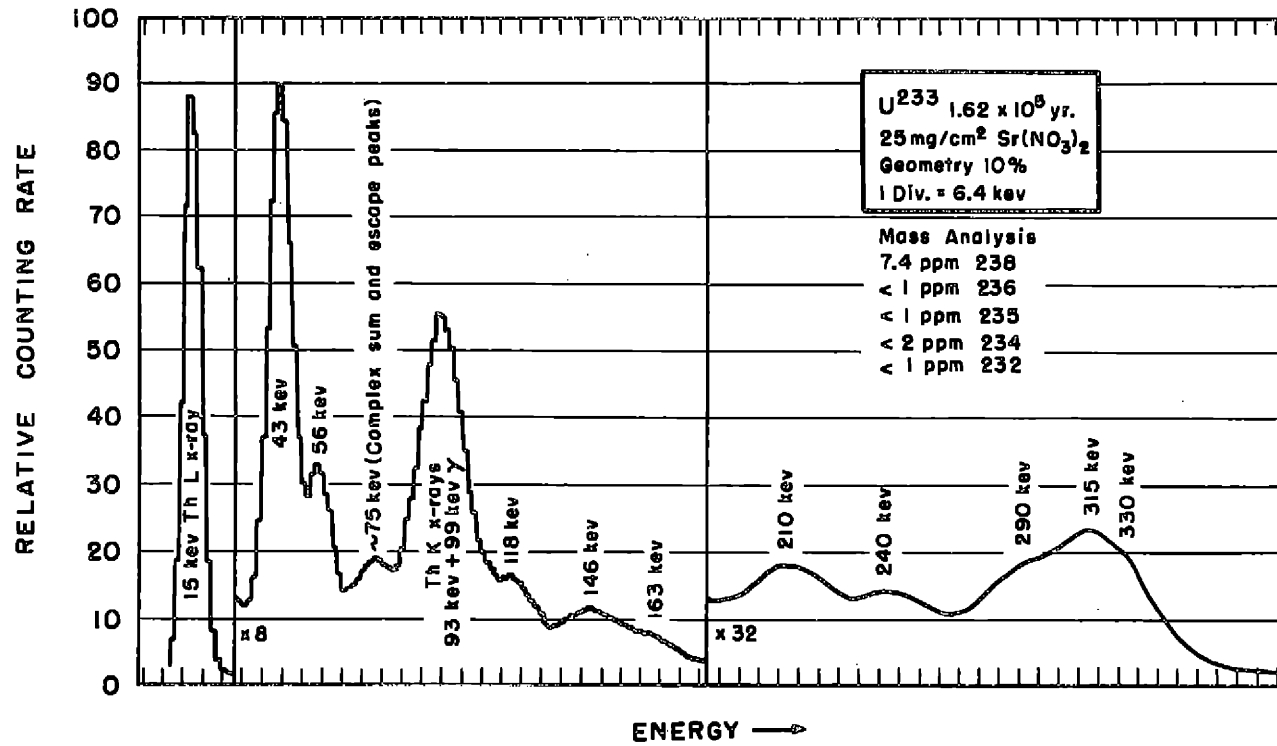


Figure 83. Gamma-ray spectrum of U²³³ obtained with a 4-inch diameter x 4-inch thick cylindrical NaI(Tl) crystal. After C. E. Crouthamel, C. Gatrousis, S. J. Gaslovich, reference 649.

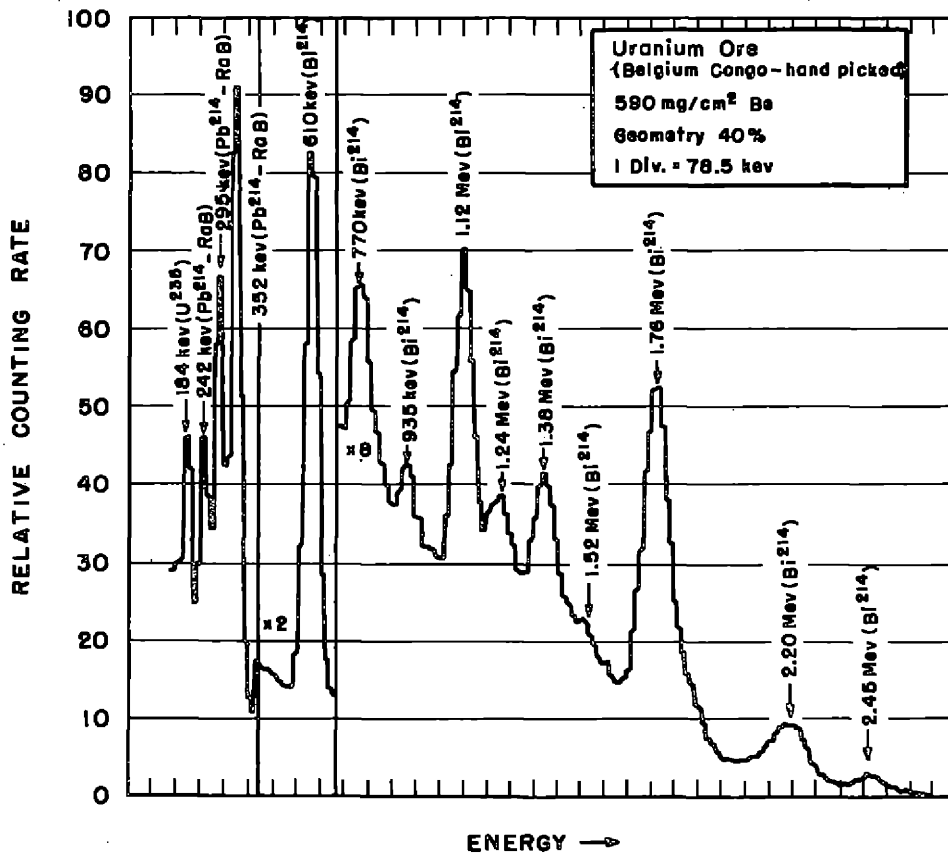


Figure 84.

Gamma-ray spectrum of uranium ore from the Belgian Congo obtained with a 4-inch diameter x 4-inch thick cylindrical NaI(Tl) crystal. After C. E. Crouthamel, C. Gatrousis, S. J. Goslovich, reference 649.

Samples of the metastable U^{235} isomer have been prepared by electrostatically collecting the recoil atoms of Pu^{239} in air.^{657,658} A negative potential of several hundred volts was applied to the metallic collection plate.

Carswell and Milsted⁶⁵⁹ have succeeded in preparing thin sources by a spraying technique. The material to be deposited is dissolved as the nitrate in an organic solvent (alcohol or acetone). The solution is drawn into a fine glass capillary

tube and sprayed onto the backing material by applying an electric field.

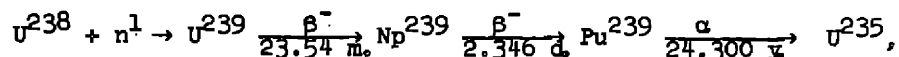
Electrodeposition is a generally satisfactory method for preparing uniform samples with quantitative or nearly quantitative yields. Uranium has been plated from a variety of solutions: acetate,^{34, 623, 628-631} formate,⁶³⁹ oxalate,^{623, 633-638} carbonate,^{34, 632} fluoride,^{623, 632, 641} and chloride.⁶⁴² A satisfactory electrolyte for the deposition of uranium is 0.4M ammonium oxalate.⁶⁵⁴ A rotating platinum anode is used to stir the solution placed in a vertical cylindrical cell. The cell is made of glass, lucite, or some other chemically inert material. The cathode on which the uranium is to be deposited is thoroughly cleansed and made the bottom of the cell. The assembled cell is placed in a hot water bath and the temperature kept at about 80°C. A current density of approximately 0.1 amp/cm² is used. The deposition is influenced strongly by the rate of stirring, current density, and presence of foreign ions.⁶⁵⁴

Vacuum sublimation provides an excellent means for the preparation of thin deposits of uranium. The sublimation of uranium acetylacetonate, $U(C_5H_7O_2)_4$, has been used. A more convenient method is the sublimation of uranium oxides. A uranium salt solution is placed on a tungsten or tantalum ribbon supported between two electrodes. The solution is dried by a heat lamp or by passing a low current through the metal ribbon. The sample backing material is suspended at a suitable height above the metal ribbon. A bell jar is placed over the assembled unit and evacuated. The uranium is volatilized by increasing the current through the metal ribbon. The uniformity of the deposit depends upon the distance between the ribbon and the backing plate. The collection efficiency also depends upon this distance but in an inverse manner to that of deposit uniformity. Usually a compromise is made between collection efficiency and sample uniformity. Much of the uranium that is

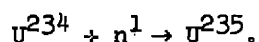
not collected on the sample plate can be recovered from masking plates and glass chimneys placed between the filament and backing material. The collection efficiency may also be improved by subliming from furnaces so that the beam of uranium molecules is directed toward the backing plate. The furnace is heated by electron bombardment or induction heating.

3. Activation Analysis. In activation analysis, a nuclide irradiated by neutrons, gamma rays, or charged particles is transformed into a radioactive nuclide more easily detected than the original one. The amount of original material may be determined either absolutely or comparatively. For an absolute determination, the cross section of the reaction, the irradiation flux, and the disintegration rate of the reaction product must be known or determined. For comparative analysis, a substance of unknown mass is irradiated simultaneously with a similar substance of known mass. The positions of these two substances are either side by side or, if separated, in positions of like flux. The reaction product activities of the two samples are compared to give the relative masses of the starting materials. The comparative technique is, in general, much easier to apply. The uncertainties of many variables are eliminated by relative measurements.

Activation with thermal neutrons may be successfully employed as a method of analysis for natural uranium, uranium-236, and the fissionable isotopes of uranium. Natural uranium consists of U^{238} (99.3%), U^{235} (0.72%), and U^{234} (0.0057%).⁶⁵¹ The principal reactions of these nuclides with thermal neutrons are:*



* Half-lives given below the arrow are taken from the "Table of Isotopes," reference 651. The value of ν for the fission of U^{235} is taken from "Neutron Cross Sections," reference 660.



The amount of natural uranium present in a sample may then be determined from the amount of U^{239} , Np^{239} , or Pu^{239} activity formed after irradiation. Measurement of the Pu^{239} activity, however, requires either a fairly large amount of U^{238} , a long irradiation period, or a combination of the two. The amount of natural uranium may also be determined from the fission of uranium-235 either by (1) fission counting the sample, (2) isolating and counting a fission product such as Ba^{140} or Te^{132} *, or (3) measuring the total gamma activity induced in the sample by a short neutron irradiation.⁶⁴⁵ Thermal neutron irradiation of U^{234} results in U^{235} and is of little value in the determination of natural uranium.

Neutron irradiation of U^{236} gives U^{237} , a beta-emitter having a half-life of 6.75 days.⁶⁵¹ It is readily identified through its beta decay, associated gamma rays, and half-life. Uranium-238 irradiated with fast neutrons also produces U^{237} , $U^{238}(n,2n)U^{237}$. The cross section for this reaction has been determined with incident neutron energies from 6 to 10 Mev and at 16 Mev.⁶⁶¹

Activation analysis by fission counting is of value only if one fissioning nuclide is present or if the amounts of other fissioning nuclides present are known and corrections can be made for them. The same is true for the isolation and determination of fission products. Uranium isotopes that are fissionable with thermal neutrons together with their thermal neutron fission cross sections are:⁶⁶⁰

U^{230}	25 ± 10 barns
U^{231}	400 ± 300 barns

*The fission product nuclides Ba^{140} and Te^{132} are chosen since they are free from interfering reactions and are produced in good yields.

U ²³²	80 ± 20 barns
U ²³³	527 ± 4 barns
U ²³⁵	582 ± 6 barns
U ²³⁹	14 ± 3 barns *

Other than U²³⁵, U²³³ is the best uranium isotope to determine by fission counting or fission product analysis. Uranium-232 may possibly be determined in this manner. The other isotopes are of such short half-life that analysis of their own radiations is a much better means for their identification.

Excitation functions have been determined for a number of reactions with charged particles or gamma rays incident on uranium isotopes. These reactions may be used for activation analysis. For absolute analysis, it should be pointed out that (1) the cross sections reported are sometimes subject to considerable error; (2) energy determinations of the incoming particle or ray are also subject to error; and (3) the reaction product can often be produced by a number of reactions. Comparative analysis appears to be a much better method for the determination of uranium. For gamma-ray (bremsstrahlung) activation, simultaneous irradiations in a like flux are fairly easy to accomplish. The two samples, unknown and standard, are mechanically rotated in the gamma-ray beam. For charged particle activation, the simultaneous irradiation of two samples in a like flux may require some ingenuity on the part of the experimenter. **

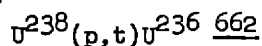
A partial list of reactions between uranium isotopes and charged particles or gamma rays for which excitation functions

* Pile neutrons.

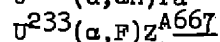
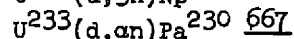
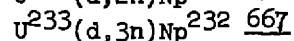
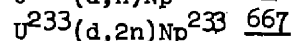
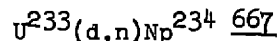
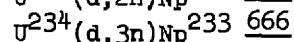
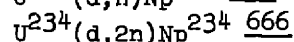
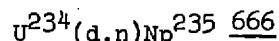
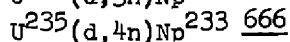
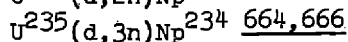
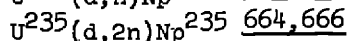
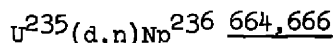
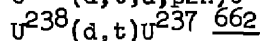
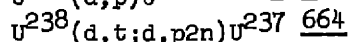
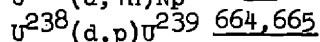
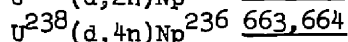
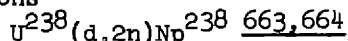
** Because of the short range of charged particles, irradiations are generally made with targets attached to or within the vacuum system of the accelerator. To maintain the system's vacuum requirements, to cool the samples properly, and to irradiate the samples simultaneously in a like flux may present some difficulty in equipment design.

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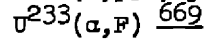
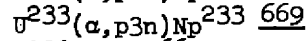
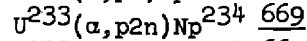
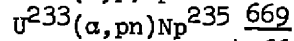
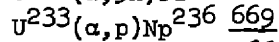
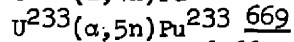
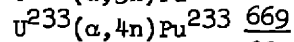
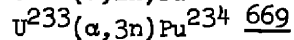
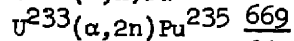
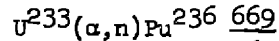
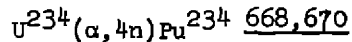
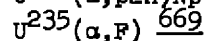
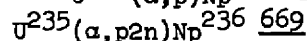
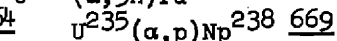
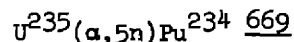
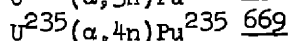
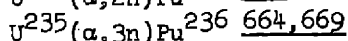
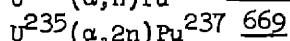
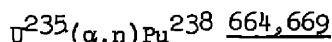
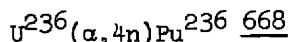
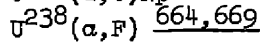
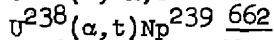
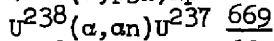
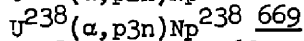
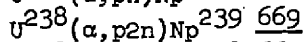
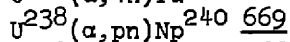
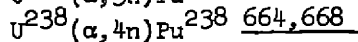
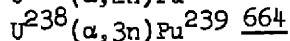
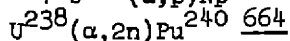
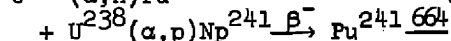
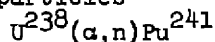
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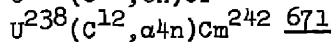
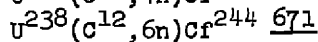
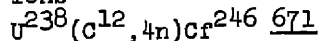
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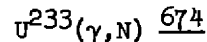
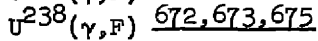
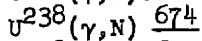
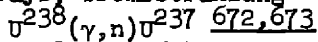
Alpha particles



Carbon ions



Gamma rays, bremsstrahlung



The analysis of uranium by activation methods is reviewed by Koch,⁶⁷⁶ who also gives references to much of the literature.

IV-F. Dissolution of Uranium Samples

1. Metallic uranium.

HNO₃. Uranium metal dissolves in nitric acid to form uranyl nitrate. With massive amounts of uranium the rate of dissolution is moderately rapid.⁶⁷⁷ The reaction between uranium turnings, powder, or sintered metal and nitric acid vapors or nitrogen dioxide may occur with explosive violence.⁶⁷⁷ Oxides of nitrogen are the principal gaseous products in the dissolution of the metal by HNO₃. The presence of oxygen in the dissolver system tends to reduce the emission of these oxides.⁶⁷⁸ The rate of dissolution of large amounts of metallic uranium may be increased by the addition of small amounts of sulfuric,⁶⁷⁹ phosphoric,^{679,680} or perchloric⁶⁸¹ acid to the nitric acid.

H₂SO₄. Hot concentrated sulfuric acid attacks uranium metal slowly forming uranium (IV) sulfate.⁶⁷⁷ Sulfuric acid-hydrogen peroxide mixtures react slowly with the metal at 75°C forming uranyl sulfate.⁶⁸² The addition of small amounts of chloride or fluoride to H₂SO₄ - H₂O₂ mixtures increases the dissolution rate.⁶⁸²

H₃PO₄. Cold 85% phosphoric acid attacks uranium metal slowly.⁶⁷⁷ Concentration of the acid by heating produces a fairly rapid reaction in which uranium (IV) phosphate is formed. If heated too long, a chemically inert, glassy substance is formed.

HClO₄. Uranium metal is inert toward cold, dilute perchloric acid. As the concentration is increased by heating a point is reached at which the reaction proceeds with violence.^{677,682} Oxidizing agents added to dilute perchloric acid dissolve the metal.⁶⁷⁷

HCl. Concentrated hydrochloric acid vigorously attacks uranium metal. Dilution of the acid diminishes the attack. But even with 4M HCl there is a rapid evolution of hydrogen.⁶⁸² A

finely divided, black precipitate soon forms after dissolution begins. This precipitate is not dissolved by heating. Only by the addition of oxidizing agents (hydrogen peroxide, bromine, chlorate, nitrate, persulfate, dichromate, or ferric ions) does the precipitate dissolve. Gaseous chlorine, aided by small amounts of iron or iodine, also oxidizes and solubilizes the uranium precipitate. The addition of small amounts of fluosilicic acid^{682,683} or large amounts of phosphoric acid⁶⁸² to the hydrochloric acid prevents formation of the black precipitate during the dissolution of uranium metal.

HF. The reaction of hydrofluoric acid with uranium metal is slow even at temperatures of 80°-90°C.⁶⁷⁷ The reaction is inhibited by the formation of insoluble UF_4 on the surface of the metal.

HBr. Hydrobromic acid attacks metallic uranium in manner similar to, but slower than, hydrochloric acid.⁶⁷⁷ The black precipitate is formed.

HI. The reaction between uranium metal and hydriodic acid is slow.⁶⁷⁷

Organic acids. Acetic, formic, propionic, and butyric acids react rapidly with uranium in the presence of hydrogen chloride.⁶⁷⁷ Benzoic acid in ether reacts with the metal, forming the benzoate.⁶⁷⁷ Acetyl chloride and acetic anhydride react to form uranous acetate.⁶⁷⁷

Miscellaneous solvents. Uranium is dissolved in a number of media other than acids:^{34,677,682} solutions of heavy metal salts (silver perchlorate,⁶⁸⁴ cupric ammonium chloride⁶⁸⁵ or acetate⁶⁸⁶), alkaline peroxide solutions ($NaOH-H_2O_2$ or $Na_2O_2-H_2O$ solutions⁶⁸⁷), solutions of bromine and ethyl acetate,^{682,688} hydrogen chloride and ethyl acetate,⁶⁸² hydrogen chloride and acetone,⁶⁸² and nitrogen dioxide and hydrogen fluoride.⁶⁸⁹

Table XXXVIII denotes qualitatively some solutions that satisfactorily dissolve uranium.⁶⁸²

Anodic dissolution. Metallic uranium may be dissolved elec-

trolytically by anode oxidation. A variety of electrolytes have been used.⁶⁹⁰⁻⁶⁹² Satisfactory dissolutions have been made with sulfuric acid,⁶⁸² nitric acid,⁶⁹¹ tartaric acid,⁶⁹¹ phosphoric acid containing nitrate,⁶⁹² and sodium bicarbonate.⁶⁹¹

2. Alloys of uranium. The ease with which uranium alloys are dissolved depends largely upon the chemical behavior of the alloying metal. Larsen⁶⁸² has reviewed the dissolution of some of the more common uranium alloys. Table XXXVIII summarizes the effect of various reagents on these alloys.

3. Compounds of uranium. Table II lists solvents for a number of uranium compounds. General Review references 2, 4, 5, and 7 (Section I) cover the chemical properties of these and other compounds more fully.

Table XXXVIII. Reagents for the Dissolution of Uranium and Its Alloys.^a

Description	S = satisfactory			N = unsatisfactory			
	HNO ₃	Aqua Regia	Nitric - HF	HCl + Ox	HCl - EtOAc	Br ₂ - EtOAc	NaOH - H ₂ O ₂
U	S	S	S	S	S	S	S
U-Zr	N	N	S	N	N	S	N
U-Nb	N	N	S	N	N	S	S
U-Fe	S	S	S	S	S	S	N
U-Cr	N	N	N	S	S	S	N
U-Ru	N	S	N	N	N	N	N
U-Mo	N	S	N	S	N	S	S
U-Fissium ^b	N	S ^c	N	N	N	N	N
U-Si ^d	S		S				
U-Pu	S ^e	S	N	S	S	S	N

^a R. P. Larsen, reference 682.

^b Alloys containing from 1 to 3% Zr, Mo, Ru, Rh, Pd, and Ce.

^c Fluoride must be added to dissolve Zr.

^d Nitric acid dissolutions leave Si residue, but nitric-hydrofluoric acid dissolutions can easily lead to volatilization of fluosilicic acid.

^e Pu itself is not readily dissolved in nitric acid, hydrochloric acid being preferable.

The dissolution of uranium oxides is of considerable interest since uranium samples prepared as accelerator targets, for neutron irradiations, or samples found in the natural state are frequently in the form of oxides. Also, many compounds of uranium may be transformed to the oxide by heating, hydrolysis, or fusion. All of the oxides, UO_3 , U_3O_8 , and UO_2 , are soluble in nitric acid, forming uranyl nitrate. UO_3 is soluble in other mineral acids. U_3O_8 and UO_2 are dissolved by fuming with perchloric acid.³⁴ They are slowly dissolved in hot concentrated sulfuric acid.³⁴ The presence of fluoride accelerates this dissolution.⁶⁹³ Alkaline peroxides react with uranium oxides to form soluble peruranates.^{34,687}

4. Meteorites, minerals, and ores. The extraction of uranium from natural deposits may be accomplished by decomposition and dissolution of the entire sample including uranium or by leaching the uranium from the sample. Grinding and roasting facilitate the recovery. Roasting removes organic material. It also helps form soluble uranium compounds.

Decomposition of the sample may be accomplished by acid attack, by fusion, or by a combination of the two. Mineral acids, individually or in combination, may be used. The presence of hydrofluoric acid generally aids in dissolution. Ores, sand, etc. may be fused with sodium carbonate, sodium hydroxide, sodium peroxide, sodium bisulfate, sodium chloride and sodium hydroxide, ammonium sulfate, potassium bifluoride, and magnesium oxide.^{1,34,694} The melt is solubilized in water or acid and the separation of uranium made by procedures outlined in Section IV-D. Rodden and Warf³⁴ have described a number of procedures in which uranium was made soluble by acid attack or by fusion methods. The recovery of uranium from monazite sands has been reported by Calkins, et al.⁶⁹⁴

Acid and alkaline leaching are used on an industrial scale for the recovery of uranium from its ores. In acid leaching,

hydrochloric, nitric, or sulfuric acid may be used.¹ Industrially, sulfuric acid is used because of its economy. Oxidizing agents (Fe(III), MnO₂, etc.) are used to convert uranium (IV) to uranium (VI). A separation of uranium and thorium with an oxalic acid-nitric acid leach solution has been reported.⁶⁹⁵

In alkaline leaching, various combinations of alkaline carbonates, hydroxides, and peroxides have been used.¹ Industrially, uranium is dissolved by alkaline carbonates as the UO₂(CO₃)₃⁴⁻ complex. Oxygen or other suitable oxidants are used to convert uranium (IV) to uranium (VI). Hydroxyl ions are formed by the dissolution of uranium in carbonate solutions. The presence of bicarbonate ion in the dissolving solution prevents the precipitation of uranium. The recovery of uranium by acid and alkaline leaching is reviewed in General Review reference 13 (Section I).

5. Biological samples. The determination of uranium in biological samples is reviewed by Steadman.⁶⁹⁶ Uranium may be extracted and determined directly from liquid samples. The sample may also be ashed, as are solid samples, prior to uranium extraction. Ashing may be carried out as a wet or dry process. Wet-ashing is commonly done with a nitric acid solution. Ashing may be completed with perchloric acid. However, extreme caution must be exercised when heating organic materials with perchloric acid. The ashed residue is dissolved in acid and the uranium determination continued from there. Wet-ashing need not be carried to completion. Analysis may be made upon the sample after it has been thoroughly digested in acid.

6. Air dust samples. Samples of air dust are commonly collected on filter papers. The uranium may be dissolved by digesting the sample in nitric acid solution or the sample may be ashed and the residue dissolved in acid.

V. Collection of Detailed Procedures

A procedure for the determination of uranium may entail one or more purification steps as outlined in the preceding sections. For example, uranium may be separated from impurities by a series of solvent extractions with one or more different solvents. These may be interspersed with precipitation and/or ion exchange methods. The procedures described herein have been gathered from project reports, the open literature, and by private communication. Only a limited number are presented. They have been selected because they represent many of the separation methods already described or because they represent different problems in handling samples: problems of dissolution, extraction in the presence of high beta-gamma activity, etc. A number of the procedures described do not make use of the radiometric determination of uranium. The method of separation in these procedures, however, is applicable to radiochemical analysis and is, therefore, included. A number of papers and reports describe, in detail, procedures for the determination of uranium. These should be noted. The work of Rodden and Warf³⁴ has frequently been mentioned in this paper. In addition to procedures for the precipitation, solvent extraction, volatilization, and electrodeposition of uranium, these authors have presented a number of selected procedures for the solution of ores and minerals and the separation and determination of uranium. Procedures for the analytical determination in naturally occurring materials have also been described by Rodden and Tregonning,³⁶⁶ Grimaldi, May, Fletcher, and Titcomb,⁶⁹⁷ Schoeller and Powell,²⁰⁹ and in the "Handbook of Chemical Determination of Uranium in Minerals and Ores."⁶⁹⁸ The recent publication by Moore⁴⁷¹ on extraction with amines contains a collection of procedures, many of which have to do with the separation of uranium.

PROCEDURE 1: Uranium-237.

Source: B. Warren, LA-1567 (1953) p. 18.

Editor's note: Uranium-237 may be separated from fission products, neptunium, and plutonium more easily by ion exchange and/or solvent extraction techniques (see, for example, Procedure 7). The following procedure is, however, an excellent example of uranium purification by precipitation methods.

1. Introduction

The significant steps in the determination of U^{237} in materials containing fission products, neptunium and plutonium are the following. Rare-earth, neptunium, and plutonium activities are removed by appropriate lanthanum fluoride scavenging steps in the presence of hydroxylamine hydrochloride. The latter reagent serves to reduce both neptunium and plutonium so that they may be carried down, and also to complex uranium and prevent its later removal in iron scavenging steps. Barium and zirconium are precipitated by barium fluozirconate scavenging. following a cycle of ferric hydroxide scavenging and ammonium diuranate precipitation steps, uranium is reduced by zinc metal in hydrochloric acid medium and precipitated, presumably as $U(OH)_4$, with ammonium hydroxide. The uranium is further purified by alternate conversions to tetrafluoride and hydroxide. $24.1d$ $Th^{234}(UX_1)$ which has grown in from U^{238} is removed by a zirconium iodate scavenge and the uranium is converted to ammonium diuranate. Uranium is finally plated from nitric acid medium onto a platinum foil. After flaming of the foil and weighing, uranium is beta-counted as U_3O_8 . Chemical yields average 50 to 65%. Quadruplicate determinations require approximately 8 hours.

2. Reagents

U^{238} carrier: 1 ml containing 10 mg of (5000/l) uranium. Preparation: Weigh out 1 gm of U metal, dissolve in conc. HNO_3 , transfer to a 100-ml volumetric flask. Make up to volume, adjusting the final solution

PROCEDURE 1 (Continued)

to 3M in HNO_3 . The carrier is standardized by pipeting 1 ml aliquots into a 000-Coors porcelain crucible, evaporating to dryness, igniting at 800° for 45 min, and weighing as U_3O_8 .

La carrier: 10 mg La/ml (added as $\text{La}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ in H_2O)

Ba carrier: 10 mg Ba/ml (added as $\text{Ba}(\text{NO}_3)_2$ in H_2O)

Zr carrier: 10 mg Zr/ml (added as $\text{ZrO}(\text{NO}_3)_2 \cdot 2\text{H}_2\text{O}$ in 1M HNO_3)

Fe carrier: 10 mg Fe/ml (added as $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ in very dilute HNO_3)

HCl: conc.

HNO_3 : 1M

HNO_3 : 8M

HNO_3 : conc.

HF: conc.

H_2SO_4 : conc.

HIO_3 : 0.35M

NH_4OH : conc.

$\text{NH}_2\text{OH} \cdot \text{HCl}$: 5M

4% aqueous $(\text{NH}_4)_2\text{C}_2\text{O}_4$

Br_2 : liquid

Zn metal: 20 mesh, granular

Methanol: anhydrous

Methyl red indicator solution: 0.1% in 90% ethanol.

3. Equipment

Fisher burner

Centrifuge

Block for holding centrifuge tubes

40-ml centrifuge tubes: Pyrex 8140 (10 per sample)

000-Coors porcelain crucibles (one per standardization)

Pt-tipped tweezers

Pipets: assorted sizes

PROCEDURE 1 (Continued)

Stirring rods

Plating assembly: 1 cell per aliquot of sample

Source of current - Fisher Powerhouse (D.C.)

with variable resistance in series with cells.

Cell - Brass base (3" x 3") for holding

Pt cathode; 5-ml Pt circular 2" diameter disk (cathode);
gasket (Koroseal-Upholstery 36681) to seal cathode and chimney;
glass chimney, 2" diameter, 4" high, with 4 ears at height of
3"; 1 1/4" steel springs for holding chimney to base; rotating
Pt anode. The cell is heated for 1 3/4 hours at 105° after
assembly to insure formation of seal between glass and Pt.

Water bath for cell - Autemp heater; 6"
crystallizing dish (for water bath); rubber pad for holding
cell.

4. Procedure

Step 1. Add 1 ml of standard U carrier to an aliquot of
sample in a 40-ml long taper centrifuge tube. Dilute to about
10 ml, heat to boiling, and precipitate $(\text{NH}_4)_2\text{U}_2\text{O}_7$ by the drop-
wise addition of conc. NH_4OH .

Step 2. Centrifuge and discard the supernate.

Step 3. Dissolve the precipitate in 1 to 2 ml of 1M HNO_3 ,
add 5.4 ml of H_2O , 3 drops of La carrier, and 10 drops of 5M
 $\text{NH}_2\text{OH} \cdot \text{HCl}$. Allow to stand for 5 min.

Step 4. Add 3 drops of conc. HF and allow to stand for 5
min. Centrifuge for 5 min, transfer supernate to a 40-ml
centrifuge tube, and discard the precipitate.

Step 5. Add 3 drops of La carrier and let stand for 5 min.
Centrifuge for 5 min, transfer supernate to a 40-ml centrifuge
tube, and discard the precipitate.

Step 6. Add 3 drops of Zr carrier and 15 drops of Ba

PROCEDURE 1 (Continued)

carrier. Centrifuge for 5 min and transfer the supernate to a 40-ml centrifuge tube, discarding the precipitate.

Step 7. Add 4 drops of conc. H_2SO_4 and centrifuge for 5 min. Transfer the supernate to a 40-ml centrifuge tube and discard the precipitate.

Step 8. Add 2 drops of Fe carrier, heat the solution to boiling, and precipitate $\text{Fe}(\text{OH})_3$ by the addition of conc. NH_4OH . Cool the tube under cold H_2O , centrifuge for 2-1/2 min, and transfer the supernate to a 40-ml centrifuge tube, discarding the precipitate.

Step 9. Add 0.4 to 0.5 ml of liquid Br_2 (Note 1) slowly to slight excess and boil the solution to a light yellow color. Add conc. NH_4OH until $(\text{NH}_4)_2\text{U}_2\text{O}_7$ precipitate forms. Cool under cold water, centrifuge, and save the precipitate.

Step 10. Add 1 to 2 ml of 1M HNO_3 and 10 ml of H_2O , heat the solution to boiling, and add conc. NH_4OH to reprecipitate $(\text{NH}_4)_2\text{U}_2\text{O}_7$. Centrifuge and save the precipitate.

Step 11. Add 1 to 2 ml of 1M HNO_3 , 10 ml of H_2O , 10 drops of 5M $\text{NH}_2\text{OH} \cdot \text{HCl}$, and 2 drops of Fe carrier. Let stand for 5 min. Heat the solution to boiling and precipitate $\text{Fe}(\text{OH})_3$ by addition of conc. NH_4OH . Cool the tube under cold H_2O , centrifuge for 2-1/2 min, and transfer the supernate to a 40-ml centrifuge tube, discarding the precipitate.

Step 12. Repeat Step 9.

Step 13. Add 1 ml of conc. HCl , 10 ml of H_2O , heat the solution to boiling and precipitate $(\text{NH}_4)_2\text{U}_2\text{O}_7$ with conc. NH_4OH . Cool the tube, centrifuge, and save the precipitate.

Step 14. Dissolve the precipitate in 1 ml of conc. HCl and 10 ml of H_2O . Add 2 gm of Zn metal (20 mesh, granular), and heat the mixture until the solution turns brown. Heat 1 additional minute.

PROCEDURE 1 (Continued)

Step 15. Let stand until the vigorous gas evolution subsides and decant into a 40-ml centrifuge tube. Discard the Zn.

Step 16. Heat the solution to boiling and precipitate $U(OH)_4$ (?) with conc. NH_4OH . (The precipitate will be greenish-black.) Centrifuge and save the precipitate.

Step 17. Dissolve the precipitate in 10 drops of conc. HCl. Add 5 ml of H_2O and 4 drops of conc. HF. Stir vigorously until UF_4 precipitates, add 7 drops of conc. NH_4OH and stir. Centrifuge 5 min and save the precipitate.

Step 18. Add 1 ml of conc. HCl, heat slightly, add 10 ml of H_2O , and heat the solution to boiling (the precipitate should dissolve). Add conc. NH_4OH and precipitate $U(OH)_4$ (?) (greenish-black precipitate).

Step 19. Repeat Step 17, except that 4 ml of H_2O are added instead of 5.

Step 20. Add 1 ml of conc. HNO_3 and heat until NO_2 ceases to be evolved. Add 10 ml of H_2O and precipitate $(NH_4)_2U_2O_7$ with conc. NH_4OH . Centrifuge, discard the supernate, and dissolve the precipitate in 1 ml of conc. HNO_3 .

Step 21. Add 10 ml of H_2O , 4 drops of Zr carrier, and 1 ml of 0.35M HIO_3 . Centrifuge, transfer the supernate to a 40-ml centrifuge tube, and discard the precipitate.

Step 22. Heat the solution to boiling and precipitate $(NH_4)_2U_2O_7$ with conc. NH_4OH . Centrifuge and discard the supernate.

Step 23. Dissolve the precipitate in 1 to 2 ml of 1M HNO_3 , dilute with 10 ml of H_2O , and centrifuge. Transfer the supernate to a 40-ml centrifuge tube and discard the precipitate.

Step 24. Reprecipitate $(NH_4)_2U_2O_7$ by boiling the solution and adding conc. NH_4OH . Centrifuge and save the precipitate.

Step 25. Add 5 drops of 8M HNO_3 and transfer to the plating cell which contains 10 ml of H_2O and 3 drops of 8M HNO_3 . Rinse the centrifuge tube with three washes each consisting of 5

PROCEDURE 1 (Continued)

drops of 8M HNO_3 and 0.5 ml of H_2O , transferring the washings to the plating cell.

Step 26. Add 10 ml of 4% $(\text{NH}_4)_2\text{C}_2\text{O}_4$ and wash the cell walls down with approximately 5 ml of H_2O . The total volume in the cell should be about 40 ml.

Step 27. Add 5 drops of methyl red solution, and conc. NH_4OH drop-wise until the solution turns yellow. Add 8M HNO_3 until the solution turns red or orange (one drop is usually required); then add 3 drops of HNO_3 in excess.

Step 28. Plate for 1-1/2 hours at 1.5 amp and 8 volts at 80 to 90°. For the first 30 min, at 10-min intervals add sufficient 8M HNO_3 to make the solution red to methyl red. At 40 min, add 3 drops of conc. NH_4OH , or enough to make the solution yellow to the indicator.

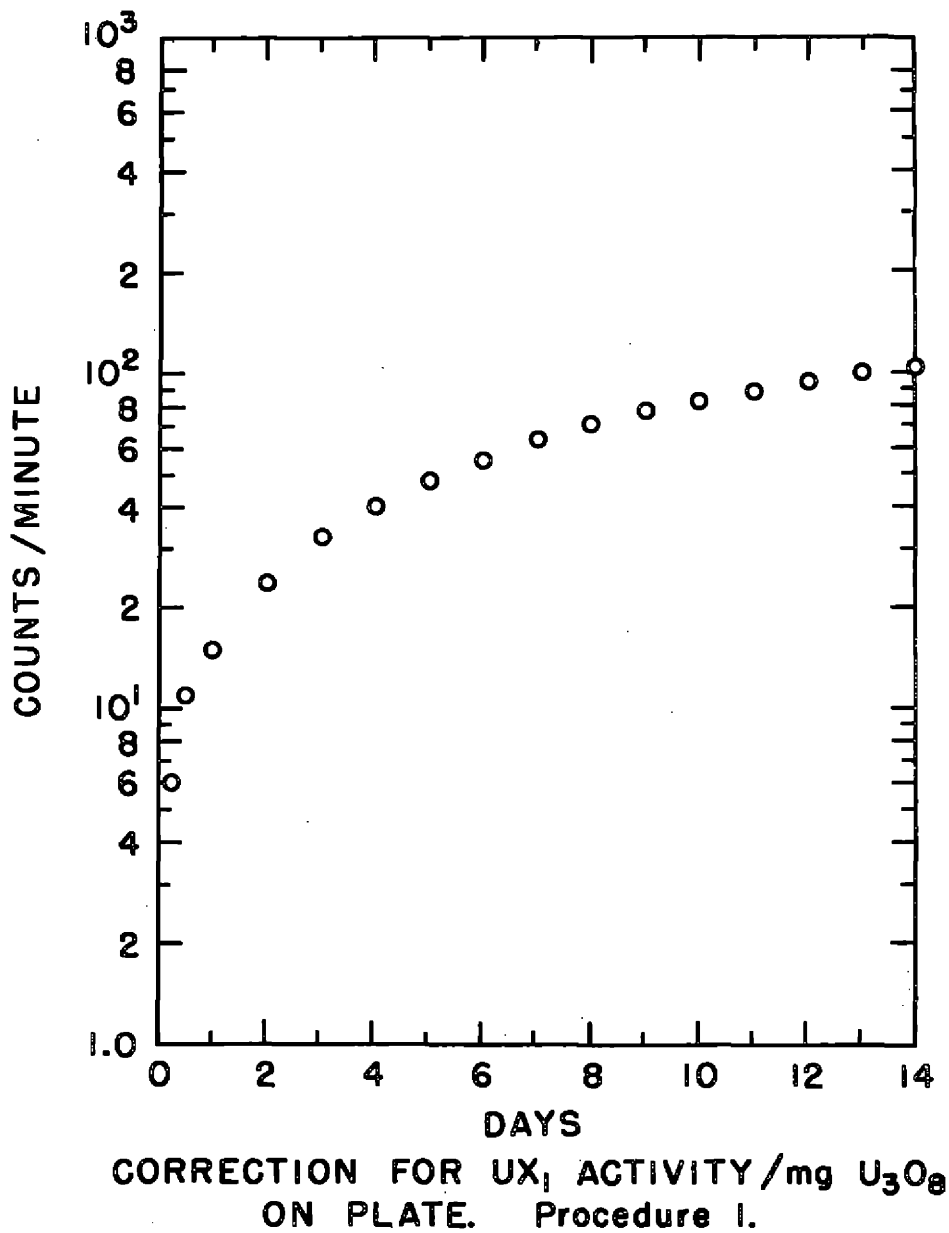
Step 29. Wash down the cell walls with H_2O to replenish that lost by evaporation, and continue electrolysis for an additional 50 min.

Step 30. Remove plate, wash with H_2O and methanol. Flame plate for 1 min. Cool, weigh as U_3O_8 , mount, and count. Correct for Th^{234} (UX_1) activity (see accompanying figure).

Notes

1. Liquid Br_2 destroys NH_2OH and also the uranium-hydroxylamine complex.

PROCEDURE 1 (Continued)



PROCEDURE 2: Purification of Uranium-240.

Source: E. K. Hyde and M. H. Studier, ANL-4182 (1948).

Editor's note: The following procedure was used to purify U^{240} formed by the second order neutron capture of U^{238} . The principal decontaminating step is the ether extraction of uranium from a reducing aqueous solution. Uranium is further purified by a number of precipitations that are not described in detail. These, however, are fairly easy to perform.

Irradiation and Chemical Procedure

Two grams of depleted uranium (1 part U^{235} per 30,000 parts U^{238}) as U_3O_8 in a small 2s aluminum capsule was irradiated in the Hanford pile for 12 hours including time for startup and shutdown. Six hours after the end of the irradiation the capsule and its contents were dissolved in nitric acid, using mercuric ion as catalyst for dissolving the aluminum. The uranium was extracted batchwise, the dissolved aluminum serving as a salting agent. The ether containing the uranium was then passed through two static wash columns packed with 3/32 inch stainless steel helices and filled with a solution 10M in ammonium nitrate, 0.1 M in nitric acid, 0.01 N in ferrous ion and 0.1 M in urea. Neptunium was reduced by the aluminum in the dissolver and by the ferrous ion in the wash columns to an unextractable oxidation state (Np IV and Np V). Additional ether was passed through the columns to strip out the uranium. These operations were carried out by remote control behind lead shielding. The initial dissolver solution measured roughly 50 roentgens per hour at 8 inches. The ether solution emerging from the second column and containing the uranium measured only about 3 mr per hour at the surface, and most of this was ether-soluble iodine fission product activity. The uranium was extracted from the ether into an aqueous ammonium sulfate solution and washed several times with ether to remove iodine activity. LaF_3 was precipitated from the uranyl nitrate solution after reduction with sulfur dioxide to remove any traces of Np^{239} which might have come through the ether extraction.

PROCEDURE 2 (Continued)

The uranium was further purified by precipitation as diuranate, sodium uranyl acetate, and peroxide and by a final ether extraction. Throughout this final series of purifications there was no detectable decrease in β -activity; this indicates that the uranium was radioactively pure.

Small aliquots of the final uranium solution were evaporated on platinum discs and ignited to U_3O_8 to study changes in activity. The remaining uranium solution was used for extraction of neptunium daughter fractions.

PROCEDURE 3: Purification of Irradiated U^{236} .

Source: S. Fried and H. Selig, Private communication.

Editor's note: The present procedure was used in an experiment designed to measure the thermal neutron fission cross section of U^{237} . The amount of U^{235} that can be tolerated in such an experiment is very small.

Two criteria were used in selecting the purification steps in the following procedure:

- 1) To obtain uranium free of fission products and other extraneous activities without introducing contaminant normal uranium in the procedure.*
- 2) The initial part should lend itself easily to remote control manipulation.

*The reagents used were carefully purified. Thus, the nitric acid and perchloric acid were redistilled in a quartz still. The NH_4NO_3 was prepared from gaseous ammonia and distilled HNO_3 . The HCl was prepared by passing HCl gas into triply distilled H_2O , etc.

PROCEDURE 3 (Continued)

Procedure:

A. In Cave

The irradiated uranium oxide (~0.3 mg) was dissolved in concentrated HNO_3 and made up to 2 M in HNO_3 with distilled H_2O to give total volume of about 15 ml. Some Fe^{++} was added to keep Pu and Np in +4 state. The solution was saturated with NH_4NO_3 and contacted four times with 10 ml portions of ether. Each contact was scrubbed twice with 2 M HNO_3 saturated with NH_4NO_3 . The combined ether extracts were back extracted three times with 5 ml portions of H_2O . The H_2O strip was evaporated to dryness and treated with HCl to destroy NH_4NO_3 carried over.

B. Outside Cave

The sample could now be handled easily outside the cave with a minimum of shielding, most of the activity being due to U^{237} . A mass spectrometric analysis showed it contained 0.5 weight % of U^{237} . A fission count showed that additional purification was necessary to remove Np^{238} formed by (n,γ) on Np^{237} which had built up during irradiation.

The sample was taken up in about 0.5 ml 6 N HCl and put on a small Dowex-1 column and washed. The Np comes off in 6 N HCl. Finally the uranium was eluted with 0.5 M HCl. The eluate was evaporated to dryness and taken up in 0.2 ml of 5 M HCl, 0.1 M KI and 0.05 M $\text{N}_2\text{H}_4\text{OH} \cdot 2\text{HCl}$. This was heated at 90° for 2 minutes, diluted to 0.5 M in HCl and TTA extracted twice for 15 minutes. The original fraction was washed twice with benzene and evaporated to dryness.

In order to clean up the uranium for a mass spectrometric analysis, it was subjected to another ether extraction as in the first step. After the NH_4NO_3 was destroyed the sample was fumed with HClO_4 to destroy any organic residue from the ether extraction.

PROCEDURE 4: Uranium and Plutonium Analysis

Source: B. F. Rider, J. L. Russell, Jr., D. W. Harris, J. P. Peterson, Jr., GEAP-3373 (1960).

Samples of dissolved irradiated fuel contain highly radioactive fission products. For this reason, uranium and plutonium are separated prior to analysis. The following procedure gives a good yield together with a good decontamination factor.

Reagents:

1. Distilled conc. HNO_3 .
2. 2 M HNO_3 - distilled conc. HNO_3 , double distilled H_2O .
3. U-233 solution, standardized.
4. Pu-236 solution, standardized.
5. KBrO_3 - Crystals, Reagent Grade. Low natural U blank.
6. 8 M NH_4NO_3 in 2 M HNO_3 - Place 200 ml distilled 16 M HNO_3 + 100 ml double distilled H_2O in a large beaker. Bubble NH_3 gas through solution until basic to pH paper. Boil off excess NH_3 (solution neutral). Transfer to mixing cylinder, add 50 ml of distilled 16 M HNO_3 , dilute to 400 ml. Check density of solution (1.31 ± 0.01 at 20°C .).
7. Hexone - distilled.
8. HCl - C.P. reagent. Low natural U blank.
9. 1 M HNO_3 - distilled conc. HNO_3 , double distilled H_2O .
10. 30% H_2O_2 - meets A.C.S. specification, low natural U blank.
11. 0.2 M T.T.A. in xylene - 4.44 gm T.T.A. dissolved in 100 ml distilled xylene.
12. Xylene - distilled.
13. Ether - distilled.
14. 0.05 M HNO_3 - distilled conc. HNO_3 , double distilled H_2O .
15. H_2O - double distilled.

Glassware:

All glassware used is Pyrex which has been soaked overnight in

PROCEDURE 4 (Continued)

50% HNO_3 and rinsed with double distilled water. Pipets are rinsed with 50% HNO_3 and double distilled water before using.

Separation and Decontamination Procedure:

1. Place the aliquot for analysis in a 15 ml cone and evaporate to about 1 ml. Add a suitable U-233 and Pu-236 spike, one drop conc. nitric acid, and several KBrO_3 crystals. Allow to stand for 1 hour to allow oxidation of Pu to PuO_2^{++} .
2. Add 1.5 ml 8 M NH_4NO_3 in 2 M HNO_3 , and evaporate to about 2 ml.
3. Prepare 2 scrub solutions in separate 15 ml cones, containing 1 ml of 8 M NH_4NO_3 in 2 M HNO_3 and about 10 mgs KBrO_3 . Preoxidize about 10 ml hexone with 2 ml of 2 M HNO_3 and KBrO_3 . Keep covered until ready for use.
4. Extract the U and Pu four times for five minutes with 2 ml portions of hexone (methyl isobutyl ketone), adding 1 drop of 16 M HNO_3 to the original solution after each extraction. Scrub each extract in turn with the two solutions prepared in step 3.
5. Strip the combined hexone extracts with five 2 ml portions of H_2O . Evaporate the combined aqueous portions to dryness, add a few drops of HNO_3 and HCl , take to dryness. Evaporate to dryness with HNO_3 under a gentle stream of pure nitrogen on a boiling water bath.
6. Prepare 3 ml of 1 M HNO_3 and 1 drop of 30% H_2O_2 , add 1 ml to the Pu and U residue from step 5 and two 1 ml portions to separate 15 ml cones.
7. Extract immediately the Pu 2 times for 20 min. with 2 ml portions of 0.2 M T.T.A. (thenoyltrifluoroacetone) in xylene. Scrub each in turn with solutions prepared in step 6. Save the aqueous phase for uranium. Combine the T.T.A. extracts and add a few crystals of trichloroacetic acid.

PROCEDURE 4 (Continued)

8. Mount the combined T.T.A. extracts on a platinum plate for alpha pulse analysis.
9. After pulse analysis, remove the Pu for mass analysis as follows: Cover disc with HF. Evaporate to dryness under a heat lamp. Again cover disc with HF and evaporate to dryness. Cover disc with conc. HNO_3 and evaporate to dryness. Repeat 3 or 4 times. Cover disc with conc. nitric, reflux a few seconds, and transfer with a pipet to a 15 ml cone. Repeat 3 or 4 times.
10. Evaporate the combined conc. HNO_3 refluxes to dryness. Treat residue with aqua regia and evaporate to dryness. Evaporate to dryness with conc. HNO_3 on a boiling water bath several times. Add 50 λ of 0.01 M HNO_3 to the evaporated sample and submit sample for mass spectrographic analysis.
11. Wash the original 1 M HNO_3 uranium fraction (Step 7) with xylene. Add 1 drop of HNO_3 and 3 drops of HCl to the washed 1 M HNO_3 and reflux for about one-half hour to destroy the organic present. Evaporate to dryness, flame gently to destroy organic matter and dissolve the residue with 2 drops HNO_3 and evaporate to dryness on a water bath.
12. Pipette three 1 ml portions of 8 M NH_4NO_3 in 2 M HNO_3 , dissolve the evaporated U fraction in one 1 ml portion. Place the other 2 portions in two 15 ml cones for scrub solutions.
13. Extract the U with four 2 ml portions of diethyl ether, adding 100 λ of conc. HNO_3 before each extraction. Scrub each extract in turn with 2 scrub solutions prepared in Step 12.
14. Evaporate the combined ether extracts over 1 ml of H_2O in a 15 ml cone. Evaporate to dryness.
15. Add 3 drops of HCl and 1 drop of HNO_3 , and evaporate to dry-

PROCEDURE 4 (Continued)

ness repeatedly until the organic is destroyed. Flame gently to expell ammonium salts. Then dissolve in HNO_3 and evaporate to dryness on a water bath. Add 50 λ of 0.05 M HNO_3 to the dry cone and submit sample for mass spectrographic analysis.

Plutonium Calculation:

To determine the amount of Pu in the original sample, it is necessary to measure in a Frisch chamber the alpha spectrum of the plate prepared in Step 8. The ratio of Pu-239 and Pu-240 activity to Pu-236 activity is calculated. If the ratio is multiplied by the original activity of Pu-236 added, the original activity of Pu-239 plus Pu-240 can be obtained. From the mass analysis a Pu-239 to Pu-240 atom ratio is obtained. The specific activity of the mixture is calculated from that of the individual isotopes. The Pu-239 plus Pu-240 activity can be converted to Pu-239 plus Pu-240 weight by dividing this activity by the specific activity of the mixture.

Uranium Calculation:

The ratio of the various U isotopes to U-233 from the mass spectrometer data is multiplied by the amount of U-233 spike originally added to the sample to obtain the amount of each uranium isotope present in the original sample.

PROCEDURE 5: Spectrophotometric Extraction Methods Specific for Uranium.

Source: W. J. Maeck, G. L. Booman, M. C. Elliott, and J. E. Rein, Anal. Chem. 31, 1130 (1959).

Abstract

Uranium as tetrapropylammonium uranyl trinitrate is quan-

PROCEDURE 5 (Continued)

titatively separated from large quantities of diverse ions by extraction into methyl isobutyl ketone (4-methyl-2-pentanone) from an acid-deficient aluminum nitrate salting solution. Milligram levels are determined by a direct absorbance measurement of the trinitrate complex in the separated organic phase at 452 m μ . Microgram amounts are determined by adding dibenzoylmethane (1,3-diphenyl-1,3-propanedione) in an ethyl alcohol-pyridine mixture to the separated organic phase and measuring the absorbance of the chelate at 415 m μ . The coefficient of variation is less than 1% at the 10-mg. and 25- γ levels. The limit of sensitivity is 0.8 γ for the dibenzoylmethane method.

Apparatus and Reagents

Absorbance measurements of the tetrapropylammonium uranyl trinitrate complex were made with a Cary Model 14 recording spectrophotometer and 1-cm. Corex cells. A Teflon 9 x 9 x 6 mm. spacer placed in the bottom of the cells permits absorbance measurements with 2 ml. of sample. Absorbance measurements of the dibenzoylmethane complex were made with a Beckman DU spectrophotometer and 5-cm. Corex cells.

Extractions were made in 125 x 15 mm. test tubes with polyethylene stoppers. A mechanical extraction device^A was used for agitation.

Reagent grade inorganic and Eastman Kodak Co. White Label organic chemicals were used without purification. Distilled water was used throughout. The uranium solutions were prepared by dissolving purified black oxide, U₃O₈, in a slight excess of nitric acid, and making to volume with water.

The dibenzoylmethane reagent is prepared by dissolving 0.1140 gram of dibenzoylmethane in 500 ml. of a 5% solution (v./v.) of ethyl alcohol in pyridine.

Salting and Scrub Solutions. A. 0.005M Tetrapropylammonium

PROCEDURE 5 (Continued)

Nitrate, 1N Acid-Deficient Salting Solution. Place 1050 grams of aluminum nitrate nonahydrate in a 2-liter beaker and add water to a volume of 850 ml. Heat, and after dissolution add 67.5 ml. of concentrated ammonium hydroxide. Stir for several minutes until the hydroxide precipitate dissolves. Cool to less than 50°C., add 10 ml. of 10% tetrapropylammonium hydroxide, and stir until dissolved. Transfer to a 1-liter volumetric flask and make to volume with water. A preliminary extraction with methyl isobutyl ketone is suggested to remove uranium contamination in which case tetrapropylammonium hydroxide will have to be re-added.

B. 0.025M Tetrapropylammonium Nitrate, 1N Acid-Deficient Salting Solution. Same as A except that 50 ml. of 10% tetrapropylammonium hydroxide is used.

C. 0.25M Tetrapropylammonium Nitrate, 1N Acid-Deficient Salting Solution. Neutralize 100 ml. of 10% tetrapropylammonium hydroxide to pH 7 with 5N nitric acid. Transfer to a large evaporating dish and let stand until a thick crystal slurry forms (which may take as long as 4 days). Place 210 grams of aluminum nitrate nonahydrate in a 400-ml. beaker and transfer the tetrapropylammonium nitrate crystals into the beaker with 20 ml. of water. Stir and add water to a volume of approximately 180 ml. Add 13.5 ml. of concentrated ammonium hydroxide and stir until dissolution is complete (which may require several hours). Transfer to a 200-ml. volumetric flask and make to volume with water.

D. Scrub Solution for Dibenzoylmethane Method. Add 940 grams of aluminum nitrate nonahydrate, 33 grams of tartaric acid, 31 grams of oxalic acid, and 64 grams of (ethylenedinitrilo)-tetraacetic acid to 100 ml. of water and 150 ml. of concentrated ammonium hydroxide. Heat with stirring until dissolved. Cool,

PROCEDURE 5 (Continued)

filter, transfer to a 1-liter volumetric flask, and make to volume with water. Remove uranium contamination by a methyl isobutyl ketone extraction.

E. Special Solutions. The following salting and scrub solutions are used in the dibenzoylmethane method for samples containing cerium(IV) or thorium.

1. Prepare an aluminum nitrate salting solution as A, but omit the tetrapropylammonium hydroxide.

2. Prepare a scrub solution by dissolving 15⁴ grams of ammonium acetate and 20 grams of the sodium salt of diethyldithiocarbamate in water to a volume of approximately 900 ml. Adjust to pH 7, filter, and make to a 1-liter volume with water.

3. Prepare a mercuric nitrate solution by dissolving 0.063 gram of mercuric nitrate in 90 ml. of 1N nitric acid and making to a 100-ml. volume with 1N nitric acid.

Procedures

Milligram Amounts of Uranium. With aqueous samples of 0.5 ml. or less and containing up to 2 meq. of acid, 0.5 to 12 mg. of uranium can be extracted from a salting solution which is 0.025M in tetrapropylammonium nitrate and 1N acid-deficient. Samples of high acidity should be neutralized to less than 2 meq. of free acid, or a salting solution which is 2N acid-deficient can be used for samples containing up to 6 meq. of acid. If cerium(IV) and thorium are present, the absorbance from uranium will be maximum if the combined uranium(VI), thorium, and cerium(IV) do not exceed 0.05 mmole in the sample aliquot. Samples that contain more than 0.05 mmole of combined uranium, cerium(IV), and/or thorium can be analyzed after dilution, provided the resulting sample aliquot contains more than 0.5 mg. of uranium. If this condition cannot be met, the 0.25M tetrapropylammonium nitrate salting solution is used, which can accommo-

PROCEDURE 5 (Continued)

date up to 0.5 mmole of combined uranium, cerium(IV), and thorium.

Pipet a sample of 0.500 ml. or less, containing from 0.5 to 12 mg. of uranium, into a test tube containing 4.0 ml. of salting solution B or C. Add 2.0 ml. of methyl isobutyl ketone, stopper, and extract for 3 minutes. Centrifuge to facilitate phase separation. Transfer as much as possible of the organic phase with a micropipet to a 1-cm. cell containing the Teflon spacer. Measure the absorbance at 452 m μ against a blank prepared by substituting 1N nitric acid for the sample.

Microgram Amounts of Uranium. Aqueous sample aliquots containing up to 2 mg. of uranium and as much as 8N in acid can be quantitatively extracted from a salting solution 0.005M in tetrapropylammonium nitrate^a. Neutralize samples of higher acidity to less than 8N before extraction.

SAMPLES WITHOUT CERIUM(IV) AND THORIUM. Pipet a sample of 0.500 ml. or less, containing from 0.8 to 75 γ of uranium, into a test tube containing 5.0 ml. of salting solution A. Add 2.0 ml. of methyl isobutyl ketone, stopper, and extract for 3 minutes. Centrifuge to facilitate phase separation. Transfer as much as possible of the organic phase to a test tube containing 5.0 ml. of scrub solution D, stopper, and mix for 3 minutes. Centrifuge to facilitate phase separation. Remove a 1.00-ml. aliquot of the organic phase and transfer to a 25-ml. flask. Add 15 ml. of the dibenzoylmethane-pyridine reagent and thoroughly mix. Allow to stand 15 minutes, transfer to a 5-cm. Corex cell, and measure the absorbance at 415 m μ compared to a blank prepared by substituting 1N nitric acid for the sample aliquot.

SAMPLES CONTAINING CERIUM(IV) OR THORIUM. Pipet a sample of 0.500 ml. or less, containing from 0.8 to 75 γ of uranium,

PROCEDURE 5 (Continued)

into a test tube containing 5.0 ml. of the salting solution E-1. Add 4.0 ml. of methyl isobutyl ketone, stopper, and extract for 3 minutes. Centrifuge to facilitate phase separation. Transfer as much as possible of the organic phase to another tube containing 5.0 ml. of scrub solution E-2, stopper, and mix for 20 minutes. Centrifuge as before. Transfer at least 3 ml. of the organic phase to a test tube containing 5.0 ml. of salting solution E-1. Add 0.5 ml. of scrub solution E-3, stopper, mix for 10 minutes, and centrifuge. Remove a 2.00-ml. aliquot of the organic phase and transfer to a 25-ml. flask. Add 15 ml. of the dibenzoylmethane-pyridine reagent and thoroughly mix. Let stand 15 minutes, transfer to a 5.0-cm. Corex cell, and measure the absorbance at 415 m μ compared to a blank prepared by substituting 1N nitric acid for the sample aliquot.

Calibration. Two different standards containing levels of uranium equivalent to approximately 0.1 and 0.7 absorbance are processed. The concentration of samples is established by the average absorptivity of these standards provided agreement within statistical limits (95% confidence level) is obtained.

^a W. J. Maeck, G. L. Booman, M. C. Elliott, J. E. Rein, Anal. Chem. 30, 1902 (1958).

PROCEDURE 6: Determination of Uranium in Uranium Concentrates.

Source: R. J. Guest and J. B. Zimmerman, Anal. Chem. 27, 931 (1955).

Abstract

A method is described for the determination of uranium in high grade uranium material. Uranium is separated from contaminants by means of an ethyl acetate extraction using aluminum nitrate as a salting agent. After the uranium has been stripped from the ethyl acetate

PROCEDURE 6 (Continued)

layer by means of water, colorimetric determination of the uranium is carried out by the sodium hydroxide-hydrogen peroxide method. The procedure is accurate, rapid, and easily adaptable to routine work.

Reagents and Apparatus

Reagents. Ethyl acetate (Merck, reagent grade).

ALUMINUM NITRATE SALTING SOLUTION. Place approximately 450 grams of reagent grade (Mallinckrodt) aluminum-nitrate [$\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$] in a 600-ml. beaker and add 25 to 50 ml. of distilled water. Cover the beaker and heat the mixture on a hot plate. If a clear solution does not result after 5 to 10 minutes of boiling, add 20 ml. of water, and continue the boiling for 5 more minutes. Repeat this step until a clear solution is obtained after boiling. Remove the cover glass and concentrate the solution by boiling until a boiling point of 130°C . is reached. Cover the beaker with a watch glass and either transfer the solution to a constant temperature apparatus or keep the solution warm, finally heating to just under boiling before use. If the solution is allowed to cool to approximately 60°C ., recrystallization of aluminum nitrate will take place. It is necessary, therefore, to dilute the salting agent solution by about one third in order to prevent recrystallization if the solution cools to room temperature. Accordingly, if the solution is to stand overnight, add 35 ml. of distilled water per 100 ml. of salting agent solution, mix well, and cover.

If the salting agent solution is to be stored, the following procedure has been found convenient. Adjust the solution to the proper concentration (boiling point, 130°C .) and transfer to a 100-ml. three-necked reaction flask set on a heating mantle. Adjust the heating so that the temperature of the solution is kept at about 110°C . In one of the necks place a water condenser, in another neck a thermometer, and in the third neck a removable ground-glass stop-

PROCEDURE 6 (Continued)

per. This third neck is utilized for pipetting the salting agent solution.

ALUMINUM NITRATE WASH SOLUTION. Add 100 ml. of aluminum nitrate salting solution (boiling point, 130°C.) to 73 ml. of distilled water and 4 ml. of concentrated nitric acid.

Apparatus. Beckman DU spectrophotometer.

Heating mantle.

Three-necked reaction flask (1000 ml.).

Water condenser.

No. 0 rubber stoppers. Boil twice in ethyl acetate before use.

Sixty-milliliter separatory funnels (Squibb, pear-shaped).

Procedure

Sample Dissolution. Place an appropriate quantity (1 to 5 grams) of the sample in a tared weighing bottle, stopper the bottle, and weigh the bottle and contents immediately. Carry out a moisture determination on a separate sample if uranium is to be calculated on a dry weight basis.

Bring the sample into solution in one of three ways: (1) nitric acid treatment, (2) multiacid treatment, or (3) sugar carbon-sodium peroxide fusion.

For the nitric acid treatment, dissolve the sample in a suitable quantity of nitric acid and transfer the solution and insoluble residue into an appropriate volumetric flask and make up to volume. Regulate the dilution so that the aliquot chosen for extraction will contain between 10 and 30 mg. of uranium oxide if the final dilution for the colorimetric finish is to be 250 ml. Adjust the acidity of the sample solution to about 5% in nitric acid.

If nitric acid treatment is not sufficient, treat the sample with hydrochloric acid, nitric acid, perchloric acid, and finally sulfuric acid. If necessary, add a few milliliters

PROCEDURE 6 (Continued)

of hydrofluoric acid. Fume the sample to dryness and leach the residue with nitric acid, finally transferring the solution and residue to an appropriate volumetric flask and adjusting to 5% in nitric acid as in the single acid treatment.

If the sample is refractory, use the sugar carbon-sodium peroxide fusion method described by Muehlberg^a. After dissolution of the sample in this manner, transfer the acidified solution to an appropriate volumetric flask and dilute so that the final solution is 5% in nitric acid.

Aliquot solution samples directly or dilute as required for an ethyl acetate extraction. If the sample is aliquoted directly for an extraction, add 5 drops of concentrated nitric acid per 5-ml. aliquot of sample and standards before extraction. Where samples are diluted before aliquots are taken for extraction, adjust the acidity so that the final volume is 5% in nitric acid.

Ethyl Acetate Extraction. Place an appropriate aliquot (usually 5 ml.) in a 60-ml. separatory funnel, the stopcock of which has been lubricated with silicone grease. Add, by means of a graduated pipet, 6.5 ml. of aluminum nitrate solution per 5 ml. of sample solution. The aluminum nitrate salting solution should be added while hot (above 110°C.). Cool the solution to room temperature and add 20 ml. of ethyl acetate. Stopper the separatory funnels with pretreated rubber stoppers. Shake the mixture for 45 to 60 seconds. Occasionally crystallization will take place in the separatory funnel near the stopcock. In such a case place the lower part of the separatory funnel in a beaker of hot water until the solidified portion dissolves.

After the layers have separated, drain off the aqueous (lower) layer. Occasionally a cloudiness will appear at the boundary of the aqueous and organic layer. This cloudy portion should not be drained off. Add 10 ml. of aluminum nitrate wash

PROCEDURE 6 (Continued)

solution to the funnel and again shake the mixture for 45 to 60 seconds. Drain off the aqueous layer, once again being careful to retain the cloudy portion at the boundary in the funnel. Rinse inside the stem of the separatory funnel with a stream of water from a wash bottle.

Water Stripping of Uranium from Ethyl Acetate Layer Followed by Sodium Hydroxide-Hydrogen Peroxide Colorimetric Finish. Add 15 ml. of water to the separatory funnel containing the ethyl acetate, stopper the flask, and shake the mixture for about 1 minute. After washing off the stopper with water, drain the aqueous layer into a volumetric flask of suitable size and wash the separatory funnel and ethyl acetate layer 4 or 5 times with 5-ml. portions of water by means of a wash bottle. Combine the aqueous fractions.

Add enough 20% sodium hydroxide solution (w./v.) to neutralize the solution and dissolve any precipitated aluminum hydroxide, then add 10 ml. in excess per 100 ml. of final volume. Add 1 ml. of 30% hydrogen peroxide per 100 ml. of final volume and make up the volume to the mark with distilled water. Read the absorbance after 20 minutes on the Beckman DU spectrophotometer at 370 m μ against a reagent blank, using 1-cm. Corex cells and a slit width of 0.2 mm. Compare the absorbances of the samples against the absorbances of standard uranium solutions which have been carried through the procedure at the same time. Choose the standards so that they cover the range into which the samples are expected to fall, using a ratio of one standard to six samples. In practice it is customary to work between the limits of 10 and 30 mg. of uranium oxide. This is arranged by estimating the required sample weights and diluting and sampling accordingly. The final volume for colorimetric reading is usually 250 ml.

PROCEDURE 6 (Continued)

Double Extraction of Uranium with Ethyl Acetate Followed by Application of Differential Colorimetry. Uranium determinations requiring the highest accuracy may be carried out by a double extraction of uranium with ethyl acetate followed by the application of differential colorimetry as described by Hiskey and others.^{b-f} In such a case it is recommended that between 100 and 150 mg. of uranium oxide be extracted, and a wave length of 400 m μ be used during the colorimetric finish. The procedure described below has been found satisfactory.

Extract an appropriate aliquot of the sample solution with 20 ml. of ethyl acetate as described above. Draw off the aqueous layer into a second separatory funnel containing 10 ml. of ethyl acetate. Stopper the funnels and shake the mixture for 45 to 60 seconds. Drain off and discard the aqueous layer. Add 10 ml. of aluminum nitrate wash solution to the first ethyl acetate extract, stopper, and shake the mixture for 45 to 60 seconds. Drain off the aqueous layer into the separatory funnel containing the second ethyl acetate extract, stopper, and shake the mixture for 45 to 60 seconds. Drain off and discard the aqueous layer. Combine the ethyl acetate fractions. Rinse the second separatory funnel with 20 ml. of water, draining the washings into the separatory funnel containing the combined ethyl acetate fractions. Shake the mixture for 1 minute. Continue the water stripping as described above, collecting the fractions in an appropriate volumetric flask. Finish colorimetrically as described previously, allowing the strongly colored solution to stand 1 to 2 hours to ensure stability before reading as a fading effect of about 0.005 absorbance (optical density) has sometimes been noted on freshly prepared samples.

Read the absorbance of the sample solution on the Beckman DU spectrophotometer at 400 m μ against a reference solution

PROCEDURE 6 (Continued)

which contains a known amount of uranium and has been carried through the extraction and color development procedure in the same manner as the sample. Also carry along other standards containing slightly higher and lower amounts of uranium than the sample. Determine the concentration of uranium in the sample either by the calibration-curve method or the correction method, as described by Neal⁶. If the amount of uranium in the sample is not known, make a test run by taking an aliquot of the sample solution and assaying for uranium by the more rapid single extraction method. The standard solutions to be used can then be chosen according to the result obtained.

Removal of Interfering Thorium. After an ethyl acetate extraction, strip the uranium in water from the ethyl acetate and collect the uranium fraction in a 250-ml. beaker. Add enough 20% (w./v.) sodium hydroxide solution to neutralize the solution and redissolve precipitated aluminum hydroxide. Then add 10-ml. excess of 20% sodium hydroxide solution and 1 ml. of 30% hydrogen peroxide per 100 ml. of final volume. Filter the solution through an 11-cm. 41H filter paper (Whatman), collecting the filtrate in a volumetric flask of suitable size. Wash the paper and precipitate once with 5 ml. of a solution of 2% sodium hydroxide containing 0.1 ml. of 30% hydrogen peroxide. Redissolve the precipitate by washing the paper with 10 ml. of 10% nitric acid solution, collecting the washings in the original beaker. Neutralize the solution with 20% sodium hydroxide solution, and add 2 ml. in excess. Add 0.5 ml. of 30% hydrogen peroxide, and filter off the precipitate on the original filter paper, washing as before and collecting the filtrates in the original volumetric flask. If the precipitate on the paper is colored yellow, repeat this step. Make the solution in the volumetric flask up to volume and read the absorbance

PROCEDURE 6 (Continued)

on the spectrophotometer. Carry standards through the same procedure as the samples.

^a W. L. Muehlberg, Ind. Eng. Chem. 17, 690 (1925).

^b C. F. Hiskey, Anal. Chem. 21, 1440 (1949).

^c C. F. Hiskey, J. Rabinowitz, and I. G. Young, Anal. Chem. 22, 1464 (1950).

^d G. W. C. Milner and A. A. Smales, Analyst 79, 414 (1954).

^e W. T. L. Neal, Analyst 79, 403 (1954).

^f I. G. Young, C. F. Hiskey, Anal. Chem. 23, 506 (1951).

PROCEDURE 7: Uranium-237.

Source: B. Warren, LA-1721 (Rev) (1956).

1. Introduction

In the carrier-free method for the determination of U^{237} , the principal decontamination step (which is preceded by a $La(OH)_3$ scavenge and partial removal of plutonium as the cupferron complex) is the extraction of uranium into 30% TBP (tertiary butyl phosphate) in benzene. Additional decontamination is effected by adsorption of uranium, first on an anion and then on a cation exchange resin. The uranium is finally electroplated on platinum. The chemical yield is 40 to 60% and is determined through the use of U^{233} tracer. The U^{237} is β -counted in a proportional counter with a 2.61-mg/cm^2 Al absorber, and from the number of counts the number of atoms of the isotope is calculated. Four samples can be run in about 6 hours.

PROCEDURE 7 (Continued)

2. Reagents

U^{233} tracer: amount determined by the α -counting technique employed

La carrier: 10 mg La/ml [added as $La(NO_3)_3 \cdot 6H_2O$]

Fe carrier: 10 mg Fe/ml [added as $Fe(NO_3)_3 \cdot 9H_2O$ in very dilute HNO_3]

HCl: 0.1M

HCl: 5M

HCl: 10M

HCl: conc.

HNO_3 : 3M

HNO_3 : 5M

HNO_3 : conc.

NH_4OH : conc.

$NH_2OH \cdot HCl$: 5M

$(NH_4)_2C_2O_4$ in H_2O : 4%

TBP (tertiary butyl phosphate): 30% by volume in benzene (Note 1)

Aqueous cupferron reagent: 6%

Methyl red indicator solution: 0.1% in 90% ethanol

Methanol: anhydrous

Chloroform

NH_3 : gas

Cl_2 : gas

3. Equipment

Centrifuge

Fisher burner

Block for holding centrifuge tubes

Pt-tipped tweezers

Steam bath

5-ml syringe and transfer pipets

Mounting plates

PROCEDURE 7 (Continued)

40-ml centrifuge tubes: Pyrex 8140 (three per aliquot of sample)

Ion exchange columns:

8 cm x 3 mm tubing attached to bottom of 15-ml centrifuge tube

Anion resin: 5 cm Dowex A2-X8, 400 mesh, (Note 2)

Cation resin: 5 cm Dowex 50-X8, 100 to 200 mesh, (Note 2)

Stirring rods

Motor-driven glass stirrers

Plating set-up: same as that used in Procedure 1 except that the Pt cathode is a 1" disk and the glass chimney has a 7/8" i.d.

4. Procedure

Step 1. To an aliquot of sample not exceeding 20 ml in a 40-ml centrifuge tube, add 1 ml of U^{233} tracer and 3 drops of La carrier, and bubble in NH_3 gas until the precipitate which forms coagulates. Digest for 15 min on a steam bath, centrifuge, and discard the supernate.

Step 2. Dissolve the precipitate in 0.6 ml of conc. HCl and dilute to 10 ml with H_2O . Add 5 drops of 5M $NH_2OH \cdot HCl$ and 2 drops of Fe carrier (if this element is not already present), and allow to stand for 10 min. Add 4 ml of chloroform, 6 ml of 6% cupferron, and extract the Pu(IV)-cupferron complex by stirring for 2 min. Remove the chloroform layer by means of a transfer pipet and discard. Extract the aqueous phase three additional times with chloroform. To the aqueous layer add 3 drops of La carrier and bubble in NH_3 gas until the precipitate formed coagulates. Digest for 15 min on a steam bath, centrifuge, and discard the supernate.

Step 3. Dissolve the precipitate in 1.6 ml of conc. HNO_3 , dilute to 5 ml with H_2O , add 2 ml of TBP solution, and stir for 2 min. Draw off the TBP layer and transfer to a clean 40-

PROCEDURE 7 (Continued)

ml centrifuge tube. Extract again with 2 ml of TBP solution and combine with the previous extract. Add 1 ml of TBP solution to the original tube, draw it off, and combine with the other extracts.

Step 4. Wash the TBP extracts with two 3-ml portions of 5M HNO_3 , discarding the washings. Bubble in Cl_2 gas for 5 min at a vigorous rate.

Step 5. Transfer the solution to a Dowex A2 anion exchange column. Permit one-half of the solution to pass through the resin under 8 to 10 lb pressure. Add 1 ml of conc. HCl to the column and allow the remainder of the solution to pass through under pressure. Wash the column twice with 2-1/2 ml of 10M HCl and then twice with 5M HCl, discarding the washings. Elute the U with two 2-1/2-ml portions of 0.1M HCl, catching the eluate in a clean 40-ml centrifuge tube.

Step 6. Dilute the eluate to 10 ml with H_2O and pass through a Dowex 50 cation exchange column under 1 to 2 lb pressure. Wash the resin three times with 2-1/2-ml portions of 0.1M HCl and discard the washings. Elute the U with two 2-1/2-ml portions of 3M HNO_3 into the plating cell.

Step 7. Add 5 ml of 4% $(\text{NH}_4)_2\text{C}_2\text{O}_4$, 3 drops of methyl red indicator solution, and make basic by the dropwise addition of conc. NH_4OH . Make the solution just red to the indicator by the dropwise addition of 6M HNO_3 , and add 3 drops in excess. Plate at 1.1 amp and 8 volts for 1½ hr at 80°C. At the end of the first 10 min, add 3 drops of methyl red solution and make acid with 6M HNO_3 . Check acidity at two additional 10-min intervals, and at the end of 40 min add 3 drops of conc. NH_4OH . At 10-min intervals thereafter check to see that the plating solution is just basic to the indicator. Remove the cell from the water bath, wash three times with methanol,

PROCEDURE 7 (Continued)

and dismantle the cell, carefully keeping the Pt disk flat. Flame the disk over a burner. α -count the U^{233} , mount, and β -count in a proportional counter with a 2.61-mg/cm² Al absorber.

Notes

1. The TBP is purified before use by washing first with 1M NaOH and then with 5M HNO₃.
2. Before use, both the anion and cation resin are washed alternately at least five times each with H₂O and HCl, and are then stored in H₂O.
3. See Procedure 1 for the correction for UX₁ activity per mg U₃O₈ on plate.

PROCEDURE 8: Radioassay of Uranium and Plutonium in Vegetation, Soil and Water.

Source: E. L. Geiger, Health Physics 1, 405 (1959).

Abstract

A method is discussed for the separation of uranium and plutonium from vegetation, soil and water. The method is based on the extraction of uranium and plutonium from 4 to 6 N nitric acid into 50% tri-n-butyl phosphate in n-tetradecane diluent. Uranium and plutonium are recovered together with sufficient reduction in total solids to allow α -counting and pulse height analysis. Data from several hundred "spiked" samples to which uranium and plutonium were added indicate a nearly equal recovery of uranium and plutonium. Average recoveries are 76 ± 14 per cent for vegetation, 76 ± 16 per cent for soil, and 82 ± 15 per cent for water. The procedure is designed for samples that may be collected and analyzed

PROCEDURE 8 (Continued)

for radioactivity as a part of a health physics regional monitoring program.

Procedures

Preparation of samples

Vegetation. Cut oven-dried vegetation into small pieces and weigh 10.0 g into a 150 ml beaker. Heat the sample at 600°C, starting with a cold muffle furnace. When only white ash remains, remove the beaker from the muffle furnace and allow to cool. Carefully add 2 ml of water, then add 10 ml of 8 N HNO_3 -0.5 M $\text{Al}(\text{NO}_3)_3$ solution. Cover the beaker with a watch glass and boil the solution for 5 min. Allow to cool, add 1 ml of 2 M KNO_2 solution and transfer the sample to a 100 ml centrifuge tube. Use 4 N HNO_3 to complete the transfer. Centrifuge and decant the supernate into a 125 ml cylindrical separatory funnel graduated at 30 ml. Wash the residue with 4 N HNO_3 , centrifuge, and decant the wash solution to the separatory funnel. The acid normality of the combined solutions at this point is 4-6 N and the total volume should not exceed 29 ml. Proceed to the extraction procedure.

Soil. Grind 5 g of oven-dried soil with a mortar and pestle until the entire sample can pass through a 200-mesh sieve. Weigh 1.0 g of the 200-mesh soil into a 50 ml platinum crucible and heat the sample at 600°C for 4 hr. Remove the sample from the muffle furnace and allow to cool. Add 3 ml of 70% HNO_3 and 10 ml of 48% HF then stir the sample for 2-3 min with a platinum rod. Heat the sample in a 200°C sand bath until all traces of moisture are removed. Repeat this HNO_3 - HF treatment being careful that the sample is completely dry before proceeding to the next step. Allow the sample to cool and then add 15 ml of 6 N HNO_3 -0.25 M $\text{Al}(\text{NO}_3)_3$ solution. Cover with a watch glass and heat in the sand bath for 5 min.

PROCEDURE 8 (Continued)

Allow to cool and decant the solution through a filter, such as Whatman no. 40, into a 125 ml cylindrical separatory funnel graduated at 30 ml. Leave as much of the residue as possible in the crucible and repeat the hot 6 N HNO_3 -0.25 M $\text{Al}(\text{NO}_3)_3$ treatment. Allow to cool and then filter. Proceed to the extraction procedure.

Water. Place 1 l. of the sample in a 1.5 l. beaker and if basic, neutralize with nitric acid. Add 15 ml of 70% HNO_3 and evaporate to 30-40 ml. Decant the solution through a filter, such as Whatman no. 40, into a 100 ml beaker. Wash the 1.5 l. beaker, the residue and the filter with 4 N HNO_3 . Evaporate the combined solution in the 100 ml beaker to 5 ml. Add 20 ml of 4 N HNO_3 , cover with a watch glass, and heat for 5 min. Transfer the sample to a 125 ml cylindrical separatory funnel graduated at 30 ml. Wash the beaker with 4 N HNO_3 and transfer to the separatory funnel, being careful that the total volume in the separatory funnel does not exceed 29 ml. Proceed to the extraction procedure.

Extraction

Add 1 ml of 2 M KNO_2 to the sample in the 125 ml cylindrical separatory funnel. Dilute to the 30 ml mark with 4 N HNO_3 and stir the solution briefly. Add 30 ml of 50% TBP in n-tetradecane. Agitate the solution vigorously for 4 min with an air-driven stirrer. Discard the acid portion (lower layer). Wash the TBP portion with 4 N HNO_3 and again discard the acid portion. Back extract with seven 15 ml portions of distilled water, collecting the strip solution in a 150 ml beaker. Evaporate the combined aqueous portions to 10-15 ml, then quantitatively transfer the solution to a flamed stainless steel planchet. Allow to dry under a heat lamp, flame the planchet to burn off organic residue, and count on an α -counter. Retain for pulse-height analysis if the α -count exceeds a specified level.

PROCEDURE 9: Separation of Uranium by Solvent Extraction with Tri-n-octylphosphine Oxide.

Source: C. A. Horton and J. C. White, Anal. Chem. 30, 1779 (1958).

Abstract

A simple, rapid method for the determination of uranium in impure aqueous solutions was developed. Uranium(VI) is extracted by 0.1M tri-n-octylphosphine oxide in cyclohexane from a nitric acid solution. A yellow color is formed in an aliquot of the organic extract by addition of dibenzoylmethane and pyridine in ethyl alcohol. Interference by cations is minimized or eliminated by selective reduction, by fluoride complexation, or by absorbance measurement at 416 m μ rather than 405 m μ , the wave length of maximum absorbance. Interference by excess fluoride or phosphate is eliminated by addition of aluminum nitrate before extraction. The range of the method is 20 to 3000 γ of uranium in the original solution, and the standard deviation is $\pm 2\%$.

Apparatus and Reagents

Absorbance measurements were made with a Beckman DU spectrophotometer, using 1.00-cm. Corex or silica cells.

Phase equilibration for most extractions was carried out in the bottom portion of the apparatus (see accompanying figure). Phase separation and removal of aliquots of the upper organic phase occurred after inverting the apparatus so that the solution was in the portion of this apparatus shown on top in the figure. Some extractions were carried out in 60- or 125-ml. separatory funnels.

STANDARD URANIUM SOLUTIONS. A stock solution of 24.0 mg. of uranium(VI) per ml. was prepared by heating 7.10 gram of uranium (IV-VI) oxide (U_3O_8), prepared from pure uranium(VI) oxide (UO_3), in 10 ml. of perchloric acid to dissolve it, and then diluting the resultant solution to 250 ml. with water. Dilutions of this solution were made as required. Another standard solution in 5% sulfuric acid was also used in checking the spectrophotometric calibration curve.

PROCEDURE 9 (Continued)

DIBENZOYLMETHANE. A solution that contained 1 gram of dibenzoylmethane (1,3-diphenyl-1,3-propanedione), obtained from Eastman Kodak Co., in 100 ml. of 95% ethyl alcohol was prepared for use as the chromogenic agent.

PYRIDINE. For most of the work, a solution prepared by mixing 1 volume of redistilled reagent grade pyridine and 1 volume of 95% ethyl alcohol was used.

TRI-n-OCTYLPHOSPHINE OXIDE. A 0.10M solution of this material, prepared in the authors' laboratory, in cyclohexane, Eastman 702 or 702S, was used. This phosphine oxide is now available commercially from Eastman (EK 7440).

Sodium bisulfite, 10 (w./v.) % in water, stored at about 10°C.

Hydroxylamine sulfate, 2M in water.

Potassium fluoride, 1M in water, stored in a polyethylene bottle.

Procedure

Preliminary Treatment. A. Samples which do not contain interfering ions. Pipet a 0.5- to 8-ml. aliquot of a solution in nitric, perchloric, or sulfuric acid, estimated to contain 15 to 2500 γ of uranium(VI), into the bottom portion of the extraction vessel. By the addition of strong 10M sodium hydroxide, nitric acid, or sodium nitrate, adjust the solution so that a total aqueous volume of 10 ml. will be 1 to 3N in hydrogen ion and 2 to 4M in nitrate ion. For almost neutral solutions, 2 ml. of concentrated nitric acid will give the correct concentrations for a 10-ml. aqueous volume. Adjust the total volume to 10 ml. Up to 12 ml. of aqueous solution can be shaken with 5 or 10 ml. of extractant in the apparatus without undue splashing. If the total aqueous volume is greater than 12 ml. after adjusting the acidity and nitrate content, perform the extraction in a separatory funnel instead of the special extraction vessel.

PROCEDURE 9 (Continued)

B. Samples containing iron(III), chromium(VI), or vanadium(V). Pipet a 0.5- to 6-ml. aliquot of a solution in dilute perchloric or sulfuric medium, estimated to contain 15 to 2500 γ of uranium(VI), into the bottom portion of the extraction vessel. Reduce the iron(III) to iron(II), the chromium(VI) to chromium(III), and the vanadium(V) to vanadium(IV or III) without reducing the uranium(VI) to uranium(IV). Sodium bisulfite is a satisfactory reductant if the solutions are boiled to remove excess sulfur dioxide. Hydroxylamine sulfate is also a satisfactory reductant, but amalgamated zinc is unsuitable. Add sufficient nitric acid or sodium nitrate and water so that the final aqueous volume of 8 to 12 ml. will be 1 to $3N$ in hydrogen ion and 2 to $4M$ in nitrate ion.

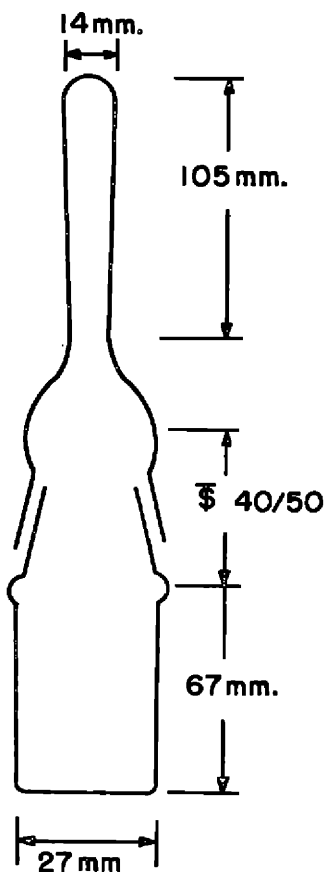
C. Samples containing titanium, thorium, hafnium, zirconium, or iron(III), but only traces of aluminum. Pipet a 0.5- to 6-ml. aliquot of a solution in dilute nitric, perchloric, or sulfuric acid, estimated to contain 15 to 2500 γ of uranium, into the bottom of an extraction vessel. Add sufficient base or acid, nitrate, and water to give a volume of about 8 ml., such that the solution is 1 to $3N$ in hydrogen ion and 2 to $4M$ in nitrate ions. Add up to a maximum of 2.5 ml. of $1M$ potassium fluoride when the concentrations of interfering ions are unknown. If high concentrations of these ions are known to be present, additional fluoride can be tolerated.

D. Samples containing excessive fluoride or phosphate. Pipet an aliquot into an extraction vessel, and adjust the acid and nitrate contents as in Treatment C. Add sufficient aluminum nitrate to complex the fluoride and phosphate estimated to be present. The total volume should be 12 ml. or less.

Extraction. For amounts of uranium estimated to be under about 1400 γ , pipet 5 ml. of $0.1M$ tri-n-octylphosphine oxide in cyclohexane into the extraction vessel containing the treated

PROCEDURE 9 (Continued)

sample. For 1400 to 3000 γ of uranium, use 10 ml. of extractant. Attach the top of the vessel and shake for 10 minutes on a reciprocating shaker. Invert the extraction apparatus for separation of the phases and removal of aliquots of the upper organic phase.



EXTRACTION APPARATUS
USED IN Procedure 9.

Color Development. Transfer by pipet a 1-, 2-, or 3-ml. aliquot of the organic extract into a 10- or 25-ml. volumetric flask such that the final solution will contain between 0.5 and 10 γ of uranium per ml. For 10-ml. volumes, add 1.0 ml. of 50%

PROCEDURE 9 (Continued)

pyridine in ethyl alcohol, 2 ml. of 1% (w./v.) % dibenzoylmethane in ethyl alcohol, and 95% ethyl alcohol to volume. For 25-ml. volumes, use 2.5 ml. of pyridine and 5 ml. of dibenzoylmethane. After 5 or more minutes, measure the absorbance at 405 m μ in 1-cm. cells, using 95% ethyl alcohol as a reference solution. For samples receiving Treatment C, also measure the absorbance at 416 m μ . A blank should be carried through the entire procedure daily. Calculate the uranium content using the factors obtained by extracting standard pure uranium solutions as directed in Treatment A, and measured at both 405 and 416 m μ .

PROCEDURE 10: Radiochemical Determination of Uranium-237.

Source: F. L. Moore and S. A. Reynolds, Anal. Chem. 31, 1080 (1959).

Abstract

A radiochemical method for the determination of uranium-237 is based on complexing the uranyl ion in alkaline solution with hydroxylamine hydrochloride, followed by scavenging with zirconium hydroxide and extraction of the uranium from hydrochloric acid solution with triisooctylamine-xylene. The technique has been applied successfully to the determination of uranium-237 in homogeneous reactor fuel solutions.

Preparation and Standardization of Uranium Carrier

Weigh out approximately 50 grams of uranyl nitrate hexahydrate. Dissolve and make to 1 liter with 2 M nitric acid. Standardize the carrier by pipetting 5-ml. aliquots into 50-ml. glass centrifuge cones and precipitating ammonium diuranate by adding concentrated ammonium hydroxide. Filter quantitatively through No. 42 Whatman filter paper and ignite in porcelain crucibles at 800°C. for 30 minutes. Weigh as U₃O₈.

PROCEDURE 10 (Continued)

Procedure

In a 40-ml. tapered centrifuge tube add 1 ml. of uranium carrier and 0.2 ml. of zirconium carrier (approximately 10 mg. per ml. of zirconium) to a suitable aliquot of the sample solution. Dilute to approximately 10 ml., mix well, and precipitate ammonium diuranate by the addition of concentrated ammonium hydroxide. Centrifuge for 2 minutes and discard the supernatant solution. Wash the precipitate once with 15 ml. of ammonium hydroxide (1 to 1).

Dissolve the precipitate in 1 to 2 ml. of concentrated hydrochloric acid solution, dilute to about 10 ml., add 1 ml. of hydroxylamine hydrochloride (5 M), and mix well. Precipitate zirconium hydroxide by the addition of concentrated ammonium hydroxide. Centrifuge for 2 minutes, add 0.2 ml. of zirconium carrier, and stir the supernatant solution, being careful not to disturb the precipitate. Centrifuge for 2 minutes. Add 0.2 ml. of zirconium carrier and repeat.

Transfer the supernatant solution to another 40-ml. centrifuge tube, add several drops of phenolphthalein, and adjust the pH to approximately 8 by adding concentrated hydrochloric acid solution dropwise. Add an equal volume of concentrated hydrochloric acid solution and extract for approximately 0.5 minute with a one-half volume portion of 5% (w./v.) triisooctylamine-xylene. Discard the aqueous phase. Wash the organic phase by mixing for 0.5 minute with an equal volume portion of 5 M hydrochloric acid solution. Repeat the wash step. Strip the uranium from the organic phase by mixing thoroughly with an equal volume portion of 0.1 M hydrochloric acid solution for 0.5 minute. Discard the organic phase.

Add 0.2 ml. of zirconium carrier, mix well, and repeat the above procedure, beginning with the precipitation of ammonium diuranate.

PROCEDURE 10 (Continued)

Finally, precipitate ammonium diuranate by the addition of concentrated ammonium hydroxide. Centrifuge for 2 minutes. Decant and discard the supernatant solution. Filter on No. 42 Whatman filter paper and ignite at 800°C. for 30 minutes.

Weigh the uranium oxide on a tared aluminum foil (0.0009 inch), fold, and place in a 10 x 75 mm. culture tube. Insert a suitable cork and count the uranium-237 gamma radioactivity in a well-type scintillation counter. Count the same day of the last chemical separation.

Apply a blank correction if very low uranium-237 levels are being determined. Determine this correction by taking the same aliquot of uranium carrier through the exact procedure described above. The blank correction is due primarily to the gamma radioactivity associated with the uranium-235 in the uranium carrier.

PROCEDURE 11: Separation of Uranium and Bismuth.

Source: A. Krishen and H. Freiser, Anal. Chem. 29, 288 (1957).

Editor's note: Uranium has been separated from a 5000-fold excess of bismuth by the following method. Uranium is, however, not completely extracted (only 98.48% at pH values greater than 6.5). If this correction is applied, uranium is quantitatively determined by polarographic means within an experimental error of ±1%.

Reagents

Acetylacetone. Commercial acetylacetone was purified by the method described by Steinbach and Freiser.^a

Procedure

Analysis. The method of Souchay and Faucherre,^b using 0.1M EDTA and 2M sodium acetate as a supporting electrolyte, was found to be suitable in the presence of dissolved acetylacetone. The half-wave potential was shifted to -0.47 volt but the wave height was not affected.

PROCEDURE 11 (Continued)

Separation. Solutions of uranyl sulfate containing 0.1 and 1.0 mg. of uranium were mixed with different amounts of bismuth trichloride solution. A solution of the disodium salt of EDTA was added to give a bismuth to EDTA ratio of 1 to 30. The pH of the mixture was raised to 7.5 by careful addition of 1N sodium hydroxide. Then approximately 10 ml. of acetylacetone was added and the mixture shaken for 10 minutes. The acetylacetone phase was separated, filtered, and made up to a volume of 10 ml., of which 2 ml. was withdrawn by a pipet into a 10-ml. borosilicate glass volumetric flask. The flask was very gently warmed until the liquid was reduced in volume to about 0.5 ml. Then the supporting electrolyte was added and the resulting solution deaerated for 5 minutes in a 10-ml. Lingane-Laitinen H-type polarographic electrolysis cell. The polarogram was then recorded and the concentration of uranium found from suitable calibration curves.

^a J. F. Steinbach, H. Freiser, Anal. Chem. 26, 375 (1954).

^b P. Souchay, J. Faucherre, Anal. Chim. Acta 3, 252 (1949).

PROCEDURE 12: Isolation and Measurement of Uranium at the Microgram Level.

Source: C. L. Rulfs, A. K. De, and P. J. Elving, Anal. Chem. 28, 1139 (1956).

Abstract

A double cupferron separation of uranium using extraction has been adapted to the micro level. Uranium(VI) does not extract in the first stage, which removes many potentially interfering elements. Uranium(IV), obtained in the residual aqueous solution by reduction at a mercury cathode, is simultaneously extracted as the cupferrate into ether, from which it can be re-extracted into

nitric acid. A relatively simple one-piece glass apparatus is used for all operations. The uranium recovery at the milligram level in an initial 30-ml. sample was determined colorimetrically as 94%. With 0.03 to 0.13 γ of radioactive uranium-233 tracer and 20 γ of natural uranium as carrier, the recovery is 86%; the latter includes the additional step of electrodeposition of the uranium onto a platinum planchet prior to measurement by alpha counting, which is only 94% complete. The decontamination possible with this procedure was checked with 0.07 γ quantities of uranium-233 in the presence of high mixed fission product activities; 85% recovery was obtained, containing only 0.9% of the fission product alpha activity (assumed to be uranium).

Apparatus

The reaction cell and simple electrical circuit used is shown in the accompanying figure.

The electrolysis vessel, C, is protected from mercury ions diffusing from the working reference calomel electrode, A, by a medium glass frit between B and C, and a fine frit backed with an agar plug between B and A. Between runs, cell C is kept filled with saturated potassium chloride solution.

The apparatus for the electrodeposition of uranium onto platinum disks or planchets and for alpha-counting measurement of the resulting uranium plates have been described.^a Beta activity was measured by a chlorine-quenched argon-filled Geiger-Müller counter (1.4 mg. per sq. cm. of window) with a Model 165 scaler; a scintillation well counter with a thallium-activated sodium iodide crystal and a Model 162 scaler was used for gamma-activity measurement of solutions (ca. 5 ml.) contained in a 13 x 150 mm. test tube. The scalers and counters are made by the Nuclear Instrument and Chemical Corp. For examination of the gamma-ray spectrum, a gamma-ray scintillation spectrometer (built

PROCEDURE 12 (Continued)

in the Department of Chemistry, University of Michigan) was used through the courtesy of W. Wayne Meinke.

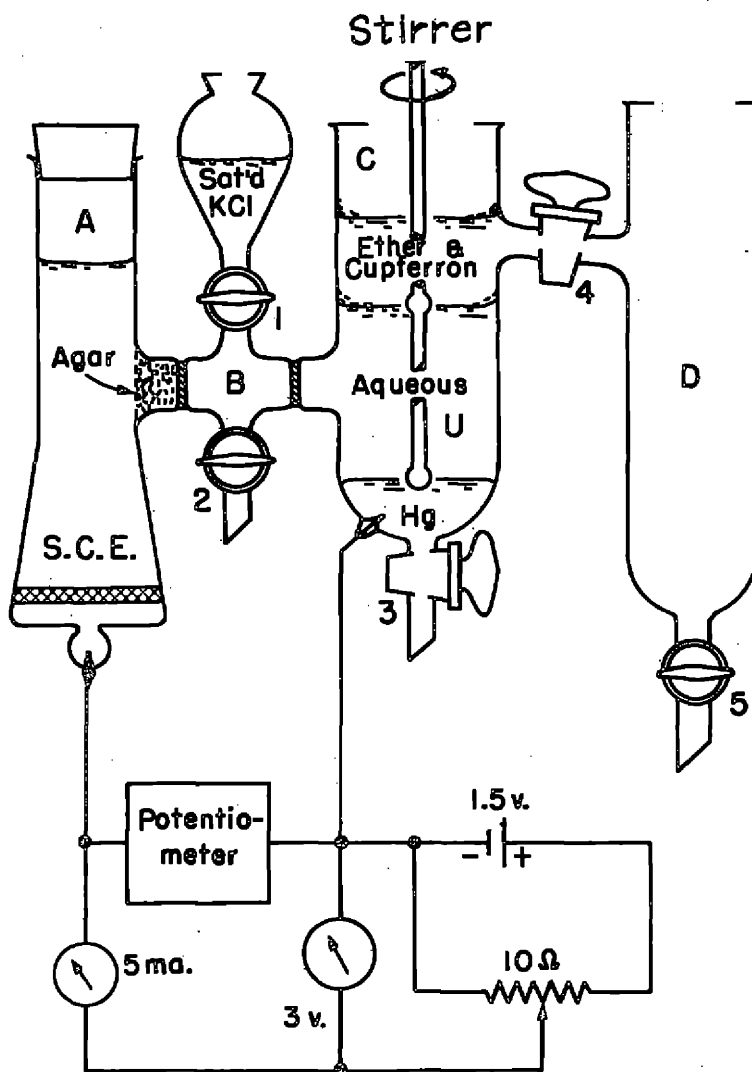
Reagents

All chemicals used were of C.P. or reagent grade unless otherwise specified. The ethereal cupferron solution used (200 to 300 mg. of cupferron per 50 ml.) was actually a hydrogen cupferrate solution; the ether and cupferron were mixed in a mixing cylinder with 5 to 10 ml. of 10 to 20% sulfuric acid and shaken until dissolution was complete.

Procedures

Reductive Extraction. At the commencement of a run, bridge B is flushed through stopcock 2 by filling B with fresh potassium chloride solution from the funnel through 1. C is drained and rinsed; 1 is left open for a time to flush the frit. With 3 closed, 4 to 5 ml. of triple-distilled mercury is placed in C. About 30 ml. of uranyl sulfate solution (0.5 to 5 mg. of uranium and 0.5 to 1.5% in sulfuric acid) is added and a potential of -0.35 volt vs. S.C.E. is applied to the mercury. About 15 to 20 ml. of the ether cupferron solution is added. Stirring is adjusted at just over the minimal rate for efficient current flow (usually about 0.2 ma. flows without stirring and 1.2 to 2.6 ma. with stirring).

Stopcock 1 is opened for about 30 seconds at approximately 5-minute intervals throughout the run to minimize any loss of uranium into the bridge. At 15- or 20-minute intervals, stirring is interrupted, the ether extract is bled through stopcock 4 into cell D, and 15 to 20 ml. of fresh ether-cupferron solution is added. Runs of 40- to 55-minute total duration appear to be adequate. Three increments of ether-cupferron solution were usually used, followed by a 5- to 10-ml. pure ether rinse at the conclusion of the run. (See Note 1.)



ELECTRICAL CIRCUIT FOR ELECTROCHEMICAL REDUCTION OF URANIUM FOR Procedure 12.

Extraction and Measurement at Microgram Uranium Level. A solution of uranium-233 (10^{-7} to 10^{-8} gram) together with about 20 γ of natural uranium (as sulfate) was submitted to reductive extraction with cupferron for about 50 to 60 minutes. The uranium (IV/III) cupferrate was then re-extracted in cell D from the ether solution into three successive 15-ml. portions of 7M

PROCEDURE 12 (Continued)

nitric acid. The combined nitric acid extract was evaporated to about 5 ml., treated with 25 to 30 ml. of concentrated nitric and 2 ml. of perchloric acid, and then evaporated to dryness. The residue was digested with 10 ml. of 0.1M nitric acid for a few minutes; the solution obtained, after addition of about 10 γ more of natural uranium (as sulfate), was used for electrodeposition of the uranium onto a platinum planchet from an oxalate medium.^a A windowless flow counter with Q-gas was used for counting the alpha emission from the electrodeposited uranium.^a

The whole operation took about 4 to 5 hours. Each measurement of alphas from the samples was calibrated by counting a uranium oxide standard (National Bureau of Standards No. 836-5).

Note 1. In some runs the current dropped to a low level soon after the requisite number of coulombs had passed for about a 3-electron reduction of the uranium present. In other cases, the current did not decrease, but discontinuance of the run beyond any point where twice the theoretical current had passed gave satisfactory uranium recovery. In the latter cases, a gray ether-insoluble, but alcohol-soluble precipitate (apparently a mercury cupferrate), was usually evident in the aqueous phase. The current efficiency for the desired process appeared to be good in most runs.

The combined ether extracts may be re-extracted in cell D by inserting a clean stirrer, or they may be transferred with rinsing into a clean separatory funnel. Three extractions with 20 to 30 ml. each of 0.5M, 4M, and 0.5M nitric acid were adequate to re-extract uranium into aqueous solution.

^a C. L. Rulfs, A. K. De, P. J. Elving, J. Electrochem. Soc. 104, 80 (1957).

PROCEDURE 13: The Determination of Uranium by Solvent Extraction.

Source: R. F. Clayton, W. H. Hardwick, M. Moreton-Smith and R. Todd, Analyst **83**, 13 (1958).

Abstract

The development of solvent-extraction methods for determining trace amounts of uranium-233 in irradiated thorium is described. Thorium and its alpha-emitting daughters are complexed with EDTA, and, when uranium-233 is extracted as its diethyldithiocarbamate complex, only bismuth-212 accompanies it. This is immaterial for colorimetric or fluorimetric finishes, but, for determination of the uranium-233 by alpha counting, the bismuth-212 must first be allowed to decay. If, however, the uranium-233 is extracted as its 8-hydroxyquinoline complex, no alpha emitter accompanies it and concentrations of uranium-233 ranging from 100 μg per ml down to 0.01 μg per ml in 0.7 M thorium solution have been determined in this way.

METHOD FOR DETERMINING URANIUM-233 IN THORIUM NITRATE SOLUTIONS
BY EXTRACTION WITH OXINE

REAGENTS--

Oxine solution A--A 10 per cent w/v solution of 8-hydroxyquinoline in isobutyl methyl ketone.

Oxine solution B--A 2.5 per cent w/v solution of 8-hydroxyquinoline in isobutyl methyl ketone.

EDTA solution--Dissolve 372.9 g of the disodium salt of ethylenediaminetetra-acetic acid in 500 ml of water containing 80 g of sodium hydroxide and make up to 1 liter.

1 ml \approx 232 mg of thorium.

Nitric acid, N.

Ammonia solution, sp. gr. 0.880.

Ammonia solution, 0.2 N.

Bromothymol blue indicator solution.

Anti-creeping solution--A 20 per cent solution of ammonium

PROCEDURE 13 (Continued)

chloride containing 2 per cent of a water-soluble glue (Stephen's Stefix was found to be suitable).

PROCEDURE FOR 0.01 TO 1 μ g OF URANIUM-233 PER ml--

With a pipette place a suitable volume of sample solution, containing not more than 600 mg of thorium, in a 40-ml centrifuge tube fitted with a glass stirrer. Add EDTA solution to give about a 10 per cent excess over the thorium equivalent and then add 3 drops of bromothymol blue indicator solution.

Add ammonia solution, sp. gr. 0.880, until the indicator turns blue. Return the color of the indicator to yellow by adding N nitric acid and then add 0.2 N ammonia solution until the color of the indicator just turns back to blue (pH 7). Add 2 ml of oxine solution A, stir for 5 minutes, spin in a centrifuge to separate the phases and then stopper the tube.

Evaporate duplicate 0.25-ml portions of the solvent phase slowly on stainless-steel counting trays that have had 1 drop of anti-creeping solution evaporated at their centers.^a Heat the trays to redness in the flame of a Meker burner, cool and count.

PROCEDURE FOR 1 TC 100 μ g OF URANIUM-233 PER ml--

With a pipette place a suitable volume of sample solution, containing about 10 μ g of uranium-233, in a 40-ml centrifuge tube and dilute to 3 ml with water. Add EDTA solution to give a 10 per cent excess over the thorium equivalent. Add 2 drops of bromothymol blue indicator solution and adjust the pH to 7 as previously described.

Add 5 ml of oxine solution B, stir for 5 minutes, spin in a centrifuge to separate the phases and then stopper the tube. Evaporate duplicate 0.1 or 0.25-ml portions of the solvent phase for counting, as before.

Note that for a fluorimetric finish to either procedure, suitable duplicate portions of the solvent phase should be

PROCEDURE 13 (Continued)

evaporated in platinum fluorimeter dishes before fusion with sodium fluoride.

METHOD FOR DETERMINING URANIUM-233 IN THORIUM NITRATE SOLUTIONS
BY EXTRACTION WITH SODIUM DIETHYLDITHIOCARBAMATE

REAGENTS--

Hexone.

Sodium diethyldithiocarbamate solution--A freshly prepared and filtered 20 per cent w/v aqueous solution.

EDTA solution--Prepared as described in reagents list, p. 377.

Ammonium nitrate solution, 2 M .

Ammonia solution, sp. gr. 0.880.

Nitric acid, concentrated and N.

Screened methyl orange indicator solution.

Anti-creeping solution--A 20 per cent solution of ammonium chloride containing 1 per cent of a water-soluble glue.

PROCEDURE FOR 1 TO 100 μ g OF URANIUM-233 PER ml--

With a pipette place a suitable volume of sample solution, containing about 10 μ g of uranium-233, in a 40-ml centrifuge tube fitted with a glass stirrer. Dilute to 4-ml with 2 M ammonium nitrate and add EDTA solution to give a 10 per cent excess over the thorium equivalent. Stir and make just alkaline to screened methyl orange by adding ammonia solution and then add 0.5 ml of sodium diethyldithiocarbamate solution.

Stir and add N nitric acid until the solution is mauve (not red). Add 5 ml of hexone, stir for 5 minutes and add more acid to maintain the mauve color if necessary. Spin in a centrifuge to separate the phases and then stopper the tube.

Evaporate suitable duplicate portions of the solvent phase on stainless-steel counting trays that have had 1 drop of anti-creeping solution evaporated at their centers. Heat the trays to redness in the flame of a Meker burner, allow the bismuth-212

PROCEDURE 13 (Continued)

to decay, and then count. Alternatively, for a fluorimetric finish, evaporate duplicate portions of the solvent phase in platinum fluorimeter dishes for fusion with sodium fluoride.

Note that greater sensitivity can be obtained by starting with a larger volume of sample or by evaporating larger portions of the solvent phase.

^a W. H. Hardwick, M. Moreton-Smith, *Analyst* **83**, 9 (1958).

PROCEDURE 14: Uranium Radiochemical Procedure Used at the University of California Radiation Laboratory at Livermore.

Source: E. K. Rulet, UCRL-4377 (1954).

Decontamination: 3×10^{12} Atoms of U^{237} isolated from a 3-day-old mixture containing 10^{14} fissions showed no evidence of contamination when decay was followed through three half lives.

Yield: 30 to 50 percent.

Separation time: About four hours.

Reagents: Dowex A-1 resin (fall rate of 5-30 cm/min).
Zinc dust, isopropyl ether.
 $2M$ $Mg(NO_3)_2$ with $1M$ HNO_3 .
 $2M$ HCl with $2M$ HF .

1. To the solution containing the activity in HCl in an Erlenmeyer flask, add uranium tracer in HCl , 1 ml of conc. formic acid and several ml of conc. HCl .
2. Boil gently until volume is ~2-3 ml, replenishing the solution with several ml of formic acid during the boiling.
3. Transfer to 20-ml Lusteroid centrifuge cone, rinsing flask twice with 1-2 ml of water and add 2-3 mg of La^{+++} . Add 1 ml of conc. HF , stir, heat, and centrifuge.

PROCEDURE 14 (Continued)

4. Transfer supernatant to 20-ml Lusteroid centrifuge cone, add 2-3 mg La^{+++} , stir, heat, and centrifuge.
5. Transfer supernatant to 20-ml Lusteroid, add 5-10 mg La^{+++} , 1/2 ml of conc. HCl and heat in a water bath. When the solution is hot, add small portions of zinc dust and stir. About three small additions of zinc dust should be made over a half-hour period with vigorous stirring after each addition. If the zinc dust tends to ball and sink to the bottom of the tube at any time, addition of more conc. HCl will solve this problem. When all the zinc from the last addition has dissolved, add 4 or 5 ml of water and 1 ml of conc. HF. Stir, cool, and centrifuge, retaining the LaF_3 precipitate. Wash precipitate with 2 ml of 2M HCl + 2M HF.
6. Dissolve the precipitate with 6M HNO_3 saturated with H_3BO_3 , stirring and heating. Transfer the solution to a 40-ml glass centrifuge cone, washing out the Lusteroid cone with water.
7. Add several drops of H_2O_2 and stir and heat. Add 2 mg Fe^{+++} and make solution basic with NH_4OH and some NaOH. Heat, stir, and centrifuge. Wash the precipitate with 2 ml of water.
8. Dissolve the precipitate in one to two drops of conc. HNO_3 and heat. Cool, add 10 ml 2M $\text{Mg}(\text{NO}_3)_2$ + 1M HNO_3 saturated with ether. Equilibrate twice with 10-15 ml of diisopropyl ether.
9. Pipet the ether phase into a clean 40-ml cone and equilibrate ether layer with 3 ml of conc. HCl. Pipet off and discard ether layer. Heat HCl for ~30 seconds and again pipet off the ether layer.
10. Pass the HCl solution through a Dowex-1 anion resin bed (2 cm x 3 mm) by pushing the solution through the column with air pressure. Rinse the centrifuge tube once with 1 ml of conc. HCl and push this solution through the column. Wash resin with ~1 ml of conc. HCl. Discard effluent.

PROCEDURE 14 (Continued)

11. To column, add $\sim 3/4$ ml of 0.5M HCl. Collect the eluate and evaporate to a small volume and plate on a platinum disc.

PROCEDURE 15: Use of Ion Exchange Resins for the Determination of Uranium in Ores and Solutions.

Source. S. Fisher and R. Kunin, Anal. Chem. 29, 400 (1957).

Abstract

The separation of uranium from the ions interfering with its analysis is accomplished by the adsorption of the uranium(VI) sulfate complex on a quaternary ammonium anion exchange resin. Interference of such ions as iron(III) and vanadium(V) is avoided by their preferential reduction with sulfurous acid so that they, as well as other cations, are not retained by the resin. Uranium is eluted for analysis by dilute perchloric acid. The method is applicable to both solutions and ores.

Ore Solution

Two methods for the opening of uranium-bearing ores were investigated in conjunction with the ion exchange separation. The first is the standard digestion with hydrofluoric and nitric acids, with subsequent evaporation to dryness followed by a sodium carbonate fusion.^a The carbonate melt is dissolved in 5% sulfuric acid to form a solution for the separation. A second method for routine analysis, designed to eliminate the need for hood facilities and platinum vessels, involves an oxidative leach with an acidic manganese(IV) oxide system. This procedure is given in detail below. Other workers^b, using the authors' separation procedure, have recommended solution of the ore by treatment with 12M hydrochloric acid plus 16M nitric acid followed by fuming

PROCEDURE 15 (Continued)

with sulfuric acid to produce a suitable uranium solution for the column influent.

Procedure. Weigh out samples of ore estimated to contain an amount of uranium oxide less than 100 mg. but sufficient to be detected by the chosen method of analysis. Add 20 ml. of 20% by volume sulfuric acid and 2 grams of manganese(IV) oxide. Heat the mixture to boiling. Allow to cool to room temperature. Dilute with approximately 50 ml. of water. Adjust to a pH between 1.0 and 1.5 by the dropwise addition of 20% sodium hydroxide. Filter through fine-pore filter paper using two 10-ml. portions of water to wash the residue on the paper.

Ion Exchange Separation

Apparatus. Tubes 0.5 inch in diameter with high-porosity sintered glass filter disks fused to the lower end are used to contain the resin. The rate of flow of solutions through the tube is regulated by a screw clamp on rubber tubing below the filter. Small separatory funnels are attached to the top of the column to feed the sample and reagents.

Procedure. Convert a portion of quaternary ammonium anion exchange resin (Amberlite XE-117, Type 2) of mesh size 40 to 60 (U.S. screens) to the sulfate form by treating a column of it with 10% sulfuric acid, using 3 volumes per volume of resin. Rinse the acid-treated resin with deionized water until the effluent is neutral to methyl red. Drain the resin so prepared free of excess water and store in a bottle. A 5-ml. portion of this resin is used for a single analysis. The resin is loaded into the filter tube and the bed so formed is backwashed with enough water to free it of air. After the resin has settled the excess water is drained off to within 1 cm. of the top of the bed prior to the passage of the sample through the bed.

Add 5 drops of 0.1% methylene blue to the partially neu-

PROCEDURE 15 (Continued)

tralized (pH 1.0 to 1.5) solution from the dissolved sodium carbonate melt or from the filtered manganese(IV) oxide leach. Add 6% sulfurous acid dropwise until the methylene blue is decolorized and then add a 5-ml. excess. Pass the reduced sample through the resin bed at a rate not exceeding 2 ml. per minute. Wash the sample container with two 10-ml. portions of water, passing the washing through the resin bed at the same flow rate. Elute the uranium with 50 ml. of 1M perchloric acid. Determine the uranium content of the perchloric acid fraction colorimetrically by the standard sodium hydroxide-hydrogen peroxide method^a or volumetrically^c. For colorimetric analysis standard uranium solutions containing perchloric acid should be used in establishing the curve.

^a F. S. Grimaldi, I. May, M. H. Fletcher, J. Titcomb, U. S. Geol. Survey Bull. 1006 (1954).

^b H. J. Seim, R. J. Morris, D. W. Frew, U. S. Atomic Energy Comm. Document UN-TR-5 (1956).

^c G. J. Rodden, "Analytical Chemistry of the Manhattan Project," McGraw-Hill, New York, 1950.

PROCEDURE 16: The Use of a "Compound" Column of Alumina and Cellulose for the Determination of Uranium in Minerals and Ores Containing Arsenic and Molybdenum.

Source: W. Ryan and A. F. Williams, Analyst 77, 293 (1952).

Abstract

A technique in inorganic chromatography, with alumina and cellulose adsorbents in the same extraction column, is described for the separation of uranium from minerals and ores. The purpose of the alumina is to retain arsenic and molybdenum, which

PROCEDURE 16 (Continued)

are not readily retained by cellulose alone when ether containing 5 per cent v/v of concentrated nitric acid, sp.gr. 1.42, is used as the extraction solvent.

METHOD FOR SAMPLES CONTAINING MOLYBDENUM OR ARSENIC OR BOTH

Solvent--Add 5 ml of concentrated nitric acid, sp.gr. 1.42, to each 95 ml of ether.

PREPARATION OF ALUMINA-CELLULOSE COLUMN--

The adsorption tube for the preparation of the column is a glass tube about 25 cm long and 1.8 cm in diameter. The upper end is flared to a diameter of about 8 cm to form a funnel that permits easy transfer of the sample. The lower end terminates in a short length of narrow tubing and is closed by a short length of polyvinyl chloride tubing carrying a screw clip. The inside surface of the glass tube is coated with a silicone in the manner described by Burstall and Wells.^a

Weigh 5 or 6 g of cellulose pulp* into a stoppered flask and cover it with ether-nitric acid solvent. Pour the suspension into the glass tube, agitate gently and then gently press down the cellulose to form a homogeneous column. Wash the column with about 100 ml of ether-nitric acid solvent and finally allow the level of the solvent to fall to the top of the cellulose. Next pour about 15 g of activated aluminium oxide† on top of the cellulose, pour on 30 ml of ether-nitric acid solvent and vigorously agitate the alumina with a glass plunger. Allow the packing to settle. Allow the level of the ether to fall to the surface of the alumina and the column is ready for use.

* Whatman's Standard Grade cellulose powder is suitable.

† Type H, Chromatographic Alumina, 200 mesh. Supplied by Peter Spence Ltd.

PROCEDURE 16 (Continued)

PREPARATION OF SAMPLE SOLUTION FROM MINERAL OR ORE--

Weigh into a platinum dish sufficient of the sample to give 100 to 150 mg of U_3O_8 , which is a convenient amount for a volumetric determination. Decompose the sample by treatment with nitric and hydrofluoric acids in the manner described by Burstall and Wells.^a Finally remove hydrofluoric acid by repeated evaporations with nitric acid and take the sample to dryness. If the addition of dilute nitric acid indicates the presence of undecomposed material, filter the insoluble residue on to a filter-paper and ignite and fuse it in a nickel crucible with a few pellets of potassium hydroxide. Then add the melt to the filtrate and take the whole to dryness.

Add 4 ml of diluted nitric acid, 25 per cent v/v, to the dry residue, gently warm to dissolve the mixture and then cool the solution, which is then ready for chromatography.

EXTRACTION OF URANIUM--

Transfer the sample on a wad of dry cellulose pulp to the top of the prepared alumina-cellulose column^a and extract the uranium with 200 to 250 ml of ether-nitric acid solvent if arsenic or molybdenum and arsenic is present in the original sample. If molybdenum alone is present, the amount of solvent can be reduced to 150 ml. Screen the column from direct sunlight. After removal of ether from the eluate, determine the uranium volumetrically.^{a, b}

^a F. H. Burstall, R. A. Wells, Analyst 76, 396 (1951).

^b F. H. Burstall, A. F. Williams, "Handbook of Chemical Methods for the Determination of Uranium in Minerals and Ores," H. M. Stationery Office, London, 1950.

PROCEDURE 17: Determination of Uranium-235 in Mixtures of Naturally Occurring Uranium Isotopes by Radioactivation.
Source: A. P. Seyfang and A. A. Smales, Analyst 78, 394 (1953).

Abstract

A method previously used^a for determining uranium in minerals by neutron irradiation followed by measurement of the separated fission-product barium has been extended to the determination of uranium-235 in admixture with uranium-234 and uranium-238.

With microgram amounts of uranium-235, short irradiations in the Harwell pile give ample sensitivity. Precision and accuracy of better than ± 2 per cent have been achieved for a range of uranium-235 contents covered by a factor of more than 10^5 .

Method

REAGENTS--

Magnesium oxide--Analytical reagent grade.

Nitric acid, sp. gr. 1.42.

Barium chloride solution--Dissolve 18 g of $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ in water and make up to 500 ml.

Lanthanum nitrate solution--A 1 per cent w/v solution of $\text{La}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$.

Ammonium hydroxide, sp. gr. 0.880.

Strontium carbonate solution--A 2 per cent w/v solution.

Hydrochloric acid-diethyl ether reagent--A mixture of 5 parts of concentrated hydrochloric acid, sp. gr. 1.18, and 1 part of diethyl ether.

Sodium tellurate solution--A 0.4 per cent w/v solution.

Zinc metal powder.

Methyl orange indicator.

Potassium iodide solution--A 1 per cent w/v solution.

Sodium hypochlorite solution--A commercial solution containing 10 per cent of available chlorine.

PROCEDURE 17 (Continued)

Hydroxylamine hydrochloride.

Ferric chloride solution--A 1 per cent w/v solution.

Sulfuric acid--A 20 per cent v/v solution.

IRRADIATION--

Solids-- Samples containing not much more uranium-235 than natural uranium (say, up to three times more or 2 per cent) may be irradiated as solid; this is usually U_3O_8 . For these cut a 5-cm length of 2-mm polythene tubing and seal one end by warming and pressing. Introduce freshly ignited analytical reagent grade magnesium oxide to form a compact layer 4 to 5 mm in height at the sealed end of the tube. Weigh the tube and contents, add about 50 mg of U_3O_8 and re-weigh. Add a further similar layer of magnesium oxide on top of the U_3O_8 and then seal the open end. Leave a free space about 1 cm long between the top of the magnesium oxide layer and the seal, for ease of opening. Treat standard and samples similarly. Place the tubes either in a special polythene bottle for irradiation in the pneumatic "rabbit" of the pile or in a 3-inch aluminum can for irradiation in the "self-serve" holes in the pile. Irradiation is carried out for any required time; usually it is about 5 minutes. After irradiating, place the containers in lead shielding for about 15 hours. After this period, tap down the contents of the polythene tube away from one end and carefully cut off the top. Empty the contents into a 50-ml centrifuge tube. (The plug of magnesium oxide serves to "rinse" the sample tube as it is emptied.) Add 2 ml of concentrated nitric acid (sp.gr. 1.42), gently warm to dissolve, and finally boil off the nitrous fumes. Add 5.00 ml of a barium solution to act as carrier (a solution of 18 g of barium chloride, $BaCl_2 \cdot 2H_2O$, in 500 ml of water).

Liquids--For more highly enriched samples or when the amount of sample available is small, solutions containing weighed

PROCEDURE 17 (Continued)

quantities of solid sample must be irradiated in small silica ampoules. The ampoules, which have a capacity of about 1 ml, are prepared from silica tubing. After one end of each has been sealed, the ampoules are weighed, the sample solution added from a fine-pointed glass dropping-tube and the ampoules re-weighed. Pack the ampoules, after sealing the open ends, in cotton wool in a 3-inch aluminium can and irradiate them in the "self-serve" position of the pile. The time of irradiation necessary can be calculated from the usual activation formula; as an example, 1 μg of uranium-235 irradiated for 24 hours in a flux of 10^{12} neutrons per sq. cm per second gives about 5000 counts per minute of barium-140 at 5 per cent counting efficiency, 24 hours after the irradiation.

After removing them from the pile, place the samples and standards in lead shielding for about 15 hours; the main activity is due to silicon-31. Transfer the ampoules to 100-ml tall-form beakers containing a few milliliters of water and 5.00 ml of barium carrier solution, carefully break off both ends of each ampoule and warm to ensure thorough mixing. Decant into centrifuge tubes and wash out the ampoules and beakers with further small portions of water.

CHEMICAL SEPARATION--

Evaporate the solution containing the irradiated uranium and barium carrier to 5 to 6 ml and add two drops of 1 per cent lanthanum nitrate solution. Warm if necessary to dissolve any barium nitrate that may have crystallized, add concentrated ammonium hydroxide dropwise until a permanent precipitate is obtained and then two drops in excess. Centrifuge and decant into another centrifuge tube. Add methyl orange indicator, and then hydrochloric acid until the solution is acid. Add 2 drops of 2 per cent strontium solution, about 25 ml of hydrochloric

PROCEDURE 17 (Continued)

acid - diethyl ether reagent, mix thoroughly, centrifuge and decant. Wash the precipitate with 5 ml of reagent, centrifuge and decant. Dissolve the barium chloride precipitate in 3 to 4 ml of water, re-precipitate it by adding 20 ml of reagent, centrifuge and decant. Wash with 5 ml of reagent, centrifuge and decant.

Dissolve the precipitate in about 5 ml of water, add 6 drops of lanthanum solution and 6 drops of the 4 per cent tellurate solution and then about 3 mg of zinc metal powder. When the effervescence ceases, make the solution just ammoniacal to methyl orange, centrifuge and decant into another tube. Add 4 drops of 1 per cent potassium iodide solution and 2 drops of sodium hypochlorite solution. Warm and set aside for 2 minutes. Acidify with about 1 ml of hydrochloric acid, and add about 0.1 g of hydroxylamine hydrochloride. Boil under a hood until all the iodine appears to be removed and the volume is reduced to 5 to 6 ml. Add 2 drops of strontium solution and 2 drops of lanthanum solution and repeat the double barium chloride precipitation and washing, as above.

Dissolve the precipitate in about 5 ml of water, add 6 drops of lanthanum solution, and 6 drops of 1 per cent ferric chloride solution. Make ammoniacal to methyl orange, add half a crumbled Whatman accelerator tablet, and heat just to boiling. Filter through a 7-cm Whatman No. 30 filter-paper into a centrifuge tube, wash twice with 2 to 3-ml portions of water. Dilute the filtrate to about 20 ml and make slightly acid with hydrochloric acid. Heat nearly to boiling and add dropwise 2 ml of 20 per cent v/v sulfuric acid. Allow the precipitate to settle, decant, wash with 10 ml of water, centrifuge, decant and repeat the washing procedure to complete removal of the excess of acid.

Transfer as much as possible of the precipitate, by means of a dropping tube and a few drops of water, to a tared aluminium

PROCEDURE 17 (Continued)

counting tray. Dry under an infra-red lamp and finally heat in a muffle furnace at 500°C for 15 minutes. Cool, weigh and reserve for counting.

COUNTING TECHNIQUE--

The counting equipment for this work consists of (i) a power unit (type 1082A or 200 is suitable), (ii) scaling unit (type 200 or 1009B), (iii) time accessory unit (type 1003B), (iv) probe unit (type 200B or 1014A). Time pulses can be obtained from a master electric clock serving several units. A mica end-window Geiger-Müller counter (2 mg per sq. cm.), of type EHM2, is suitable; it is mounted in a lead castle with a Perspex lining and shelves.

Check the counting equipment in the normal fashion with a suitable beta-emitter, such as natural uranium oxide in equilibrium with UX_1 and UX_2 . Place the sample to be counted in a Perspex carrier and insert it in a shelf at a suitable distance from the Geiger-Müller tube to attain a counting rate of 2000 to 3000 counts per minute. Count for a sufficient time to obtain at least 10,000 counts for each barium sulfate precipitate, counting the precipitates one after another without undue delay. Correction for growth of lanthanum-140 is unnecessary if samples and standards are counted within, say, 60 minutes of each other provided the barium sulfate precipitations are carried out on each nearly simultaneously.

CALCULATION OF RESULTS--

Correct all counts for background, coincidence loss and chemical yield and express as the results in counts per minute.

$$\text{Then } \frac{\text{Weight of } U^{235} \text{ in standard}}{\text{Weight of } U^{235} \text{ in sample}} = \frac{\text{Corrected count of standard}}{\text{Corrected count of sample}}$$

$$\text{and } \frac{\text{Weight of } U^{235} \text{ in sample}}{\text{Weight of sample}} \times 100 = \text{percentage of uranium-235 in sample.}$$

^a A. A. Smales, Analyst 77, 778 (1952).

PROCEDURE 18: Determination of Microgram and Submicrogram
Quantities of Uranium by Neutron Activation Analysis.

Source: H. A. Mahlman and G. W. Ledicotte, Anal. Chem. 27, 823
(1955).

Abstract

Microgram and submicrogram quantities of uranium have been determined in synthetic samples, ores, and soils by neutron radioactivation analysis. The principles of the activation analysis method used in this determination and the processing of irradiated samples are discussed. This method of analysis is a sensitive and specific method for determining uranium in concentrations as small as 0.1 γ per gram with a probable relative standard error of 10%. Concentrations of uranium in quantities as small as 0.0001 γ per gram can be determined by neutron activation analysis.

Radioactivation Analysis of Samples that Contain
Uranium

Nuclear Irradiation of Sample. Weighed portions of the samples and the comparative standard are put into small quartz tubes. The tubes are closed with cork stoppers that are wrapped in aluminum. They are then irradiated in the reactor. After irradiation, the samples are allowed to decay about 4 hours and are then chemically processed as described below. The synthetic samples used in this laboratory had been processed by a filter paper partition chromatography technique. After the separation, the paper was conveniently irradiated in short pieces of quartz tubing whose openings were plugged by means of cork stoppers.

Chemical Separation of Neptunium-239. In most neutron activation analyses, a chemical separation is made to isolate the radioactivity of the element from all other radioactive species in the sample. Usually an "isotopic carrier"--a known amount of the natural inactive element--is added to the solutions

PROCEDURE 18 (Continued)

of both the irradiated specimen and the comparison samples. The solutions are then processed chemically to isolate the carrier and desired radioelement from other elements and contaminant radioactivities. Small amounts of other elements are added as holdback or scavenging carriers to assist in the decontamination process.

Although neptunium-239 has a convenient half life, it does not have a stable isotope that can be used as an isotopic carrier. However, Seaborg^a has shown that trace quantities of neptunium-239 can be quantitatively carried on a nonisotopic carrier, such as cerium. The method of analysis reported below uses lanthanum as a nonisotopic carrier for the neptunium-239 radioactivity. (See Note 1.)

Chemical Separation Procedure. PREPARATION. The irradiated ore and soil specimens are dissolved by digestion in a mixture of concentrated nitric, hydrofluoric, perchloric, and sulfuric acids. (Additional hydrofluoric acid can be added if a residue of silica remains in the bottom of the crucible.) After dissolution, the sample is concentrated to heavy sulfuric acid fumes, cooled, and transferred to a 15-ml. centrifuge tube. If a residue (sulfate salts) remains after the transfer, the solution is centrifuged for 5 minutes, the supernatant transferred to another tube, and the residue washed with 1 ml. of 1M nitric acid. The wash is added to the supernatant and the residue discarded. (Centrifugation is always for the stipulated time and at full speed.) The sample is then further processed by the procedure reported herein.

The irradiated synthetic samples (paper chromatograms) are processed by carefully igniting the paper contained in a porcelain crucible in a muffle furnace. The residue is dissolved in about 0.5 ml. of concentrated nitric acid. After dissolution, the

PROCEDURE 18 (Continued)

sample is transferred to a 15-ml. centrifuge tube and the processing continued with the procedure reported herein.

PROCEDURE. Three (3.0) milligrams of lanthanum and 0.250 ml. of 5M hydroxylamine hydrochloride solution are added to the supernatant solution and the mixture digested for 5 minutes with occasional stirring. The solution is cautiously neutralized with concentrated ammonium hydroxide to precipitate lanthanum hydroxide, after which the mixture is centrifuged and the supernatant liquid discarded.

The precipitate of lanthanum hydroxide is dissolved in 2 ml. of 2M hydrochloric acid, and 1.0 mg. of strontium (added as a solution of strontium nitrate to serve as a holdback or scavenging carrier) and 0.250 ml. of 5M hydroxylamine hydrochloride solution are added to the solution. The solution is again digested for 5 minutes with intermittent stirring, and 0.200 ml. of concentrated hydrofluoric acid is added dropwise to the solution to precipitate lanthanum fluoride. After centrifugation, the supernatant liquid is discarded and the precipitate washed with 0.5 ml. of 1M hydrofluoric acid-1M nitric acid solution.

After washing, the lanthanum fluoride precipitate is dissolved in 0.5 ml. of saturated boric acid solution and 1.0 ml. of 6M nitric acid. One (1.0) milliliter each of 10% potassium permanganate solution and water are added to this solution, and the resulting mixture is agitated well and digested for 5 minutes. Lanthanum fluoride is again precipitated with 0.250 ml. of concentrated hydrofluoric acid; the solution is centrifuged and the supernatant liquid transferred to another centrifuge tube. The precipitate is washed with 0.5 ml. of 1M hydrofluoric acid-1M nitric acid solution and the wash combined with the supernatant liquid. The precipitate is discarded.

Three milligrams of lanthanum are added to the supernatant

PROCEDURE 18 (Continued)

liquid, and the solution is digested for 5 minutes and centrifuged. An additional 3.0 mg. of lanthanum are added to the supernatant liquid and the solution agitated and digested for 5 minutes without disturbing the first precipitate on the bottom of the tube; then the solution is centrifuged and the supernatant liquid transferred to another centrifuge tube. The precipitate is washed with 0.5 ml. of 1M hydrofluoric acid-1M nitric acid solution; centrifuged, and the wash combined with the supernatant liquid. The precipitate is discarded.

One milligram of zirconium (added as a solution of zirconium nitrate to serve as a holdback or scavenging carrier) and 0.250 ml. of 5M hydroxylamine hydrochloride are added to the solution and the mixture agitated and digested 5 minutes. Three (3.0) milligrams of lanthanum and 2 ml. of 2M hydrofluoric acid are added to the solution, and the solution is digested for 20 minutes and then centrifuged. The supernatant liquid is discarded. The precipitate is washed with 0.5 ml. of 1M hydrofluoric acid-1M nitric acid solution, and the resulting mixture is centrifuged. The wash solution is discarded after the centrifugation.

The precipitate is slurried in a small amount of 1M nitric acid (about 0.5 ml.) and transferred to a small borosilicate glass culture tube by means of a transfer pipet. The centrifuge cone is rinsed with three 0.5-ml. portions of 1M nitric acid and the rinses transferred to the culture tube. The tube is stoppered with a cork stopper and the γ radioactivity measured by a well-type gamma scintillation counter.

The standard sample of uranium oxide (U_3O_8) is dissolved in nitric acid and an aliquot of the solution processed under the same conditions as the specimen samples. The uranium content of the sample in question is determined by equating the ratio of the corrected neptunium-239 radioactivity count in the unknown

PROCEDURE 18 (Continued)

and the corrected neptunium-239 radioactivity count in the standard sample.

Note 1. Hamaguchi and co-workers^b have used Np²³⁷ tracer to determine the chemical yield.

^a G. T. Seaborg and co-workers, Metallurgical Project Rept. CN-2689,41 (Feb. 15, 1945) (classified).

^b H. Hamaguchi, G. W. Reed, A. Turkevich, Geochim. et Cosmochim. Acta 12, 337 (1957).

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